

Headgroup-Ordered Monolayers of Uncharged Glycolipids Exhibit Selective Interactions with Ions

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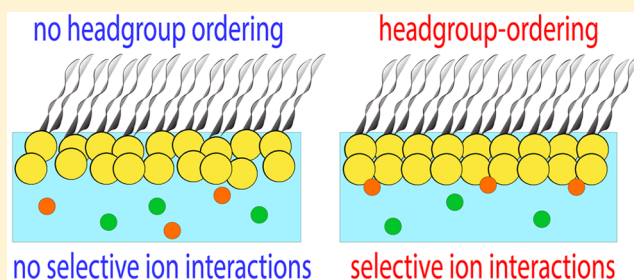
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Supporting Information

ABSTRACT: Selective interactions of ions with charge-neutral saccharides can have far-reaching consequences in biological and wet-technological contexts but have so far been observed only indirectly. Here, we directly quantify by total-reflection X-ray fluorescence the preferential accumulation of ions near uncharged saccharide surfaces in the form of glycolipid Langmuir monolayers at air/water interfaces exhibiting different levels of structural ordering. Selective interactions with ions from the aqueous subphase are observed for monolayers featuring crystalline ordering of the saccharide headgroups, as determined by grazing-incidence X-ray diffraction. The attracted ion species depend on the structural motifs displayed by the ordered saccharide layer. Our results may constitute a basis to understand the salt-specific swelling of wood materials and various phenomena in membrane biophysics.



Glycolipids are essential constituents of biological membranes. Due to their structural diversity and strategic localization in various functional membrane systems, the glycolipids are considered the third alphabet of life, the sugar code.¹ Glycolipids exhibit highly specific interactions with other saccharides² and proteins.³ Interestingly, ions are believed to promote the specific interaction of saccharide headgroups even if the latter are neither charged nor of zwitterionic character. For example, the strength of the homotypic interaction between lipid-anchored LewisX trisaccharides was reported to increase in the presence of calcium.^{4,5} Additional indirect evidence of selective ion interactions with saccharides is the observation of ion-specificity in the swelling of wood materials in salt solutions.⁶

Despite its great relevance in biology as well as for the pharmaceutical, food, cosmetic, and paper industries,^{7,8} our knowledge about the interaction of salts with neutral saccharides has remained limited and is based on indirect observations. In the present work, preferential interactions of ions with uncharged saccharide surfaces in the form of glycolipid Langmuir monolayers at air/water interfaces are directly quantified. Depending on the choice of the glycolipids in terms of headgroup chemistry and alkyl chain saturation, saccharide surfaces of various characteristics are realized. To this end, the highly abundant glycolipids mono- and digalactosyldiacylglycerol⁹ are investigated in their chain-saturated and natural unsaturated forms (*MGDG-sat*, *MGDG-unsat*, *DGDG-sat*, and *DGDG-unsat*; see Figure 1A,F). In

addition, a chain-saturated glycolipid with a lactose headgroup (*N*-palmitoyl-lactosylceramide, *LacCer-sat*, Figure 2A), a chain-saturated glycolipid with a trihexose headgroup (*Trihexo-sat*, Figure S4), and a glucosylated sterol (β -D-glucosyl sitosterol, *Glu-sitosterol*, Figure 2F) are studied. Grazing-incidence X-ray diffraction (GIXD) reveals the structural ordering of the monolayers down to an Angstrom level. Complementary total-reflection X-ray fluorescence (TRXF) measurements enable quantification of the monolayers' preferential interactions with ions. The aqueous subphases contain $c_0 = 1$ mM KI, CsBr, or CaBr₂. These salts have been chosen because of their pronounced effects on wood swelling⁶ and their good detectability by TRXF.

