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## Magnesium Fluctuations Modulate RNA Dynamics in the SAM-I Riboswitch

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### Abstract

Experiments demonstrate that  $Mg^{2+}$  is crucial for structure and function of RNA systems, yet the detailed molecular mechanism of  $Mg^{2+}$  action on RNA is not well understood. We investigate the interplay between RNA and  $Mg^{2+}$  at atomic resolution through ten 2 microsecond explicit solvent molecular dynamics simulations of the SAM-I riboswitch with varying ion concentrations. The structure, including three stemloops, is very stable on this timescale. Simulations reveal that outer sphere coordinated  $Mg^{2+}$  ions fluctuate on the same time scale as the RNA, and that their dynamics couple. Locally,  $Mg^{2+}$  association affects RNA conformation through tertiary bridging interactions; globally, increasing  $Mg^{2+}$  concentration slows RNA fluctuations. Outer sphere  $Mg^{2+}$  ions responsible for these effects account for 80% of  $Mg^{2+}$  in our simulations. These ions are transiently bound to the RNA, maintaining interactions, but shuttled from site to site. Outer sphere  $Mg^{2+}$  are separated from the RNA by a single hydration shell, occupying a thin layer 3–5 Å from the RNA. Distribution functions reveal outer sphere  $Mg^{2+}$  are positioned by electronegative atoms, hydration layers, and have a preference for the major groove. Diffusion analysis suggests transient outer sphere  $Mg^{2+}$  dynamics are glassy. Since outer sphere  $Mg^{2+}$  ions account for most of the  $Mg^{2+}$  in our simulations, these ions may change the paradigm of  $Mg^{2+}$ -RNA interactions.

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#### Supporting Information Available

Supporting information includes the following: details of normal mode calculation and comparison of the normal modes to the principal components; additional time scale and autocorrelation data; 3-D representation of association sites and residence times; association site correlations with two additional normal modes; RMSF of RNA by residue; the  $K^+$  populations as a function of distance from RNA; methods for autocorrelation and correlation computations; details about the predicted chelated  $Mg^{2+}$  ion; and a movie of the equilibration procedure. This material is available free of charge via the Internet at <http://pubs.acs.org>.

Rather than a few inner sphere ions anchoring the RNA structure surrounded by a continuum of diffuse ions, we observe a layer of outer sphere coordinated  $\text{Mg}^{2+}$  that is transiently bound but strongly coupled to the RNA.

## 1 Introduction

Magnesium ( $\text{Mg}^{2+}$ ) has a profound effect on RNA structure and function.<sup>1,2</sup> This is not surprising in light of the high negative charge density along the RNA backbone. Although the importance of  $\text{Mg}^{2+}$  is experimentally undeniable, the molecular mechanism by which  $\text{Mg}^{2+}$  acts on RNA is not well understood in atomistic detail. Some  $\text{Mg}^{2+}$  effects can be explained by site-specific interactions<sup>3-5</sup> (e.g.,  $\text{Mg}^{2+}$  mediated tertiary contacts), yet there is also evidence that  $\text{Mg}^{2+}$  acts collectively on the RNA.<sup>6-8</sup>  $\text{Mg}^{2+}$ -RNA interactions have been previously categorized in terms of hydration layers in the pioneering works of Draper and co-workers.<sup>9,10</sup> Chelated ions make two or more direct site-specific contacts with the RNA, while diffuse ions further from the RNA are thought to be responsible for the collective action on RNA. Important molecular dynamics simulations of monovalent ion interactions with RNA have recently been performed.<sup>11</sup> Molecular dynamics simulations are also useful for further elucidating the mechanism of  $\text{Mg}^{2+}$ -RNA interactions in atomic detail.

Riboswitches constitute a suitable class of RNAs for studying the effect of  $\text{Mg}^{2+}$  by simulation. These RNAs are sufficiently large to possess the intricate tertiary structures required for complex interactions with  $\text{Mg}^{2+}$ , yet sufficiently small to be tractable for systematic studies (typically 50-100 nucleotides). Since many RNA systems, including ribozymes, the spliceosome, and the ribosome, contain tertiary interactions and  $\text{Mg}^{2+}$  dependent effects, the molecular mechanisms of  $\text{Mg}^{2+}$  interaction we discuss here should be widely relevant to other RNA systems.

Riboswitch RNAs control gene expression with high precision in response to the presence of specific metabolites.<sup>12,13</sup> Riboswitches occupy the 5' untranslated region of mRNA and contain two partially overlapping domains: the aptamer, which binds to the metabolite; and the expression platform, which controls gene expression. The S-adenosylmethionine riboswitch I (SAM-I) controls SAM synthetase, a key player in sulfur metabolism (Figure 1). In the absence of SAM, a long terminator helix is formed, which causes transcriptional termination. In the presence of SAM, a compact aptamer is formed at the expense of the terminator helix, allowing gene expression to proceed.<sup>14</sup> As one of the first riboswitches discovered, the SAM-I riboswitch is well studied and is an excellent model system.<sup>15</sup> In this work, we use SAM-I to examine the effect of the ionic environment, and specifically  $\text{Mg}^{2+}$ , on RNA.

RNA is strongly charged, so the ionic environment is of crucial importance in attenuating electrostatic repulsion as the tertiary structure of RNA packs it into a small volume.<sup>9,10</sup> In order to balance its strong negative charge and maintain electroneutrality, RNA attracts positive ions out of bulk solution and repel negative ions. The number of excess ions an RNA molecule attracts into its local environment is called the preferential interaction coefficient ( $\Gamma_i$ ) for a species  $i$ , and depends strongly on available ionic concentrations.<sup>16</sup> Local enrichment of  $\text{Mg}^{2+}$  and  $\text{K}^+$ , along with a local depletion of anions, are responsible for balancing most of the RNA charge.

$\text{Mg}^{2+}$  is especially important because RNA rarely forms tertiary structure in the absence of  $\text{Mg}^{2+}$ .<sup>10</sup> First, it is entropically cheaper to balance the charge with divalent  $\text{Mg}^{2+}$  because it requires only half as many condensed ions as monovalent ions to balance the charge. Furthermore, since  $\text{Mg}^{2+}$  is divalent, it can attract multiple phosphates at once, which results in attractive bridging interactions between what would otherwise be repulsive

phosphates.<sup>17,18</sup> Lastly, due to its small size,  $\text{Mg}^{2+}$  is able to get closer to RNA, which enables it to outcompete other divalent ions.<sup>19,20</sup>

The treatment of ions in theoretical models of RNA is an active area of research. Continuum approximations range from Debye-Hückel<sup>21</sup> to more sophisticated non-linear Poisson-Boltzmann (NLPB) approaches.<sup>22–24</sup> These reduced description treatments are necessary in implicit solvent simulations or when atomistic calculations are impractical. Since these models treat the ionic environment as a continuum, molecular details of the ionic environment are necessarily lost, including ion-ion correlations and discrete ion effects.<sup>6</sup> The importance of these discrete ion effects on the molecules with which they interact is an open question. In Manning counterion condensation theory, when the linear charge density of a polyelectrolyte chain is high enough, as it normally is in RNA, ions condense from aqueous solution to shield the phosphate backbone, effectively renormalizing its charge.<sup>25</sup> If the distance between these Manning condensed counterions is less than or equal to the Debye screening length (both are approximately 6.5 Å in our simulations), these counterions strongly interact with each other and behave as a strongly correlated ionic liquid or glass.<sup>26</sup>

Here we carefully examine the interaction between RNA and its ionic environment through molecular dynamics simulations. We observe a previously underappreciated class of transient, outer-sphere coordinated  $\text{Mg}^{2+}$  with glassy dynamics, which couples to and modulates RNA kinetics. Outer-sphere ions are by far the largest group observed in our simulations (80%), and the focus of this paper. This class of  $\text{Mg}^{2+}$  is poorly represented by the continuum approaches we have outlined. While the thermodynamics of this class are captured with varying degrees of accuracy, these models are incapable of capturing the rich internal dynamics or the control over RNA kinetics by outer-sphere  $\text{Mg}^{2+}$ .

