

Quantification of OH and HO₂ radicals during the low-temperature oxidation of hydrocarbons by Fluorescence Assay by Gas Expansion technique

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•OH and •HO₂ radicals are known to be the key species in the development of ignition. A direct measurement of these radicals under low-temperature oxidation conditions (T = 550–1,000 K) has been achieved by coupling a technique named fluorescence assay by gas expansion, an experimental technique designed for the quantification of these radicals in the free atmosphere, to a jet-stirred reactor, an experimental device designed for the study of low-temperature combustion chemistry. Calibration allows conversion of relative fluorescence signals to absolute mole fractions. Such radical mole fraction profiles will serve as a benchmark for testing chemical models developed to improve the understanding of combustion processes.

HOx radicals | autoignition

In the context of a needed significant reduction of the emission of greenhouse gases, understanding the autoignition chemistry of hydrocarbons and biofuels is of critical importance to allow the development of new combustion strategies for clean and efficient internal combustion engines (1). Advanced combustion concepts that rely on compression self-ignition (1, 2), as well as the improvement of safety in oxidation processes (3), demand an improved understanding of the reaction kinetics governing the detailed mechanisms of organic compounds, particularly in the low-temperature oxidation regime. The chemistry of oxidation and autoignition is governed by species, with unpaired electrons, called free radicals, which are present in trace amounts. Among free radicals important in oxidation are H atoms, alkyl radicals, and hydroxyl (•OH) or hydroperoxy (•HO₂) radicals, the last two being the most important. It is then of critical importance for a better understanding of this chemistry to develop methods to measure radicals under all possible reaction conditions governing autoignition, including low-temperature oxidation (below 1,000 K).

Because •OH radicals are present in relatively high concentrations at high temperatures, they have been measured for a long time in flames by molecular beam mass spectrometry (4) and by laser-based methods such as laser-induced fluorescence (LIF) or absorption spectroscopy (5, 6). The detection of •HO₂ radicals has recently been achieved in a flame using a photo-fragmentation laser-induced fluorescence technique (7) as well as during the oxidation of dimethylether in a laminar flow reactor using midinfrared Faraday rotation spectroscopy (8). A method to measure the ratio between •HO₂ and •RO₂ radicals based on electron spin resonance spectroscopy was proposed by Carlier and Sochet (9) but has never been applied to real hydrocarbons and low-temperature oxidation.

H atoms (10) or •OH radicals (11) were also followed under high-temperature (T > 1,200 K) shock tube conditions. Finally, planar LIF using a laser sheet has been used in optical access engines to image OH radical concentrations to map the hot combustion zones (12). This technique does not directly give

access to a real quantitative measurement of the radical concentration due to differences in quenching and line broadening caused by local gradients of fuel/air equivalence ratio and temperature (13).

In this work, we present the coupling of an optical technique named Fluorescence Assay by Gas Expansion (FAGE), initially designed for the measurement of absolute concentrations of •OH and •HO₂ radicals in the atmosphere (14), to a combustion device called a jet-stirred reactor (JSR), with the goal of quantifying both radicals under reaction conditions similar to those found before development of autoignition (2). A JSR allows studying gas-phase reactions at temperatures up to 1,000 K without external radical production. Such measurements will be particularly valuable for improving knowledge of the chemistry and detailed kinetic models in the temperature zone (650–900 K), in which reaction mechanisms are the most complex and the kinetic data are the least accurate (15).

The schematic reaction mechanism currently accepted for the low-temperature combustion of many typical hydrocarbon fuel molecules is shown in Fig. 1, in which the key reactions are numbered to facilitate the following discussion. The starting point of the chain reaction leading to ignition is the reaction of a hydroxyl radical (•OH), the main chain carrier, with the fuel molecule, named hereafter RH. Abstraction of an H atom by the OH radical (reaction 1) leads to formation of a radical R•. At low temperatures (below around 750 K), R• rapidly yields a peroxy radical (ROO•) after a barrierless reaction with an oxygen molecule (reaction 2). This ROO• radical leads then to the formation of hydroperoxides and an •OH radical, mainly by a complex mechanism involving two isomerizations and addition of a second oxygen molecule, not described here. The obtained

Significance

The design of internal combustion engines relies on a good understanding of the kinetic mechanism of the autoignition of hydrocarbons. •OH and •HO₂ radicals are known to be the key species governing all stages of the development of ignition. A direct measurement of these radicals under low-temperature oxidation conditions has been achieved by coupling the fluorescence assay by gas expansion technique, an experimental technique designed for the quantification of these radicals in the free atmosphere, to a jet-stirred reactor, an experimental device designed for the study of low-temperature combustion chemistry.

