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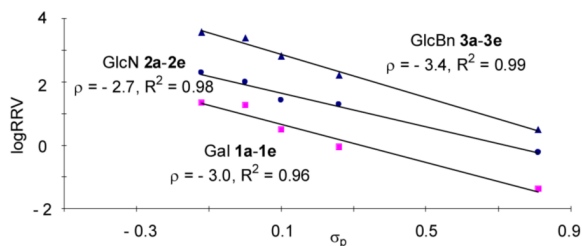
Thio-arylglycosides with Various Aglycon Para-Substituents, a Probe for Studying Chemical Glycosylation Reactions

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Summary

Three series of thioglycosyl donors differing only in their respective aglycon substituents within each series have been prepared as representatives of typical glycosyl donors. The relative anomeric reactivities of these donors were quantified under competitive glycosylation conditions with various reaction time, promoters, solvents and acceptors. Over three orders of magnitude reactivity difference were generated by simple transformation of the para-substituent on the aglycon with methanol as the acceptor, while chemoselectivities became lower with carbohydrate acceptors. Excellent linear correlations were attained between relative reactivity values of donors and σ_p values of the substituents in the Hammett plots. This indicates that the glycosylation mechanism remains the same over a wide range of reactivities and glycosylation conditions. The negative slopes of the Hammett plots suggested that electron donating substituents expedite the reactions and the magnitudes of slopes can be rationalized by neighboring group participation as well as electronic properties of the glycon protective groups. Within the same series of donors, less nucleophilic acceptors gave smaller slopes in their Hammett plots. This is consistent with the notion that acceptor nucleophilic attack onto the reactive intermediate is part of the rate limiting step of the glycosylation reaction.



Excellent linear Hammett correlations were obtained between relative reactivity values of three series of donors differing only in their aglycon substituents and σ_p values of the substituents.

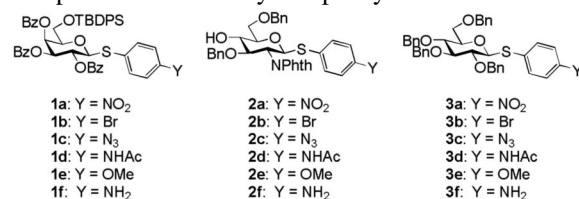
Introduction

With the increasing recognition of the biological importance of carbohydrates, carbohydrate chemistry is gaining more prominence.¹ While much efforts has been spent on synthesis of specific oligosaccharides, relatively little attention has been devoted to systematic exploration of the general aspects of the glycosylation reaction. Such investigations can, however, be very fruitful.

It is known that protective groups and aglycon leaving group structures have profound effects on the anomeric reactivities of glycosyl donors.² Donors bearing electron donating protective groups have much higher anomeric reactivities towards a promoter than those bearing electron withdrawing and conformational restricting protective groups. This phenomenon was first observed with the glycosyl halides,³ which has been subsequently found to be general for a variety of donors including pentenyl glycosides,⁴ thioglycosides,^{2, 5, 6} glycosyl fluorides,⁷ glycosyl thioimidates⁸ and glycosyl sulfoxides.⁹ The understanding of protective group effect on glycosylation rate led to the development of reactivity based chemoselective glycosylation method, where a thioglycoside donor with higher anomeric reactivity (armed) is mixed together with a less reactive (disarmed) thioglycosyl acceptor.^{2, 6} Upon addition of a limiting amount of promoter, the armed donor is preferentially activated glycosylating the bifunctional thioglycoside acceptor. The newly formed thioglycosyl disaccharide, in turn can be further activated as a donor leading to a more convergent synthesis of oligosaccharides. Multiple sequential chemoselective glycosylations can be carried out in the sequence of decreasing order of anomeric reactivities in one single reaction flask.^{2, 5, 6, 10, 11}

In order to aid in building block selection, Ley and coworkers quantified the relative reactivities of thioglycosyl donors bearing various protective groups towards a carbohydrate acceptor.¹² Similarly, a database was established by the Wong group tabulating Relative Reactivity Values (RRVs) of thiotolyl building blocks glycosylating methanol.^{2, 11} RRV_{methanol} values have been employed as a quantitative measure to guide the design of building blocks with requisite anomeric reactivities for reactivity based chemoselective glycosylation, as applied in assemblies of several complex oligosaccharides.^{2, 10, 11, 13-15} Generally, minimum ten-fold difference in $RRVs_{\text{methanol}}$ between the two competing building blocks is desired so that theoretically greater than 90% glycosylation yield can be obtained. However, the effects of different acceptors on RRVs were not quantitatively evaluated.

Aglycon of thioglycosides can also significantly influence the anomeric reactivity. Roy and colleagues reported that the *p*-acetamidophenylthioglycoside could be selectively activated and used to glycosylate the electronically deactivated *p*-nitrophenylthioglycosides.¹⁶ The aglycon of the resulting disaccharide was then reduced and acetylated to yield another *p*-acetamidophenylthioglycosyl donor for convenient chain elongation. Besides electronic effects, the size of the aglycon can play an important role as well. The Boons group took advantage of the slow activation of the bulky dicyclohexylmethylthioglycosides allowing the selective activation of a sterically less hindered ethylthioglycoside and its subsequent chemoselective glycosylation.¹⁷ Oscarson and colleagues quantified the effects of size and electronic properties on the anomeric reactivities of a series of thioglycosides bearing different aglycons and applied the knowledge gained to a trisaccharide synthesis.¹⁸ Following this trend of anomeric reactivity tuning, our group has developed a new one pot method utilizing a series of glycosyl donors (**1a-1e**, **2a-2e**) differing only in their aglycon substituents, which were readily acquired through post-synthetic aglycon modification of the common intermediates **1f** and **2f** (Scheme 1), allowing rapid access to glycosyl donors with multiple levels of anomeric reactivities.¹⁹ These building blocks were then used in construction of several tetrasaccharides and pentasaccharides by one pot syntheses.



As the donor series **1a-1f** and **2a-2f** only differed in their aglycon substituents, this provided us unique opportunities to investigate the glycosylation using Hammett analyses. Hammett

relationship has been a useful tool for probing reaction mechanisms,²⁰ and in relation to glycoscience, glycosidase mechanisms.²¹⁻²⁷ There are several possible scenarios with the Hammett plots for chemical glycosylations: 1) linear Hammett relationship for a wide range of leaving group σ_p values, which will indicate that the glycosylation mechanism remains the same for all substituents; 2) bent Hammett relationship consisted of linear correlations with different slopes suggesting a change in mechanism. Such phenomena have been observed previously with several glycosidases,^{22, 28, 29} which resulted from rate limiting step change with different substituents; 3) Hammett relationship with poor linear correlations and badly scattered points implying factors other than substituent electronic effects influencing the reaction. Herein, we report the quantification of anomeric reactivities of three donor series **1a-1e**, **2a-2e** and **3a-3e**. Hammett analyses were then performed on these donors by plotting reaction rates against σ_p values under a variety of glycosylation conditions to probe the glycosylation mechanism.

Results

Preparation of Three Series of Donors

Three donor series **1a-1f**, **2a-2f** and **3a-3f** were selected as representative glycosyl donors for our investigation. Thiogalactosides **1a-1f** (Gal series) contain multiple electron withdrawing protective groups on the glycon ring with a participating neighboring group (OBz) on C2. Thioglucosamine series **2a-2f** (GlcN series) bear a participating neighboring group phthalimido (NPhth) on C2 and several electron donating groups, while glucosides **3a-3f** (GlcBn series) only contain electron donating protective groups with no participating neighboring groups. **1a-1f** and **2a-2f** are prepared through post-synthetic modifications of aglycons as previously described.¹⁹ Instead of spending five synthetic steps on average to synthesize each donor starting from individual aryl thiol, it only took one to two steps to acquire building blocks **1b-1e** and **2b-2e** divergently from common synthetic intermediates **1f** and **2f**.

With the simpler protective group pattern on glucose series donors **3a-3e**, we started the synthesis from per-acetylated glucose **4** and the corresponding thiophenols for **3a**,^{16, 30} **3b** and **3e** (Scheme 1a). Borontrifluoride etherate mediated glycosylation produced thioglycosides **5-7**^{31, 32} in excellent yields. Replacement of the four acetates on the glycon ring with benzyl groups was accomplished by treatment of **5-7** with NaOMe followed by benzylation. The yield for formation of **3a** was lower due to competitive reduction of the nitro moiety by NaH.

Initial attempt to synthesize acetamido substituted donor **3d** from 4-acetamido thiophenol and glucoside **4** failed due to complexation of tin (IV) chloride by the acetamido moiety. As an alternative, we performed stannous chloride reduction of **3a** to produce aminophenyl glucoside **3f**, which was found to be very prone to air oxidation. Instead, a one-pot procedure of stannous chloride reduction of **3a** followed by immediate acetylation in the same reaction flask led to donor **3d** in 84% yield (Scheme 1b). Stannous chloride reduction of 4-nitrophenyl glucoside **5** with subsequent diazonium salt formation and *in situ* displacement by sodium azide produced the azido bearing glucoside **9** in 59% yield for the two steps. Removal of acetates followed by benzylation gave donor **3c** in 78% yield (Scheme 1c).

Determination of the Rate of Glycosylation by NMR

In order to quantitatively determine donor reactivities, we first investigated the possibility of real time monitoring of glycosylation by NMR. The reaction of **1b** and methanol using *N*-iodosuccinimide (NIS) and catalytic amount of trifluoromethane sulfonic acid (TfOH) as the promoter system³³ was performed in CD₂Cl₂ with a ¹H-NMR spectrum of the reaction mixture acquired every eight seconds. As the reaction proceeded, the amount of donor decreased as

determined from integrations of the anomeric proton (Figure 1a). Interestingly, the rate of reaction calculated from the speed of donor consumption rose during the first few minutes of the reaction before decreasing with time in a manner expected for a reaction exhibiting positive-order kinetics in substrate concentration (Figure 1b). Similar phenomenon was also observed when dimethyl(methylthio) sulfonium triflate was used as the promoter. The cause of such an induction period³⁴ at the onset of glycosylation is unclear. As the glycosylation progressed passing the induction period, the observed pseudo-linear decrease of the reaction rate indicated that glycosylation is apparent first order to donor under this condition (Figure 1b,c). However, it was difficult to accurately determine the absolute rate via this method, as subtle variation of the reaction condition could cause substantial rate fluctuation, which led us to focus on measurements of relative rate of glycosylation.

Determination of Relative Reactivity Values using Methanol as the Acceptor

To measure the relative glycosylation rates, we adapted a competitive glycosylation condition where two different glycosyl donors are mixed together with methanol.^{11, 12, 18, 35} Upon addition of a limiting amount of the promoter, the two donors competed for promoter activation, with the rate of glycosylation determined through the decrease of donor concentration. The measurement of relative reactivities allowed us to avoid the need to measure acceptor concentration as well as rate constant dependence on the identity of the promoter.

