

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## A third monoclinic polymorph of 3,4,5-trihydroxybenzoic acid monohydrate

Güneş Demirtaş,\* Necmi Dege and Orhan Büyükgüngör

Department of Physics, Arts and Sciences Faculty, Ondokuz Mayıs University, 55139 Samsun, Turkey

Correspondence e-mail: gunesd@omu.edu.tr

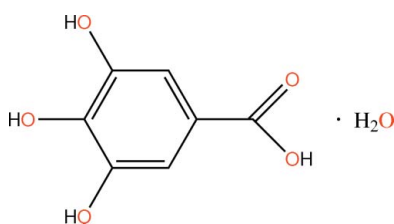
Received 30 April 2011; accepted 17 May 2011

Key indicators: single-crystal X-ray study;  $T = 296$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.039;  $wR$  factor = 0.104; data-to-parameter ratio = 11.8.

The title compound,  $\text{C}_7\text{H}_6\text{O}_5 \cdot \text{H}_2\text{O}$ , is a new polymorph of the structures reported by Jiang *et al.* (2000) [*Acta Cryst.* **C56**, 594–595] and Okabe *et al.* (2001) [*Acta Cryst.* **E57**, o764–o766]. The gallic acid molecule is essentially planar (r.m.s. deviation = 0.550 Å). An intramolecular  $\text{O}-\text{H} \cdots \text{O}$  hydrogen bond occurs in the gallic acid molecule, which is linked to the water molecule by a further  $\text{O}-\text{H} \cdots \text{O}$  hydrogen bond. In the crystal, the components are linked by  $\text{O}-\text{H} \cdots \text{O}$  hydrogen bonds. The hydrogen-bonding pattern differs from those reported for the previous polymorphs.

## Related literature

For the biological activity of gallic acid, see: Lu *et al.* (2006); Madlener *et al.* (2007). For the previously reported polymorphs, see: Jiang *et al.* (2000); Okabe *et al.* (2001). For a related structure, see: Genç *et al.* (2004).



## Experimental

## Crystal data

 $\text{C}_7\text{H}_6\text{O}_5 \cdot \text{H}_2\text{O}$  $M_r = 188.13$ Monoclinic,  $P2_1/c$  $a = 9.7943$  (7) Å $b = 3.6122$  (2) Å $c = 21.5905$  (15) Å $\beta = 91.268$  (6)° $V = 763.66$  (9) Å<sup>3</sup> $Z = 4$ Mo  $K\alpha$  radiation $\mu = 0.15$  mm<sup>-1</sup> $T = 296$  K

0.61 × 0.28 × 0.09 mm

## Data collection

Stoe IPDS 2 diffractometer

Absorption correction: integration

 $(X\text{-RED32; Stoe \& Cie, 2002})$  $T_{\min} = 0.948, T_{\max} = 0.986$ 

4558 measured reflections

1502 independent reflections

1262 reflections with  $I > 2\sigma(I)$  $R_{\text{int}} = 0.042$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.039$  $wR(F^2) = 0.104$  $S = 1.04$ 

1502 reflections

127 parameters

3 restraints

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\max} = 0.15$  e Å<sup>-3</sup> $\Delta\rho_{\min} = -0.23$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
$\text{O2}-\text{H3} \cdots \text{O3}$	0.82	2.26	2.7006 (18)	114
$\text{O5}-\text{H6} \cdots \text{O6}$	0.82	1.85	2.6542 (17)	167
$\text{O1}-\text{H2} \cdots \text{O6}^i$	0.82	2.03	2.7539 (19)	147
$\text{O2}-\text{H3} \cdots \text{O3}^{ii}$	0.82	2.52	3.1721 (19)	137
$\text{O3}-\text{H4} \cdots \text{O4}^{iii}$	0.82	1.94	2.7154 (19)	158
$\text{O6}-\text{H6A} \cdots \text{O2}^{iv}$	0.83 (2)	2.03 (2)	2.8237 (19)	161 (3)
$\text{O6}-\text{H6B} \cdots \text{O1}^v$	0.82 (2)	2.00 (2)	2.814 (2)	171 (5)

Symmetry codes: (i)  $-x + 1, -y + 2, -z + 1$ ; (ii)  $-x + 2, y - \frac{1}{2}, -z + \frac{3}{2}$ ; (iii)  $-x + 1, y - \frac{1}{2}, -z + \frac{3}{2}$ ; (iv)  $x - 1, y + 1, z$ ; (v)  $x - 1, y, z$ .

Table 2

Other crystal structures of gallic acid monohydrate (Å, °).

