



Published in final edited form as:

Chem Commun (Camb). 2016 March 21; 52(23): 4389–4392. doi:10.1039/c6cc00959j.

Polymorphism in Phenobarbital: Discovery of a New Polymorph and Crystal Structure of Elusive form V

Saikat Roy^{a,c}, N. Rajesh Goud^a, and Adam J. Matzger^{a,b}

Adam J. Matzger: Matzger@umich.edu

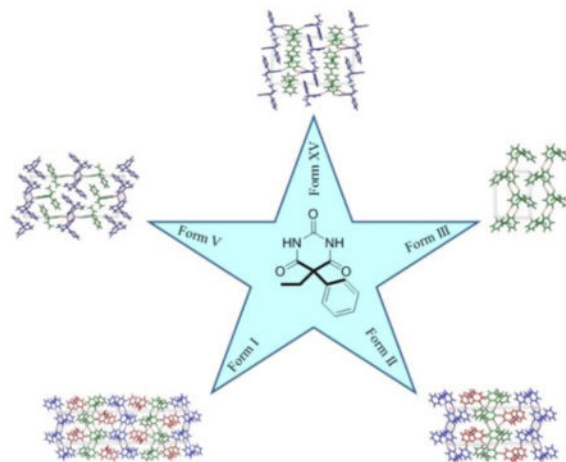
^aDepartment of Chemistry, University of Michigan, Ann Arbor, Michigan, 48109-1055, USA

^bMacromolecular Science and Engineering, University of Michigan, 48109-1055, USA

Abstract

This report highlights the discovery of a new polymorph of the anticonvulsant drug phenobarbital (PB) using polymer-induced heteronucleation (PIHn) and unravelling the crystal structure of the elusive form V. Both forms are characterized by structural, thermal and VT-Raman spectroscopy methods to elucidate phase transformation behavior and shed light on stability relationships.

Graphical Abstract



This report highlights the discovery of new polymorph ‘form XV’, and the crystal structure of the elusive form V of the anticonvulsant drug Phenobarbital.

Polymorphism, the ability of molecules to adopt more than one crystal structure, has attracted significant attention in the pharmaceutical industry due to its intellectual property and regulatory issues.^{1–3} Meticulous analysis of the polymorphic landscape provides valuable insights into structure-property relations which extend beyond pharmaceuticals into

Correspondence to: Adam J. Matzger, Matzger@umich.edu.

^cPresent address: Tata Chemicals Ltd. Innovation Centre, Pune, India.

Electronic Supplementary Information (ESI) available: Experimental and characterization. CCDC numbers: 1450610, 1450611 See DOI: 10.1039/x0xx00000x

other areas such as explosives, pigments etc. On the other hand, polymorphism poses a major challenge to emerging crystal structure prediction methods and indeed represents a frontier for computational chemistry.^{4–6} The inability to predict the propensity of a molecule towards polymorphism, much less the details of structure and stability for these phases, is an “ongoing scandal” in solid state chemistry that is closely related to the much discussed problem of reliable crystal structure prediction from the knowledge of molecular structure.⁷ Although a recent review⁸ suggests that one in three molecules in CSD are polymorphic, the majority of them have one or two structurally characterized polymorphs. Molecules with three or more polymorphs are relatively rare. At the extreme end, our group was successful in tapping the polymorphic potential of flufenamic acid (FFA) through eight structurally characterized polymorphs.⁹ More recently, Zeidan and co-workers unraveled a case of dodecamorphism in aripiprazole (APZ) which now holds a new record with nine structurally characterized polymorphs.¹⁰ A CSD study and literature search revealed that other than FFA and APZ, a pharmaceutical precursor 5-methyl-2-[(2-nitrophenyl)amino]-3-thiophenecarbonitrile (ROY)¹¹ has seven forms elucidated structurally and to the best of our knowledge only eight other pharmaceuticals with five polymorphs (Fig. S1, ESI[†]). The polymorphic propensity of these pharmaceuticals may be attributed to various factors such as synthon diversity, flexible molecular conformation and packing differences in their crystal structures.^{12–25}