The monolayer of *MGDG-sat* (Figure 1A) is defined by three diffraction peaks in the wide-angle region (at high Q_{xy}) above the horizon ($Q_z > 0$), characterizing an oblique lattice structure of tilted chains ($t = 31^\circ$, Figure 1B–D and Supporting Information tables), as often encountered for optically active compounds.^{10,11} Neither lateral compression to higher surface pressure Π nor the type of subphase significantly change the lattice structure (Figures 3 and S6). The small cross-sectional chain area of only $A_0 = 18.6 \text{ \AA}^2$ indicates very tight packing with no rotational freedom. The in-plane

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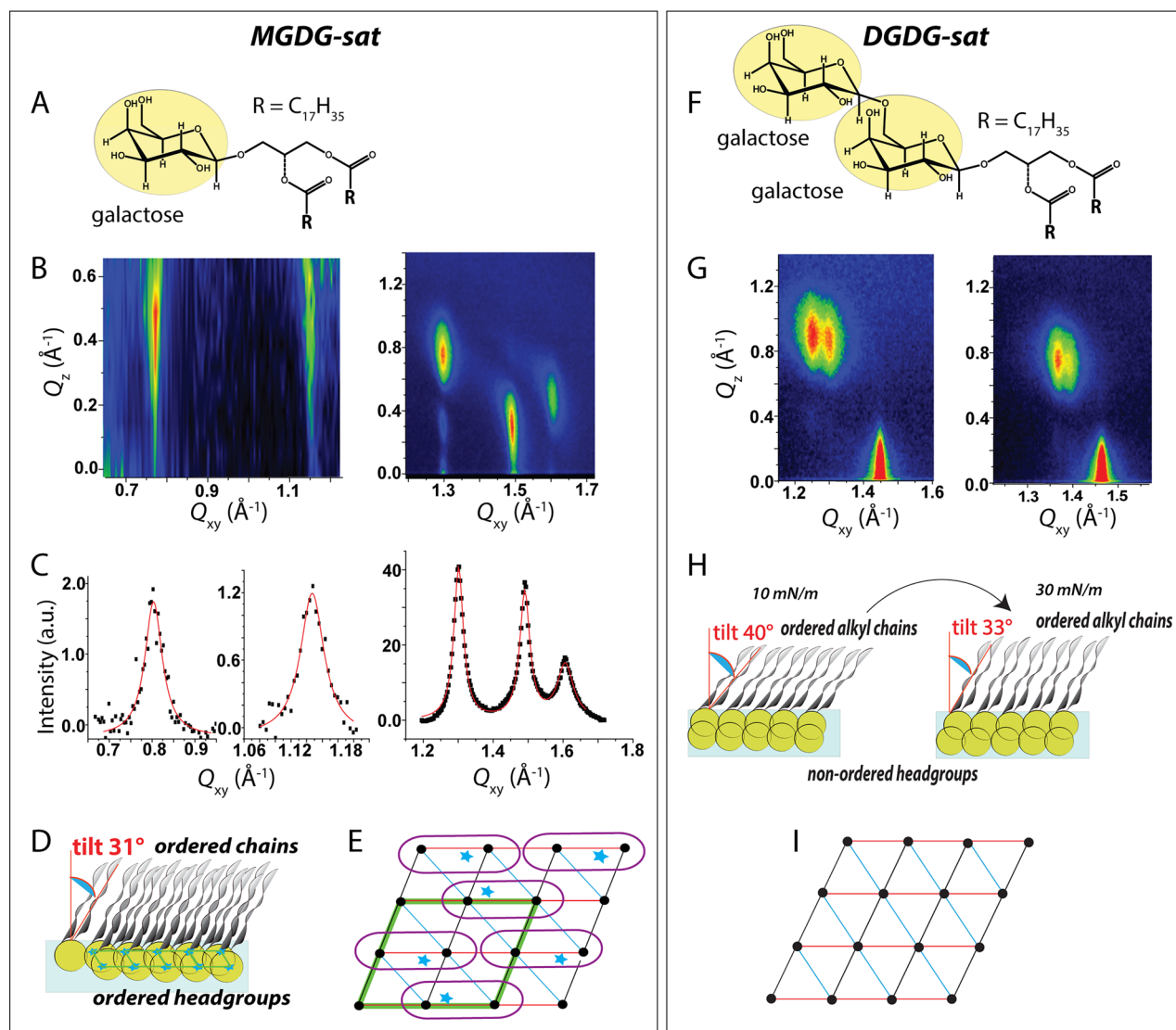