Outer-sphere  $\text{Mg}^{2+}$  ions comprise one of four regimes of  $\text{Mg}^{2+}$  behavior we observe in our simulations. These regimes are largely delineated by the distance from the RNA, and are divided into the following classes: inner-sphere, outer-sphere, diffuse, and free. *Inner-sphere* ions, which include chelated ions as a subset, shed part of their hydration shell and bind directly to the RNA via inner-sphere contacts. *Outer-sphere* ions are coordinated with the RNA via outer-sphere contacts, and interact through a single hydration shell. *Diffuse* ions have multiple hydration shells and are poorly ordered, but are enriched above bulk concentration. Beyond the diffuse ions are the *free* ions, where the enrichment of ions ceases and the RNA charge has been screened out.

Our results show that outer-sphere  $\text{Mg}^{2+}$  local density fluctuations, residence times, and RNA fluctuations occupy similar timescales and kinetically couple. We observe that  $\text{Mg}^{2+}$  alters RNA kinetics in two ways. First, increased  $\text{Mg}^{2+}$  concentration tends to slow the kinetics of RNA fluctuations, and second, local  $\text{Mg}^{2+}$  density changes due to individual ion association events are correlated with RNA conformation. In order to better understand  $\text{Mg}^{2+}$  interactions, we characterize the behavior of  $\text{Mg}^{2+}$  in the presence of RNA. The  $\text{Mg}^{2+}$  distribution is controlled by electronegative RNA atoms, hydration shells, and a preference for the major groove. The differing distance classes of  $\text{Mg}^{2+}$  have different diffusive behavior, and careful analysis of the diffusion data suggests that the outer-sphere  $\text{Mg}^{2+}$  near RNA behaves like a supercooled strong ionic liquid. Because outer-sphere  $\text{Mg}^{2+}$  ions account for 80% of excess  $\text{Mg}^{2+}$  in our simulations, and because of their rich dynamics, they may change the paradigm of  $\text{Mg}^{2+}$ -RNA interactions. Rather than a few inner sphere ions anchoring the RNA structure surrounded by a continuum of diffuse ions, our results reveal a layer of outer-sphere coordinated  $\text{Mg}^{2+}$  strongly coupled to the RNA.

## 2 Model and Methods

### 2.1 Simulation Details

Molecular dynamics simulations of the SAM-I riboswitch were performed with and without SAM in the presence of five varying  $\text{Mg}^{2+}$  concentrations. Individual simulations were 2  $\mu\text{s}$ , for a total of 20  $\mu\text{s}$  of sampling. Simulations were carried out using the Gromacs v4.5.1 molecular dynamics package.<sup>27</sup> The Amber 99 forcefield<sup>28</sup> was used, with modified  $\text{K}^+$  ion parameters and SPC/E water to prevent crystal formation.<sup>29,30</sup> Because of its importance for stability, we parameterized the SAM metabolite for AMBER. In order to parameterize the SAM, we used GAMESS quantum mechanics software<sup>31</sup> and the RED software package<sup>32</sup> to compute the charges on each atom using the Restrained ElectroStatic Potential method.<sup>33,34</sup> The remaining forcefield parameters were taken from GAFF (the Generalized Amber ForceField).<sup>35</sup> Simulations were run with a time step of 2 fs and bond lengths for hydrogen atoms were fixed using LINCS.<sup>36</sup> Particle mesh Ewald electrostatics<sup>37</sup> were used with an Ewald radius of 15 Å, and a Van der Waals cutoff of 15 Å.

Since the maximum linear dimension of the SAM-I aptamer is 70.7 Å, a water box of 100 Å was used to allow a careful examination of the ionic environment. Simulations contained approximately 98,000 atoms. The simulations had target concentrations of 100 mM KCl, and 0, 0<sup>†</sup>, 1, 6, and 10 mM  $\text{MgCl}_2$  in bulk. The 0 mM  $\text{Mg}^{2+}$  simulation contained no  $\text{Mg}^{2+}$  ions, while the 0<sup>†</sup> simulation contained a single poorly resolved  $\text{Mg}^{2+}$  ion taken from the crystal structure (near A10 and U64) which we predict to be chelated (see below for details). The presence of a  $\text{Mg}^{2+}$  near this predicted chelation site has been suggested to be important for stability.<sup>4,38</sup> This ion was included in all remaining simulations. Since the RNA is strongly charged, it balances this charge by attracting  $\text{Mg}^{2+}$  and  $\text{K}^+$  ions into and repelling  $\text{Cl}^-$  ions from its local environment. These excess ions do not count towards the bulk concentration. Therefore, the ions in a simulation are a combination of excess ions which balance the RNA charge and bulk ions. The number of excess ions is equal to the preferential interaction coefficient. Each bulk ion contributes 1.7 mM for our box size. Since preferential interaction coefficients vary with concentration, several preliminary equilibrations were run to determine the correct number of ions to include so that the desired concentrations would remain in bulk. For example, in the 10 mM simulation, the predicted chelated ion and 32 hydrated  $\text{Mg}^{2+}$  ions were included, but 27 of these were excess ions and only 6 contributed to the bulk concentration. See Table 1 for more details.

### 2.2 Equilibration

Equilibration began from the crystal structure (Protein Data Bank accession code 2GIS<sup>14</sup>), using our well established protocol.<sup>39,40</sup> The RNA was frozen and placed in a waterless box. Ions were placed randomly with larger Van der Waals radii to prevent them from condensing onto the RNA without an appropriate hydration shell. For the sole purpose of establishing the starting configuration of ions for explicit solvent simulations, the ions were equilibrated using stochastic dynamics and a dielectric constant of 80 to mimic water for 10 ns until the electrostatic energy converged. Then water was added to the box, and annealed to 300 K over 200 ps with the RNA and ions frozen. The ions were released and allowed to equilibrate at constant volume for 2 ns. Then the RNA was gradually released over 10 ns, spending 2 ns each with position restraints of 1000, 100, 10, 1, and 0 kcal/mol/nm<sup>2</sup>, at constant pressure. This amounted to 10 ns of waterless equilibration followed by 12.2 ns of explicit solvent equilibration. From this point on, unrestrained explicit solvent simulations were performed including explicit solvent and explicit ions with full particle mesh Ewald electrostatics.

Upon the addition of water, the bulk concentration and the radial distribution functions of the  $K^+$  and  $Cl^-$  ions converged quickly, on a time scale of approximately 0.5 ns. The  $K^+$  distribution function changed due to a few  $K^+$  ions partially shedding their hydration shells, which was not allowed before the addition of water. Since the equilibration was much longer than 0.5 ns,  $K^+$  and  $Cl^-$  were well equilibrated. Post simulation analysis of relevant time scales, (see Section 3.1.2), validated that equilibration was reasonable for  $Mg^{2+}$  as well. While some  $Mg^{2+}$  sites had association times in excess of the simulation time, the largest density fluctuation times for  $Mg^{2+}$  of these sites were approximately 20 ns, and bulk  $Mg^{2+}$  concentration fluctuated with a time scale between 5 and 8 ns depending on the  $Mg^{2+}$  abundance. These time scales are all less than or comparable to the 12.2 ns of explicit solvent equilibration. Thus, this equilibration procedure allowed the RNA to equilibrate with the ionic environment while keeping it close to the crystal structure.

### 3 Results and Discussion

In Section 3.1, the simulations of the SAM-I riboswitch show that outer-sphere  $Mg^{2+}$  and RNA fluctuate on the same time scales and couple. Both collective effects (the bulk concentration of  $Mg^{2+}$ ) and local effects (individual  $Mg^{2+}$  association events) affect RNA dynamics. In Section 3.2, characterization of the  $Mg^{2+}$  diffusion and distribution reveals molecular mechanisms by which  $Mg^{2+}$  influences RNA: a layer of  $Mg^{2+}$  ions separated from the RNA by a single hydration layer constitutes the majority of the excess ions ( $\Gamma_{Mg^{2+}}$ ). Ions in this layer are positioned by electronegative RNA atoms, hydrations shells, and a preference for the major groove.

#### 3.1 $Mg^{2+}$ -RNA Interactions and Time Scales

**3.1.1 Global RNA Dynamics**—In our simulations, the  $Mg^{2+}$  concentration has a measurable effect on global fluctuations (Figure 2a). To characterize the global dynamics of the RNA, global coordinates are needed. The explicit solvent molecular dynamics simulation trajectories were projected onto normal modes.<sup>41–43</sup> In this context, normal mode analysis is advantageous relative to principal component analysis<sup>44</sup> because normal mode analysis produces a consistent measure between all simulations (see SI for more details). A structure-based potential<sup>45,46</sup> was used to compute normal modes, rather than the typical Tirion potential.<sup>47</sup> This is the first time this potential has been used for normal modes to our knowledge.

