



# Metal–organic frameworks for H<sub>2</sub> and CH<sub>4</sub> storage: insights on the pore geometry–sorption energetics relationship

Mohamed H. Alkordi, Youssef Belmabkhout, Amy Cairns and Mohamed Eddaoudi\*

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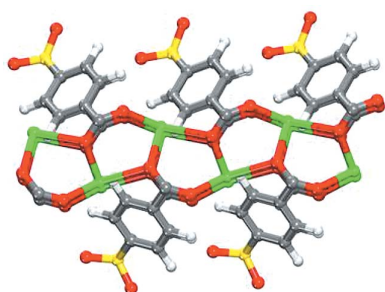
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Division of Physical Sciences and Engineering, Advanced Membranes and Porous Materials Center, Functional Materials Design, Discovery and Development Research Group (FMD<sup>3</sup>), King Abdullah University of Science and Technology (KAUST), Thuwal 4700, Saudi Arabia. \*Correspondence e-mail: [mohamed.eddaoudi@kaust.edu.sa](mailto:mohamed.eddaoudi@kaust.edu.sa)

This study aims to assess the possibility of improving H<sub>2</sub> and CH<sub>4</sub> binding affinity to the aromatic walls of a designed new Metal–Organic Framework (MOF) through simultaneous dispersive interactions. It is suggested here that desirable H<sub>2</sub> and CH<sub>4</sub> storage media at low pressures require narrow uniform pores associated with large surface area, a trade-off that is challenging to achieve.

## 1. Introduction

Metal–Organic Frameworks (MOFs), an emerging class of functional solid-state materials, continue to receive wide scientific interest due to their potential applications in hydrogen storage, gas separation, carbon dioxide capture, enhanced catalysis and drug delivery (Eddaoudi *et al.*, 2001; Moulton & Zaworotko, 2001; Kitagawa *et al.*, 2004; Férey *et al.*, 2005; Horcajada *et al.*, 2006; Cho *et al.*, 2007). MOFs possess unique structural attributes including dual composition, crystallinity and a modular pore system (Rowsell & Yaghi, 2006; Belof *et al.*, 2007; Cho *et al.*, 2007; Dincă *et al.*, 2007; Hayashi *et al.*, 2007; Liu *et al.*, 2007; Alkordi *et al.*, 2008; Banerjee *et al.*, 2008; Llewellyn *et al.*, 2008; Nugent *et al.*, 2013; Belmabkhout *et al.*, 2014; Shekhah *et al.*, 2014). Markedly, these attributes are ideal for the assessment and the establishment of the requisite structure–function relationship toward the construction of made-to-order MOFs for a targeted application. In particular, porous MOFs are regarded as prospective adsorbents that can offer practical solutions to the enduring challenges pertaining to the safe storage and efficient use of H<sub>2</sub> in mobile applications. Conceivably, MOFs are widely investigated for hydrogen storage due to the ability to control their pore system (functionality and volume) and subsequently impact the H<sub>2</sub>–MOF interactions and the total H<sub>2</sub> uptake (Rowsell & Yaghi, 2005; Collins & Zhou, 2007; Lin *et al.*, 2007; Chen *et al.*, 2008; Dincă & Long, 2008; Nouar *et al.*, 2008; Kishan *et al.*, 2010; Zhou *et al.*, 2012). Our group, among others, continue to explore the modularity of MOFs in order to gain better insights on the structure–property relationship and subsequently construct a made-to-order MOF with the ideal gas–MOF interactions and suitable gas uptake for given gas separation/storage applications (Nugent *et al.*, 2013; Belmabkhout *et al.*, 2014; Shekhah *et al.*, 2014). Our study on the **soc**-MOF platform with the underlying square-octahedral (soc) topology (Belof *et al.*, 2007; Liu *et al.*, 2007; Alezi *et al.*, 2015; Cairns *et al.*, 2016) indicated that a made-to-order MOF