As glycosylation was apparent first order to donor under the experimental condition, the rate of glycosylation could be described by $k_{obs}[D]$, where k_{obs} is the apparent first order rate constant and $[D]$ is the donor concentration. Thus, the relative reactivity between these two donors in a competition reaction could be derived according to equation (1)

$$\frac{k_{obs1}}{k_{obs2}} = \frac{\log([D1]_t) - \log([D1]_0)}{\log([D2]_t) - \log([D2]_0)} \quad (1)$$

(0 is designated as the starting time and t is designated as the finishing time). Because ¹H-NMR of such competition reactions could be fairly complex, we quantified the compounds from chromatogram integration following HPLC separation (Figure 2).^{11, 18} The concentrations of compounds were calculated from peak integration areas. An internal reference R inert to glycosylation was added to the reaction mixture in order to normalize the amounts of compounds between different experiments and to minimize the errors caused by injected sample volume changes. Thus, equation 1 was revised to equation 2

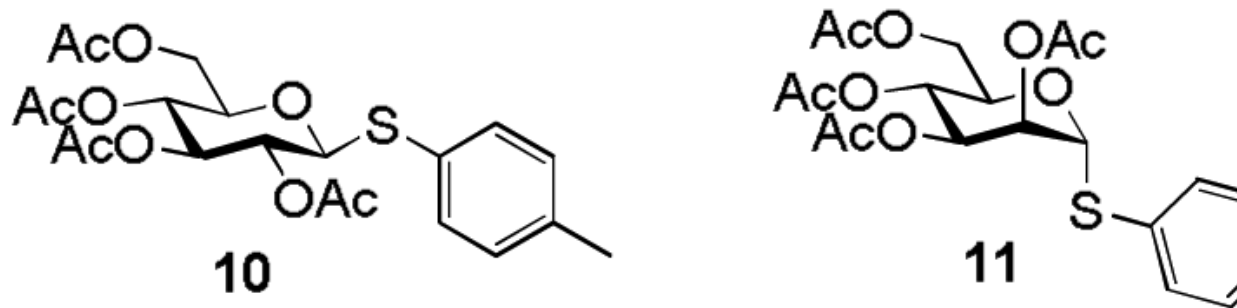
$$\frac{k_{obs1}}{k_{obs2}} = \frac{\log\left(\frac{D1_t \cdot R_0}{R_t}\right) - \log(D1_0)}{\log\left(\frac{D2_t \cdot R_0}{R_t}\right) - \log(D2_0)} \quad (2)$$

(D and R are the amounts of compounds calculated from the HPLC peak integration area).

In all cases, baseline separation of desired peaks in HPLC chromatogram was achieved. The reactivity differentials between the two competing donors were kept below 20 folds to ensure accuracy of the experiments. When there were greater than 20 folds of donor reactivity differential or chromatogram overlap, a third donor was used to bridge the two donors of interest. For each pair of donors, the reaction was repeated at least three times. Usually, less than 10% differential was obtained among all measurements, with the average value taken as the relative reactivity value between the two donors, an example of which is shown in Table 1.

The reaction time did not affect the relative reactivity much. For example, k_{2a}/k_{10} values of 4.80, 4.28 and 4.49 were obtained when the competitive reactions of donor **2a** against donor **10** were quenched at 3, 5 and 120 minutes respectively. Next we examined the effect of promoter. Relative reactivities between donor pair of **1b** and **10** were measured with two promoter systems, NIS/TfOH and *p*-TolSCI/AgOTf.³⁶ The observed k_{1b}/k_{10} values using the two systems were similar (3.02 for NIS/TfOH and 2.86 for *p*-TolSCI/AgOTf). This suggested that relative reactivities were independent of promoter, consistent with our hypothesis and observations by Wong and coworkers.¹¹ Greater than 92% yields were obtained for the two methyl glycoside products based on consumption of **1b** and **10**, confirming that consumed donors were converted into products.

With the experimental conditions established, the RRVs for three series of donors (Gal **1a-1e**, GlcN **2a-2e**, GlcBn **3a-3e**) were obtained using MeOH as the acceptor, with the reactivity of thiomannoside **11** set as 1¹¹ (Table 2). Although donors **2a-2e** contain a free hydroxyl group, due to the much higher reactivity of methanol and excess methanol used,³⁷ chemoselective glycosylation of **2a-2e** with methanol was achieved with little self couplings of donors. It is evident that the para-substituents on the aglycon of thioaryl glycoside donors significantly impact anomeric reactivities with over three orders of magnitude difference obtained by a single substitution (**3e** vs **3a**). This effect is much larger than that resulting from changing one protective group on the glycon ring, which is generally less than 10 fold.² The much bigger influence of aglycon on anomeric effect is due to the direct conjugation of para-substituents with the anomeric sulfur atom.



For our glycosylations, Hammett plots were constructed by plotting $\log(\text{RRV})$ values of the three series of donors against σ_p values of the para-substituents^{20, 38} (Figure 3). Excellent linear correlations were observed within each donor series, suggesting that the same glycosylation mechanism is operative over a large range of reactivities.^{25, 39} Furthermore, the slopes of the Hammett plots (ρ values) were all negative indicating that the reactive intermediates were cationic in nature and electron donating substituents such as methoxy expedited the reaction compared with electron withdrawing substituents such as nitro.

Prediction of Donor RRVs by Hammett Equation and Correlation of ¹H-NMR Anomeric Proton Chemical Shifts with σ_p Values

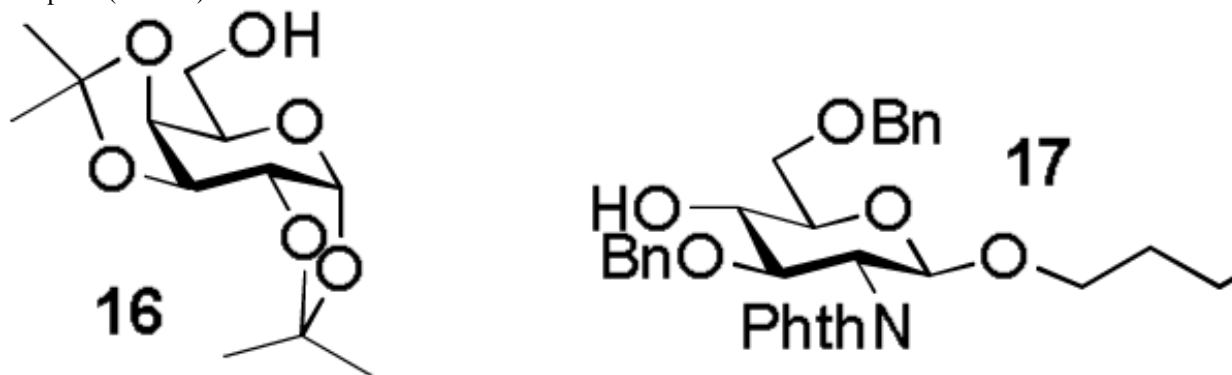
Building block preparation is a time-consuming process and it requires much experience to design and fine tune the protective groups in order to achieve exact anomeric reactivities required by the armed-disarmed chemoselective glycosylations.^{2, 11, 40} It will be highly desirable if reactivities of the building blocks can be predicted prior to the actual synthesis. Anomeric reactivities have been estimated by calculating the energy difference between the ground state and the oxocarbenium ion.⁴¹ With the large number of atoms in typical fully protected glycosyl donors, global minima of the structures can be difficult to obtain by calculation. Furthermore, the need to include solvation for accurate determination of the oxocarbenium ion energy renders this approaches challenging.⁴¹ With the excellent linear correlations between $\log(\text{RRV})$ against σ_p and large number of σ_p values available,²⁰ σ_p can

be employed as a predictor of RRV. For example, $\log(\text{RRV})$ of galactoside **12** was calculated to be 1.38 from the Hammett plot (Figure 3, $\sigma_p = -0.14$ for methyl group), while the experimentally determined value was 1.28.² The same was applicable to thioglycosides **13** and **14** (Table 3).

Since the only structural differences among our donor series are in their aglycon substituents, we envision that the chemical shifts of their anomeric protons should only be affected by the electronic properties of para-substituents on aglycons. Excellent linear correlations between chemical shifts of the anomeric protons (Table 4) and σ_p values were obtained in all donor series examined (Figure 4), which supports previous observations that chemical shifts could be used to predict relative reactivities.^{12, 19}

Effect of Acceptor and Solvent on Relative Reactivity Values

Prior competitive glycosylation studies have been mainly focused on systematic investigations of the effect of donor structural change on glycosylation.^{2, 12, 18} It is known that acceptors can also drastically affect glycosylation.⁴²⁻⁴⁵ In order to quantify the effect of acceptor, we performed competitive glycosylation reactions of Gal donors **1a-1e** using galactoside **16** and glucosamine **17** respectively as representatives of primary and secondary carbohydrate acceptors (Table 5).



Excellent linear correlations were obtained in the Hammett plots for reactions of donors **1a-1e** with acceptors **16** and **17** in CH_2Cl_2 with ρ values of -2.7 and -2.0 respectively (Figure 5). Interestingly, compared with methanol acceptor for the same donor series, the reactivity differential became smaller (k_{1e}/k_{1a} 62.5 with acceptor **17** vs 516.7 with methanol acceptor) and the reaction rate was less sensitive to changes in para-substituents of the aglycon as reflected by the smaller absolute ρ value.

In the programmable reactivity based chemoselective glycosylation, $\text{RRV}_{\text{methanol}}$ were utilized as a quantitative measure to guide building block selection.^{2, 5, 6, 10, 11, 14, 15} However, in some cases, glycosylation yields were lower than those indicated based on $\text{RRV}_{\text{methanol}}$ differentials,^{11, 40} which could be due to the decrease of chemoselectivity resulting from glycosylating a less reactive carbohydrate acceptor.