In order to determine time scales of RNA motion, trajectories were projected onto the first 50 normal modes. After subtracting a linear fit to remove drift along unstable normal modes, autocorrelations were computed as a function of time (Eq. (S2)). The autocorrelations generally had long non-exponential tails, suggesting the presence of multiple time scales. This is to be expected because processes in RNA range over many orders of magnitude, up to the ligand dissociation time scale, which is too slow to observe in the simulations. To obtain the dominant time scale for each normal mode, the time when the autocorrelation fell below  $1/e$  was taken. This dominant time scale represents the duration of the largest structural changes that occur within the 2  $\mu$ s simulation. The peak of the distribution of times changes with  $Mg^{2+}$  concentration. The median time scale of the autocorrelations increased from 2.03 ns for 0<sup>+</sup> mM  $Mg^{2+}$  to 3.55 ns for 10 mM  $Mg^{2+}$ , showing that increasing  $Mg^{2+}$  concentration slows structural fluctuations.

It has been shown that while  $Mg^{2+}$  is much more effective than  $K^+$  at stabilizing RNA,  $Mg^{2+}$  also tends to slow transition rates by reducing the entropy of transition state ensembles.<sup>17,48</sup> The observed lengthening of autocorrelation time scales for RNA fluctuations in the presence of  $Mg^{2+}$  provides a mechanism for such a slowing in the

presence of  $\text{Mg}^{2+}$ . The  $\text{Mg}^{2+}$  stabilizes the compact structure, so the RNA must wait longer before a sufficiently large fluctuation occurs to induce a transition.

**3.1.2 Local  $\text{Mg}^{2+}$  Dynamics**—Before it is possible to quantify the effects of individual  $\text{Mg}^{2+}$  association events on RNA, it is necessary to identify sites where  $\text{Mg}^{2+}$  associates with RNA. Association sites were identified by finding the average position of a  $\text{Mg}^{2+}$  ion with respect to the RNA between times when its diffusion fell below  $10 \mu\text{m}^2/\text{s}$  and rose back above  $100 \mu\text{m}^2/\text{s}$ . (The experimental diffusion of aqueous  $\text{MgCl}_2$  is  $1100 \mu\text{m}^2/\text{s}$ .) Diffusion as a function of time was computed by fitting  $\langle x^2 \rangle = 6Dt$  to the mean squared deviation of  $\text{Mg}^{2+}$  positions (fitted to the RNA structure). We identified 223 sites where  $\text{Mg}^{2+}$  associates with RNA.  $\text{Mg}^{2+}$  dwell times in these sites ranged from a few ns to several hundred ns.

To quantify  $\text{Mg}^{2+}$  association with RNA,  $\text{Mg}^{2+}$  density near these association sites was computed as a function of time by counting the number of  $\text{Mg}^{2+}$  within a cutoff distance of the site. The cutoff used was typically  $10 \text{ \AA}$ , which is large enough to accommodate two or three  $\text{Mg}^{2+}$ , but values of  $8 \text{ \AA}$ ,  $6 \text{ \AA}$ , and  $4 \text{ \AA}$  were also tried with similar results. This density measure showed whether an association site was effectively occupied or unoccupied at any given time.

In order to determine the time scales of local  $\text{Mg}^{2+}$  density fluctuations, autocorrelations of the local  $\text{Mg}^{2+}$  density time traces were computed (Eq. (S2)). Most (approximately 85%) of the local density autocorrelations were generally exponential. Characteristic times were taken from when the autocorrelation crossed  $1/e$ , and binned (Figure 2b). Nearly all sites have fluctuation times between 1 and 10 ns. A site may tend to be occupied for much longer, but in that case the fluctuation times correspond to transient periods when the site is empty, and these unoccupied times are generally less than 10 ns. Shell sizes other than  $10 \text{ \AA}$  give comparable timescales (Figure S2), even though a  $4 \text{ \AA}$  shell just captures if a  $\text{Mg}^{2+}$  is in that site, and  $10 \text{ \AA}$  captures a broader picture of the local density.

Since residence times need not correspond with local  $\text{Mg}^{2+}$  density fluctuation times, it is informative to examine residence times as well. To determine residence times for each site, a modified autocorrelation approach was used. For each site and each  $\text{Mg}^{2+}$  a time-dependent boolean variable was assigned to one if the  $\text{Mg}^{2+}$  ion was within  $4 \text{ \AA}$  of the association site, and zero otherwise.  $4 \text{ \AA}$  was chosen to ensure only ions in that association site were considered associated, rather than ions in neighboring sites or more transient ions which might come within  $10 \text{ \AA}$ . This boolean variable was then autocorrelated without subtracting the mean. These autocorrelations represent how long the same ion remains associated rather than how long a density fluctuation from the mean lasts. For each association site, the averages in (Eq. (S2)) were taken over time and over all  $\text{Mg}^{2+}$  ions. Due to the noisy tails of the autocorrelation functions, residence times were taken from when autocorrelations fell below  $1/e$ .

Residence times occupy a much broader range, including 8 that are longer than the simulation time (Figure 2c). Residence times correlate well with the  $4 \text{ \AA}$  shell density fluctuation times (Figure S2). Sites with longer residence times tend to have longer fluctuations, though the fluctuations are generally fast compared with the residence time.  $10 \text{ \AA}$  density fluctuation times correlate poorly with the residence times due to the presence of faster ions further out in the shell. It is likely that long term outer-sphere association sites such as the 8 identified here contribute substantially to the enhanced stability of RNA in the presence of  $\text{Mg}^{2+}$ , and modulate kinetics by anchoring their local RNA environments against fluctuations for long times while associated.

**3.1.3 Local Mg<sup>2+</sup> Events Affect Global Dynamics**—Since RNA fluctuations occupy the same time scales as Mg<sup>2+</sup> density fluctuations and Mg<sup>2+</sup> residence times, they can couple. We now show that Mg<sup>2+</sup> association in particular sites affects SAM-I conformation and dynamics. Mg<sup>2+</sup> bridging interactions, in which a single Mg<sup>2+</sup> mutually attracts phosphates, have been suggested as a mechanism for Mg<sup>2+</sup> stabilization of tertiary structure.<sup>17</sup> Temporal correlations (Eq. (S4)) between individual Mg<sup>2+</sup> association events and phosphate-phosphate (P-P) distances in the simulations provide evidence for this mechanism (Figure 3). Mg<sup>2+</sup> association events were modeled by the local ion density about association sites using 10 Å shells as described above.

Correlating time traces of P-P distances with time-dependent local Mg<sup>2+</sup> densities shows several regions are especially susceptible to Mg<sup>2+</sup> binding. Mg<sup>2+</sup> binding within the P4 major groove (C8 to G82 phosphates) shows substantial correlation. When Mg<sup>2+</sup> enters the P4 major groove, the groove partially closes (Figure 3c), resulting in negative (red) correlations between Mg<sup>2+</sup> density within the groove and P-P distance. Several other regions also show strong correlations (Figure 3a), including the P3 groove (C41 to G55 phosphates) (Figure 3d), demonstrating that individual Mg<sup>2+</sup> association events affect local RNA conformation.

The K<sup>+</sup> correlations in Figure 3b provide a control for the conclusions drawn from the Mg<sup>2+</sup> correlations in Figure 3a. K<sup>+</sup> fluctuates more than an order of magnitude more quickly than Mg<sup>2+</sup>, so K<sup>+</sup> cannot couple effectively with RNA motion. The observed correlations between local K<sup>+</sup> density (at the Mg<sup>2+</sup> sites) and P-P distances are indeed weaker than for Mg<sup>2+</sup>. They represent K<sup>+</sup> passively responding to the RNA conformation rather than actively modulating it through bridging interactions as Mg<sup>2+</sup> does.