In this study, we highlight conformational polymorphism^{3,26} of the anticonvulsant drug phenobarbital (PB, 5-ethyl-5-phenylpyrimidine-2,4,6(1*H*,3*H*,5*H*)-trione), Fig. 1. It is recorded on the list of *essential medicines*²⁷ by the World Health Organization (WHO). PB is used widely for treating epilepsy and neonatal seizures. The polymorphic propensity of PB has intrigued researchers over the last six decades.^{28–37} In total, eleven polymorphs (I–XI), a hydrate (form XII), a hemihydrate (form XIII) and an amorphous phase (form XIV) are reported in the literature, making it one of the benchmark pharmaceuticals for studying polymorphism and crystal structure prediction.^{38,39} However, thus far only three polymorphs were amenable to structure elucidation, hinting at a molecule resistant to forming single crystals suitable for structure determination.^{35, 37}

Discovery of polymorphs for such well-studied pharmaceuticals pose a major challenge to polymorph discovery techniques, such as polymer-induced heteronucleation (PIHn).^{40, 41} PIHn is a crystallization technique through which a single solvent/temperature condition leads to novel form discovery, enabled by a diverse array of insoluble polymers acting as heterogeneous growth sites.^{6, 13, 42–44} Initial crystallization utilizing PIHn resulted in forms I–III. Normal melt crystallization of PB yields an amorphous solid, which over the time converts to form I and II in the same experimental condition. However, melt crystallization of PB mixed with 4-(benzyloxy) benzonitrile modified polystyrene resulted in a novel phase, termed as form XV. Further, controlled sublimation of PB resulted in the crystallization of the elusive form V as very thin needles (SI. 2, experimental section, ESI[†]). Optical microscopy (Fig. S2, ESI[†]) and Raman spectroscopy (Fig. S3, ESI[†]) identified these five distinct polymorphs with characteristic Raman spectra in two main spectral regions at 3400–2900 and 1800–1400 cm⁻¹ (Fig. 2). Powder X-ray diffraction (PXRD) patterns clearly reveal the difference between form V and XV and distinguish them from other forms (Fig. S4, ESI[†]). Form V exhibits unique 2θ peaks at 7.0, 8.0, 13.2, 14.1, 14.9 and 17.2, whereas novel

form XV was identified from the peaks at 6.8, 7.3, 7.6, 11.9, 13.7, 14.7, 15.3, 19.6, 2θ (Table S1, ESI[†]).

Single crystal X-ray diffraction was carried out on PB form V and XV. Form V is known since 1951 and unit cell dimensions were reported by Williams in 1973.³⁶ Since then several researchers^{30,31} were able to produce only microcrystalline form V either by melt crystallization or sublimation. Through sublimation experiments using a hostage connected control processor, we were able to obtain diffraction quality form V crystals suitable for structure determination. Both form V and Form XV crystallized in monoclinic $P2_1/n$ space group with two molecules in the asymmetric unit (Fig. S5, Table S2, ESI[†]). In the five crystal structures of PB, eleven symmetry independent molecules represent diverse conformational flexibility consistent with conformational polymorphism (Fig. 3). Such diversity is a major challenge for structure prediction by computational methods. In fact, a crystal structure prediction exercise was attempted by Day *et al* to predict form V.³⁹ Though none of the predicted structures with $Z'=2$ matches with form V or form XV, structural analysis revealed that the predicted hydrogen bonding motifs of one of the structures matched with form V (Vide infra).

Structural analysis of the five PB forms display the presence of N–H...O hydrogen bonded $R_2^2(8)$ dimers (Fig. 4a). These dimers propagate through various symmetry elements (inversion, glide or screw axes) to form 1D ribbon chains in different forms. In form I and form II (both with $Z'=3$) two independent chains propagate through two symmetry independent molecules (blue and green molecules in Fig. 5a and 5b) connected by a third molecule (red) acting as bridge between the two chains. Overall both form I and form II exhibit very similar crystal packing (Fig. 5a and 5b). In form III ($Z'=1$), 2_1 screw axis extend these ribbon-like chains which are further supported by C–H...O hydrogen bonds (Fig. 5c). In Form V ($Z'=2$), one type of symmetry independent molecule (blue) extend as a molecular chain, while the other molecule (green) connects this chain through N–H...O hydrogen bonds. The N–H...O catemer chains of crystallographically inequivalent molecules interact with each other forming an $R_3^3(12)$ tape motif (Fig. 4b), completing the 3D packing (Fig. 5d). Presence of this hydrogen bonded tape motif was predicted in the PB crystal structure prediction exercise by Day *et al*.³⁹ The hydrogen bond dimer chains are notably different in form XV, where N–H...O dimers involve two symmetry independent molecules (Fig. 4c). These dimers further extend through N–H...O hydrogen bonds enabling a significantly different packing for form XV in comparison to other polymorphs (Fig. 5e).