	1	2
Unit-cell parameters	$a = 5.794$ (4) $b = 4.719$ (5) $c = 28.688$ (5) $\beta = 95.08$ (3) $V = 781.4$ (3)	$a = 14.15$ (1) $b = 3.622$ (9) $c = 15.028$ (10) $\beta = 97.52$ (7) $V = 764$ (1)
Space group	Monoclinic, $P2_1/c$	Monoclinic, $P2_1/n$
Reference	Jiang <i>et al.</i> (2000)	Okabe <i>et al.</i> (2001)

Data collection: *X-AREA* (Stoe & Cie, 2002); cell refinement: *X-AREA*; data reduction: *X-RED32* (Stoe & Cie, 2002); program(s) used to solve structure: *WinGX* (Farrugia, 1997) and *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999) and *PLATON* (Spek, 2009).

The authors thank the Ondokuz Mayıs University Research Fund for financial support of this project (project No. PYO.FEN.1904.09.006).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BX2352).

## References

- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.  
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.  
Genç, S., Dege, N., Çetin, A., Cansız, A., Şekerci, M. & Dinçer, M. (2004). *Acta Cryst.* **E60**, o1580–o1582.  
Jiang, R.-W., Ming, D.-S., But, P. P. H. & Mak, T. C. W. (2000). *Acta Cryst.* **C56**, 594–595.

Lu, Z., Nie, G., Belton, P. S., Tang, H. & Zhao, B. (2006). *Neurochem. Int.* **48**, 263–274.

Madlener, S., Illmer, C., Horvath, Z., Saiko, P., Losert, A., Herbacek, I., Grusch, M., Elford, H. L., Krupitza, G., Bernhaus, A., Fritzer-Szekeres, M. & Szekeres, T. (2007). *Cancer Lett.* **245**, 156–162.

Okabe, N., Kyoyama, H. & Suzuki, M. (2001). *Acta Cryst.* **E57**, o764–o766.

Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.

Stoe & Cie (2002). *X-AREA* and *X-RED32*. Stoe & Cie, Darmstadt, Germany.

**supplementary materials**

*Acta Cryst.* (2011). E67, o1509-o1510 [ doi:10.1107/S1600536811018848 ]

## A third monoclinic polymorph of 3,4,5-trihydroxybenzoic acid monohydrate

G. Demirtas, N. Dege and O. Büyükgüngör

### Comment

Gallic acid and its derivatives are a group of naturally occurring polyphenol antioxidants which have recently been shown to have potential healthy effects as an excellent free radical scavenger (Lu *et al.*, 2006). In addition, gallic acid block DNA synthesis in leukemia cells (Madlener *et al.*, 2007). Up to now, the different crystal structure of gallic acid monohydrate form have been reported by (Jiang *et al.*, 2000; Okabe *et al.*, 2001). The unit-cell parameters belong to these crystal structures have given at Table 2.

We report here the crystal structure of the title compound,  $C_7H_6O_5 \cdot H_2O$ . The asymmetric unit of title compound,  $C_7H_6O_5 \cdot H_2O$ , consists of one 3,4,5-trihydroxybenzoic acid molecule and one water molecule (Fig. 1). The C—C bond distances range from 1.378 (2) Å to 1.480 (2) Å. The longest C—C bond distance is between C6 and C7 with 1.480 Å. The C—O bond distance are range from 1.212 (2) Å to 1.3723 (19) Å. The shortest C—O bond distance is between C7 and O4 with 1.212 (2) Å. The C—O bond distance for different crystal structure was given 1.359 (2) Å by (Genç *et al.*, 2004).

The crystal structure has two intramolecular hydrogen bonds and five intermolecular hydrogen bonds (Table 1). The O2—H3 $\cdots$ O3 and O5—H6 $\cdots$ O6 hydrogen bonds present in the asymmetric unit. In the crystal structure, the gallic acid molecules and water molecules are linked by the O6—H6A $\cdots$ O2, O6—H6B $\cdots$ O1, O5—H6 $\cdots$ O6 and O1—H2 $\cdots$ O6 hydrogen bonds. While the C—H $\cdots$ O hydrogen bond is present in crystal structure reported by (Jiang *et al.*, 2000), there isn't present in our crystal structure. The *via* hydrogen atom of water molecule makes bifurcated hydrogen bond in crystal structure reported by (Jiang *et al.*, 2000), but this kind hydrogen bond is not present in our crystal structure. In crystal structure reported by (Okabe *et al.*, 2001), the hydroxyl groups make two intramolecular hydrogen bonds.