The effect of Mg<sup>2+</sup> association events on local RNA conformation in turn affects the global conformation. Correlations were computed between the local density around association sites and the first 25 normal modes. The presence or absence of Mg<sup>2+</sup> can account for 30% to 40% of RNA fluctuations along many normal modes (Figure 4a). Several of the lowest and most correlated normal modes are shown with the directions of motion along the normal mode. Normal mode 1 responds to bridging interactions in the P4 major groove (Figure 4b). When Mg<sup>2+</sup> enters the blue region (positive correlation), progress along that coordinate increases, and the major groove of P4 surrounding those phosphates closes. (Correlations are opposite from Figure 3c because positive motion along the structural reaction coordinate is in opposite directions.) Normal mode 4 responds to competition between bridging interactions in the P3 and P4 major grooves (Figure S5c). Positive motion along normal mode 4 opens the P4 major groove and closes the P3 major groove, and the resulting correlations are as expected. Normal mode 8 responds to competing bridging interactions on opposite ends of the P4 major groove (Figure S5d). In normal mode 8, the P4 major groove opens on the loop end while closing on the P1 end. Correlations with the association sites in this groove reflect that difference: the loop end displays negative correlations; the P1 end shows positive correlations. Our results show that individual Mg<sup>2+</sup> association events affect the global conformation by changing local conformation through bridging interactions between phosphates.

## 3.2 Characterization of Outer-Sphere Mg<sup>2+</sup>

**3.2.1 Ionic Concentrations and Preferential Interaction Coefficients**—Since RNA is strongly charged, it attracts excess ions to balance its charge. The excess number of each species of ion *i* that an RNA molecule attracts into its local environment is called the preferential interaction coefficient  $\Gamma_i$ .<sup>16</sup> Preferential interaction coefficients are of interest because they are experimentally measurable,<sup>49</sup> and allow connection between the simulations and experiment. Electroneutrality requires the preferential interaction

coefficients to balance the net RNA charge according to  $Z = \sum q_i \Gamma_i$ , where  $Z$  is the magnitude of the RNA charge and  $q_i$  is the charge of each ionic species  $i$ .<sup>50,51</sup> Within this constraint,  $\Gamma_i$  for each species can vary with bulk ionic concentrations due to the availability of various ions.

Simulations were run at several varying concentrations. Table 1 shows the number of ions included in simulations as well as the concentrations and preferential interaction coefficients ( $\Gamma_i$ ) for each species  $i$ . The number of ions in the box is the sum of ions contributing to the concentration and the excess ions  $\Gamma$ . Since  $\Gamma_i$  varies with concentration, the number of ions in the box had to be chosen through repeated trials to obtain the desired bulk concentrations. Concentrations as a function of time are computed by counting the number of “bulk” ions  $B_{Mg^{2+}}$  and water molecules  $B_{H_2O}$  more than 20 Å from any RNA atom and taking their ratio. While this technically gives the molality rather than the molarity, we note that the molarity and molality are very close for low ionic concentrations in water. The concentration of an ionic species is given by the time average of the ratio of these molecule counts multiplied by the molarity of pure water.

$$[Mg^{2+}]^* = 55.51 \text{ M} \frac{B_{Mg^{2+}}}{B_{H_2O}} \quad (1)$$

Since only 95% of the charge is balanced out by 20 Å in 100 mM KCl, the computed concentrations will not sum to zero because some charge density must remain in the rest of the box to balance the residual RNA charge. However, unlike the case near the RNA where ions interact differently with the RNA due to large electrostatic potentials, at 20 Å, the electrostatic potential is a smooth, small perturbation. Ion densities in a small potential well will respond linearly with their concentration, charge, and the well depth. The corrected bulk concentrations (which must be electroneutral) are given by

$$[i] = [i]^* - q_i [i]^* \frac{\sum_j q_j [j]^*}{\sum_j q_j^2 [j]^*} \quad (2)$$

$$[Mg^{2+}] = [Mg^{2+}]^* - 2[Mg^{2+}]^* \frac{2[Mg^{2+}]^* + [K^+]^* - [Cl^-]^*}{4[Mg^{2+}]^* + [K^+]^* + [Cl^-]^*} \quad (3)$$

where  $q_i$  is the charge,  $[i]^*$  denotes the raw concentration, and  $[i]$  denotes the corrected concentration. If the arbitrary definition of bulk ions as more than 20 Å from the RNA is varied, the corrected concentrations converge much more quickly than the raw concentrations. Once corrected concentrations are obtained, preferential interaction coefficients may be computed. Preferential interaction coefficients ( $\Gamma$ ) are the ions in excess of the ions expected from the corrected concentration

$$\Gamma_{Mg^{2+}} = N_{Mg^{2+}} - N_{H_2O} \frac{[Mg^{2+}]}{55.51 \text{ M}} \quad (4)$$

where  $N_{Mg^{2+}}$  is the total number of  $Mg^{2+}$  ions and  $N_{H_2O}$  is the total number of water molecules.

Preferential interaction coefficients ( $\Gamma$ ) are experimentally measurable.<sup>49</sup> For a purine riboswitch with 72 residues at 1 mM  $Mg^{2+}$  and 50 mM  $K^+$ ,  $\Gamma_{Mg^{2+}}$  varies from 13 to 18, depending on mutations and the switch state.<sup>50,52</sup> We observe a  $\Gamma_{Mg^{2+}}$  of 10.5 at this  $Mg^{2+}$  concentration. A lower prediction is expected in the simulation since it has twice as much



$K^+$  competing to balance the RNA charge. There are additional factors that complicate the comparison between experiment and simulation. NMR experiments generally show nearly all  $K^+$  remains fully hydrated, while in most molecular dynamics simulations (including ours) approximately half of the excess  $K^+$  partially dehydrates and contacts RNA directly.<sup>53</sup> Our simulations show 40% to 50% of  $\Gamma_{K^+}$  is partially dehydrated (Figure S11).

We emphasize that our ion parameters produce very stable 2 microsecond trajectories. Removal of this population and proportional increase of the remaining populations of excess ions would raise  $\Gamma_{Mg^{2+}}$  for 1 mM  $Mg^{2+}$  from 10.5 to 14.7. We note that most of the  $Mg^{2+}$  contributing to  $\Gamma_{Mg^{2+}}$  are in the outer-sphere layer of ions, between the inner-sphere and diffuse regimes.

**3.2.2 Distance Classes and Diffusion of  $Mg^{2+}$  Ions**— $Mg^{2+}$  ions can be divided into classes based on distance from RNA<sup>9,10</sup> (Figure 5a). *Inner-sphere* ions are the closest and make direct contact with the RNA.  $Mg^{2+}$  typically has an octahedral geometry and coordinates with six waters. Each inner-sphere contact with RNA replaces one of these waters. Inner-sphere ions can be further divided into monodentate and chelated ions based on whether they make one or multiple inner-sphere contacts with RNA. The next layer is *outer-sphere*. These  $Mg^{2+}$  retain all their inner-sphere waters, and are only separated from the RNA by this single hydration layer. This layer accounts for most of the  $Mg^{2+}$  in our simulations, 80% to 85% of  $\Gamma_{Mg^{2+}}$  (Table 2). Ions in this layer fluctuate on time scales comparable with RNA fluctuations, and couple to the RNA. These  $Mg^{2+}$  ions cannot condense closer than this layer without paying substantial dehydration costs. There are several hints that this layer behaves as an ionic glass (Section 3.2.4). The next layer is *diffuse*. Diffuse ions are separated from RNA by two or more hydration shells, are poorly ordered, and are well described by continuum models. The furthest group of ions is *free*. The divide between diffuse and free ions is ambiguous. For the present analysis we simply subtract inner-sphere and outer-sphere ions from  $\Gamma_{Mg^{2+}}$ . Population enrichment (Figure 5a) and reduced diffusion (Figure 5b) suggest between 10 and 20 Å as a reasonable boundary for these two classes.

Diffuse ions are well represented by NLPB distributions due to their low density and rapid dynamics. Outer-sphere ions may be analogous to Manning counterions that condense as close to the RNA as possible and screen its charge. They are dense, have internal correlations, and couple strongly with the RNA. Reasonable thermodynamic agreement with experiment has been obtained by combining a few chelated sites with a NLPB representation of diffuse ions.<sup>7,8,54</sup> The Tightly Bound Ion method treats outer-sphere ions thermodynamically resulting in improved calculations of stability.<sup>6</sup> Our results show that  $Mg^{2+}$  also exerts substantial kinetic control. In contexts where RNA kinetics are of interest, the kinetic effects of outer-sphere  $Mg^{2+}$  must be carefully considered.