Hirshfeld surface⁴⁵ analysis of the polymorphs of PB (Fig. 6) revealed that form I, II and III exhibit greater contribution from O...H contacts (35.9%, 35.7% and 34.7%) than form V (32.7%) and form XV (33.3%). However, form III has more N...H (4%) and C...H (19%) contacts than other polymorphs. Overall, the analysis suggests that form V and XV have comparatively lower contribution from strong hydrogen bonding contacts than the other three polymorphs (Fig. S6, ESI[†]).

Differential scanning calorimetry (DSC) confirmed that form I exhibits the highest melting point (176 °C), while all the other polymorphs convert to form I during heating (Fig. S7, ESI[†]). This thermal conversion was monitored through variable temperature Raman

spectroscopy. Form I does not show any phase transformation on heating whereas form II converts to form I before melting at 170 °C (Fig. S8a, ESI[†]). Form V converts completely to form III at 120 °C (Fig. 7), which in turn transforms to form II and I above 165 °C (Fig. S8b, ESI[†]). Form XV converts to form II and I above 80 °C (Fig. S8c, ESI[†]). In order to determine the relative free energy relationship, slurry conversion experiments were carried out with PB polymorphs in ethanol. This experiment resulted in the complete conversion of all polymorphs to form I. Thus at room temperature form I is the most stable modification of PB. Equilibrium solubility experiments in heptane suggested that form I and form II are similar in energy, whereas form XV is the least stable modification among forms I, II, III and XV (Fig. S9, ESI[†]).

In conclusion, new polymorphic landscape has been accessed for a compound that is well studied over the last six decades. This study highlights the importance of crystal growth techniques like PIHn in expanding the polymorphic propensity of a molecule that is previously known to exhibit several polymorphs. Moreover we were able to solve the half a century old puzzle of elucidating form V crystal structure. The presence of diverse hydrogen bonding motifs, multiple molecules in the asymmetric unit, and a high degree of conformational flexibility, project PB as a challenging benchmark molecule for crystal structure prediction. Unraveling the crystal structures of other polymorphs of PB, augmented by better crystal growth methods and more powerful diffraction techniques may further clarify the intricate relationship between molecular conformation and polymorphism.

Supplementary Material

Refer to Web version on PubMed Central for supplementary material.

Acknowledgments

This work was supported by the National Institutes of Health Grant Number GM106180. We thank Jeff W. Kampf for his help with structure determination of form V.

References

1. Byrn, SR.; Pfeiffer, RR.; Stowell, JG. *Solid-State Chemistry of Drugs*. SSCI; West Lafayette, IN: 1999.
2. Hilfiker, R. *Polymorphism in the Pharmaceutical Industry*. Wiley-VCH, Weinheim; Weinheim: 2006.
3. Bernstein, J. *Polymorphism in Molecular Crystals*. Clarendon: Oxford; 2002.
4. Day GM, Cooper TG, Cruz-Cabeza AJ, Hejczyk KE, Ammon HL, Boerrigter SXM, Tan JS, Della Valle RG, Venuti E, Jose J, Gadre SR, Desiraju GR, Thakur TS, van Eijck BP, Facelli JC, Bazterra VE, Ferraro MB, Hofmann DWM, Neumann MA, Leusen FJJ, Kendrick J, Price SL, Misquitta AJ, Karamertzanis PG, Welch GWA, Scheraga HA, Arnautova YA, Schmidt MU, van de Streek J, Wolf AK, Schweizer B. *Acta Crystallogr Sect B: Struct Sci*. 2009; 65:107.
5. Neumann MA, Leusen FJJ, Kendrick J. *Angew Chem Int Ed*. 2008; 47:2427.
6. Roy S, Matzger AJ. *Angew Chem Int Ed*. 2009; 48:8805.
7. Maddox J. *Nature*. 1988; 335:201.
8. Cruz-Cabeza AJ, Reutzel-Edens SM, Bernstein J. *Chem Soc Rev*. 2015; 44:8619. [PubMed: 26400501]
9. Lopez-Mejias V, Kampf J, Matzger AJ. *J Am Chem Soc*. 2012; 134:9872. [PubMed: 22690822]

