

Published in final edited form as:

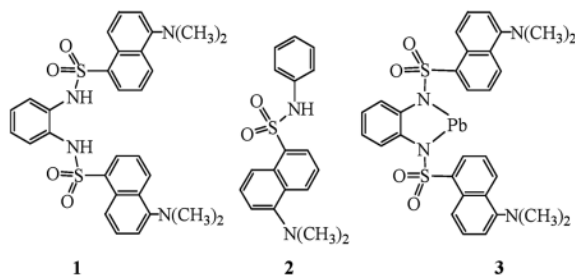
*J Am Chem Soc.* 2005 May 11; 127(18): 6514–6515. doi:10.1021/ja050296e.

## Fluorescent Sensing and Selective Pb(II) Extraction by a Dansylamide Ion-Exchanger

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Lead poisoning, which is especially prevalent in children, is still the most common environmentally caused disease nationally and worldwide.<sup>1</sup> Development of extractants and fluorescent chemosensors<sup>2</sup> for Pb(II) is of importance for monitoring EPA mandated levels and for a variety of applications related to lead toxicity, biodistribution, and removal. For such applications, it is essential to achieve high selectivity for Pb(II) against both alkaline earth and other common metal ions present in biological systems, such as Cu(II) and Zn(II).<sup>3</sup> Exploiting the unique coordination properties of Pb(II)<sup>3,4</sup> provides opportunities for the design of selective sensors<sup>5</sup> and extractants.<sup>6</sup> Successful fluorescent Pb(II) sensor designs are based on complexation by peptides,<sup>7</sup> ionizable chelates,<sup>8</sup> DNAzymes,<sup>9</sup> or macrocycles<sup>10</sup> with covalently linked fluorophores. Macrocycles can be tedious to synthesize and typically exhibit a wide variety of polydentate Pb(II) coordination patterns. On the other hand, ion exchange extraction from water to an organic solvent by simpler chelates can lead to improved selectivity by taking advantage of the unique preference of Pb(II) for low-coordinate hemidirected<sup>3</sup> geometries with a stereochemically expressed lone pair. Moreover, simpler chelates tend to be easier to synthesize, an advantage that could potentially be exploited for sensor discovery by using rapid screening methods.<sup>11</sup> As part of our effort to design low-coordinate Pb(II) extractants and sensors, we now wish to report efficient and selective ion-exchange extraction of Pb(II) from water into 1,2-dichloroethane (DCE) with concurrent fluorescence quenching using as an ion-exchanger, the sulfonamide fluorophore **1**. This simple system does not require a secondary co-ligand in order to extract Pb and shows remarkable extraction selectivity against other metals with  $D_{\text{Pb}} > 130 D_{\text{Cu}}$  and  $D_{\text{Pb}} > 1400 D_{\text{Zn}}$ .



Sulfonamide **1** is available in good yields from *o*-phenylenediamine and 1-(dimethylamino)-5-naphthalene sulfonyl chloride (dansyl chloride).<sup>12</sup> The dansyl group is a well-known

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**Supporting Information Available:** Synthesis, procedures, extraction plots, and X-ray crystallographic details of **1**. This material is available free of charge via the Internet at <http://pubs.acs.org>.

fluorophore that is sensitive to its chemical environment.<sup>13</sup> Godwin et al.<sup>7</sup> and Leray et al.<sup>10a</sup> have reported elegant examples of Pb(II) fluorescence detection in aqueous matrices based on the dansyl fluorophore. Reinhoudt et al.<sup>8b</sup> have reported quenching of the dansyl fluorescence by Pb(II) on a self-assembled monolayer (SAM)-modified glass surface. The X-ray crystal structure of **1** shows the S=O groups located in an anti conformation with self-association via N–H···O=S hydrogen bonds, giving dimers, instead of the 1D chains observed before for analogous disulfonamides.<sup>14</sup>

Ligand **1** was shown to extract Pb(II) from water into DCE via ion-exchange when used together with NH(*i*-Pr)<sub>2</sub>. One clear advantage of **1** over the analogous sulfonamide extractants, 1,2-C<sub>6</sub>H<sub>4</sub>(NH<sub>2</sub>SO<sub>2</sub>C<sub>6</sub>H<sub>5</sub>)<sub>2</sub> and 1,2-C<sub>6</sub>H<sub>4</sub>(NH<sub>2</sub>SO<sub>2</sub>C<sub>6</sub>H<sub>4</sub>-*p*-Bu<sup>t</sup>)<sub>2</sub><sup>15</sup>, is that the formed binary complex, Pb[1,2-C<sub>6</sub>H<sub>4</sub>(NSO<sub>2</sub>C<sub>10</sub>H<sub>6</sub>-5-N(CH<sub>3</sub>)<sub>2</sub>)<sub>2</sub>] (**3**), is soluble in DCE, and therefore, there is no need for the use of a co-ligand. <sup>1</sup>H NMR of the organic phases after extraction indicated the presence of **3**. The negative mode electrospray ionization mass spectrum of a methanol solution of **3** showed the formation of the **3**-CH<sub>3</sub>O<sup>-</sup> ion at *m/z* = 811.