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suitable for hydrogen storage at relatively moderate pressures has to be highly porous (high surface area) and concomitantly encompass narrow pores (< 1 nm) and a high localized charged density (polarizable field charges). Notably, a large number of studies on CH<sub>4</sub> and H<sub>2</sub> storage by MOFs delineated the requirement of high surface area and high heat of adsorption for gas storage (Zhou *et al.*, 2012). Hydrogen interactions with metal complexes, clusters or ions, within the inorganic part of a given MOF, are dominated by electrostatic forces between the quadrupole moment of the hydrogen molecule and the inorganic complex. Specifically, such H<sub>2</sub>-MOF interactions' strengths play a major role in determining the H<sub>2</sub> uptake characteristics and hence are the subject of considerable theoretical and experimental investigations (Rowell & Yaghi, 2005; Collins & Zhou, 2007; Hirscher & Panella, 2007; Lin *et al.*, 2007; Dincă & Long, 2008; Nouar *et al.*, 2008; Murray *et al.*, 2009; Zhou *et al.*, 2012). In particular, the weaker, dispersive interactions between H<sub>2</sub> molecules and the organic linkers in MOFs, best represented by benzene ring moieties, have been theoretically investigated (Hübner *et al.*, 2004; Bhatia & Myers, 2006; Lochan & Head-Gordon, 2006; Düren *et al.*, 2009; Han *et al.*, 2009) and experimentally documented (Rosi *et al.*, 2003). Recent studies demonstrate that such interactions could, in principle, be enhanced through chemical modifications of the organic linkers, providing a prospective strategy for a material designer to fine-tune the organic building blocks and subsequently enhance H<sub>2</sub> sorption characteristics of the MOF (Lochan & Head-Gordon, 2006). Nevertheless, to the best of our knowledge, no experimental synthetic studies have been published to address the potential to improve the H<sub>2</sub> and CH<sub>4</sub> binding affinity to the walls of MOFs through simultaneous dispersive interactions, acting additively, between the gas molecules and multiple aromatic rings placed at optimal interaction distance(s) within a specific geometry. Therefore, we opted to explore this approach separately, regardless of the degree of porosity, which could potentially pave the way for the rational design of MOF adsorbents as suitable and effective gas storage media. Computational studies revealed moderate binding affinities for the H<sub>2</sub> molecule towards benzene and various aromatic rings (Hübner *et al.*, 2004; Bhatia & Myers, 2006; Lochan & Head-Gordon, 2006; Düren *et al.*, 2009; Han *et al.*, 2009). Such interactions are mostly dispersive and within the range of 3.4–4.0 kJ mol<sup>-1</sup> for a H<sub>2</sub> molecule interacting with the benzene ring of terephthalic acid (Lochan & Head-Gordon, 2006). The aforementioned binding enthalpy is remotely below the estimated and debated target for efficient H<sub>2</sub> storage materials, range of 15–20 kJ mol<sup>-1</sup> (room temperature at pressures up to 30 bar; Han *et al.*, 2009) or 21–32 kJ mol<sup>-1</sup> (–20°C and pressure range 1–100 bar; Lochan & Head-Gordon, 2006). Herein, we set to investigate if such interactions could be additive and hence can lead to enhanced interactions between a H<sub>2</sub> molecule and multiple aromatic rings in a tailored MOF adsorbent. As a test model, we envision a molecular square constructed of four benzene rings interacting simultaneously with a single H<sub>2</sub> molecule, residing in the center of the square, as a potential model for a material with enhanced H<sub>2</sub> binding affinity.