Figure 1. (A,F) Chemical structures of *MGDG-sat* and *DGDG-sat*. (B,G) GIXD contour plots displaying the scattered intensity versus the in-plane and the out-of-plane components of the scattering vector, Q_{xy} and Q_z , respectively, obtained for *MGDG-sat* ($\Pi = 30$ mN/m) and *DGDG-sat* ($\Pi = 10$ and 30 mN/m) monolayers on 1 mM CsBr. (C) Bragg peaks obtained for *MGDG-sat*. (D,H) Schematic side-view of *MGDG-sat* and *DGDG-sat* monolayers, respectively, on the water surface. (E) Schematic top-view representation of the lattice formed by *MGDG-sat*. The positions of chains and headgroups are indicated with black dots and blue stars, respectively. Red, black, and blue triangles indicate the repeating unit cell of the alkyl chains. The unit cell of the molecule lattice is indicated with a green parallelogram. Violet line: delimitation of the molecules. (I) Schematic top-view representation of the lattice formed by the alkyl chains of *DGDG-sat*.

molecular area, $A = 2A_0/\cos(t) = 43.4 \text{ \AA}^2$, allows the galactose moiety of *MGDG-sat* to orient parallel to the interface (the in-plane area of the sugar headgroup is reported to be 35.4 \AA^2),^{12,13} in agreement with previously reported data.¹⁴ Interestingly, two additional Bragg peaks are seen in the midangle region, i.e., at lower Q_{xy} (Figure 1B,C). These peaks indicate an ordering of weakly hydrated galactose moieties,¹⁵ in good agreement with previous SAXS and WAXS data.^{16,17} A supercell indicating the ordering of entire molecules is identified (Figure 1E and Table S3). It is induced by strong intermolecular hydrogen bonds between the sugar headgroups, similar to the previously reported monolayer structure of a GPI fragment,^{18–20} although the existence of a superlattice in principle does not require crystalline ordering of the entire headgroups. This supercell (green parallelogram), reminiscent of subgel phases in bulk,^{21–23} is commensurate with the hydrocarbon chain lattice ($a' = 2 \times a_{\text{chains}}$, $b' = 2 \times b_{\text{chains}}$, $\gamma =$

110.1°) and, with an area of 86.8 \AA^2 , contains two *MGDG-sat* molecules. The rigid network of hydrogen bonds between galactose headgroups dictates the packing order of the chains (no change upon lateral compression). The full width at half-maximum (fwhm) of the Bragg rods (SI) agrees well with the length of an extended C18 alkyl chain in all-trans conformation, confirming that the interfacial layer is a monolayer at all investigated surface pressures.^{24,25}

MGDG-unsat does not form ordered monolayers. Gel phases in 3D systems have been found only at extremely low temperatures ($-30 \text{ }^\circ\text{C}$).²⁶ Obviously, the highly ordered structure in *MGDG-sat* monolayers is a synergistic result of concomitant headgroup and chain interactions.

DGDG-sat does not exhibit diffraction peaks in the mid-to-wide-angle region, indicating the absence of headgroup order. Only the three diffraction peaks defining the alkyl chain lattice are observed in the wide-angle region (Figure 1G). The bulky