The kinetic effects are highlighted by the different diffusive behavior between  $Mg^{2+}$  distance classes (Figure 5b). The diffusion of  $Mg^{2+}$  ions was computed as follows. The RNA was first fit to a reference structure, because otherwise the lower  $Mg^{2+}$  diffusion values were washed out by the bulk RNA diffusion, which was on the order of  $10 \mu m^2/s$ . The distance to the nearest RNA heavy atom was found for every frame. The squared displacement for 2 ns periods starting every 200 ps was computed, and then averaged within bins based on starting distance from the RNA. The mean squared displacement was then fit within each distance bin to  $\langle x^2 \rangle = 6Dt$  to obtain diffusion. Instantaneous diffusion could be determined by fitting the mean squared displacement before averaging within distance bins.

The  $Mg^{2+}$  ion in the inner sphere class has the lowest diffusion. This ion does not move relative to the RNA during the simulation so its diffusion is nearly zero, and the plotted

value of  $1 \mu\text{m}^2/\text{s}$  indicates our method's precision. The outer-sphere  $\text{Mg}^{2+}$  ions have a diffusion coefficient ( $20$  to  $30 \mu\text{m}^2/\text{s}$ ) roughly two orders of magnitude below bulk. As mentioned before, outer-sphere ions account for most of the  $\text{Mg}^{2+}$  associated with the RNA. Depending on the site, they can remain associated from a few ns to several thousand ns.

Due to heterogeneous sites, glassy dynamics, and escape to regions with higher diffusion, the diffusion of the inner-sphere regime is complex. Rare escape events dominate the average diffusion, while the median of the instantaneous diffusion is nearly an order of magnitude lower than the average.

A very steep increase in diffusion occurs between the outer-sphere and diffuse regimes. The diffuse  $\text{Mg}^{2+}$  helps with charge balancing, but its diffusion is higher because it is not associated very strongly with any particular site. Values for the free  $\text{Mg}^{2+}$  come close to the experimental value of  $1100 \mu\text{m}^2/\text{s}$  for  $\text{MgCl}_2$  diffusion. The disconnect between diffusion rates emphasizes the uniqueness of the outer-sphere regime, which occupies a region of low mobility between the essentially immobile inner sphere  $\text{Mg}^{2+}$  ion and the more rapidly diffusing diffuse regime.

In the X-ray structure, there is density nearby A10 and U64 assigned to a  $\text{Mg}^{2+}$  ion. Due to the resolution of this structure, it is difficult to assign certainty to this density, much less ascertain its chelation state. Our starting initial configuration includes this ion, which we call inner sphere due to its close proximity to the RNA. We note that the use of the term inner sphere itself is approximate due to the resolution of the X-ray structure. In our simulations, the inner sphere  $\text{Mg}^{2+}$  ion connects the partially buried non-bridging phosphate oxygens of residues A10 and U64 for the duration of the simulation, suggesting that the ion may be chelated. The chelation site is consistent with the crystal structure, and has been suggested to be important for stability.<sup>4,38</sup> Our prediction that this ion is chelated is also based on the need for a positive ion to bridge the repulsion of two phosphate oxygens separated by  $4.24 \text{ \AA}$ , the need to offset the cost of partial burial of the phosphate on U64, and the destabilization and disordering of the riboswitch, especially the P4 helix, in simulations lacking an ion in this position. Whether we are correct about magnesium chelation at this site remains to be seen experimentally.

In bulk solution,  $\text{Mg}^{2+}$  ions exchange their inner-shell contacts on a time scale of  $1.5 \mu\text{s}$ ,<sup>55,56</sup> so we did not expect this  $\text{Mg}^{2+}$  to unbind from the RNA. During the course of our simulations, none of the  $\text{Mg}^{2+}$  exchanged any of their waters, whereas we may have expected some exchange events to occur. Simulations of even longer durations will be able to more directly address this. We note that this is not necessarily a major limitation for our system, since dehydration of  $\text{Mg}^{2+}$  is energetically costly,<sup>10</sup> and RNAs of this size have few electrostatic wells deep enough to pay the substantial dehydration cost for  $\text{Mg}^{2+}$ . Such sites, which frequently occur at buried backbone phosphates, are more common in larger RNAs. For RNAs of this size at physiological concentrations, inner-sphere  $\text{Mg}^{2+}$  is likely only a small contributor to the preferential interaction coefficient  $\Gamma_{\text{Mg}^{2+}}$  or energetic stabilization ( $\Delta G_{\text{Mg}^{2+}}$ ) compared to outer-sphere and diffuse  $\text{Mg}^{2+}$ .<sup>3</sup>

**3.2.3 Spatial Characterization of Outer-Sphere  $\text{Mg}^{2+}$** —The distribution functions of  $\text{Mg}^{2+}$  about the RNA reveal three molecular principles of  $\text{Mg}^{2+}$ -RNA association. First,  $\text{Mg}^{2+}$  shows a strong preference for the backbone and major groove. Second,  $\text{Mg}^{2+}$  is positioned by hydration layers. Finally,  $\text{Mg}^{2+}$  is positioned by an attraction for the most electronegative RNA atoms: non-bridging phosphate oxygens (NBPOs), and atoms Purine N7, Uracil O4, and Guanine O6 of the nucleotide bases.

Radial distribution functions (RDF) reveal where  $\text{Mg}^{2+}$  resides during the simulations (Figure 6). The top panel shows the RDF for the central phosphorus of the phosphate group. The inner peak at 3.2 Å is the distance for the inner sphere  $\text{Mg}^{2+}$  ion which did not exchange. The double peak near 5 Å corresponds to hydrated, outer-sphere ions. The double peak suggests there is anisotropic substructure unexplained by this RDF. Most of the ions in the second and third hydration peaks are more closely associated with other phosphates. The substructure of the outer-sphere layer can be attributed to interactions with electronegative atoms of the RNA, in particular, base atoms Purine N7, Uridine O4, Guanine O6, and especially the non-bridging phosphate oxygens (NBPO) (Figure 6b). All of these RDFs have a single peak at 4.2 Å from the first hydration layer, revealing it is the electronegative atoms, not the phosphate centers, which control the  $\text{Mg}^{2+}$  distribution.

These radial distributions are consistent with previous results. A previously computed phosphate RDF averaged over four phosphates<sup>57</sup> shows qualitative agreement with the phosphate RDF and double peak. The RDF about specific electronegative atoms that explains this double peak has not been previously computed, but shows reasonable similarity to the RDF of water about  $\text{Mg}^{2+}$ .<sup>58,59</sup> The main difference is that the inner-sphere peak in Figure 6b is much lower than the inner-sphere peak for water due to the high free energy cost of dehydrating  $\text{Mg}^{2+}$  ions.

The spatial distribution functions (SDF) provide a more detailed picture of the  $\text{Mg}^{2+}$  distribution and corroborate the three principles controlling it (Figure 7). As noted in previous crystallographic studies,<sup>60</sup> the SDFs show that  $\text{Mg}^{2+}$  has a strong preference for the major groove.

Figure 7a shows the average SDF around phosphate groups. The SDF about the single phosphate shows two lobes (i) around the NBPOs as would be expected from Figure 6b. (The dual peak in Figure 6a arises because different parts of these lobes are different distances from the central phosphorus.) These lobes are modulated in density by preference for the major groove and by other electronegative atoms. As a result, the belt (ii) joining the lobes is particularly strong, as is the lobe around the major groove NBPO. The lobe around the minor groove NBPO is weaker, and polarized towards the 3' terminus, due to the major groove NBPO of the neighboring residue. Secondary shells separated by two and three hydration layers are visible where  $\text{Mg}^{2+}$  can interact more closely with other electronegative atoms, especially the bases in the major groove.

Figure 7b (the P1 helix) shows typical ways that  $\text{Mg}^{2+}$  associates with helices. Again the preference for the major groove is strongly visible. There are sites along the backbone where  $\text{Mg}^{2+}$  sits in the belt (ii) between two NBPOs or in a notch (i) between neighboring phosphates, as suggested by Figure 7a. Several association sites within the major groove where  $\text{Mg}^{2+}$  interacts with the bases are visible. Electronegative atoms are again key in positioning these sites. One particularly strong site (iii) is at the G-U wobble pictured at the top of this helix. G-U wobbles have been previously noted to form  $\text{Mg}^{2+}$  association sites.<sup>5,60</sup> The molecular basis for this is immediately apparent: the G-U wobble forms a depression in the major groove surrounded by an N7 and an O6 on the guanine and an O4 on the uracil. The canonical base pairs only have two of these three electronegative atoms. A preliminary analysis of our data does show some clustering of  $\text{K}^+$  ions at similar N7 positions as observed in previous studies of monovalent ion interactions with RNA by Chen, *et al.*<sup>11</sup>

**3.2.4 Glassy Character of Outer-Sphere  $\text{Mg}^{2+}$** —Outer-sphere  $\text{Mg}^{2+}$  ions are not quite bound and not quite un-bound. These ions are localized to the RNA, but hop from site to site in a glass-like manner. That is, the energy landscape for these ions exhibits a high degree of

frustration, where one site of localization for a  $Mg^{2+}$  ion may be as favorable as another site, causing the ion to migrate from site to site, without being trapped in one specific site. Our measured diffusion and distribution of the outer-sphere  $Mg^{2+}$  suggest they behave in a glassy manner.