**Table 1**  
Selected geometric parameters (Å, °).

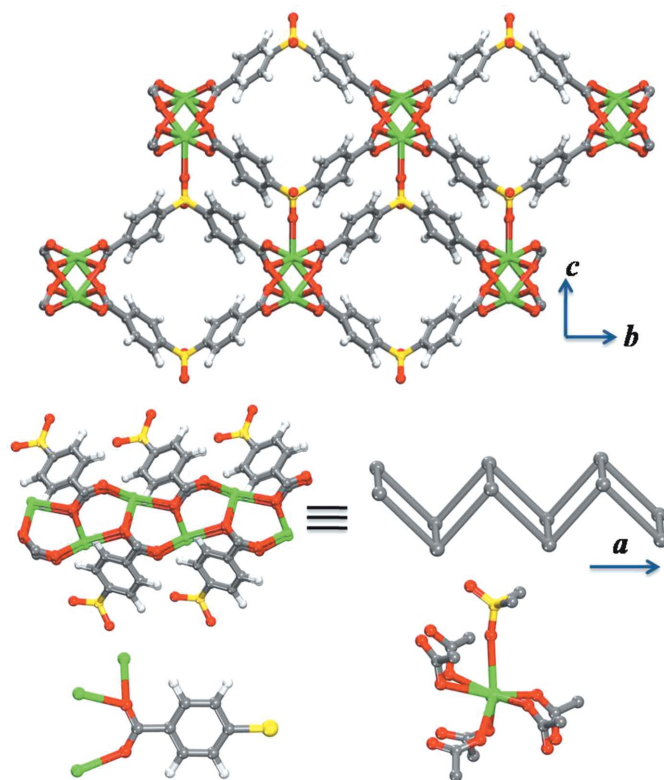
Pb1–O3 <sup>i</sup>	2.434 (6)	Pb1–O4 <sup>i</sup>	2.441 (6)
O3–Pb1–O4 <sup>i</sup>	72.0 (2)	O2–S3–C6 <sup>ii</sup>	108.1 (3)
O3 <sup>i</sup> –Pb1–O4 <sup>i</sup>	117.1 (2)	O1–S3–C6 <sup>ii</sup>	107.6 (4)

Symmetry codes: (i)  $x, -y + \frac{1}{2}, z$ ; (ii)  $x, -y - \frac{1}{2}, z$ .

## 2. Experimental

The solvothermal reaction of Pb(NO<sub>3</sub>)<sub>2</sub> and 4,4'-sulfonyldibenzoic acid in *N,N*-dimethylformamide (DMF) yields colorless crystals of **1** (Fig. 1). The as-synthesized compound was characterized and formulated by single-crystal X-ray diffraction studies as [Pb<sub>2</sub>(C<sub>14</sub>H<sub>8</sub>O<sub>6</sub>S)<sub>2</sub>·DMF (**1**). The phase purity of **1** was confirmed by similarities between its calculated and experimental powder X-ray diffraction patterns (PXRD, supporting information).

A sample of **1** was activated for sorption studies by solvent exchange in acetonitrile, where complete removal of the DMF guest molecules was confirmed by IR spectroscopy, see the supporting information. The activated sample was found to be stable up to 400°C as confirmed by TGA studies, see the supporting information.



**Figure 1**  
Crystal structure of **1** (top), the Pb–CO<sub>2</sub> rod-shaped infinite SBU (middle), and the coordination mode of the carboxylic linker and the Pb(II) ion (below). Pb (green), C (gray), S (yellow), O (red), H (white).

### 3. Results and discussion

The crystal structure of **1** revealed square-like channels running through the *a*-axis. The distinctive shape of the ditopic ligand molecule (dihedral Ph–SO<sub>2</sub>–Ph angle of 103.93°) complemented by the coordination sphere around Pb(II) (CO<sub>2</sub>–Pb–CO<sub>2</sub> dihedral angle of 78.73°) facilitated the construction of the MOF containing square-like channels (Table 1). In the crystal structure of **1**, infinite CO<sub>2</sub>–Pb(II) secondary building units (SBUs) (Rosi *et al.*, 2005) are observed and resulting from coordination of the carboxylate linkers in the bridging bis-bidentate mode to Pb(II) ions. Each carboxylate group is coordinated to three Pb(II) ions, enabling the formation of the CO<sub>2</sub>–Pb(II) infinite coordination chains, along the *a*-axis. Each Pb(II) ion is coordinated to six oxygen atoms from bridging carboxylate groups (O–Pb bond distances of 2.434–2.815 Å). Additionally, each Pb(II) ion is coordinated to an oxygen atom from the nearest sulfone group (O–Pb bond distance of 2.865 Å).