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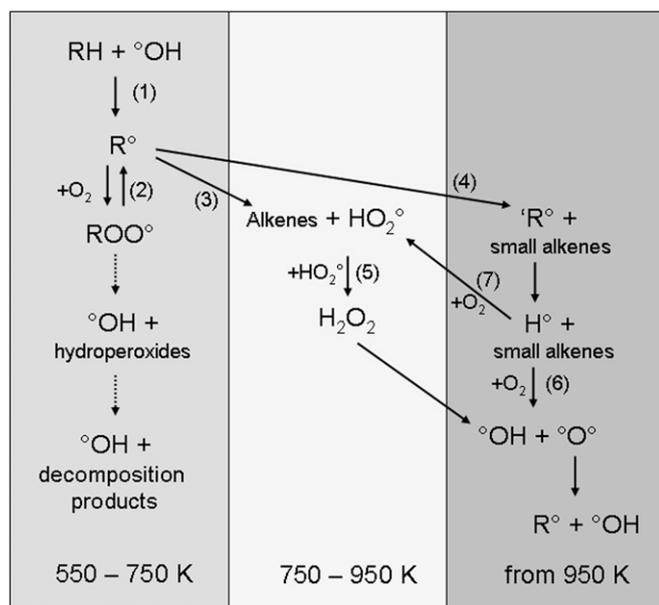


Fig. 1. Simplified scheme of the mechanism of oxidation of hydrocarbons (dotted arrows represent a succession of several elementary steps).

hydroperoxides can easily decompose giving two radicals, including a second $\bullet\text{OH}$ radical. This degenerate chain-branching reaction explains the high reactivity of hydrocarbon/oxygen mixtures at low temperatures.

An increase in temperature enhances the reversibility of reaction 2, thus hindering the formation of peroxy radicals and consequently of $\bullet\text{OH}$ radicals. At moderate temperatures (between around 750 and 950 K), the radical $\text{R}\bullet$ reacts with oxygen, yielding an unsaturated hydrocarbon (named alkenes in Fig. 1) and an $\bullet\text{HO}_2$ radical (reaction 3), or, if temperature is high enough, decomposes into a small alkene and another radical $\text{R}\bullet$ (reaction 4). $\bullet\text{HO}_2$ radicals are much less reactive than $\bullet\text{OH}$ radicals and mainly react in a termination step leading to hydrogen peroxide (H_2O_2) (reaction 5). This reduced importance of degenerate chain-branching reactions in favor of chain-terminating reactions leads to a slowdown of the reactivity and explains the occurrence of a feature specific to mixtures of organic compounds with oxygen (16): the commonly called “negative temperature coefficient” (NTC) zone; i.e., a zone (usually around 650 K under atmospheric pressure) where the reactivity decreases with increasing temperature.

With further increased temperature, the decomposition of H_2O_2 leading to two $\bullet\text{OH}$ radicals becomes faster and is the dominant degenerate chain-branching reaction, again promoting high reactivity. Above around 950 K, small hydrocarbon radicals such as obtained by reaction 4 can in turn decompose, leading to alkenes and $\bullet\text{H}$ atoms. The reaction of an $\bullet\text{H}$ atom with oxygen at high temperatures leads to an $\bullet\text{OH}$ radical and to an $\bullet\text{O}$ atom (reaction 6), which in turn regenerates an $\bullet\text{OH}$ radical by H abstraction from the fuel. Reaction 6 is a true branching step, ensuring the full development of ignition and complete combustion. However, the pressure-dependent reaction 7 can compete with reaction 6, depending on temperature and pressure, and leads to the formation of the less reactive $\bullet\text{HO}_2$ radicals.

This scheme shows clearly the critical importance of $\bullet\text{OH}$ and $\bullet\text{HO}_2$ radicals in all stages linked to the development of ignition. The purpose of this study is therefore to present an experimental device allowing the quantification of both radicals under conditions observed during the reaction period before the auto-ignition of hydrocarbons. *n*-Butane has been studied as it is the

smallest hydrocarbon structure that exhibits an oxidation behavior representative to species commonly present in gasoline and diesel fuels.