The glycosylation solvent can also significantly impact the chemoselectivity. When glycosylations of acceptor **16** by donors **1a-1e** were carried out in diethyl ether using *p*-TolSCl/AgOTf promoter system, good linear correlation between $\log(\text{RRV})$ and σ_p was retained (Figure 6 and Table 5 entry 4). However, the RRV difference in diethyl ether was smaller than that in CH_2Cl_2 as indicated by the smaller absolute ρ value. Oscarson and coworkers reported that using NIS/AgOTf as the promoter, better selectivity was observed in diethyl ether compared with CH_2Cl_2 for competitive glycosylation of two thioglycosyl building blocks.⁴⁶ This discrepancy may be attributed to the low solubility of NIS in diethyl ether, thus reducing

effective concentration of promoter available for glycosylation leading to higher selectivity.
46, 47

Discussions

Thioglycoside glycosylation reaction can be generally viewed as a three-step process.^{11, 41} The glycosyl donor is first activated by electrophilic addition of the promoter onto the anomeric sulfur atom, generating a glycosyl sulfonium ion (Scheme 2, step A), which is converted to other reactive intermediates such as glycosyl triflate or dioxolenium ion (Scheme 2, step B). Nucleophilic attack of the reactive intermediate by the acceptor leads to the glycoside product (Scheme 2, step C). Crich and coworkers have demonstrated in a series of ground breaking low temperature NMR studies that glycosyl triflates were the dominant intermediates following activation of thioglucosides and thiomannosides without a participating neighboring group on O-2.⁴⁸⁻⁵⁰ Recently, our group monitored the activation process of thioglycosyl donors bearing participating neighboring group on O-2 by low temperature NMR.⁵¹ Upon addition of the promoter, the disappearance of the glycosyl donor was almost instantaneous. The glycosyl sulfonium ion was not observed during the activation process. Rather, the major intermediates were the glycosyl triflate and the dioxolenium ion formed due to neighboring group participation. Glycosylation of the acceptor by the reactive intermediate occurred at a slower rate, often requiring more than one hour to complete at low temperature with carbohydrate acceptors. Although in these NMR studies donor activation was performed in the absence of the acceptor, these results suggested that steps A and B in thioglycoside glycosylation are rapid with the glycosyl sulfonium ion being consumed as it is forming, and step C is the slow step for glycosylation. The absence of glycosyl sulfonium ion indicated that the rate of the reverse reaction for step B must be small compared to the forward reaction.

Acceptors had a significant impact on the glycosylation rate as shown in Hammett analyses (Figure 5), suggesting that acceptor attack is involved in the rate limiting step, which is consistent with the results from previous low temperature NMR studies. The dependence of ρ value on acceptor could be explained as the acceptor nucleophilicity decreased from methanol to galactoside **16**, the acceptor reacted slower with the reactive intermediate, allowing its accumulation. Therefore, rate differences resulting from differential aglycons became less pronounced leading to a smaller absolute ρ value (2.7 vs 3.0, Figure 5). Secondary hydroxyl group in glucosamine **17** was less reactive⁵² than the primary alcohol in **16**, causing further decrease in the absolute value of ρ (2.0 vs 2.7).

In the Hammett relationship studies, ρ values were negative demonstrating building blocks with electron donating groups were more reactive with acceptors. Since the donors used in Hammett plots differed only in their aglycon substituents, the reactive intermediates generated within the same donor series should have the same structure as aglycons have already been cleaved. Therefore, the rate constants for step C for the same donor series should be the same. The fact that building blocks with electron withdrawing aglycons reacted slower could be understood by assuming step A is a fast equilibrium. As the sulfur atoms in *p*-nitrophenyl glycosides are less electron rich, their activation by an electrophilic promoter will be slower than the corresponding *p*-methoxyphenyl glycosides. This will generate lower concentrations of the reactive intermediates from *p*-nitrophenyl glycosides, leading to lower overall glycosylation rates.

The absolute ρ value of Gal donor series **1a-1e** was larger than that of the GlcN series **2a-2e** (3.0 vs 2.7, Figure 3), suggesting that Gal series of donors were more sensitive to the substituents. This can be rationalized by the presence of more electron withdrawing protective groups on the glycon ring in the Gal donor series, imparting more cationic character to the reactive intermediates.

Since OH is an equally potent electron donating moiety as benzyl group,¹¹ GlcN series **2a-2e** contain one more electron withdrawing protective group on the glycon ring as compared to GlcBn series **3a-3e** donors. Yet, the GlcN series were less sensitive to aglycon changes as shown by a smaller absolute ρ value in its Hammett plot (2.7 vs 3.4, Figure 3). This presumably resulted from the participating neighboring group NPhth on C-2 of GlcN donors. With NPhth participation, the positive charge on the corresponding reactive intermediate from GlcN donors becomes more dispersed; hence the electron donating para-substituent on aglycon has lesser stabilizing effect leading to a smaller absolute ρ value (Figure 7). Similar phenomena of smaller absolute ρ value due to neighboring group participation have been observed in several nucleophilic displacement reactions.^{53, 54}

Conclusions

We have demonstrated that simple modification of the para-substituents of thioaryl glycosides can render over three orders of magnitude difference in anomeric reactivities of glycosyl donors. The reactivity differential was found to be dependent upon the identity of the acceptor with the less reactive carbohydrate acceptors leading to lower chemoselectivities. Excellent linear correlations were obtained in Hammett relationship studies demonstrating that the glycosylation mechanism remains unchanged over the wide range of donor reactivities and reaction conditions. The negative slopes of the Hammett plots suggest that electron donating substituents expedite the reactions and the magnitudes of slopes can be rationalized by neighboring group participation and electronic properties of the glycon protective groups. The knowledge gained from our results can provide valuable insights into further improvement of the powerful armed-disarmed chemoselective glycosylation methodology and the understanding of chemical glycosylation in general.

Experimental Section

General procedure for competitive glycosylation and HPLC measurements

Two donors of interest (0.05 mmol each) were dried together with a reference compound (50 mg) under high vacuum overnight. They were then dissolved in anhydrous CH_2Cl_2 (5 mL). Aliquots of this solution (1 mL each) were transferred to three pre-dried vials. The exact ratio of two donors in each vial was determined by HPLC analysis of an aliquot (10 μL) of the solution. Each vial was measured at least three times. Anhydrous MeOH (2.04 μL , 5 eq), acceptor **16** or acceptor **17** was then added to each vial followed by 0.5 M NIS solution in acetonitrile (20 μL) and 0.1 M TfOH in Et_2O (10 μL). The reactions were left at room temperature for 2 hours and quenched by triethylamine (20 μL). The reaction mixture was diluted with CH_2Cl_2 (2 mL), washed with saturated aqueous sodium thiosulfate solution containing 10% sodium hydrogen bicarbonate, dried over Na_2SO_4 and evaporated to dryness. The solid obtained was then dissolved in CH_2Cl_2 (1 mL). The concentration of remaining donors in each vial were determined by HPLC (~10 μL each injection; 3 injections for each vial). RRVs are calculated according to equation 1 as the average of all measurements. All HPLC analyses were performed using a HP 1050 series HPLC system with a SUPELCO normal phase analytical HPLC column (25 cm * 4.6 mm ID) and UV detection at 270 nm.

For glycosylations using the *p*-TolSCI/AgOTf promoter system, all procedures are the same as above except that 0.5 M AgOTf solution in acetonitrile (40 μL) and *p*-TolSCI (1.5 μL) were added to each vial instead of NIS and TfOH.

***p*-Azidophenyl 3,6-di-O-benzyl-2-deoxy-2-phthalimido-1-thio- β -D-glucopyranoside (2c)**—Compound **2f** (597 mg, 1 mmol) was dissolved in a mixture solvent consisted of 8 N acetic acid (20 mL) and THF (20 mL) at 0 °C. A 5% aqueous solution of NaNO_2 (2.6 mL) was added to this solution and the mixture was stirred for 10 minutes. A 5%

aqueous solution of NaN_3 (6.5 ml) was then added and the reaction mixture was allowed to warm up to room temperature in 10 minutes. The reaction was again cooled to 0°C , neutralized with 10 % aqueous NaOH , extracted with EtOAc , washed with water and brine, and dried over anhydrous Na_2SO_4 . After filtration, removal of all solvents under reduced pressure, compound **2c** was obtained by flash chromatography in 87% yield (542 mg). ^1H NMR (400 MHz, CDCl_3) δ 2.95 (s, 1H, 4-OH), 3.66-3.71 (m, 1H, H_5), 3.76-3.88 (m, 3H, H_2 , H_6 , H_6'), 4.18 (t, 1H, $J = 10.2$ Hz, H_3), 4.23 (t, 1H, $J = 10.2$ Hz, H_4), 4.52 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.57 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.62 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.72 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 5.47 (d, 1H, $J = 10.2$ Hz, H_1), 6.82-7.84 (m, 18H); ^{13}C NMR (100 MHz, CDCl_3) δ 54.6, 70.7, 74.0, 74.1, 74.8, 78.2, 79.9, 83.6 (C_1), 119.6, 127.8-128.8, 135.1, 137.9, 138.1, 140.4, 167.8, 168.4; MS (ESI) m/z calcd. For $\text{C}_{34}\text{H}_{30}\text{N}_4\text{NaO}_6\text{S}$ [$\text{M} + \text{Na}$] $^+$ 645.2; found 645.3; HRMS m/z calcd. For $\text{C}_{34}\text{H}_{30}\text{N}_4\text{NaO}_6\text{S}$ [$\text{M} + \text{Na}$] $^+$ 645.1784; found 645.1790.

***p*-N-Acetamidophenyl 2-deoxy-phthalimido-3,6-di-O-benzyl-1-thio- β -D-glucopyranoside (2d)**—Compound **2f** (59.7 mg, 0.1 mmol) was dissolved in MeOH (10 mL) at 0°C . A solution of acetic anhydride (47 μL , 0.5 mmol) was added to this solution and the mixture was stirred for 10 minutes at 0°C , then the reaction mixture was neutralized with solid NaHCO_3 , extracted with EtOAc , washed with water and brine, and dried over anhydrous Na_2SO_4 . After filtration and removal of all solvents under reduced pressure, compound **2d** was obtained by flash chromatography in 97% yield (62.0 mg). ^1H NMR (400 MHz, CDCl_3) δ 2.12 (s, 3H, NHCOCH_3), 3.19 (s, 1H, 4-OH), 3.64-3.68 (m, 1H, H_5), 3.77-3.86 (m, 3H, H_2 , H_6 , H_6'), 4.18 (t, 1H, $J = 9.6$ Hz, H_3), 4.22 (t, 1H, $J = 9.6$ Hz, H_4), 4.51 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.54 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.59 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 4.72 (d, 1H, $J = 12.4$ Hz, PhCH_2O), 5.46 (d, 1H, $J = 9.6$ Hz, H_1), 6.91-7.83 (m, 18H), 7.57 (bs, 1H, NHAc); ^{13}C NMR (100 MHz, CDCl_3) δ 29.8 (NHCOCH_3), 54.6, 70.7, 74.0 (2 X OCH_2Ph , C_6), 74.2, 74.8, 78.1, 79.8, 84.0 (C_1), 120.2, 123.6, 123.7, 128.0-128.8, 134.2-134.4, 137.9, 138.1, 138.4, 167.7, 168.4, 168.8; MS (ESI) m/z calcd. for $\text{C}_{36}\text{H}_{34}\text{N}_2\text{NaO}_7\text{S}$ [$\text{M} + \text{Na}$] $^+$ 661.2; found 661.4; HRMS m/z calcd. for $\text{C}_{36}\text{H}_{34}\text{N}_2\text{NaO}_7\text{S}$ [$\text{M} + \text{Na}$] $^+$ 661.1984; found 661.1976.