The resultant connectivity of the Pb ions by the organic linkers facilitated the construction of parallelogram, square-like, shaped channels running along the *a*-axis, held together in the *bc*-plane through the sulfone–Pb coordination. Such orthogonal bridging interactions resulted in square-like, guest-

accessible, channels in **1**. The surface area of **1** as probed by N<sub>2</sub> and Ar at 77 K and 87 K (supporting information) were estimated to be 165 and 169 m<sup>2</sup> g<sup>−1</sup>, respectively. The resulting square-like channels in **1** encompass a periodic array of aromatic rings with a relatively short interplanar distance between opposing rings (8.448 Å, centroid-to-centroid). Essentially, the periodically aligned aromatic rings delimiting the pore system dictate the pore aperture size and its maximum opening to be around ~4 Å (excluding the nearest van der Waals surfaces). These special structural features encountered in **1** (narrow one-dimensional channels aligned with a periodic array of aromatic rings) inspired us to explore and further investigate the potential effect(s) of the pore system (size, geometry and functionality) on the H<sub>2</sub> interactions with the aromatic walls. Indeed, the observed H<sub>2</sub> adsorption properties of the present MOF are remarkable and unique. Of special note are the observed H<sub>2</sub> adsorption isotherms with a sharp steepness (type I isotherm shape particularly at 77 K), which, to the best of our knowledge, are scarce for physical adsorbents. Equally interesting is the observed steady H<sub>2</sub> isosteric heat of adsorption (*Q*<sub>st</sub>) at 9 kJ mol<sup>−1</sup> (Fig. 2). In fact, the H<sub>2</sub> adsorption isotherms for **1** demonstrate rapid saturation at early dosing stages and nearly linear behaviour for heats of adsorption throughout the entire

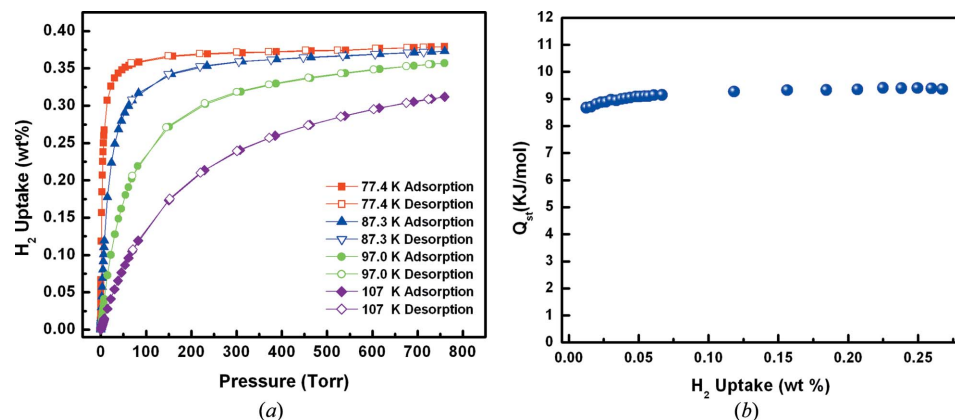


Figure 2 (a) Variable-temperature H<sub>2</sub> adsorption isotherms and (b) *Q*<sub>st</sub> of H<sub>2</sub> adsorption in **1**.

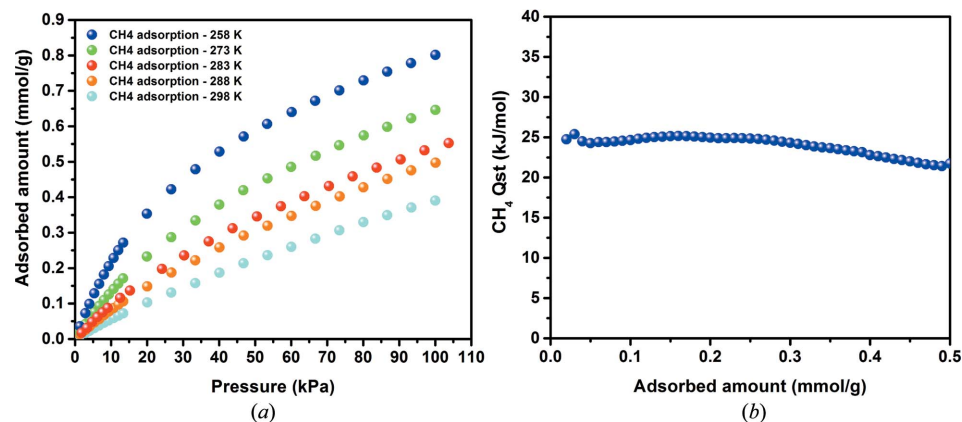


Figure 3 (a) Variable-temperature CH<sub>4</sub> adsorption isotherms and (b) *Q*<sub>st</sub> of CH<sub>4</sub> adsorption in **1**.