The presence of several ordered peaks in the RDF between  $Mg^{2+}$  and NBPOs in the major and the minor grooves of the RNA (Figure 6b) is consistent with glass-like behavior. Furthermore, the RDF between  $Mg^{2+}$  and the phosphorous atoms (Figure 6a) indicates a split peak at 4.2 Å and 5.1 Å, that is followed by peaks at 7.5 Å and 9.5 Å, and a lack of noticeable peaks for distances larger than 14 Å. This indicates that the system is locally amorphous, and likely an ionic glass.<sup>61</sup>

If the distance between these outer-sphere ions is less than or equal to the Debye screening length, these counterions strongly interact with each other and behave as a strongly correlated ionic liquid,<sup>26</sup> with corresponding diffusion and shear properties.<sup>62</sup> In our simulations the average distance between outer-sphere ions is on the same order as the Debye length (both are approximately 6.5 Å), suggesting that the ions are strongly correlated. The diffusion coefficient of the outer-sphere ions is around one hundred fold smaller than the diffusion coefficient of the free ions (Figure 5b), also indicating that these ions have the properties of a strongly-coupled liquid consistent with glass-like behavior.

## 4 Conclusion

$Mg^{2+}$  has a strong effect on RNA stability and dynamics. Through molecular dynamics simulation of a folded RNA riboswitch, we demonstrated that the dynamics of  $Mg^{2+}$  and RNA can couple because  $Mg^{2+}$  and RNA fluctuations occur on the same time scale. This coupling was substantiated by the correlation of the RNA's motion with  $Mg^{2+}$  association events. Increasing  $Mg^{2+}$  concentration was shown to slow fluctuations of the RNA. In fact, preliminary results suggest outer-sphere  $Mg^{2+}$  behaves in a glassy manner. Because  $Mg^{2+}$  couples to the RNA, phase changes in the glassy  $Mg^{2+}$  ions may result in changes in RNA kinetics. Further work is needed to fully characterize the glassy dynamics of outer-sphere  $Mg^{2+}$  ions.

To form a deeper physical understanding of the molecular basis for  $Mg^{2+}$ -RNA interactions, we have focused on the spatial distribution and diffusion of  $Mg^{2+}$  itself. Over 80% of excess  $Mg^{2+}$  ions are in the outer-sphere regime, separated from the RNA by a single hydration layer, rather than in the diffuse regime. These outer-sphere ions are positioned by electronegative RNA atoms, hydration layers, and a preference for the major groove. The 100-fold slower kinetics within the outer-sphere regime demonstrates how different these ions are from the canonical diffuse  $Mg^{2+}$ . The outer sphere  $Mg^{2+}$  occupy a region of intermediate mobility between inner sphere and diffuse, which allows them to couple to the RNA.

Previously, the interplay between  $Mg^{2+}$  and RNA has frequently been modeled by chelation sites and a diffuse ion cloud. Diffuse ions are well represented by NLPB distributions due to their low density and rapid dynamics. In contrast, outer-sphere ions may be more reminiscent of Manning counterions that condense as close to the RNA as possible. Our simulations include explicitly discrete ion effects, internal correlations, and kinetic coupling between outer-sphere  $Mg^{2+}$  and the RNA. Since outer-sphere  $Mg^{2+}$  are the dominant population and have a substantial effect on RNA dynamics, we propose a new view of  $Mg^{2+}$ -RNA interactions where the outer-sphere  $Mg^{2+}$  class, with glass-like internal correlations, both thermodynamically stabilizes and exerts kinetic control on RNA. Future

reduced treatments of  $Mg^{2+}$ -RNA interactions need to consider the importance of the inherently discrete ion effects of the outer-sphere  $Mg^{2+}$ .

## Supplementary Material

Refer to Web version on PubMed Central for supplementary material.

## Acknowledgments

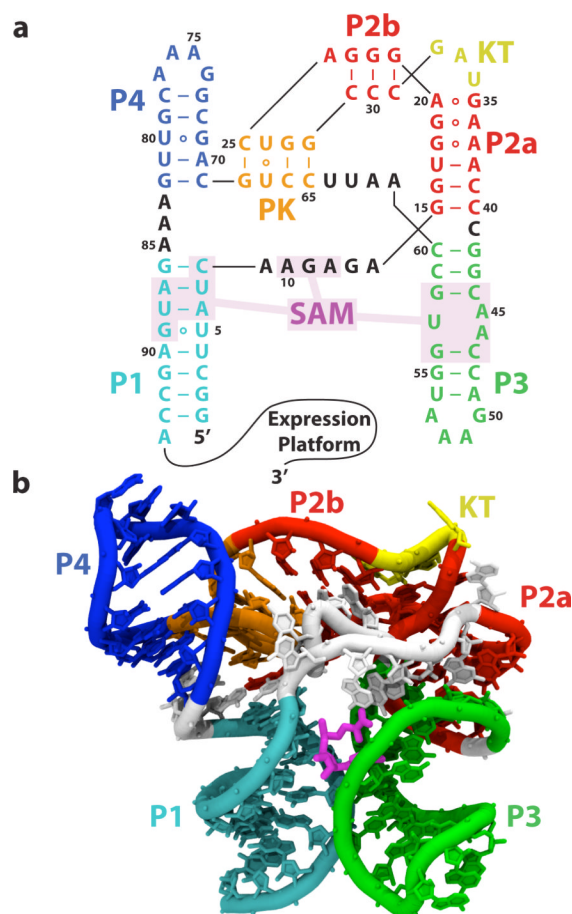
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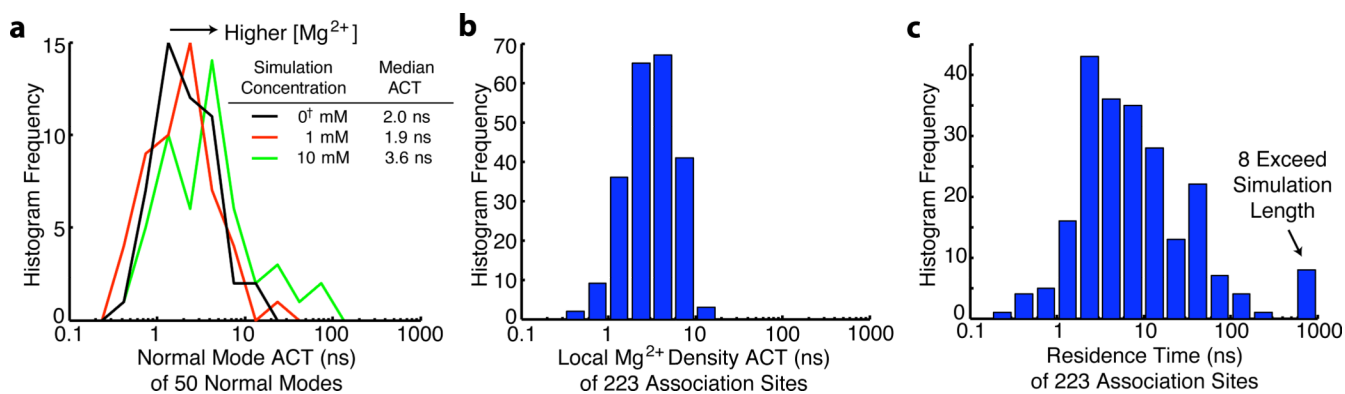
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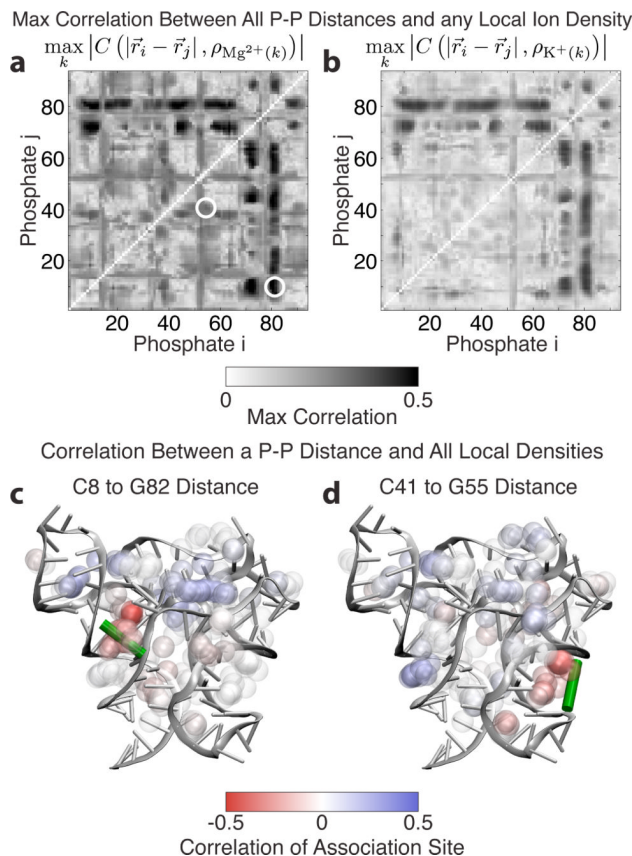