***p*-Nitrophenyl 2,3,4,6-tetra-O-benzyl-1-thio- β -D-glucopyranoside (3a)**—To a 50 mL round-bottom flask were added the tetraacetate **5** 13 (1.11 g, 2.28 mmol) and a 2:3 mixture of $\text{MeOH-CH}_2\text{Cl}_2$ (25 mL). The mixture was cooled to 0°C , treated with a 0.5 M NaOMe solution in MeOH (1.50 mL, 0.75 mmol) and stirred for 10 minutes followed by warming up to room temperature and further stirring for 2 hours. The reaction was neutralized with acetic acid to pH around 7. The suspension was concentrated under reduced pressure, diluted with CH_2Cl_2 and filtered to afford the white solid tetraol (0.84 g). The crude tetraol was dried under vacuum in a 25 mL round-bottom flask overnight and azeotropically distilled in toluene to remove any residue water. It was then suspended in anhydrous DMF (15 mL) with freshly activated $\text{MS } 4 \text{ \AA}$ (1 g). The suspension was stirred for 30 minutes at room temperature, then cooled to 0°C in an ice-bath followed by slow addition of 95% NaH (0.35 g, 13.8 mmol). The resulting suspension was stirred for 30 minutes at 0°C then treated with BnBr (1.42 mL, 11.9 mmol). The mixture was warmed to room temperature and stirred for another 4 hours. Then the mixture was cooled back to 0°C and quenched with acetic acid to pH=7 and diluted with EtOAc . The organic phase was collected and washed with saturated NaHCO_3 and dried over Na_2SO_4 . The residue was purified by flash chromatography to give the compound **3a** in 45% yield for two steps (0.65 g). ^1H NMR (600 MHz, CDCl_3) δ 3.57 (t, 1H, $J_{1,2} = 9.6$ Hz, H_2), 3.58-3.61 (m, 1H, H_5), 3.65-3.70 (m, 2H, 2 X H_6), 3.75 (t, 1H, $J = 9.6$ Hz, H_3), 3.78 (t, 1H, $J = 9.6$ Hz, H_4), 4.52 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 4.57 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 4.60 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 4.76 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 4.82 (q, 2H, $J = 11.4$ Hz, PhCH_2O), 4.82 (d, 1H, $J_{1,2} = 9.6$ Hz, H_1), 4.87 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 4.90 (d, 1H, $J = 11.4$ Hz, PhCH_2O), 7.19-7.36 (m, 20H), 7.59-7.61 (m, 2H), 7.99-8.01 (m, 2H); ^{13}C NMR (150 MHz, CDCl_3) δ 69.1, 73.7, 75.4, 75.9, 76.1, 77.9, 79.4, 80.9, 86.0, 86.8 (C_1), 124.1,

128.0-129.6, 137.7, 138.0, 138.1, 138.4, 144.6, 146.4; MS (ESI) m/z calcd. for $C_{40}H_{39}NNaO_7S$ $[M + Na]^+$ 700.2; found 700.4; HRMS m/z calcd. for $C_{40}H_{39}NNaO_7S$ $[M + Na]^+$ 700.2345; found 700.2350.

***p*-Bromophenyl 2,3,4,6-tetra-*O*-benzyl-1-thio- β -D-glucopyranoside (3b)—**

Compound **3b** was prepared as described in the synthesis of compound **3a** starting from compound **6** in 92% yield. 1H NMR (600 MHz, $CDCl_3$) δ 3.48 (t, 1H, $J_{1,2} = 9.6$ Hz, H_2), 3.50-3.52 (m, 1H, H_5), 3.63 (t, 1H, $J = 9.6$ Hz, H_3), 3.70 (t, 1H, $J = 9.6$ Hz, H_4), 3.69-3.71 (m, 1H, H_6), 3.77 (dd, 1H, $J = 1.8, 10.8$ Hz, H_6'), 4.53 (d, 1H, $J = 12.0$ Hz, $PhCH_2O$), 4.58 (d, 2H, $J = 12.0$ Hz, $PhCH_2O$), 4.62 (d, 1H, $J_{1,2} = 9.6$ Hz, H_1), 4.73 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 4.81 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 4.83 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 4.84 (d, 1H, $J = 12.0$ Hz, $PhCH_2O$), 4.88 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 7.18-7.20 (m, 2H), 7.27-7.37 (m, 20H), 7.43-7.44 (m, 2H); ^{13}C NMR (150 MHz, $CDCl_3$) δ 69.2, 73.6, 75.3, 75.7, 76.1 (4 X OCH_2Ph , C_6), 78.0, 79.2, 80.9, 86.9, 87.6 (C_1), 122.0, 127.9-128.7, 132.2, 133.0, 133.7, 138.1, 138.4, 138.5; MS (ESI) m/z calcd. for $C_{40}H_{39}BrNaO_5S$ $[M + Na]^+$ 733.2; found 733.4; HRMS m/z calcd. for $C_{40}H_{39}BrNaO_5S$ $[M + Na]^+$ 733.1599; found 733.1604.

***p*-Azidophenyl 2,3,4,6-tetra-*O*-benzyl-1-thio- β -D-glucopyranoside (3c)—**

Compound **3c** was prepared as described in the synthesis of compound **3a** starting from compound **9** in 78% yield. 1H NMR (600 MHz, $CDCl_3$) δ 3.46 (t, 1H, $J_{1,2} = 9.6$ Hz, H_2), 3.47-3.50 (m, 1H, H_5), 3.63 (t, 1H, $J = 9.6$ Hz, H_3), 3.70 (t, 1H, $J = 9.6$ Hz, H_4), 3.71-3.73 (m, 1H, H_6), 3.78 (d, 1H, $J = 10.2$ Hz, H_6'), 4.53 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.58 (d, 1H, $J_{1,2} = 9.6$ Hz, H_1), 4.58-4.60 (m, 2H, $PhCH_2O$), 4.73 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.81 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 4.84 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.85 (d, 1H, $J = 10.8$ Hz, $PhCH_2O$), 4.88 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 6.82-6.83 (m, 2H), 7.19-7.39 (m, 20H), 7.56-7.57 (m, 2H); ^{13}C NMR (150 MHz, $CDCl_3$) δ 69.2, 73.6, 75.3, 75.7, 76.1 (4 X OCH_2Ph , C_6), 78.0, 79.2, 80.9, 86.9, 87.6 (C_1), 119.7, 127.9-128.7, 129.7, 134.3, 138.1, 138.2, 138.4, 138.5, 139.9; MS (ESI) m/z calcd. for $C_{40}H_{39}N_3NaO_5S$ $[M + Na]^+$ 696.3; found 696.4; HRMS m/z calcd. for $C_{40}H_{39}N_3NaO_5S$ $[M + Na]^+$ 696.2508; found 696.2501.

***p*-*N*-Acetamidophenyl 2,3,4,6-tetra-*O*-benzyl-1-thio- β -D-glucopyranoside (3d)—**

Compound **3d** was prepared as described in the synthesis of compound **15** starting from compound **3a** in 84% yield. 1H NMR (600 MHz, $CDCl_3$) δ 2.17 (s, 3H, $NHAc$), 3.46 (t, 1H, $J_{1,2} = 9.6$ Hz, H_2), 3.47-3.49 (m, 1H, H_5), 3.62 (t, 1H, $J = 9.6$ Hz, H_3), 3.69 (t, 1H, $J = 9.6$ Hz, H_4), 3.70 (dd, 1H, $J = 4.2, 10.8$ Hz, H_6), 3.75 (dd, 1H, $J = 1.2, 10.8$ Hz, H_6'), 4.53 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.57 (d, 1H, $J_{1,2} = 9.6$ Hz, H_1), 4.56-4.60 (m, 2H, $PhCH_2O$), 4.73 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.81 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.83 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.87 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.88 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 7.07 (bs, 1H, $NHAc$), 7.17-7.19 (m, 2H), 7.26-7.40 (m, 20H), 7.54-7.55 (m, 2H); ^{13}C NMR (150 MHz, $CDCl_3$) δ 24.9 ($NHCOCH_3$), 73.6, 75.3, 75.6, 76.0, 78.0, 79.3, 81.0, 87.0, 87.9, 89.5 (C_1), 120.2, 127.9-128.7, 133.9, 137.8, 138.2, 169.0; MS (ESI) m/z calcd. for $C_{42}H_{43}NNaO_6S$ $[M + Na]^+$ 712.3; found 712.6; HRMS m/z calcd. for $C_{42}H_{43}NNaO_6S$ $[M + Na]^+$ 712.2709; found 712.2720.

***p*-Methoxyphenyl 2,3,4,6-tetra-*O*-benzyl-1-thio- β -D-glucopyranoside (3e)—**

Compound **3e** was prepared as described in the synthesis of compound **3a** starting from compound **7³²** in 97% yield. 1H NMR (600 MHz, $CDCl_3$) δ 3.44 (t, 1H, $J_{1,2} = 9.6$ Hz, H_2), 3.46-3.47 (m, 1H, H_5), 3.62 (t, 1H, $J = 9.6$ Hz, H_3), 3.69 (t, 1H, $J = 9.6$ Hz, H_4), 3.72-3.76 (m, 1H, H_6), 3.75 (s, 3H, OCH_3), 3.78 (dd, 1H, $J = 1.2, 10.8$ Hz, H_6'), 4.53 (d, 1H, $J_{1,2} = 9.6$ Hz, H_1), 4.54 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.59 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.60 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.74 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.82 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.84 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.89 (d, 1H, $J = 10.2$ Hz, $PhCH_2O$), 4.90 (d, 1H, $J = 10.2$

Hz, PhCH₂O), 6.72-6.76 (m, 2H), 7.19-7.42 (m, 20H), 7.53-7.55 (m, 2H); ¹³C NMR (150 MHz, CDCl₃) δ 55.5 (OCH₃), 69.3, 73.6, 75.3, 75.6, 76.1, 78.0, 79.2, 80.9, 87.0, 88.1(C₁), 114.6, 123.6, 127.8-128.7, 135.4, 138.3-138.6, 139.9, 159.9; MS (ESI) *m/z* calcd. for C₄₁H₄₂NaO₆S [M + Na]⁺ 685.3; found 685.4; HRMS *m/z* calcd. for C₄₁H₄₂NaO₆S [M + Na]⁺ 685.2600; found 685.2614.