The torsion angles at ring belong to gallic acid molecule are range from 0.0 (3)° to 1.9 (3)°. Therefore, the gallic acid molecule close to planar. The torsion angles for the C2—C1—C6—C7 and C4—C5—C6—C7 are 179.51 (16)° and -179.08 (17)°, respectively. The O4—C7—O5 angle is 122.82 (15)°.

### Experimental

A hot solution (60 °C) of gallic acid (0.002 mol, 0.340 g) in distilled water (approximately 20 ml) was gradually added to a hot stirring solution of magnesium sulfate heptahydrate ( $MgSO_4 \cdot 7H_2O$ ) (0.001 mol, 0.246 g) in distilled water (approximately 20 ml). The obtained mixture was stirred on a hot plate with slow evaporation to 30 ml at 60° C. The mixture was left for crystallization and in a few days, the crystals were obtained by slow evaporation from the solution at room temperature.

### Refinement

H6A and H6B atoms were located in a difference map and were refined isotropically, with O—H distances in the range 0.830 (18)-0.824 (19) Å, and the H $\cdots$ H distance of 1.320 (2) Å. The other H atoms were positioned geometrically and refined using a riding model, with C—H = 0.930 Å, O—H = 0.820 Å, and  $U_{iso}(H) = 1.2U_{eq}(C)$  and  $1.5U_{eq}(O)$ .

## Figures

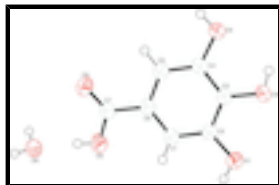


Fig. 1. The asymmetric unit of the title compound, showing the atomic numbering scheme. Displacement ellipsoids are drawn at the 50% probability level.

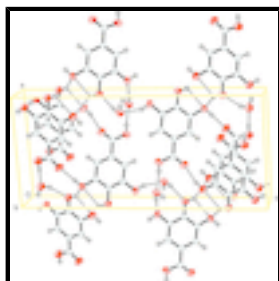


Fig. 2. Packing diagram for (I), showing the three-dimensional hydrogen bonding array. Hydrogen bonds are shown as dashed lines.

## 3,4,5-trihydroxybenzoic acid monohydrate

### Crystal data

$C_7H_6O_5 \cdot H_2O$

$M_r = 188.13$

Monoclinic,  $P2_1/c$

Hall symbol: -P 2ybc

$a = 9.7943$  (7) Å

$b = 3.6122$  (2) Å

$c = 21.5905$  (15) Å

$\beta = 91.268$  (6)°

$V = 763.66$  (9) Å<sup>3</sup>

$Z = 4$

$F(000) = 392$

$D_x = 1.636$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 6472 reflections

$\theta = 1.9$ – $27.6$ °

$\mu = 0.15$  mm<sup>-1</sup>

$T = 296$  K

Needle, pale brown

$0.61 \times 0.28 \times 0.09$  mm

### Data collection

Stoe IPDS 2  
diffractometer

Radiation source: fine-focus sealed tube  
graphite

w-scan rotation

Absorption correction: integration  
(*X-RED32*; Stoe & Cie, 2002)

$T_{\min} = 0.948$ ,  $T_{\max} = 0.986$

4558 measured reflections

1502 independent reflections

1262 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.042$

$\theta_{\max} = 26.0$ °,  $\theta_{\min} = 1.9$ °

$h = -9 \rightarrow 12$

$k = -4 \rightarrow 4$

$l = -26 \rightarrow 26$

### Refinement

Refinement on  $F^2$

Secondary atom site location: difference Fourier map

Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.039$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.104$	$w = 1/[\sigma^2(F_o^2) + (0.0466P)^2 + 0.2493P]$
$S = 1.04$	where $P = (F_o^2 + 2F_c^2)/3$
1502 reflections	$(\Delta/\sigma)_{\max} < 0.001$
127 parameters	$\Delta\rho_{\max} = 0.15 \text{ e } \text{\AA}^{-3}$
3 restraints	$\Delta\rho_{\min} = -0.23 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: <i>SHELXL97</i> (Sheldrick, 2008), $F_c^* = kF_c[1+0.001xF_c^2\lambda^3/\sin(2\theta)]^{-1/4}$
	Extinction coefficient: 0.023 (5)