The distribution ratios ( $D_{\text{Pb}} = [\text{Pb}]_{\text{o}}/[\text{Pb}]_{\text{aq}}$ ) were determined by inductively coupled plasma mass spectrometry (ICP-MS) of the aqueous phases after extraction (for [Pb]<sub>aq</sub>) and after stripping the organic phases with 0.1 M HNO<sub>3</sub> (for [Pb]<sub>o</sub>). Double stripping of the organic phases showed no increase in [Pb]<sub>o</sub>. Specifically, when **1** (3.5 mM) was used together with 2.2 equiv of NH(*i*-Pr)<sub>2</sub>, it extracted 99.5% of Pb from a 3.5 mM Pb(NO<sub>3</sub>)<sub>2</sub> solution. A plot of [Pb]<sub>o</sub> versus [1]<sub>t</sub> showed saturation for [1]<sub>t</sub> > 3.5 mM, indicating 1:1 complexation. Control experiments, using **1** or NH(*i*-Pr)<sub>2</sub> only, showed that the presence of both components is essential for extracted. The monodansylamide **2**<sup>16</sup> did not extract any Pb(II) under the same conditions. A comparison of **1** with the ion-pair extractant 18-crown-6 under identical conditions gave a  $D_{\text{Pb}} = 0.950$  for extraction by **1** versus  $D_{\text{Pb}} = 1.26 \times 10^{-3}$  by the crown ether, which is also higher than the extraction by synergistic combinations of 2,2'-bipy and analogous disulfonamide ion-exchangers.<sup>15</sup> Ligand **1** was also shown to be effective in the extraction of micromolar Pb(II) concentrations, with [1]<sub>t</sub> = 2.97 mM, [NH(*i*-Pr)<sub>2</sub>]<sub>t</sub> = 6.53 mM, and [Pb(II)]<sub>t</sub> between 3.75 and 284 μM ([1]<sub>t</sub>/[Pb(II)]<sub>t</sub> ranging from 10.5 to 792); ligand **1** extracted 100% of Pb(II) ([Pb]<sub>aq</sub> < 0.01 μM).

The extraction of various dicationic metal nitrates was investigated at constant total metal concentration ([M(II)]<sub>t</sub>) and varying ligand concentrations (0.60–2.99 mM). Ligand **1** was found to extract Pb(II) selectively over Zn(II), Co(II), Cd(II), Ni(II), and Cu(II) with  $D_{\text{Pb(II)}}/D_{\text{M(II)}} = 1410$  (for Zn), 1380 (for Co), 829 (for Cd), 794 (for Ni), and 133 (for Cu) (Figure 1). Neither Ca(II) nor Na(I) were extracted to any appreciable extent. We ascribe the observed selectivity to the formation of the low-coordinate complex **3**, which is apparently disfavored for other competing metals. The crystal structure of the binary Pb complex of the analogous bis-(phenylsulfonyl) derivative of phenylenediamine, which is insoluble in DCE, has showed the formation of a coordination polymer via S=O–Pb axial coordination and a stereochemically active lone pair.<sup>15</sup> The fact that the dansyl complex **3** is soluble in DCE indicates that the S=O–Pb axial coordination is no longer significant and is most likely replaced by weak coordination from solvent molecules or water. The formation of a low-coordinate complex with a stereochemically active lone pair could possibly explain the high extraction selectivity for Pb(II) over other dicationic metals.