***p*-Bromophenyl 2,3,4,6-tetra-*O*-acetyl-1-thio-β-D-glucopyranoside (6)**—Glucose pentaacetate **4** (5 g, 12.9 mmol) and *p*-bromo thiophenol (2.9 g, 15.5 mmol) were dissolved in CH₂Cl₂ under nitrogen at 0 °C. Borontrifluoride etherate (2.4 mL, 19.4 mmol) was added and the reaction mixture was stirred for 18 hours at room temperature. After dilution with CH₂Cl₂, the organic layer was washed with saturated NaHCO₃ and brine. The organic layer was dried over Na₂SO₄, the solvent was removed and the residue was purified by column chromatography on silica gel to afford **6** (6.2 g) in 90% yield. ¹H NMR (600 MHz, CDCl₃) δ 2.00 (s, 3H, COCH₃), 2.02 (s, 3H, COCH₃), 2.09 (s, 3H, COCH₃), 2.10 (s, 3H, COCH₃), 3.70-3.73 (m, 1H, H₅), 4.18 (dd, 1H, *J*=2.4, 12.0 Hz, H₆), 4.21 (dd, 1H, *J*=5.4, 12.0 Hz, H_{6'}), 4.65 (d, 1H, *J*= 10.2 Hz, H₁), 4.94 (t, 1H, *J* = 10.2 Hz, H₂), 5.02 (t, 1H, *J* = 10.2 Hz, H₃), 5.22 (t, 1H, *J* = 10.2 Hz, H₄), 7.37-7.38 (m, 2H), 7.44-7.46 (m, 2H); ¹³C NMR (150 MHz, CDCl₃) δ 20.8-21.0 (4 X COCH₃), 31.2 (*p*-OCH₃), 62.2 (C₆), 68.2, 69.9, 74.1, 76.0, 85.5 (C₁), 119.6, 126.7, 136.0, 141.1, 169.5-170.1 (4 X COCH₃); MS (ESI) *m/z* calcd. for C₂₀H₂₃BrNaO₉S [M + Na]⁺ 541.0; found 541.0; HRMS *m/z* calcd. for C₂₀H₂₃BrNaO₉S [M + Na]⁺ 541.0144; found 541.0124.

***p*-Azidophenyl 2,3,4,6-tetra-*O*-acetyl-1-thio-β-D-glucopyranoside (9)**—The mixture of **5** (181 mg, 0.37 mmol) and SnCl₂.H₂O (336 mg, 1.49 mmol) in EtOH (6 mL) and CH₂Cl₂ (2 mL) was refluxed at 78 °C for 6 hours, and then cooled down to room temperature. EtOH was removed under vacuum followed by addition of EtOAc (20 mL). The resulting mixture was filtered through Celite, extracted with saturated aqueous NaHCO₃ (20 mL) and the organic phase was collected. Purification by flash chromatography gave *p*-aminophenyl 2,3,4,6-tetra-*O*-acetyl-1-thio-β-D-glucopyranoside **8** in 100% yield (170 mg). ¹H NMR (400 MHz, CDCl₃) δ 2.01 (s, 3H, COCH₃), 2.05 (s, 3H, COCH₃), 2.09 (s, 3H, COCH₃), 2.12 (s, 3H, COCH₃), 3.64-3.69 (m, 1H, H₅), 3.82 (bs, 2H, NH₂), 4.18-4.22 (m, 2H, H₆, H_{6'}), 4.51 (d, 1H, *J*_{1,2}=10.0 Hz, H₁), 4.88 (dd, 1H, *J*=9.6, 10.0 Hz, H₂), 5.00 (t, 1H, *J*=9.6 Hz, H₃), 5.19 (t, 1H, *J*=9.6 Hz, H₄), 6.60-6.62 (m, 2H), 7.27-7.31 (m, 2H). Upon isolation, compound **8** (455 mg, 1 mmol) was immediately dissolved in a mixture solvent consisted of 8 N acetic acid (20 mL) and THF (20 mL) at 0 °C. A 5% aqueous solution of NaNO₂ (2.6 mL) was added to this solution and the mixture was stirred for 10 minutes. A 5% aqueous solution of NaN₃ (6.5 mL) was then added and the reaction mixture was allowed to warm up to room temperature in 10 minutes. The reaction was again cooled to 0 °C, neutralized with 10 % aqueous NaOH, extracted with EtOAc, washed with water and brine, and dried over anhydrous Na₂SO₄. After filtration, removal of all solvents under reduced pressure, compound **9** was obtained by flash chromatography in 59% yield (280 mg). ¹H NMR (600 MHz, CDCl₃) δ 1.99 (s, 3H, COCH₃), 2.02 (s, 3H, COCH₃), 2.09 (s, 3H, COCH₃), 2.10 (s, 3H, COCH₃), 3.70-3.73 (m, 1H, H₅), 4.18 (dd, 1H, *J*=1.8, 12.0 Hz, H₆), 4.21 (dd, 1H, *J*=4.8, 12.0 Hz, H_{6'}), 4.62 (d, 1H, *J*_{1,2}=9.6 Hz, H₁), 4.91 (t, 1H, *J*=9.6 Hz, H₂), 5.01 (t, 1H, *J*=9.6 Hz, H₃), 5.21 (t, 1H, *J*=9.6 Hz, H₄), 6.98-6.99 (m, 2H), 7.49-7.51 (m, 2H); ¹³C NMR (150 MHz, CDCl₃) δ 20.8-21.0, 31.2 (4 X COCH₃), 62.2 (C₆), 68.2, 69.9, 74.1, 76.0, 85.5 (C₁), 119.6, 126.7, 136.0, 141.1, 169.5-170.1 (4 X COCH₃); MS (ESI) *m/z* calcd. for C₂₀H₂₃N₃NaO₉S [M + Na]⁺ 504.1; found 504.1; HRMS *m/z* calcd. for C₂₀H₂₃N₃NaO₉S [M + Na]⁺ 504.1053; found 504.1028.

***p*-N-Acetamidophenyl 2,3,4,6-tetra-*O*-acetyl-1-thio-β-D-glucopyranoside (15)**—The mixture of **5** (1.46 g, 3.00 mmol) and SnCl₂.H₂O (1.38 g, 6.12 mmol) in EtOH (30 mL) and CH₂Cl₂ (10 mL) was refluxed at 78 °C for 6 hours, and then cooled down to room

temperature. A solution of acetic anhydride (0.94 mL, 10.0 mmol) was added to this solution and the mixture was stirred for 2 hours at room temperature, then the reaction mixture was neutralized with solid NaHCO₃, extracted with EtOAc, washed with water and brine, and dried over anhydrous Na₂SO₄. After filtration and removal of all solvents under reduced pressure, compound **15** was obtained by flash chromatography in 63% yield (0.94 g). ¹H NMR (400 MHz, CDCl₃) δ 1.98 (s, 3H, COCH₃), 2.01 (s, 3H, COCH₃), 2.08 (s, 3H, COCH₃), 2.09 (s, 3H, COCH₃), 2.18 (s, 3H, NHCH₃), 3.67-3.71 (m, 1H, H₅), 4.11-4.18 (m, 2H, 2 X H₆), 4.59 (d, 1H, *J*=9.6 Hz, H₁), 4.91 (t, 1H, *J*=9.6 Hz, H₂), 5.01 (t, 1H, *J*=9.6 Hz, H₃), 5.20 (t, 1H, *J*=9.6 Hz, H₄), 7.38 (m, 1H, NHAc), 7.44-7.49 (m, 4H); ¹³C NMR (100 MHz, CDCl₃) δ 20.8-21.0 (4 X COCH₃), 24.9 (NHCOCH₃), 62.3, 68.3, 70.1, 74.2, 76.0, 86.0 (C₁), 120.1, 126.0, 133.7, 135.1, 138.8, 169.5, 169.6, 170.4 170.9 (4 X COCH₃); MS (ESI) *m/z* calcd. for C₂₂H₂₇NNaO₁₀S [M + Na]⁺ 520.1; found 520.2; HRMS *m/z* calcd. for C₂₂H₂₇NNaO₁₀S [M + Na]⁺ 520.1253; found 520.1233.

3-Bromopropanyl-2-phthalimido-3,6-di-O-benzyl-2-deoxy-β-D-glucopyranoside (17)—To a mixture of **2a** (626 mg, 1 mmol), 3-bromopropanol (0.26 mL, 3 mmol), and MS 4Å (800 mg) in CH₂Cl₂ (20 mL) was added NIS (587 mg, 3.3 mmol) under nitrogen. The mixture was stirred for 30 minutes at -40 °C. TfOH (0.4 mL, 0.1 M in Et₂O) was added. The reaction mixture was stirred for 1 hour, during which the temperature was raised to room temperature. The reaction mixture was then diluted with CH₂Cl₂ and filtered through celite. The solution was washed with saturated aqueous Na₂S₂O₃/NaHCO₃, brine, and dried over Na₂SO₄. The residue was purified by flash chromatography to give the compound **17** in 74% yield (451 mg). ¹H NMR (400 MHz, CDCl₃) δ 1.81-1.89, 1.96-2.04 (m, 2H, BrCH₂CH₂CH₂O), 3.03 (d, 1H, *J* = 2.4 Hz, 4-OH), 3.14-3.26 (m, 2H, BrCH₂CH₂CH₂O), 3.50-3.56, 3.63-3.68 (m, 2H, BrCH₂CH₂CH₂O), 3.77-3.91 (m, 4H), 4.14 (dd, 1H, *J* = 8.4, 9.6 Hz), 4.25 (dd, 1H, *J* = 8.4, 9.6 Hz), 4.54 (d, 1H, *J*=12.0 Hz, PhCH₂O), 4.59 (d, 1H, *J*=12.0 Hz, PhCH₂O), 4.66 (d, 1H, *J*=12.0 Hz, PhCH₂O), 4.75 (d, 1H, *J*=12.0 Hz, PhCH₂O), 5.14 (d, 1H, *J* = 8.4 Hz, H₁), 6.94-7.81 (m, 14H); ¹³C NMR (100 MHz, CDCl₃) δ 30.4, 32.4, 55.6 (BrCH₂CH₂CH₂), 67.1, 70.7, 74.0, 74.1, 74.3, 74.6, 78.9, 98.7 (C₁), 123.6, 127.7-128.8, 131.8, 134.2, 138.0, 138.4, 167.8, 168.4; MS (ESI) *m/z* calcd. for C₃₁H₃₂BrNNaO₇ [M + Na]⁺ 632.1; found 632.2; HRMS *m/z* calcd. for C₃₁H₃₂BrNNaO₇ [M + Na]⁺ 632.1260; found 632.1283.

Supplementary Material

Refer to Web version on PubMed Central for supplementary material.

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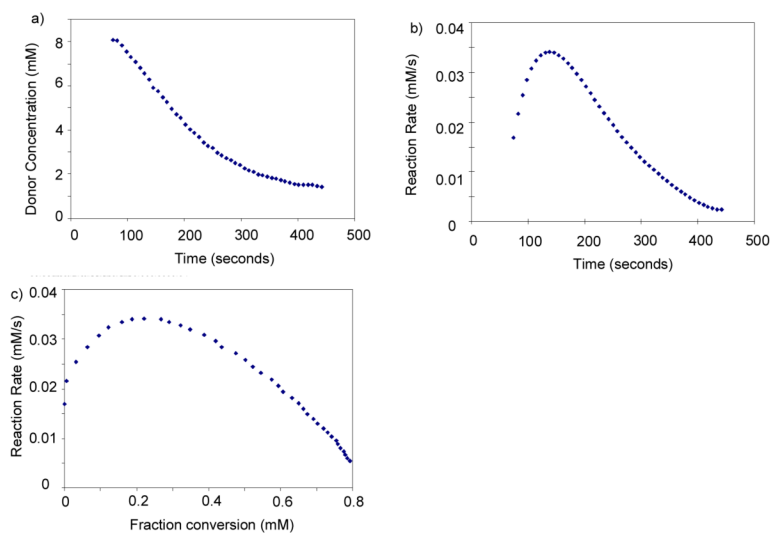


Figure 1. Changes of a) donor concentration, b) reaction rate vs. time; and c) reaction rate vs. donor fraction conversion as measured by $^1\text{H-NMR}$.