H<sub>2</sub> adsorption loading, two highly desirable features for H<sub>2</sub> storage applications. The sharp step in the H<sub>2</sub> adsorption isotherm can be translated to sorption sites saturation by H<sub>2</sub> molecules at moderate pressures. This behaviour could be attributed to the equal distribution of H<sub>2</sub> sorption sites with uniform binding affinities within the framework, most probably on the surfaces of aromatic rings present in **1**. Interestingly, such uniform interactions were also observed for methane adsorption in **1** (Fig. 3), with a relatively high and steady *Q*<sub>st</sub> of 25 kJ mol<sup>−1</sup> over the entire loading range, thus confirming the interesting structural aspects of **1** for enhanced gas–solid adsorbent interactions.

Furthermore, high-pressure gas adsorption studies conducted on **1** revealed interesting adsorption behaviour of selected gases (Fig. 4). The recorded uptakes for O<sub>2</sub> and N<sub>2</sub> in **1** were comparable. This is in contrast to commonly observed preferential N<sub>2</sub> uptake, compared with O<sub>2</sub>, in numerous examples of zeolites (Talu *et al.*, 1996; Hutson *et al.*, 1999; Agha *et al.*, 2005). In the case of most MOFs and zeolites,

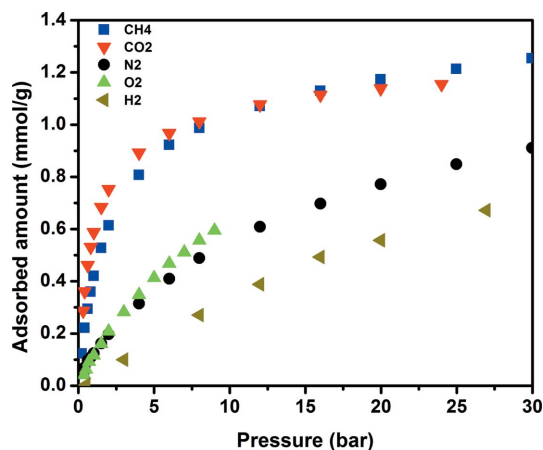


Figure 4  
High-pressure sorption isotherms for different gases at 298 K in **1**.

preferential  $\text{N}_2$  sorption is attributed to stronger interactions between  $\text{N}_2$  molecules and the material surface due to a larger quadrupole moment of  $\text{N}_2$ . The interaction/adsorption sites in **1** are dominated by the periodic array of aromatic rings aligned in the one-dimensional channels, thus explaining the weak quadrupole interactions between the gas molecules and the framework and the subsequent comparable  $\text{N}_2$  and  $\text{O}_2$  uptakes. Similarly, the nature of the gas/framework interactions is also reflected in the observed comparable uptake of  $\text{CH}_4$  and  $\text{CO}_2$  (two distinct molecules with and without quadrupole moment) in **1** (Fig. 4), emphasizing the dominance of dispersive interactions between adsorbed gas molecules and the aromatic walls of the MOF porous material. It is noteworthy that porous materials having similar uptakes for  $\text{CO}_2$  and  $\text{CH}_4$  in a wide range of pressure is uncommon behaviour in adsorption on porous materials (Nugent *et al.*, 2013). Additionally, the adsorption of  $\text{C}_2\text{H}_6$  and  $\text{C}_3+$  on **1** (supporting information) showed to some extent the same behaviour, uncommon for most MOF materials. The aforementioned results demonstrate that molecules with various degrees of high polarizabilities probe the surface of **1** in the same way. Specifically, the interaction potential of **1** with different molecules, having different chemical–physical properties like  $\text{CO}_2$ ,  $\text{CH}_4$ ,  $\text{C}_2\text{H}_6$  and  $\text{C}_3+$ , is governed mainly by dispersive (non-electrostatic) interactions.