The extraction of Pb(II) from water into DCE by **1** or the control **2** was also investigated by fluorescence spectroscopy. Emission spectra of the organic phases were collected before and after contact with Pb(NO<sub>3</sub>)<sub>2</sub>(aq) under identical conditions to the ICP-MS distribution experiments. Fluorescence quenching was observed at 516 nm after contact of a solution of **1** and NH(*i*-Pr)<sub>2</sub> in DCE with Pb(NO<sub>3</sub>)<sub>2</sub>(aq). Specifically, the fluorescence intensity was reduced by as much as 29% upon contact of **1** (10 mM) and NH(*i*-Pr)<sub>2</sub> (22 mM) with 5.5 mM

Pb(NO<sub>3</sub>)<sub>2</sub> (Figure 2). On the other hand, a solution of **2** and NH(*i*-Pr)<sub>2</sub> in DCE showed no changes in its fluorescence upon contact with Pb(NO<sub>3</sub>)<sub>2</sub> (Figure 3). This is consistent with the ICP–MS results, which showed no extraction by **2**, and indicates that bidentate coordination is necessary for Pb(II) extraction, and that the observed quenching is a direct result of Pb(II) complexation by **1**. Overall, the Pb selectivity for this remarkably simple extraction-based system could be considered comparable, or even superior to previously reported designs.<sup>7, 10a</sup> However, in our case, Pb sensing is based on quenching and not on ratiometric fluorescence enhancement.

In conclusion, it has been demonstrated that a simple (bis)-dansylamide ion-exchanger selectively extracts Pb(II) from water into DCE, with concurrent fluorescence quenching. We are planning extensive structural and spectroscopic studies with **1** and analogous chelates, as well as application of the ion-exchange extraction methodology for rapid screening of potential extractants and sensors for Pb(II) and other toxic metals.

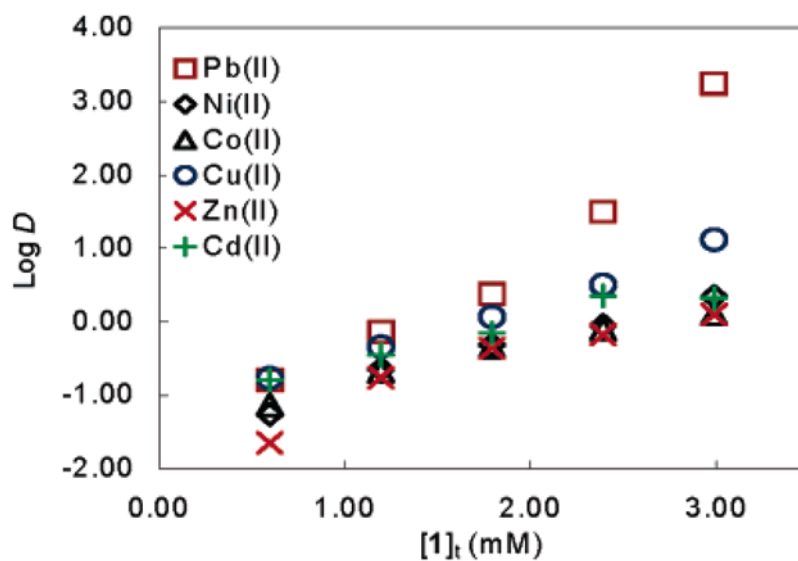
## Acknowledgements

We thank Drs. R. Alvarado and W. Zhang for experimental help. This project was supported in part by NIH-NIEHS/ARCH (S11 ES11181), NIH-MBRS/RISE (R25 GM061347), and the University of Miami (diffractometer fund).