10. Zeidan TA, Trotta JT, Tilak PA, Oliveira MA, Chiarella RA, Foxman BM, Almarsson Ö, Hickey MB. *CrystEngComm*.
11. Chen S, Guzei IA, Yu L. *J Am Chem Soc*. 2005; 127:9881. [PubMed: 15998094]
12. Blagden N, Davey RJ, Lieberman HF, Williams L, Payne R, Roberts R, Rowe R, Docherty R. *J Chem Soc Faraday Trans*. 1998; 94:1035.
13. Lopez-Mejas V, Kampf J, Matzger AJ. *J Am Chem Soc*. 2009; 131:4554. [PubMed: 19334766]
14. Drebuschak TN, Chukanov NV, Boldyreva EV. *Acta Crystallogr Sect E: Struct Rep Online*. 2006; 62:o4393.
15. Drebuschak TN, Chukanov NV, Boldyreva EV. *Acta Crystallogr Sect C: Struct Chem*. 2007; 63:o355.
16. Drebuschak TN, Chukanov NV, Boldyreva EV. *Acta Crystallogr Sect C: Struct Chem*. 2008; 64:o623.
17. Braun DE, Gelbrich T, Kahlenberg V, Tessadri R, Wieser J, Griesser UJ. *J Pharm Sci*. 2009; 98:2010. [PubMed: 19009597]
18. Arlin JB, Price LS, Price SL, Florence AJ. *Chem Commun*. 2011; 47:7074.
19. Grzesiak AL, Lang MD, Kim K, Matzger AJ. *J Pharm Sci*. 2003; 92:2260. [PubMed: 14603511]
20. Chekal BP, Campeta AM, Abramov YA, Feeder N, Glynn PP, McLaughlin RW, Meenan PA, Singer RA. *Org Process Res Dev*. 2009; 13:1327.
21. Surov AO, Solanko KA, Bond AD, Perlovich GL, Bauer-Brandl A. *Cryst Growth Des*. 2012; 12:4022.
22. Mishra MK, Desiraju GR, Ramamurty U, Bond AD. *Angew Chem Int Ed*. 2014; 53:13102.
23. Thirunahari S, Aitipamula S, Chow Pui Shan, Tan RBH. *J Pharm Sci*. 2010; 99:2975. [PubMed: 20091821]
24. Nath NK, Nangia A. *CrystEngComm*. 2011; 13:47.
25. Karanam M, Dev S, Choudhury AR. *Cryst Growth Des*. 2012; 12:240.
26. Nangia A. *Acc Chem Res*. 2008; 41:595. [PubMed: 18348538]
27. WHO. Model List of Essential Medicines. 19. World Health Organization (WHO); May. 2015
28. Huang TY. *Acta Pharm Intern*. 1951; 2:95.
29. Huang TY. *Acta Pharm Intern*. 1951; 2:43.
30. Zencirci N, Gelbrich T, Apperley DC, Harris RK, Kahlenberg V, Griesser UJ. *Cryst Growth Des*. 2010; 10:302.
31. Zencirci N, Gelbrich T, Kahlenberg V, Griesser UJ. *Cryst Growth Des*. 2009; 9:3444.
32. Mesley RJ, Clements RL. *J Pharm Pharmacol*. 1968; 20:341. [PubMed: 4385436]
33. Mesley RJ, Clements RL, Flaherty B, Goodhead KJ. *J Pharm Pharmacol*. 1968; 20:329. [PubMed: 4385435]
34. Mesley RJ. *Spectrochim Acta*. 1970; 26A:1427.
35. Williams PP. *Acta Crystallogr Sect B: Struct Sci*. 1974; 30:12.
36. Williams PP. *Acta Crystallogr Sect B: Struct Sci*. 1973; 29:1572.
37. Platteau C, Lefebvre J, Hemon S, Baetz C, Danede F, Prevost D. *Acta Crystallogr Sect B: Struct Sci*. 2005; 61:80.
38. Otsuka M, Nakanishi M, Matsuda Y. *Drug Dev Ind Pharm*. 1999; 25:205. [PubMed: 10065354]
39. Day GM, Motherwell WDS, Jones W. *Phys Chem Chem Phys*. 2007; 9:1693. [PubMed: 17396181]
40. Lang MD, Grzesiak AL, Matzger AJ. *J Am Chem Soc*. 2002; 124:1484.
41. Price CP, Grzesiak AL, Matzger AJ. *J Am Chem Soc*. 2005; 127:5512. [PubMed: 15826189]
42. Grzesiak AL, Uribe FJ, Ockwig NW, Yaghi OM, Matzger AJ. *Angew Chem, Int Ed Engl*. 2006; 45:2553. [PubMed: 16534819]
43. Grzesiak AL, Matzger AJ. *J Pharm Sci*. 2007; 96:2978. [PubMed: 17567888]
44. Foroughi LM, Kang YN, Matzger AJ. *Cryst Growth Des*. 2011; 11:1294.
45. Speckman MA, Jayatilaka D. *CrystEngComm*. 2009; 11:19.