#### 4. Conclusions

In conclusion, we present an unprecedented experimental illustration of uptake–energetics relationships for  $\text{H}_2$  and  $\text{CH}_4$  adsorption in a novel MOF with narrow one-dimensional channels aligned with a periodic array of aromatic rings. The newly synthesized material represents a model material for pinpointing the interplay between porosity/storage capacity and the strength of host–guest interactions. This approach supports the impact of narrow channels with a periodic array of aromatic rings, controlling the access to prospective larger pores within a targeted porous material, for the effective gas adsorption. The present results pave the way to additional

experimental and theoretical investigations in order to further assess the extent of additive dispersive interactions in enhancing  $\text{H}_2$  and  $\text{CH}_4$  binding affinity in gas storage materials, in general, and MOFs, in particular. Noticeably, the distinct sharp step at relatively low pressures in the  $\text{H}_2$  adsorption isotherm at 77 K and the relatively high  $Q_{\text{st}}$  for  $\text{CH}_4$  storage was achieved despite the detrimental reduction in the MOF overall porosity. Conceivably, from the present study, the optimal combination of narrow pores/windows ( $< 1$  nm diameter) with a suitable and uniform charge density in the pores (coordinatively unsaturated metal sites and polar functional groups) operating synergistically could play a significant role in promoting the storage of  $\text{H}_2$  and  $\text{CH}_4$  at moderate pressures and ambient temperatures.

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#### References

- Agha, R. K., De Weireld, G. & Frère, M. (2005). *Adsorption* **11**, 179–182.
- Alezi, D., Belmabkhout, Y., Suyetin, M., Bhatt, P. M., Weseliński, L. J., Solovyeva, V., Adil, K., Spanopoulos, I., Trikalitis, P. N., Emwas, A. H. & Eddaoudi, M. (2015). *J. Am. Chem. Soc.* **137**, 13308–13318.
- Alkordi, M. H., Liu, Y., Larsen, R. W., Eubank, J. F. & Eddaoudi, M. (2008). *J. Am. Chem. Soc.* **130**, 12639–12641.
- Banerjee, R., Phan, A., Wang, B., Knobler, C., Furukawa, H., O’Keeffe, M. & Yaghi, O. M. (2008). *Science*, **319**, 939–943.
- Belmabkhout, Y., Mouttaki, H., Eubank, J. F., Guillerm, V. & Eddaoudi, M. (2014). *RSC Adv.* **4**, 63855–63859.
- Belof, J. L., Stern, A. C., Eddaoudi, M. & Space, B. (2007). *J. Am. Chem. Soc.* **129**, 15202–15210.
- Bhatia, S. K. & Myers, A. L. (2006). *Langmuir*, **22**, 1688–1700.
- Cairns, A. J., Eckert, J., Wojtas, L., Thommes, M., Wallacher, D., Georgiev, P. A., Forster, P. M., Belmabkhout, Y., Ollivier, J. & Eddaoudi, M. (2016). *Chem. Mater.* **28**, 7353–7361.
- Chen, B., Zhao, X., Putkham, A., Hong, K., Lobkovsky, E. B., Hurtado, E. J., Fletcher, A. J. & Thomas, K. M. (2008). *J. Am. Chem. Soc.* **130**, 6411–6423.
- Cho, S. H., Gadzikwa, T., Afshari, M., Nguyen, S. T. & Hupp, J. T. (2007). *Eur. J. Inorg. Chem.* **2007**, 4863–4867.
- Collins, D. J. & Zhou, H.-C. (2007). *J. Mater. Chem.* **17**, 3154–3160.
- Dincă, M., Han, W. S., Liu, Y., Dailly, A., Brown, C. M. & Long, J. R. (2007). *Angew. Chem. Int. Ed.* **46**, 1419–1422.
- Dincă, M. & Long, J. R. (2008). *Angew. Chem. Int. Ed.* **47**, 6766–6779.
- Düren, T., Bae, Y.-S. & Snurr, R. Q. (2009). *Chem. Soc. Rev.* **38**, 1237–1247.
- Eddaoudi, M., Moler, D. B., Li, H., Chen, B., Reineke, T. M., O’Keeffe, M. & Yaghi, O. M. (2001). *Acc. Chem. Res.* **34**, 319–330.
- Férey, G., Mellot-Draznieks, C., Serre, C. & Millange, F. (2005). *Acc. Chem. Res.* **38**, 217–225.
- Han, S. S., Mendoza-Cortés, J. L. & Goddard, W. A. III (2009). *Chem. Soc. Rev.* **38**, 1460–1476.
- Hayashi, H., Côté, A. P., Furukawa, H., O’Keeffe, M. & Yaghi, O. M. (2007). *Nat. Mater.* **6**, 501–506.
- Hirscher, M. & Panella, B. (2007). *Scr. Mater.* **56**, 809–812.
- Horcajada, P., Serre, C., Vallet-Regí, M., Sebba, M., Taulelle, F. & Férey, G. (2006). *Angew. Chem.* **118**, 6120–6124.
- Hübner, O., Glöss, A., Fichtner, M. & Klöpper, W. (2004). *J. Phys. Chem. A*, **108**, 3019–3023.