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.67337 (16)	0.6332 (5)	0.58435 (8)	0.0369 (4)
H1	0.6383	0.7058	0.5459	0.044*
C2	0.81149 (16)	0.5670 (5)	0.59282 (8)	0.0349 (4)
C3	0.86369 (15)	0.4587 (5)	0.65029 (8)	0.0356 (4)
C4	0.77763 (17)	0.4238 (5)	0.70009 (8)	0.0395 (4)
C5	0.63957 (17)	0.4863 (5)	0.69214 (8)	0.0414 (4)
H5	0.5814	0.4596	0.7253	0.050*
C6	0.58767 (16)	0.5895 (5)	0.63421 (8)	0.0370 (4)
C7	0.43925 (16)	0.6603 (5)	0.62739 (8)	0.0408 (4)
O1	0.90229 (12)	0.5986 (4)	0.54561 (6)	0.0480 (4)
H2	0.8613	0.6599	0.5137	0.072*
O2	1.00108 (11)	0.3881 (4)	0.65588 (6)	0.0469 (4)
H3	1.0200	0.3265	0.6916	0.070*
O3	0.83973 (12)	0.3290 (5)	0.75521 (6)	0.0546 (4)
H4	0.7823	0.3169	0.7822	0.082*
O4	0.35878 (13)	0.5981 (5)	0.66792 (6)	0.0583 (4)
O5	0.40210 (12)	0.7959 (5)	0.57298 (6)	0.0567 (4)
H6	0.3193	0.8292	0.5719	0.085*
O6	0.14365 (13)	1.0125 (5)	0.56257 (7)	0.0499 (4)

## supplementary materials

H6A	0.120 (3)	1.147 (8)	0.5916 (12)	0.102 (11)*
H6B	0.077 (4)	0.879 (14)	0.555 (2)	0.24 (3)*

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.0338 (8)	0.0434 (9)	0.0333 (9)	0.0031 (7)	-0.0022 (7)	0.0016 (7)
C2	0.0310 (8)	0.0418 (9)	0.0321 (9)	0.0009 (7)	0.0036 (6)	0.0008 (7)
C3	0.0264 (7)	0.0457 (9)	0.0346 (9)	0.0013 (7)	-0.0009 (6)	-0.0010 (7)
C4	0.0337 (8)	0.0530 (11)	0.0316 (9)	0.0013 (7)	-0.0021 (6)	0.0037 (8)
C5	0.0322 (8)	0.0572 (11)	0.0349 (9)	0.0039 (7)	0.0042 (7)	0.0038 (8)
C6	0.0299 (8)	0.0444 (9)	0.0367 (9)	0.0041 (7)	0.0002 (7)	-0.0001 (7)
C7	0.0319 (8)	0.0525 (10)	0.0381 (10)	0.0049 (7)	0.0003 (7)	-0.0008 (8)
O1	0.0361 (6)	0.0738 (10)	0.0342 (7)	0.0052 (6)	0.0051 (5)	0.0085 (6)
O2	0.0286 (6)	0.0775 (10)	0.0345 (7)	0.0051 (6)	-0.0026 (5)	0.0035 (6)
O3	0.0334 (6)	0.0984 (12)	0.0321 (7)	0.0070 (7)	-0.0005 (5)	0.0138 (7)
O4	0.0335 (6)	0.0953 (12)	0.0465 (8)	0.0101 (7)	0.0074 (6)	0.0062 (8)
O5	0.0330 (6)	0.0882 (11)	0.0488 (8)	0.0137 (7)	-0.0020 (5)	0.0158 (7)
O6	0.0358 (7)	0.0697 (9)	0.0440 (8)	0.0126 (6)	-0.0016 (5)	0.0016 (7)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

C1—C2	1.382 (2)	C5—H5	0.9300
C1—C6	1.388 (2)	C6—C7	1.480 (2)
C1—H1	0.9300	C7—O4	1.212 (2)
C2—O1	1.372 (2)	C7—O5	1.317 (2)
C2—C3	1.388 (2)	O1—H2	0.8200
C3—O2	1.3723 (19)	O2—H3	0.8200
C3—C4	1.387 (2)	O3—H4	0.8200
C4—O3	1.368 (2)	O5—H6	0.8200
C4—C5	1.378 (2)	O6—H6A	0.830 (18)
C5—C6	1.391 (2)	O6—H6B	0.824 (19)
C2—C1—C6	118.99 (15)	C4—C5—H5	120.2
C2—C1—H1	120.5	C6—C5—H5	120.2
C6—C1—H1	120.5	C1—C6—C5	120.83 (15)
O1—C2—C1	122.41 (14)	C1—C6—C7	120.83 (15)
O1—C2—C3	117.10 (14)	C5—C6—C7	118.32 (15)
C1—C2—C3	120.48 (15)	O4—C7—O5	122.82 (15)
O2—C3—C4	121.78 (14)	O4—C7—C6	123.32 (16)
O2—C3—C2	118.17 (14)	O5—C7—C6	113.86 (15)
C4—C3—C2	120.05 (14)	C2—O1—H2	109.5
O3—C4—C5	124.44 (16)	C3—O2—H3	109.5
O3—C4—C3	115.54 (14)	C4—O3—H4	109.5
C5—C4—C3	120.01 (15)	C7—O5—H6	109.5
C4—C5—C6	119.61 (16)	H6A—O6—H6B	105 (3)
C6—C1—C2—O1	178.97 (16)	O3—C4—C5—C6	178.56 (18)
C6—C1—C2—C3	0.0 (3)	C3—C4—C5—C6	-0.9 (3)
O1—C2—C3—O2	-0.5 (2)	C2—C1—C6—C5	1.0 (3)