The goal of this work was to reproduce the chemistry leading to this phenomenon rather than to observe unsteady phenomena directly such as autoignition. For this purpose, a heated JSR has been used, a device that has already been used many times for studying the chemistry of organic compounds before ignition (15). Note that this reactor has been recently used to follow the formation of alkylhydroperoxides (17) and H_2O_2 (18) through its coupling with a reflectron time-of-flight mass spectrometer and cw-Cavity Ring Down Spectroscopy (CRDS) cell, respectively. CRDS is known to be a very sensitive technique to measure $\bullet\text{HO}_2$ radicals, but surprisingly these radicals could not be detected in these experiments, even though concentrations much above the detection limit of this technique (19, 20) were expected (18).

Results and Discussion

FAGE is a highly sensitive method for quantifying $\bullet\text{OH}$ and $\bullet\text{HO}_2$ radicals, originally designed for the quantification of these radicals in the free atmosphere. Mole fractions of $\bullet\text{OH}$ and $\bullet\text{HO}_2$ radicals found in the atmosphere are typically 0.2 and 8 ppt, respectively (21). Today eight research groups in the world have operational FAGE systems (22–28), but all of them are used in research solely linked to atmospheric chemistry. To our knowledge, no other work has been published where the advantages of radical detection by FAGE technique are exploited in the field of combustion research.

FAGE is based on LIF after expansion of the gas mixture (in general, the ambient atmosphere) through a small orifice (few 100 μm to 1 mm) to pressures of around 0.5 torr. $\bullet\text{HO}_2$ radicals, which do not fluoresce, are converted to $\bullet\text{OH}$ radicals by their fast reaction with NO. Indirect detection of $\bullet\text{HO}_2$ after conversion to $\bullet\text{OH}$ radicals through addition of NO to the reaction system is a common technique (29), leading possibly to complications due to the reaction of NO with other intermediates. In the case of FAGE, however, NO is added only after the expansion of the gas mixture into the FAGE cell; i.e., when the reaction system is already at low temperature and pressure. Recent experiments have shown that, even under these conditions, $\bullet\text{RO}_2$ radicals might also be converted to $\bullet\text{OH}$ radicals in a two-step reaction sequence with the conversion efficiency depending strongly on the NO concentration (30). In the present experiments, the NO concentration was kept very low (*SI Text*) and no interference has been detected.

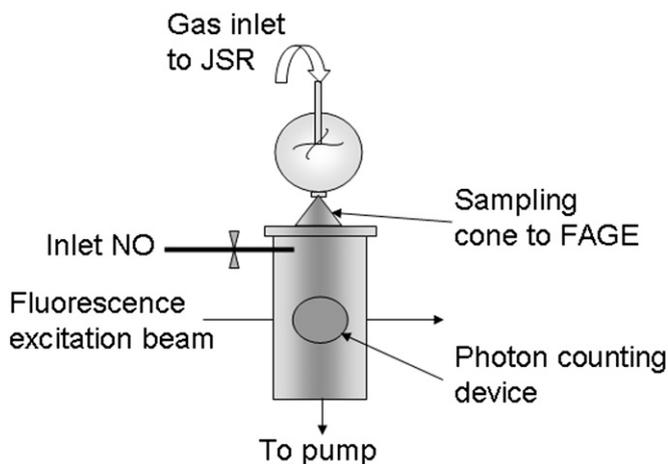


Fig. 2. Simplified scheme of the experimental device (not to scale). The upper part (JSR) is at atmospheric pressure and heated at high temperatures, and the lower part (FAGE) is at low pressure (<1 torr) and ambient temperature.

radicals, the model underpredicts the concentration at nearly all temperatures; only at temperatures above 900 K is the measured OH concentration lower than the predicted.

Summary

We have presented here direct measurements of $\bullet\text{OH}$ and $\bullet\text{HO}_2$ radical profiles under combustion conditions combining the advantages of a radical detection by FAGE to a chemical reactor well adapted for studying low-temperature combustion phenomena, a JSR. Although the absolute concentrations are somewhat uncertain due to calibration and sampling issues, the relative temperature profiles are well defined. Such measurements will certainly lead to a big step forward in refining current

models of combustion chemistry, as these radical profiles will be a powerful benchmark for testing models that have so far been developed mostly on the basis of concentration profiles for stable species only.

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