- Hutson, N. D., Rege, S. U. & Yang, R. T. (1999). *AIChE J.* **45**, 724–734.
- Kitagawa, S., Kitaura, R. & Noro, S. (2004). *Angew. Chem. Int. Ed.* **43**, 2334–2375.
- Lin, X., Jia, J., Hubberstey, P., Schröder, M. & Champness, N. R. (2007). *CrystEngComm*, **9**, 438–448.
- Liu, Y., Eubank, J. F., Cairns, A. J., Eckert, J., Kravtsov, V. C., Luebke, R. & Eddaoudi, M. (2007). *Angew. Chem. Int. Ed.* **46**, 3278–3283.
- Llewellyn, P. L., Bourrelly, S., Serre, C., Vimont, A., Daturi, M., Hamon, L., De Weireld, G., Chang, J.-S., Hong, D.-Y., Kyu Hwang, Y., Hwa Jhung, S. & Férey, G. (2008). *Langmuir*, **24**, 7245–7250.
- Lochan, R. C. & Head-Gordon, M. (2006). *Phys. Chem. Chem. Phys.* **8**, 1357–1370.
- Moulton, B. & Zaworotko, M. J. (2001). *Chem. Rev.* **101**, 1629–1658.
- Murray, L. J., Dincă, M. & Long, J. R. (2009). *Chem. Soc. Rev.* **38**, 1294–1314.
- Nouar, F., Eubank, J. F., Bousquet, T., Wojtas, L., Zaworotko, M. J. & Eddaoudi, M. (2008). *J. Am. Chem. Soc.* **130**, 1833–1835.
- Nugent, P., Belmabkhout, Y., Burd, S. D., Cairns, A. J., Luebke, R., Forrest, K., Pham, T., Ma, S., Space, B., Wojtas, L., Eddaoudi, M. & Zaworotko, M. J. (2013). *Nature*, **495**, 80–84.
- Radha Kishan, M., Tian, J. K., Thallapally, P. K., Fernandez, C. A., Dalgarno, S. J., Warren, J. E., McGrail, B. P. & Atwood, J. L. (2010). *Chem. Commun.* **46**, 538–540.
- Rosi, N. L., Eckert, J., Eddaoudi, M., Vodak, D. T., Kim, J., O’Keeffe, M. & Yaghi, O. M. (2003). *Science*, **300**, 1127–1129.
- Rosi, N. L., Kim, J., Eddaoudi, M., Chen, B., O’Keeffe, M. & Yaghi, O. M. (2005). *J. Am. Chem. Soc.* **127**, 1504–1518.
- Rowell, J. L. & Yaghi, O. M. (2005). *Angew. Chem. Int. Ed.* **44**, 4670–4679.
- Rowell, J. L. & Yaghi, O. M. (2006). *J. Am. Chem. Soc.* **128**, 1304–1315.
- Shekhah, O., Belmabkhout, Y., Chen, Z., Guillerm, V., Cairns, A., Adil, K. & Eddaoudi, M. (2014). *Nat. Commun* **5**, 4428.
- Talu, O., Li, J., Kumar, R., Mathias, P. M., Moyer, J. D. & Schork, J. M. (1996). *Gas Sep. Purif.* **10**, 149–159.
- Zhou, H.-C., Long, J. R. & Yaghi, O. M. (2012). *Chem. Rev.* **112**, 673–674.