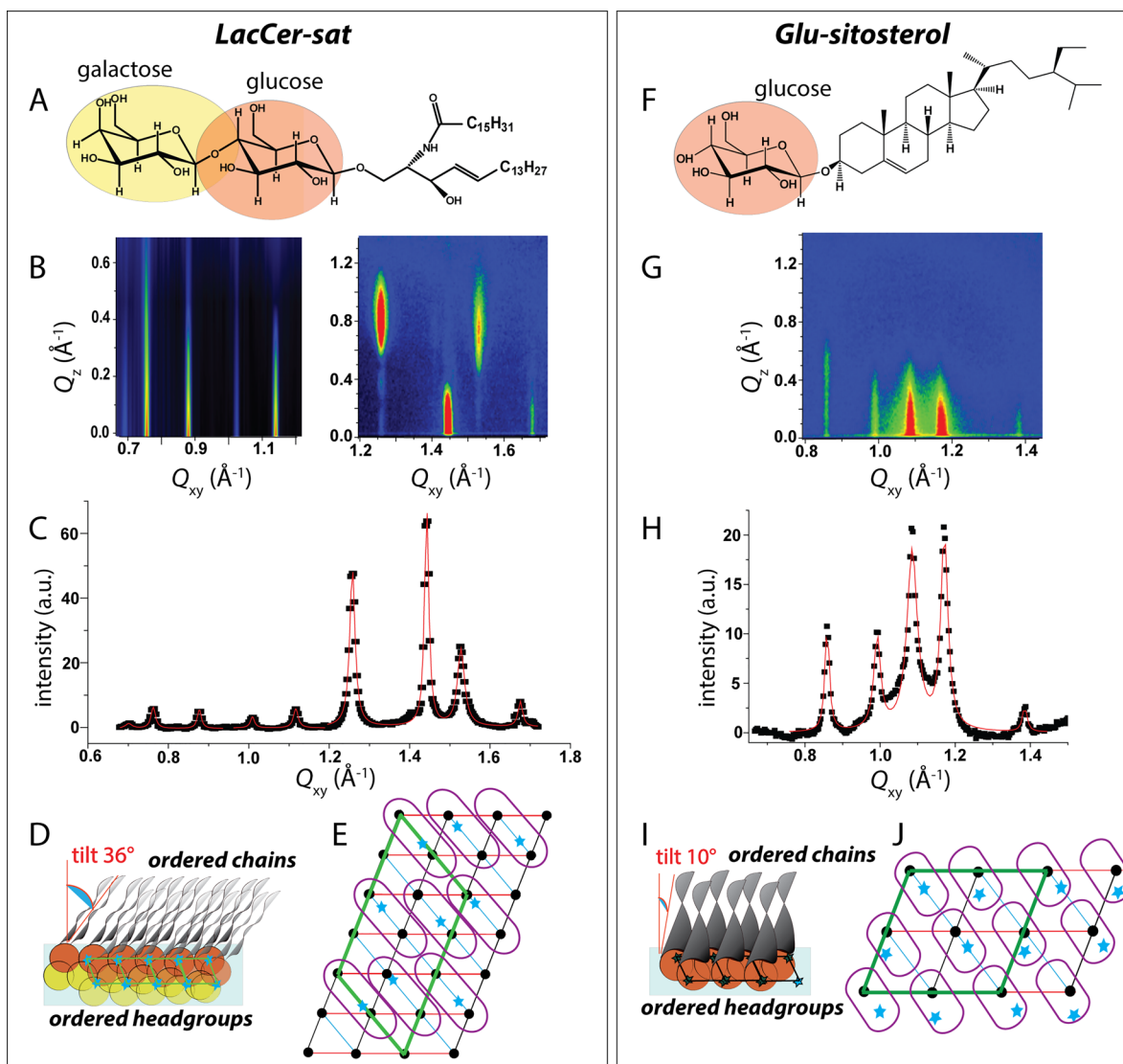


Figure 2. (A,F) Chemical structures of *LacCer-sat* and *Glu-sitosterol*. (B,G) GIXD contour plots obtained for *LacCer-sat* (10 mN/m, 1 mM CsBr) and *Glu-sitosterol* (30 mN/m, 1 mM KI) monolayers, respectively. (C,H) Bragg peaks obtained for *LacCer-sat* (10 mN/m) and *Glu-sitosterol* (30 mN/m) monolayers, respectively. (D,I) Schematic side-view of *LacCer-sat* and *Glu-sitosterol* monolayers, respectively. The positions of chains and headgroups are indicated with black dots and blue stars, respectively. Red, black, and blue triangles indicate the repeating unit cell of the alkyl chains. The repeating unit cell of the headgroup/chain superlattice is indicated with a green parallelogram. Violet line: delimitation of the molecules.

digalactose moiety of *DGDG-sat* seems to disturb the packing of the headgroups, presumably due to a higher hydration degree, and offers higher flexibility to the molecules (noticeable decrease of the tilt angle during compression). The larger area requirement mismatch between the headgroup and chains leads to a higher tilt angle of the chains (Figure 1H and Tables S7 and S9). The lack of strong H-bonded and structured headgroups of *DGDG-sat* is in agreement with previous reports showing that a disaccharide sugar headgroup in glycerol-based glycolipids dramatically lowers the phase transition temperature compared to the corresponding molecule with a monosaccharide headgroup.^{15,27,28} Because the *DGDG-sat* in-plane molecular area is only slightly larger than that of *MGDG-sat*, the two sugar moieties cannot arrange parallel to the interface but more likely adopt a perpendicular or tilted arrangement. This is in agreement with previous infrared reflection–absorption spectroscopy (IRRAS) studies reporting a tilt angle of the digalactosyl headgroups of $\approx 40^\circ$

(ref 14). GIXD data recorded for monolayers of *DGDG-unsat* display no diffraction peaks at all, indicating disordered chains and headgroups.