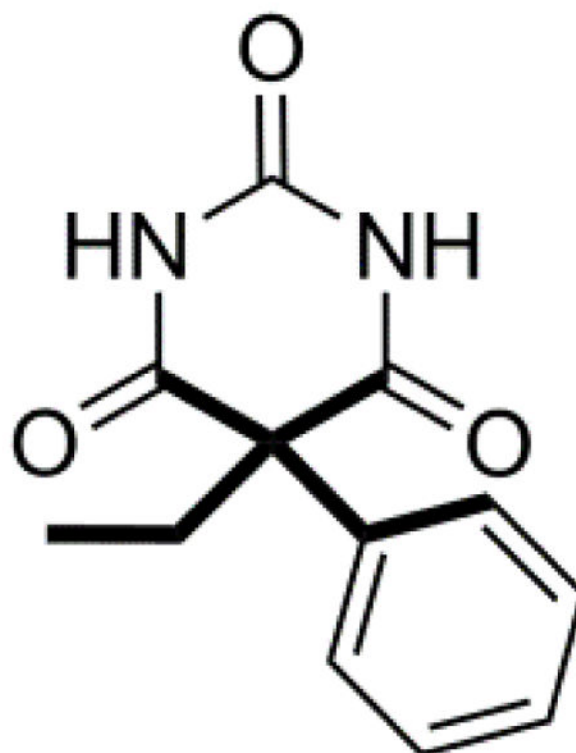


Figure 1.
Chemical structure of Phenobarbital (PB).

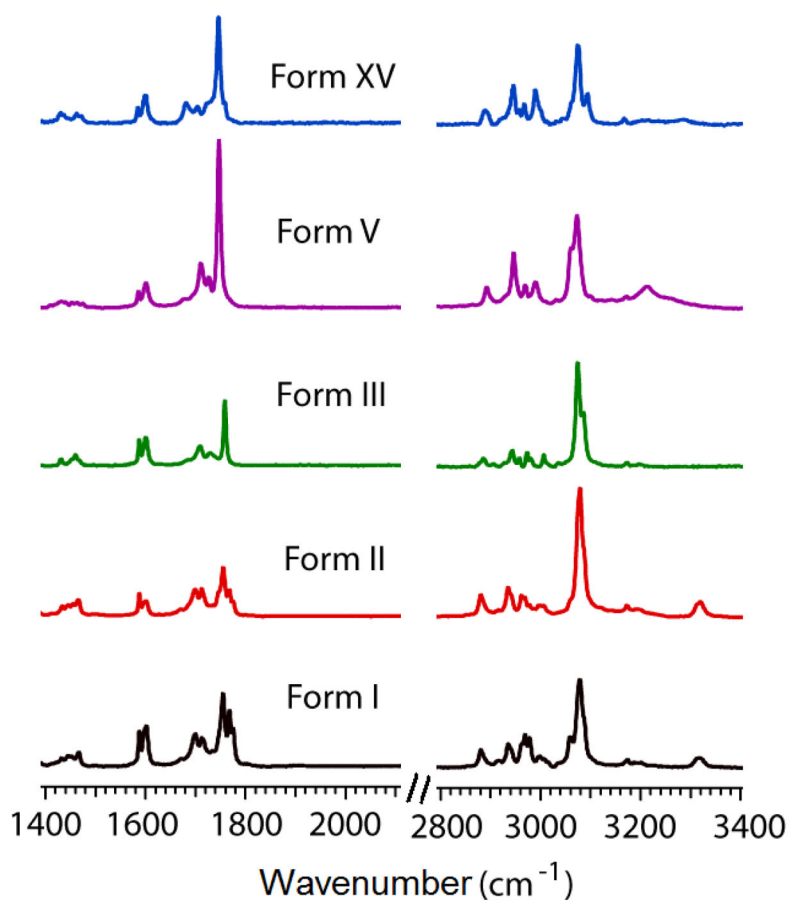


Figure 2.
Raman spectra of PB polymorphs I, II, III, V and XV.