**Figure 1.** The SAM-I riboswitch aptamer domain. (a) The secondary structure with the SAM binding pocket (nucleotides within 5 Å of SAM) shaded. Nucleotides are color-coded according to secondary structure element. Helices P1, P2a, P2b, P3, P4 are labeled, along with the kink-turn (KT) and pseudo-knot (PK). (b) The crystallographic structure of the SAM-I riboswitch aptamer domain solved by Batey and co-workers, colored using the same color scheme as (a). Magenta, SAM ligand.

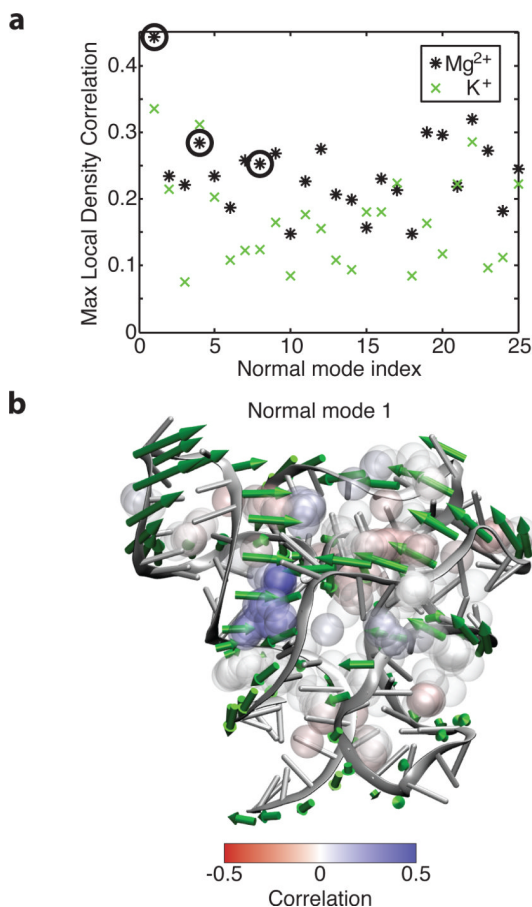




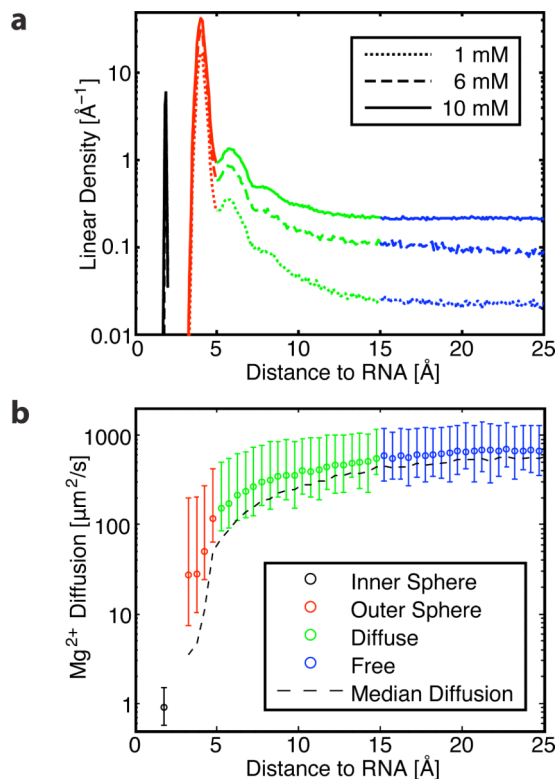
**Figure 2.**  $Mg^{2+}$  affects RNA kinetics. (a) The distribution of autocorrelation times (ACT) for the MD trajectories projected onto the lowest 50 normal modes. The time scales of RNA fluctuations increase with  $Mg^{2+}$  concentration in accordance with experiment. This occurs because  $Mg^{2+}$  tightens RNA tertiary contacts and stabilizes compact conformations. (b) The distribution of autocorrelation times for the local  $Mg^{2+}$  density around several association sites. The time scales of RNA fluctuations are comparable to the time scale of the RNA fluctuations, which allows them to kinetically couple. (c) The residence time for the association sites.

**Figure 3.**

Locally, individual  $\text{Mg}^{2+}$  association events pull phosphates together by bridging interactions. (a) The correlations between phosphate-phosphate (P-P) distance and the local  $\text{Mg}^{2+}$  density at the most strongly correlated association site (10 mM simulation). The pairs in panels (c) and (d) are circled. (b) The same plot for  $\text{K}^+$  shows lower correlations (0 $\dagger$  mM simulation). These correlations represent the passive changes in ionic concentration due to changing RNA structure. (c-d) Two representative P-P distances across a major groove (illustrated by the green bar). The  $\text{Mg}^{2+}$  density in the groove is negatively correlated (red) because when density in these regions increases, the P-P distance decreases. Balls represent association sites. The density about an association site is computed within a 10 Å shell, but the sites are plotted with a radius of 2.5 Å for clarity, and colored according to their correlation with the P-P distance.

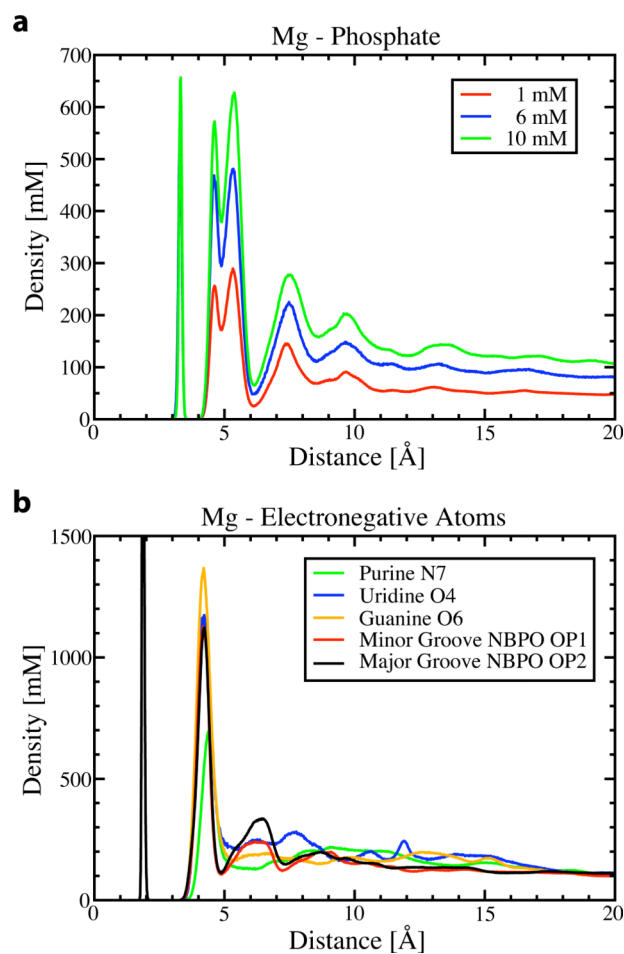


**Figure 4.**  $Mg^{2+}$  association and resulting local bridging interactions also affect the global RNA conformation. (a) Correlations between normal mode projections and the most strongly correlated local  $Mg^{2+}$  density for each normal mode. Several of the first few normal modes are strongly correlated. Circled modes are shown in (b) and in Figure S5. The positive direction of motion along the normal mode is plotted as green arrows. Densities in regions which close upon positive motion along the normal mode are generally positively correlated (blue), while densities in regions which open are generally negatively correlated (red). We note that in the case of the P4 major groove described earlier, positive motion along the normal mode corresponds to a decrease in P-P distance. Thus, a negative correlation between the  $Mg^{2+}$  density and the P-P distance corresponds to a positive correlation between the  $Mg^{2+}$  density and positive motion along the normal mode. Balls represent association sites, and are plotted one quarter scale as in Figure 3.