## References

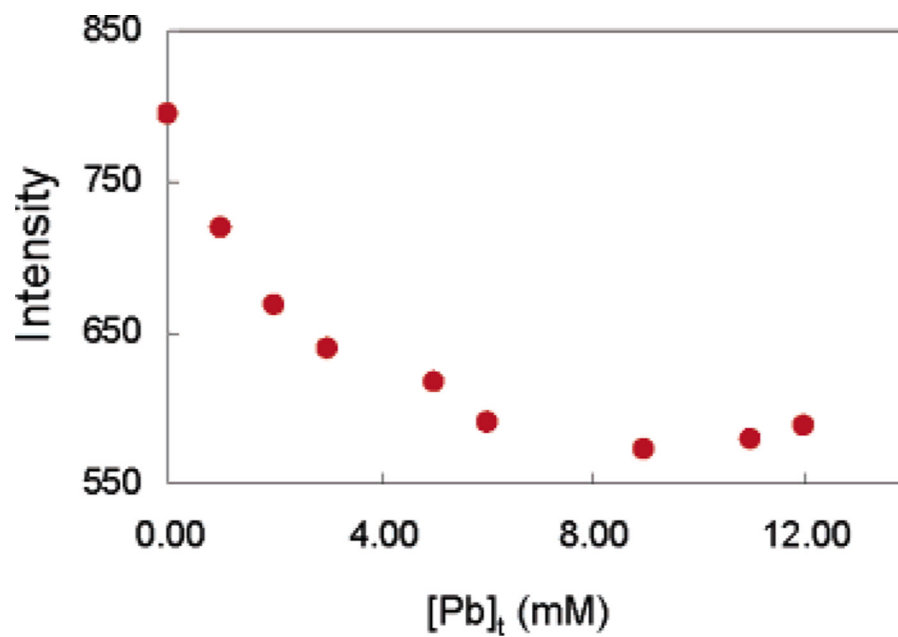
1. Lin-Fu, JS. Lead Poisoning, A Century of Discovery and Rediscovery. In: Needleman, HL., editor. Human Lead Exposure. Lewis Publishing; Boca Raton, FL: 1992.
2. (a) EV, Anslyn. Synthetic Receptors as Sensors. Tetrahedron 2004;60:11051. Czarnik, AW., editor. Fluorescent Chemosensors for Ion and Molecule Recognition. American Chemical Society; Washington, DC: 1993. ACS Symposium Series 358 (c) de Silva AP, Gunaratne HQN, Gunnaugsson T, Huxley AJM, Mccoy CP, Rademacher JT, Rice TE. Chem Rev 1997;97:1515. [PubMed: 11851458] (d) Fabbri L, Poggi A. Chem Soc Rev 1995;24:197.
3. Claudio ES, Godwin HA, Magyar JS. Prog Inorg Chem 2003;51:1.
4. Harrison, PG. Comprehensive Coordination Chemistry. Wilkinson, G.; Grillard, RD.; McCleverty, JA., editors. 3. Pergamon Press; Oxford, U.K.: 1987. p. 183 (b) Parr J. Polyhedron 1997;16:551. (c) Powers RE, Fuller WL, Raymond KN. Comp Supramol Chem 1996;10:537.
5. (a) Valeur B, Leray I. Coord Chem Rev 2000;205:3. (b) de Silva AP, Fox DB, Huxley AJM, Moody TS. Coord Chem Rev 2000;205:41. (c) Liu J, Lu Y. J Am Chem Soc 2004;126:12298. [PubMed: 15453763] (d) Sun Y, Wong MD, Rosen BP. J Biol Chem 2001;276:14955. [PubMed: 11278706]
6. (a) Hancock RD, Maumela H, De Sousa AS. Coord Chem Rev 1996;148:315. (b) Yordanov AT, Roundhill DM. Coord Chem Rev 1998;170:93. (c) Talanova GG, Hwang HS, Talanov VS, Bartsch RA. Chem Commun 1998:419.
7. Deo S, Godwin H. J Am Chem Soc 2000;122:174.
8. (a) Chae MY, Yoon J, Czarnik AD. J Mol Recognit 1996;9:297. [PubMed: 9131471] (b) Crego-Calama M, Reinhoudt DN. Adv Mater 2001;13:1171.
9. (a) Li J, Lu Y. J Am Chem Soc 2000;122:10466. (b) Liu J, Lu Y. Anal Chem 2003;75:6666. [PubMed: 14640743]
10. (a) Métivier R, Leray I, Valeur B. Chem Commun 2003:996. Chem Eur J 2004;10:4480. (b) Hayashita T, Qing D, Minagawa M, Lee JC, Ku CH, Teramae N. Chem Commun 2003:2160. (c) Chen CT, Huang WP. J Am Chem Soc 2002;124:6246. [PubMed: 12033846] (d) Xia WS, Schmehl RH, Li CJ, Mague JT, Luo CP, Guldi DM. J Phys Chem B 2002;106:833. (e) Beeby A, Parker JD, Williams JAG. J Chem Soc, Perkin Trans 2 1996:1565. (f) Addleman RS, Bennet J, Tweedy SH, Elshani S, Wai CM. Talanta 1998;46:573. [PubMed: 18967179] (g) Padilla-Tosta ME, Lloris JM, Martínez-Mañez R, Marcos MD, Miranda MA, Pardo T, Sancenon F, Soto J. Eur J Inorg Chem 2001:1475.
11. Choudhary S, Morrow JR. Angew Chem, Int Ed 2002;41:4096.
12. Amundsen LH. J Am Chem Soc 1937;59:146.