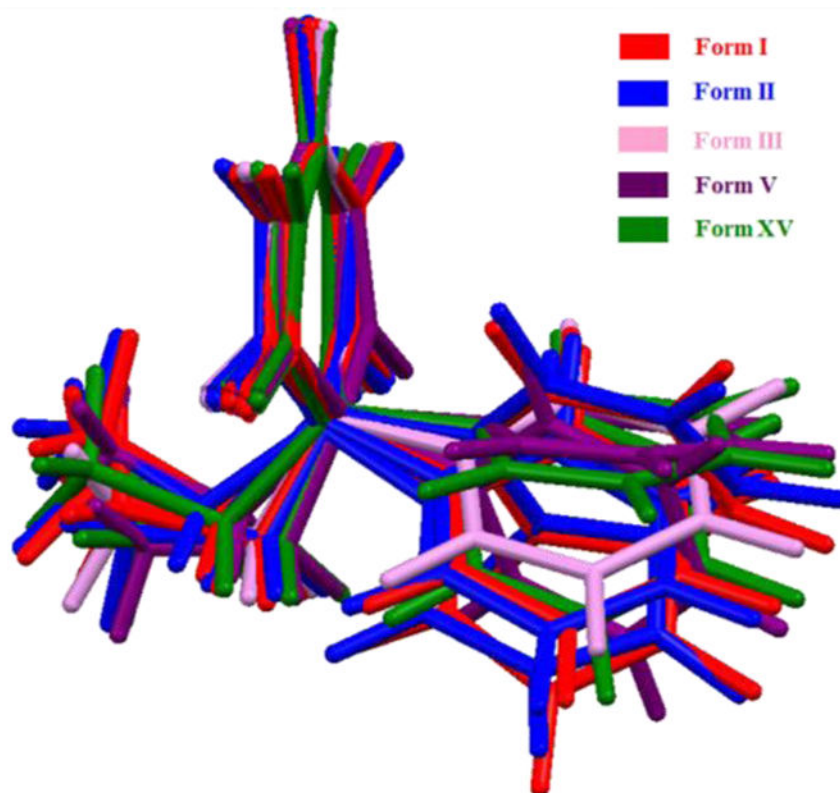


Figure 3.
Overlay of 11 symmetry-independent molecules present in form I, II, III, V and XV of PB.

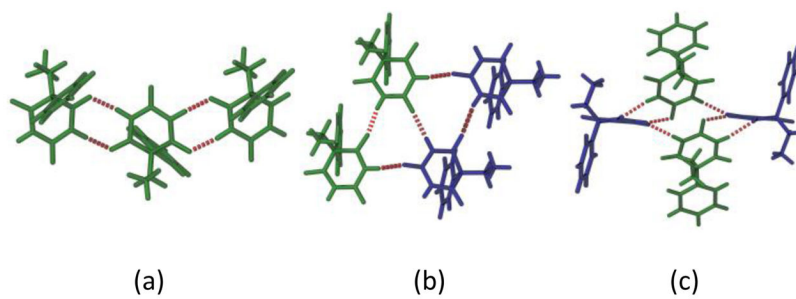


Figure 4. Hydrogen bonding motifs observed in PB polymorphs (a) $R_2^2(8)$ chain motif (b) $R_3^3(12)$ tape motif and (c) ladder.