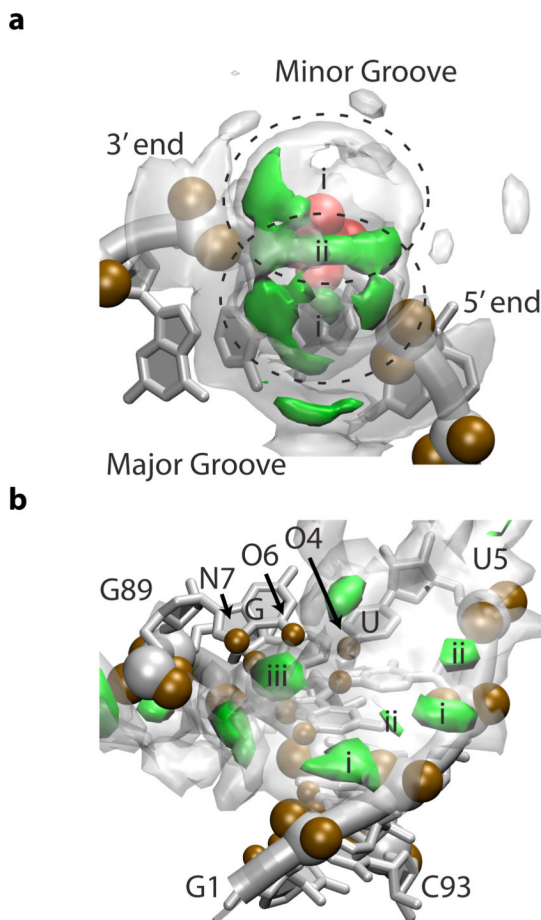


**Figure 5.**

Four classes of  $\text{Mg}^{2+}$  based on RNA distance emerged during our analysis: *inner-sphere* (black) at 2  $\text{\AA}$  *outer-sphere* (red) at 3 to 5  $\text{\AA}$ , and *diffuse* (green) at 5 to 15  $\text{\AA}$ , blending into *free* (blue). (a) The population of  $\text{Mg}^{2+}$  as a function of distance from the RNA. The outer-sphere population, which has previously received little attention, composes 80% to 85% of  $\Gamma_{\text{Mg}^{2+}}$  (Table 2). (b) The diffusion of  $\text{Mg}^{2+}$  as a function of distance from the RNA. Closer  $\text{Mg}^{2+}$  ions show decreased values of mobility. Bars reflect the variance of instantaneous diffusion above and below the mean. The uncertainty from bootstrap analysis averages approximately 3%. The median instantaneous diffusion is plotted as a dashed line. While the median is comparable to the mean in most regimes, it is substantially below the mean for outer-sphere  $\text{Mg}^{2+}$  because rare dissociation events dominate the mean. Values are from the 10 mM simulation.



**Figure 6.** Plotted are several radial distribution functions for Mg<sup>2+</sup>. (a) The RDF with the phosphorous atom at several concentrations. The dual peak at 5 Å indicates substructure. This substructure is clarified in (b) the RDF with several electronegative RNA atoms (10 mM simulation), including the phosphate oxygens. See Figure 7b for a 3D representation of electronegative atoms.



**Figure 7.** Spatial distribution functions (SDF) of  $\text{Mg}^{2+}$  about SAM-I. In our study, simulations were performed with 0 mM, 1 mM, 6 mM and 10 mM magnesium concentration. Here, we show magnesium density isosurfaces, displaying localized regions of high magnesium concentration. Specifically,  $\text{Mg}^{2+}$  density surfaces of 1.4 M and 7 M are shown in transparent white and solid green respectively (10 mM simulation with SAM). The density of the outer-sphere  $\text{Mg}^{2+}$  distribution is expressed in molarity to highlight the more than 700-fold enrichment of  $\text{Mg}^{2+}$  above the bulk concentration of 10 mM. See SI Section 8 for more details. (a) The average  $\text{Mg}^{2+}$  SDF about a phosphate (red). A representative backbone is shown for reference. Several hydration layers and a preference for the major groove are visible. Regions (i) and (ii) correspond to the two peaks in Figure 6a (b) The  $\text{Mg}^{2+}$  SDF about helix P1. Electronegative atoms controlling the distribution are labeled and shown in brown. Sites occur along the backbone (i&ii) or within the major groove near the brown electronegative atoms (iii).

Table 1

Contains (from top to bottom) number of ions in each simulation, raw concentrations determined 20 Å beyond RNA (asterisk), concentrations corrected with a small potential perturbation approximation, and preferential interaction coefficients ( $\Gamma$ ) obtained from the corrected concentrations.

	0 mM	0 <sup>+</sup> mM	1 mM	6 mM	10 mM
$N_{Mg^{2+}}$	0	1	11	22	33
$N_{K^+}$	137	133	116	100	84
$N_{Cl^-}$	45	43	46	52	58
$[Mg^{2+}]^*$ <sup>a</sup>	0.0 ± 0.0 mM	0.0 ± 0.0 mM	1.1 ± 0.9 mM	6.7 ± 2.6 mM	11.2 ± 1.6 mM
$[K^+]^*$	115.3 ± 2.6 mM	114.2 ± 1.6 mM	113.8 ± 1.9 mM	110.0 ± 4.9 mM	105.5 ± 2.7 mM
$[Cl^-]^*$	93.1 ± 1.8 mM	90.0 ± 1.1 mM	94.8 ± 1.2 mM	106.5 ± 1.4 mM	115.3 ± 0.7 mM
$[Mg^{2+}]^b$	0.0 ± 0.0 mM	0.0 ± 0.0 mM	0.9 ± 0.7 mM	5.8 ± 2.2 mM	10.2 ± 1.4 mM
$[K^+]$	103.0 ± 2.0 mM	100.6 ± 1.2 mM	102.5 ± 1.7 mM	102.3 ± 4.6 mM	100.5 ± 2.5 mM
$[Cl^-]$	103.0 ± 2.0 mM	100.6 ± 1.2 mM	104.2 ± 1.1 mM	113.9 ± 1.9 mM	120.8 ± 0.8 mM
$\Gamma_{Mg^{2+}}$	0.0 ± 0.0	1.0 ± 0.0	10.5 ± 0.4	18.7 ± 1.3	27.2 ± 0.8
$\Gamma_{K^+}$	78.0 ± 1.1	75.4 ± 0.7	57.3 ± 1.0	41.4 ± 2.6	26.5 ± 1.5
$\Gamma_{Cl^-}$	-14.0 ± 1.1	-14.6 ± 0.7	-13.7 ± 0.6	-13.2 ± 1.1	-11.2 ± 0.5

<sup>a</sup>Raw concentration

<sup>b</sup>Corrected concentration  $\propto N - \Gamma$

**Table 2**

The number and percentage of ions in the distance classes in Figure 5. Predicted chelated ions are a subset of the inner-sphere class. Note that the outer-sphere population contains 80% to 85% of the excess  $\text{Mg}^{2+}$  ions  $\Gamma_{\text{Mg}^{2+}}$ .

	1 mM	6 mM	10 mM
$\Gamma_{\text{Mg}^{2+}}$	10.5	18.7	27.2
Number of $\text{Mg}^{2+}$			
Inner-Sphere	1.0	1.0	1.0
Outer-Sphere	8.6	16.0	22.3
Diffuse	0.9	1.7	3.9
Percent of $\Gamma_{\text{Mg}^{2+}}$			
Inner-Sphere	9.5	5.4	3.7
Outer-Sphere	81.8	85.8	81.8
Diffuse	8.7	8.9	14.5