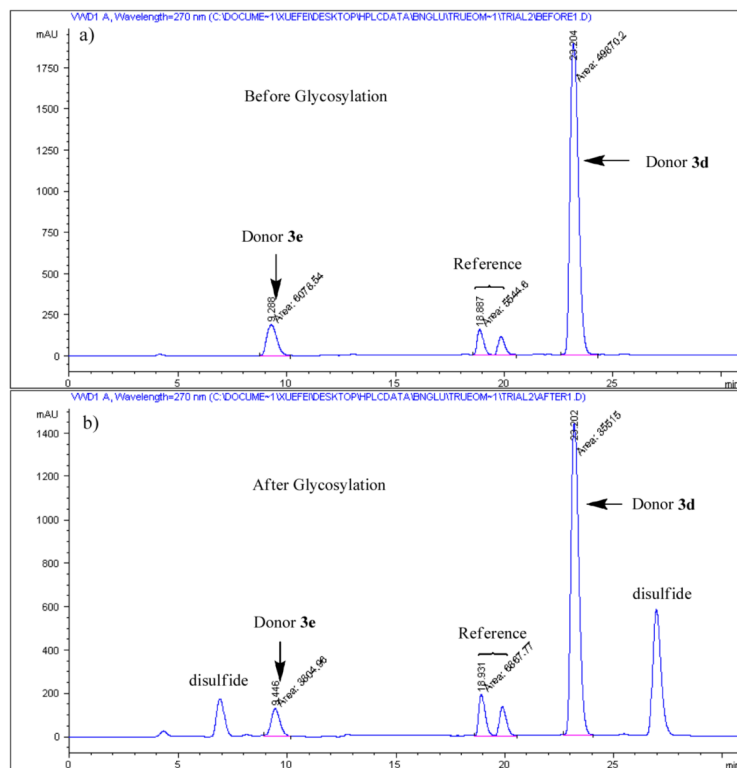


Figure 2. HPLC chromatograms a) before and b) after competitive glycosylation between donors **3d** and **3e** (1,3,4,6-tetra-*O*-acetyl-2-deoxy-2-*N*-phthalimido- α/β -D-glucopyranosides were used as the reference for these reactions. Same amounts of donors **3d** and **3e** were added before glycosylation. The large disparity in UV absorbance is due to the difference in extinction coefficients. *p*-Tolyl *p*-methoxyphenyldisulfide and *p*-tolyl *p*-acetamidophenyldisulfide were formed and detected in the chromatogram after glycosylation).

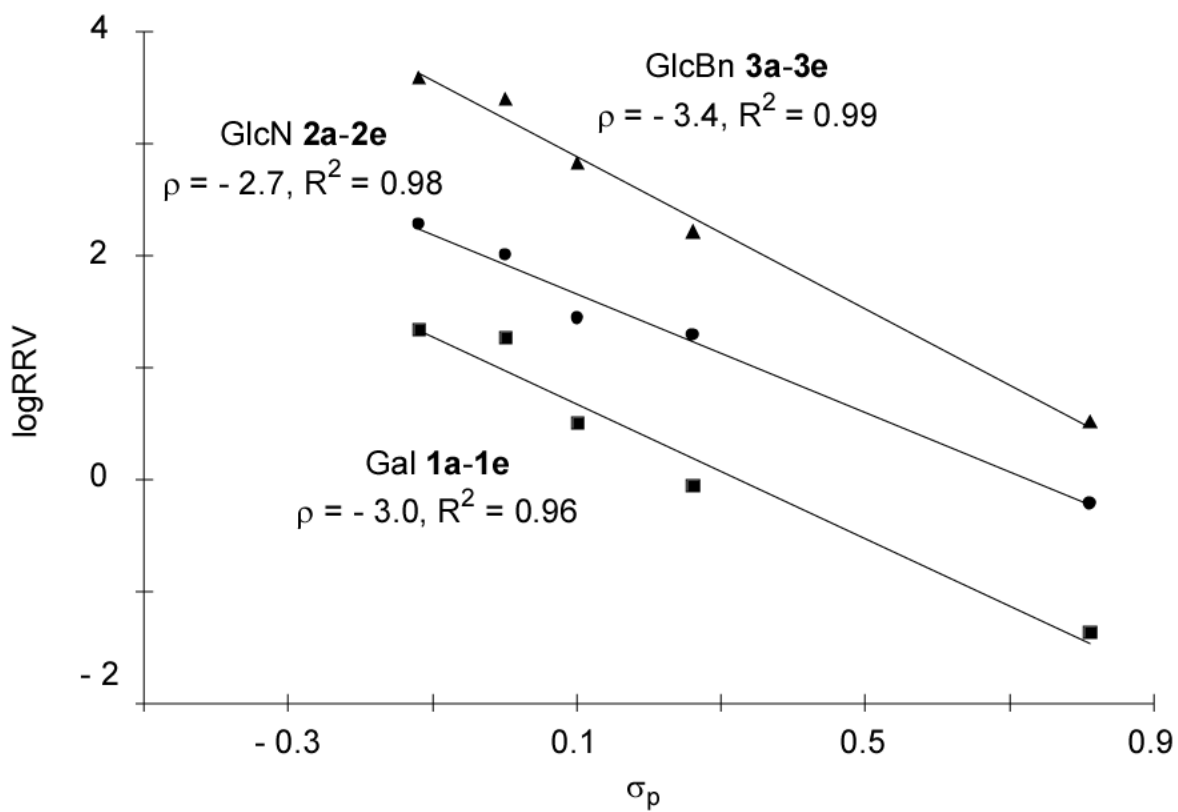


Figure 3. Hammett plots of log(RRV) vs σ_p using methanol as the acceptor

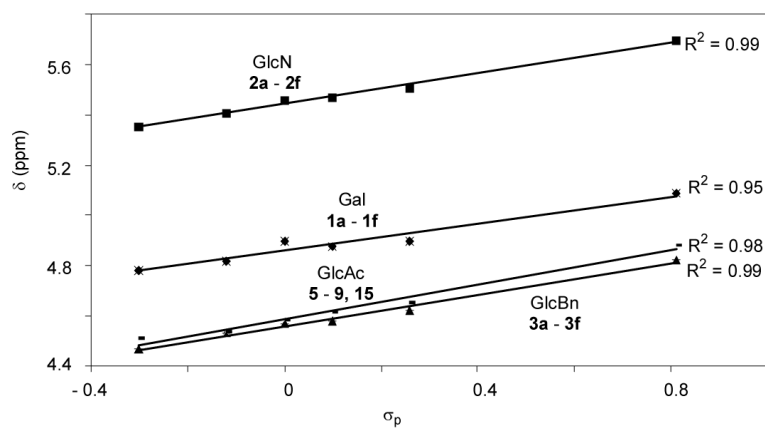


Figure 4. Correlations of Anomeric Proton Chemical Shifts of the Four Donor Series with σ_p Values.

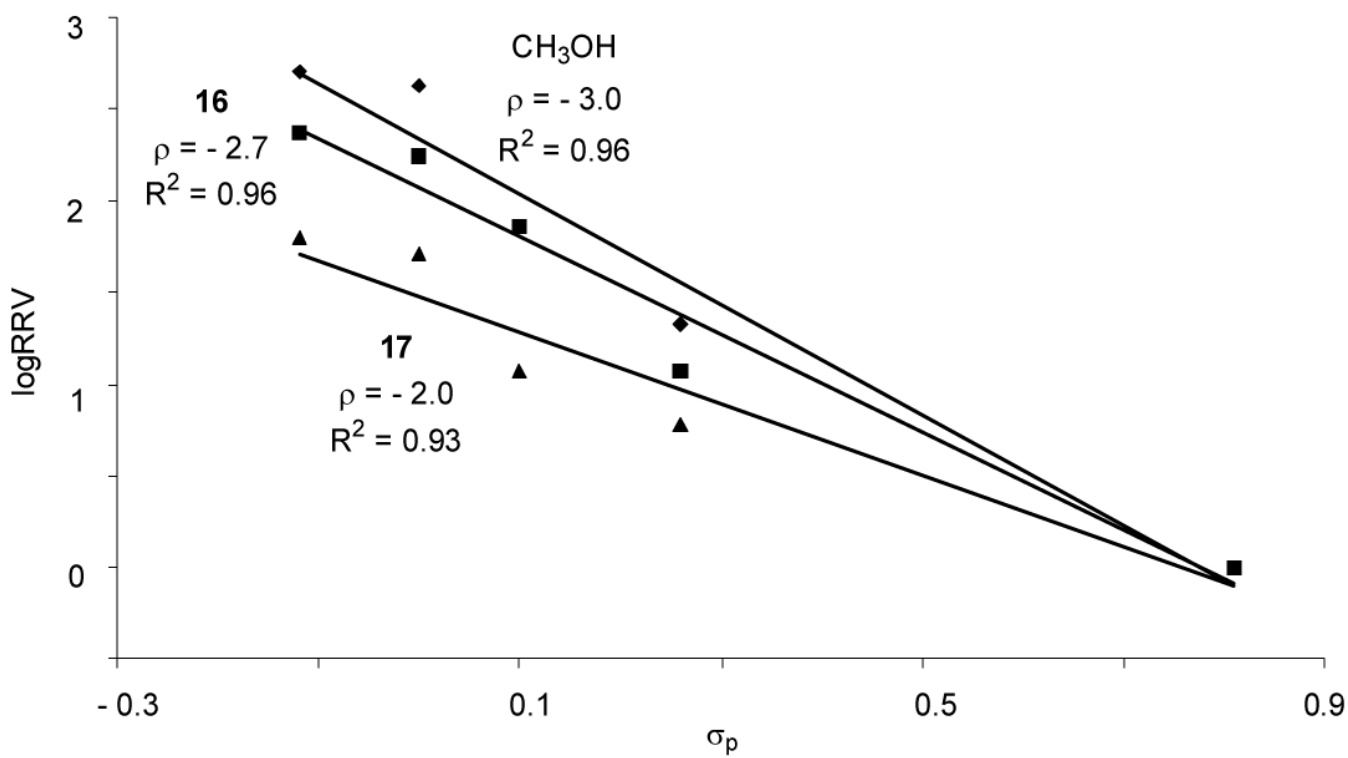


Figure 5. Hammett plots of log(RRV) of donors **1a-1e** vs σ_p (methanol, galactoside **16** and glucosamine **17** as the respective acceptor and CH₂Cl₂ as the glycosylation solvent)