C1—C2—C3—O2	178.51 (16)	C2—C1—C6—C7	179.51 (16)
O1—C2—C3—C4	179.56 (16)	C4—C5—C6—C1	-0.5 (3)
C1—C2—C3—C4	-1.4 (3)	C4—C5—C6—C7	-179.08 (17)
O2—C3—C4—O3	2.4 (3)	C1—C6—C7—O4	173.92 (19)
C2—C3—C4—O3	-177.66 (17)	C5—C6—C7—O4	-7.5 (3)
O2—C3—C4—C5	-178.04 (17)	C1—C6—C7—O5	-5.8 (3)
C2—C3—C4—C5	1.9 (3)	C5—C6—C7—O5	172.79 (18)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O2—H3...O3	0.82	2.26	2.7006 (18)	114
O5—H6...O6	0.82	1.85	2.6542 (17)	167
O1—H2...O6 <sup>i</sup>	0.82	2.03	2.7539 (19)	147
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Symmetry codes: (i)  $-x+1, -y+2, -z+1$ ; (ii)  $-x+2, y-1/2, -z+3/2$ ; (iii)  $-x+1, y-1/2, -z+3/2$ ; (iv)  $x-1, y+1, z$ ; (v)  $x-1, y, z$ .

**Table 2**

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	$c = 28.688$ (5)	$c = 15.028$ (10)
	$\beta = 95.08$ (3)	$\beta = 97.52$ (7)
	$V = 781.4$ (3)	$V = 764$ (1)
Space group	Monoclinic, P2 <sub>1</sub> /c	Monoclinic, P2/n
Reference	Jiang <i>et al.</i> (2000)	Okabe <i>et al.</i> (2001)



Fig. 1

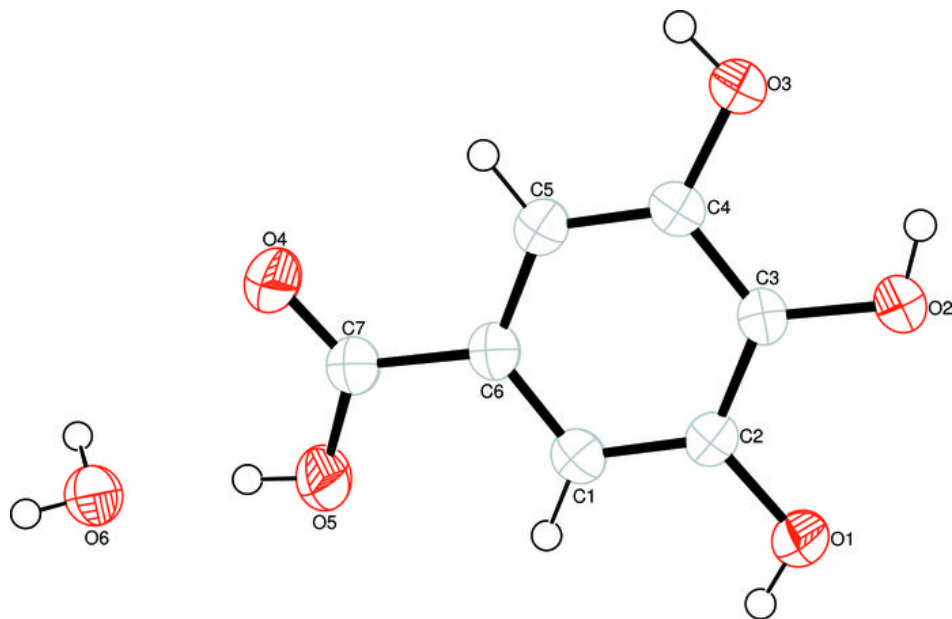


Fig. 2