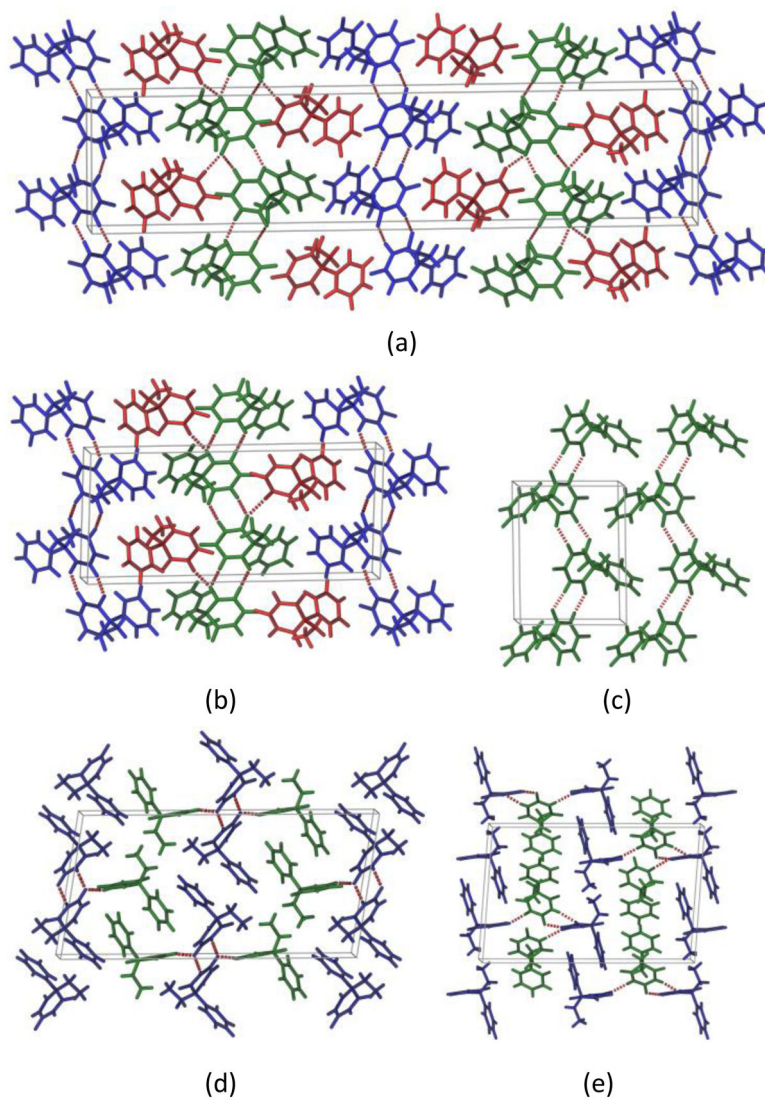


Figure 5. Packing of (a) form I (b) form II (c) form III view along ab plane (d) form V and (e) form XV view along ac plane.

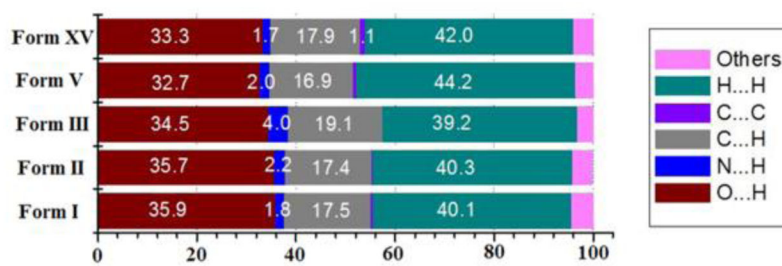


Figure 6. Average relative contribution to the Hirshfeld surface for the various intermolecular contacts for the polymorphs of PB.

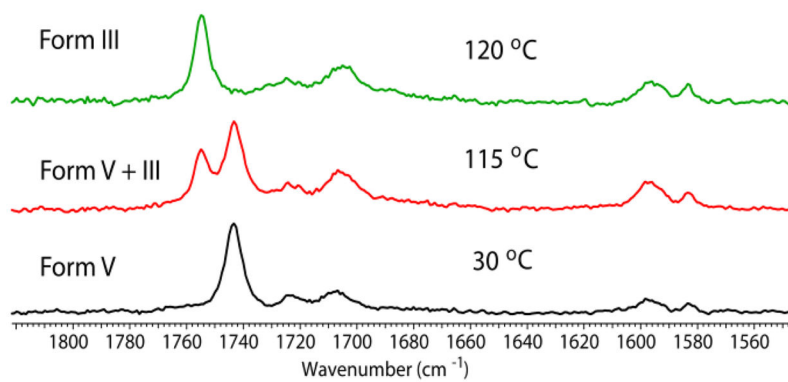


Figure 7. Variable temperature Raman spectroscopy shows partial conversion of form V to form III at 115 °C, with complete conversion occurring at 120 °C.