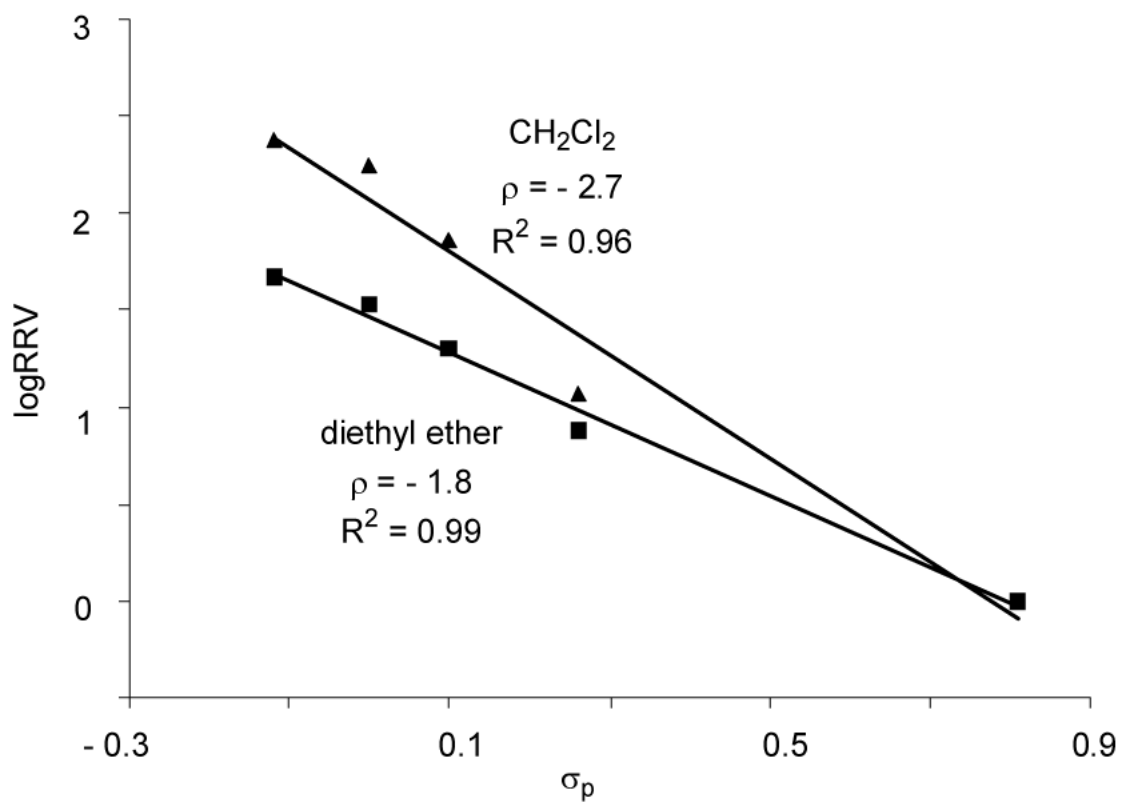


Figure 6. Hammett plots of $\log(RRV)$ of donors **1a-1e** vs σ_p (galactoside **16** as the acceptor; promoter *p*-TolSCI/AgOTf; CH_2Cl_2 or diethyl ether as the glycosylation solvent).

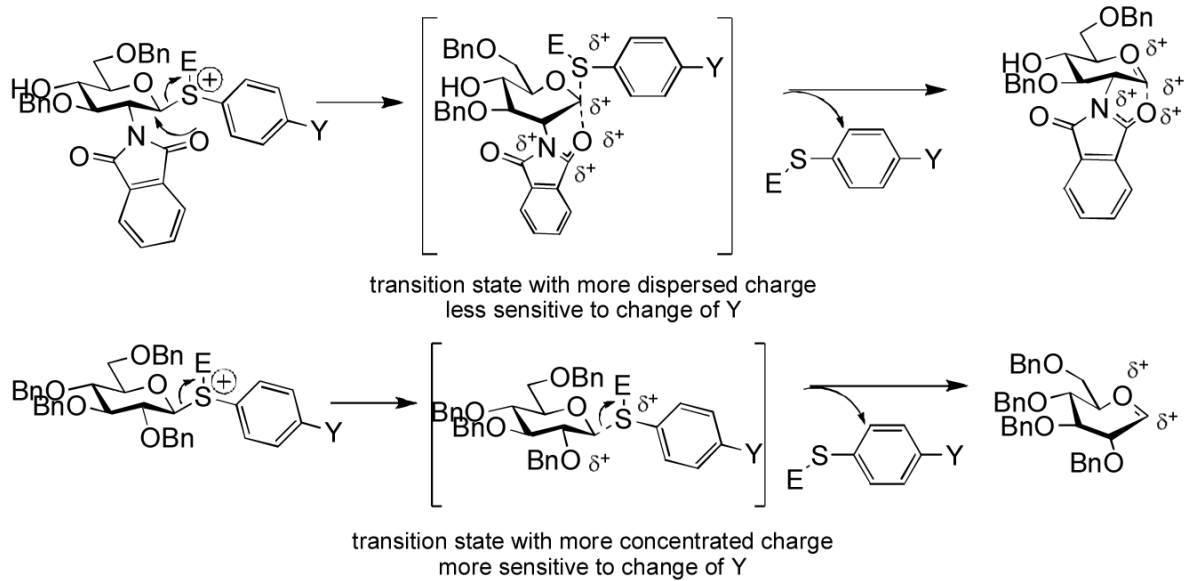
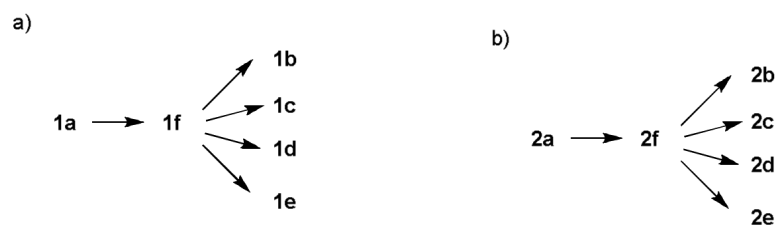
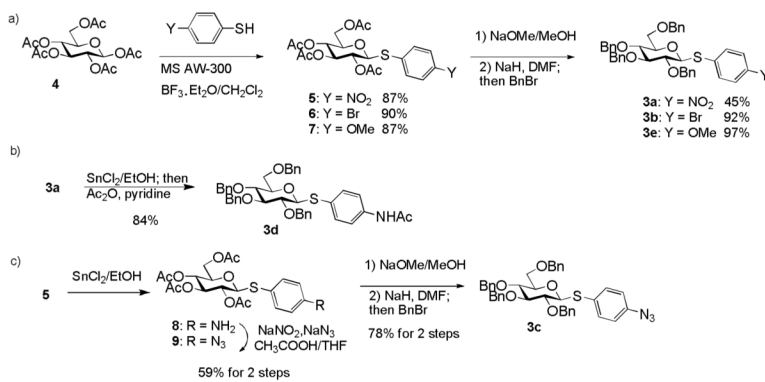


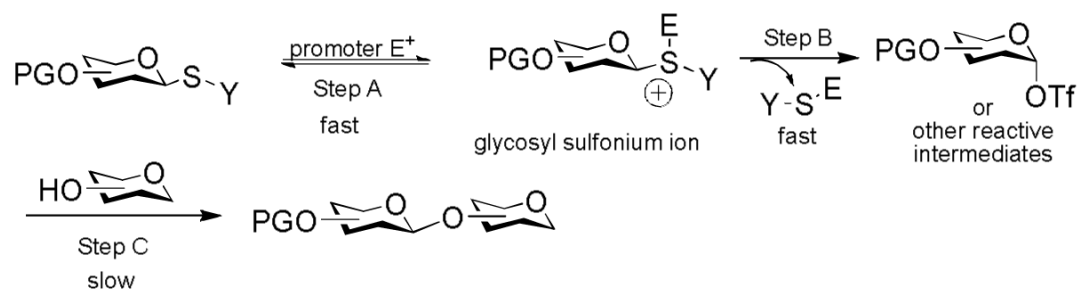
Figure 7.
The smaller absolute ρ values for GlcN donor series can be rationalized by neighboring group participation of the phthalimido group.

**Scheme 1.**

Post-synthetic modification of aglycons rapidly generated a panel of building blocks with multiple levels of anomeric reactivities.



Scheme 1.
Synthesis of Thioglucoside Donors **3a** - **3e**.



Scheme 2.
General Mechanism for Glycosylation of Thioglycosides

Table 1
Relative Reactivities of GlcN Donor Series **2a-2e** Measured with Methanol as the Acceptor

<i>Entry number</i>	<i>Donor pair (more reactive/less reactive)</i>	<i>Experiment #1</i>	<i>Experiment #2</i>	<i>Experiment #3</i>	<i>Average value</i>
1	10/2a	4.49	4.58	4.37	4.49
2	2b/10	7.06	7.04	7.51	7.21
3	2d/2b	5.27	5.32	4.95	5.18
4	2d/2c	3.97	3.35	3.98	3.74
5	2e/2d	1.88	1.89	1.93	1.90

Table 2
 Relative Reactivity Values (RRVs) of Three Donor Series Measured with Methanol as the Acceptor (Reactivity of thiomannoside **11** was set as 1, CH₂Cl₂ as the glycosylation solvent)

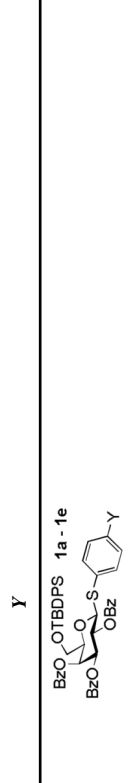
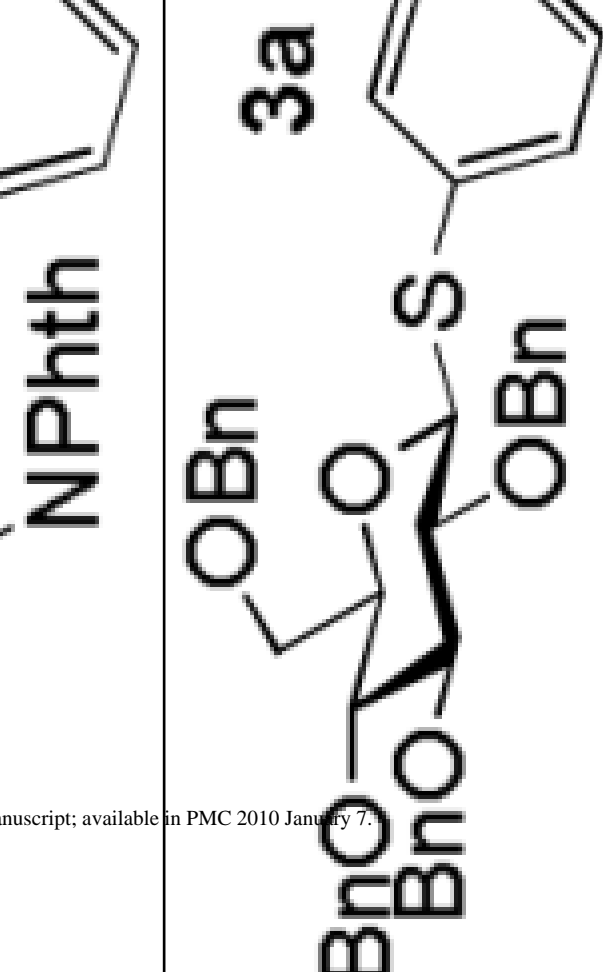
	<i>NO₂</i>	<i>Br</i>	<i>N₃</i>	<i>NHAc</i>	<i>OMe</i>
 <p style="text-align: center;">2a - 2e</p>	0.042 (1a)	0.89 (1b)	3.12 (1c)	18.1 (1d)	21.7 (1e)
	0.601 (2a)	19.2 (2b)	27.1 (2c)	101 (2d)	191 (2e)
 <p style="text-align: center;">3a - 3e</p>	3.19 (3a)	161 (3b)	657 (3c)	2482 (3d)	3772 (3e)