13. (a) Walkup GK, Imperiali B. *J Am Chem Soc* 1998;120:609. (b) Li YH, Chan LM, Tyer L, Moody RT, Himel CM, Hercules DM. *J Am Chem Soc* 1975;97:3118. (c) Zheng Y, Orbulescu J, Ji X, Andreopoulos F, Pham SM, Leblanc RM. *J Am Chem Soc* 2003;125:2680. [PubMed: 12603155] (d) Prodi L, Bolletta F, Montalti M, Zaccheroni N. *Eur J Inorg Chem* 1999:455.
14. (a) Bryan JC, Rosenberg JM, Kavallieratos K. *Acta Crystallogr* 2005;E61:o396. (b) Eagle CT, Kavallieratos K, Bryan JC. *J Chem Crystallogr* 2002;32:165.
15. Kavallieratos K, Rosenberg JM, Bryan JC. *Inorg Chem* 2005;44:2573. [PubMed: 15819538]
16. Webr G, Teale FWS. *Trans Faraday Soc* 1958;54:640.

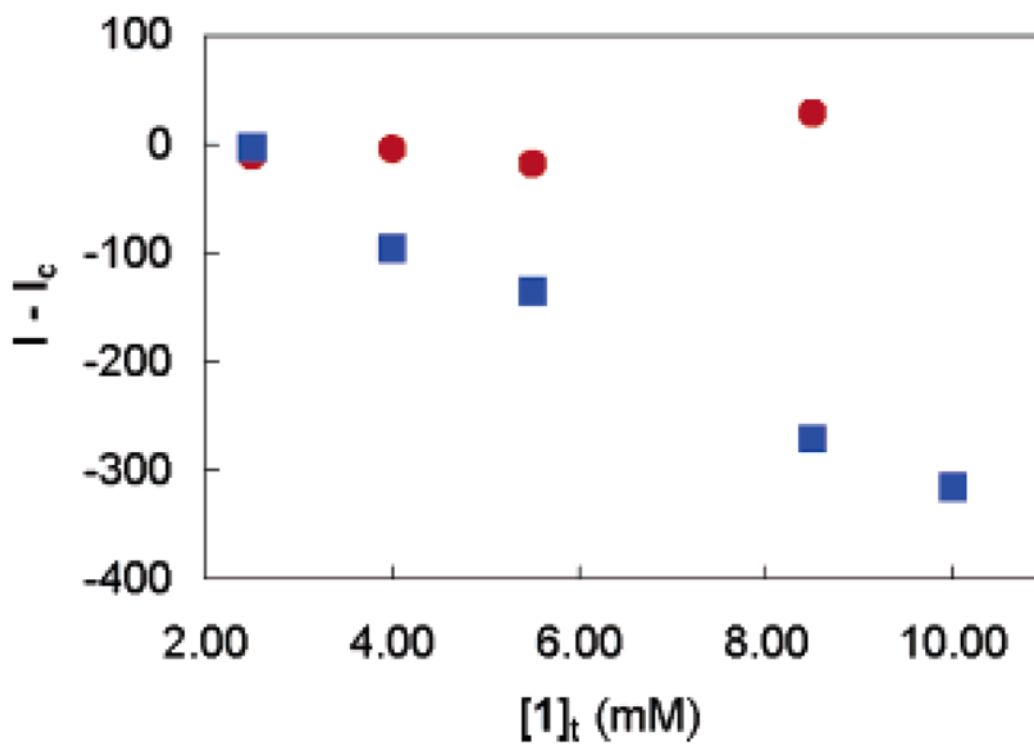


**Figure 1.**

Plot of  $\log [M(II)]_t$  versus  $[1]_t$ :  $[Pb]_t = 2.72$  mM,  $[Ni]_t = 2.72$  mM,  $[Co]_t = 2.67$  mM,  $[Cu]_t = 2.80$  mM,  $[Zn]_t = 2.53$  mM,  $[Cd]_t = 2.59$  mM.



**Figure 2.** Fluorescence emission at 516 nm upon Pb coordination and extraction by **1**:  $[1]_t = 7.40$  mM,  $[NH(i-Pr)_2] = 16.30$  mM;  $\lambda_{exc} = 340$  nm.



**Figure 3.** Fluorescence intensity difference at 516 nm for **1**/diisopropylamine (■) and **2**/diisopropylamine (●) after contact with  $\text{Pb}(\text{NO}_3)_2(\text{aq})$  ( $I$ ), compared to contact with blank ( $I_c$ ) versus  $[\text{L}]_t$ .  $[\text{Pb}]_t = 5.5$  mM;  $\lambda_{\text{exc}} = 340$  nm.