Table 3
Comparison between Calculated and Experimental log(RRV) values

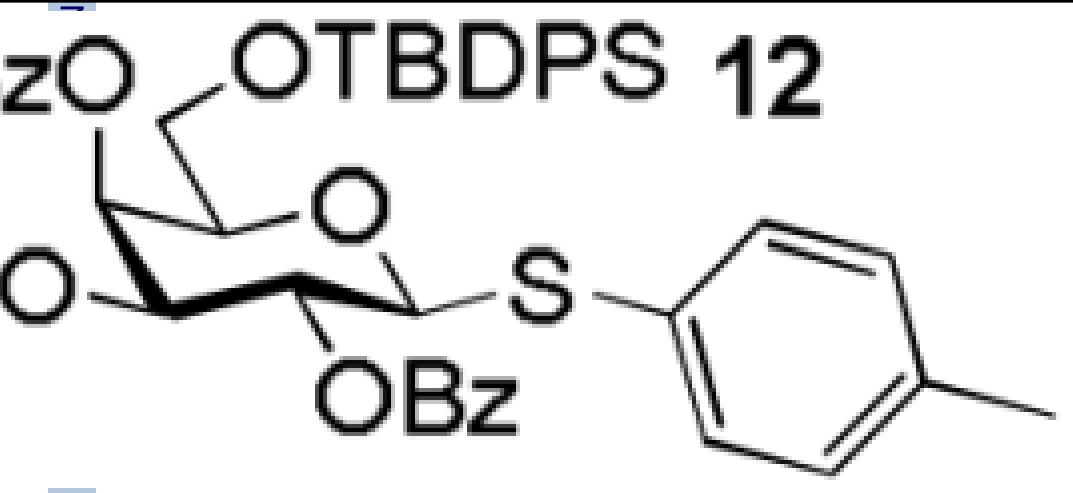
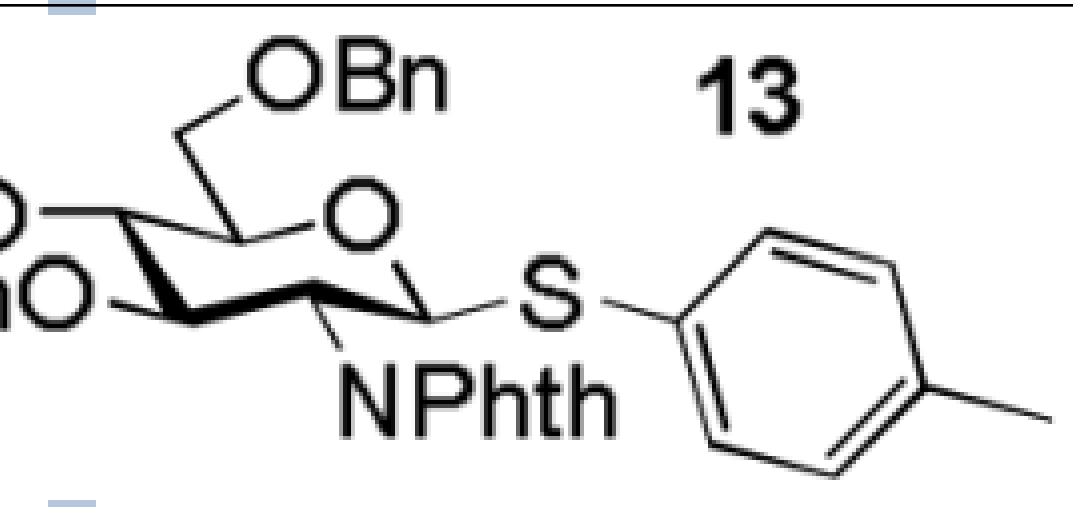
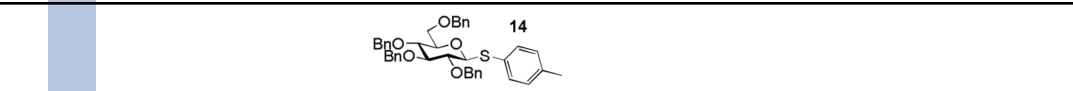
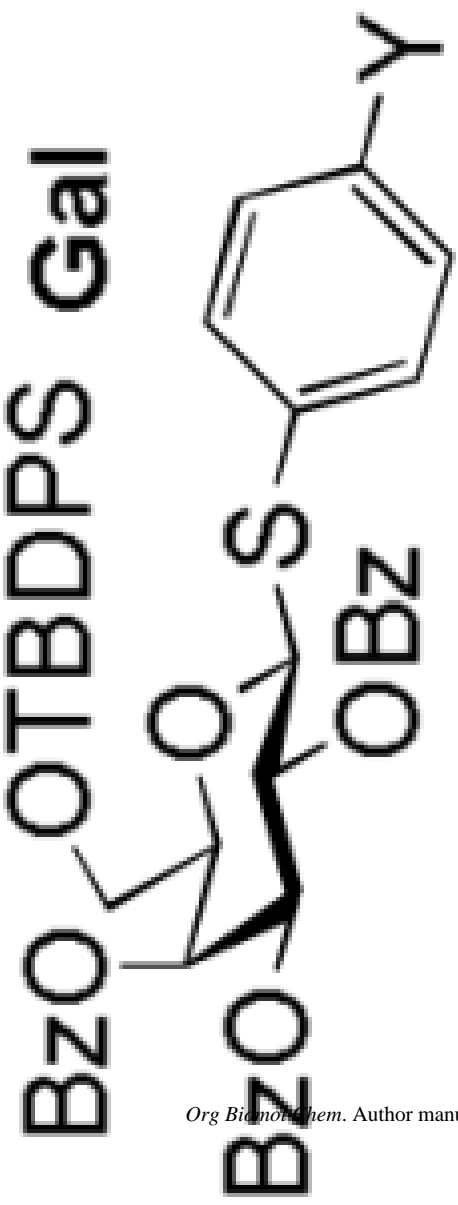
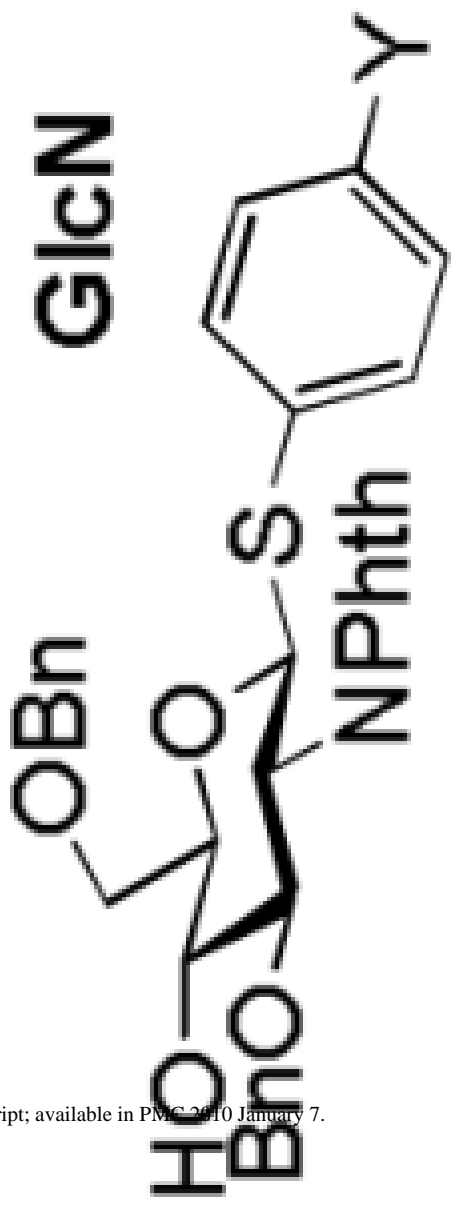
<i>Donor</i>	<i>Calculated log(RRV) value</i>	<i>Experimental log(RRV)</i>
 <p>12</p>	1.38	1.28
 <p>13</p>	2.28	2.41
 <p>14</p>	3.68	3.42

Table 4
Chemical Shifts of Anomeric Protons of Four Series of Donors in CDCl₃

	<i>NO</i> ₂	<i>Br</i>	<i>N</i> ₃	<i>NHAc</i>	<i>OMe</i>	<i>NH</i> ₂
<p>BZO OTBDPS Gal</p> 	5.09 (1a)	4.90 (1b)	4.88 (1c)	4.90 (1d)	4.82 (1e)	4.78 (1f)
<p>BZO GlcN</p> 	5.69 (2a)	5.50 (2b)	5.47 (2c)	5.46 (2d)	5.40 (2e)	5.35 (2f)

	NO_2	<i>Br</i>	N_3	<i>NHAc</i>	<i>OMe</i>	NH_2
<p style="text-align: center;">Y</p> <p style="text-align: center;">GlicBn</p>	4.82 (3a)	4.62 (3b)	4.58 (3c)	4.57 (3d)	4.53 (3e)	4.47 (3f)
<p style="text-align: center;">GlicAC</p>	4.88 (5)	4.65 (6)	4.62 (9)	4.59 (15)	4.54 (7)	4.51 (8)

Table 5
Relative Reactivities of Donors **1a-1e** Using Methanol, **16** and **17** as Acceptors (reactivity of donor **1a** was set as 1)

Entry	Acceptor (reaction solvent)	Y=NO ₂ (1a)	Y=Br (1b)	Y=N ₃ (1c)	Y=NHAc (1d)	Y=OMe (1e)
1	Methanol (CH ₂ Cl ₂)	1	21.2	74.3	431.0	516.7
2	Acceptor 16 (CH ₂ Cl ₂)	1	11.6	72.3	177.5	234.2
3	Acceptor 17 (CH ₂ Cl ₂)	1	5.95	11.8	51.4	62.5
4	Acceptor 16 (diethyl ether)	1	7.63	20.0	34.3	47.1