To tackle the problem of sugar specificity for headgroup interactions, *LacCer-sat* is investigated. As *DGDG-sat*, *LacCer-sat* features two sugar moieties but of a different nature (galactose and glucose units) and forms condensed monolayers at room temperature on the surface of 1 mM aqueous salt solutions. Yet, the GIXD data revealed a much more complex diffraction pattern with a multitude of peaks (Figure 2B,C). Contrary to *MGDG* and *DGDG*, a stronger influence of the subphase (Figure S6) and the formation of different polymorphs are observed. On a 1 mM CsBr subphase, the three intense diffraction peaks in the wide-angle region can be attributed to the alkyl chain order, while the five additional weaker peaks in the mid-to-wide-angle region indicate headgroup ordering. The existence of a headgroup order in monolayers²⁹ and in bulk crystals³⁰ of synthetic glycolipids

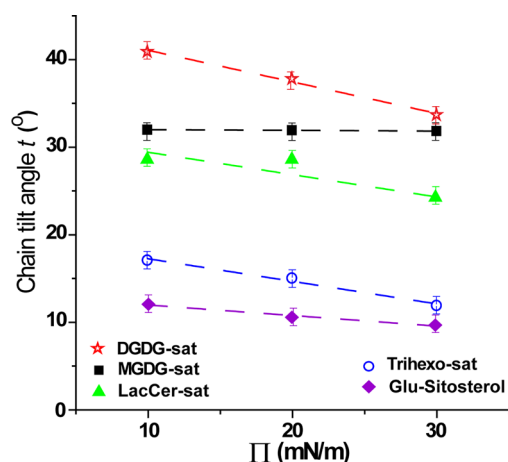


Figure 3. Variation of the tilt angle with the lateral pressure of the structured glycolipid monolayers on 1 mM CaBr_2 .

bearing lactose units has already been reported. Thus, it seems that lactose headgroups are more prone to be engaged in intermolecular interactions³¹ and layer structuring than the digalactosyl headgroups. The supercell (Figure 2E, green parallelogram) defines an area corresponding to four *LacCer-sat* molecules (Tables S10–S12). The H-bonding for *LacCer-sat* seems to be more complex due to the possible additional contribution of amide–sugar interactions.^{32,33} Different polymorphs (Tables S13–S16 and Figure S3) have been observed but will not be discussed here in detail. Literature agrees well with the tendency of *LacCer-sat* in forming different polymorphs in 3D systems.^{34,35} The Debye–Scherrer rings, seen only in the diffraction pattern at high surface pressures (Π

= 30 mN/m, SI Figure S3), indicate the formation of 3D crystals coexisting with a monolayer at the air/water interface. Such strong lactose–lactose interactions could be responsible for the formation of *LacCer*-enriched microdomains in biological systems (cell surface plasma membranes of mouse neutrophils, microdomains presenting a high specificity for antibodies).³⁶

Monolayers of *Trihexo-sat*, which features the bulkiest headgroup, are found to be characterized by ordered alkyl chains and nonordered headgroups. Interestingly, this compound exhibits the lowest alkyl chain tilt, $t \approx 20^\circ$ at 10 mN/m, which further decreases to $t \lesssim 15^\circ$ at 30 mN/m, depending on the subphase (Tables S17–22 and Figure S6).

The role played by the glucose unit in headgroup interactions is investigated in *Glu-sitosterol* monolayers (monoglucose-based glycolipids with saturated alkyl chains are not commercially available). The GIXD data reveal a monolayer structure defined by five Bragg peaks (Figure 2G,H). Three of them describe the order of the cholesterol moieties, and the additional Bragg peaks indicate headgroup order. The observed superstructure with an area of 156.3 \AA^2 corresponds to four *Glu-sitosterol* molecules (Table S25). The cross-sectional area per molecule of 38.4 \AA^2 is in good agreement with values of 37.7 \AA^2 obtained for pure cholesterol.²⁵ This in-plane area allows the sugar headgroup to adopt a parallel orientation to the interface. Complex GIXD patterns have been often described in the literature and even for Langmuir monolayers of cholesterol, attributed to the formation of multilayers.²⁵ In the present case, we refute such a scenario based on stable compression isotherms (Figure S5) and on the fwhm value of the Bragg rods (Table S23) corresponding to a monolayer. Similar to *LacCer-sat*, the *Glu-*

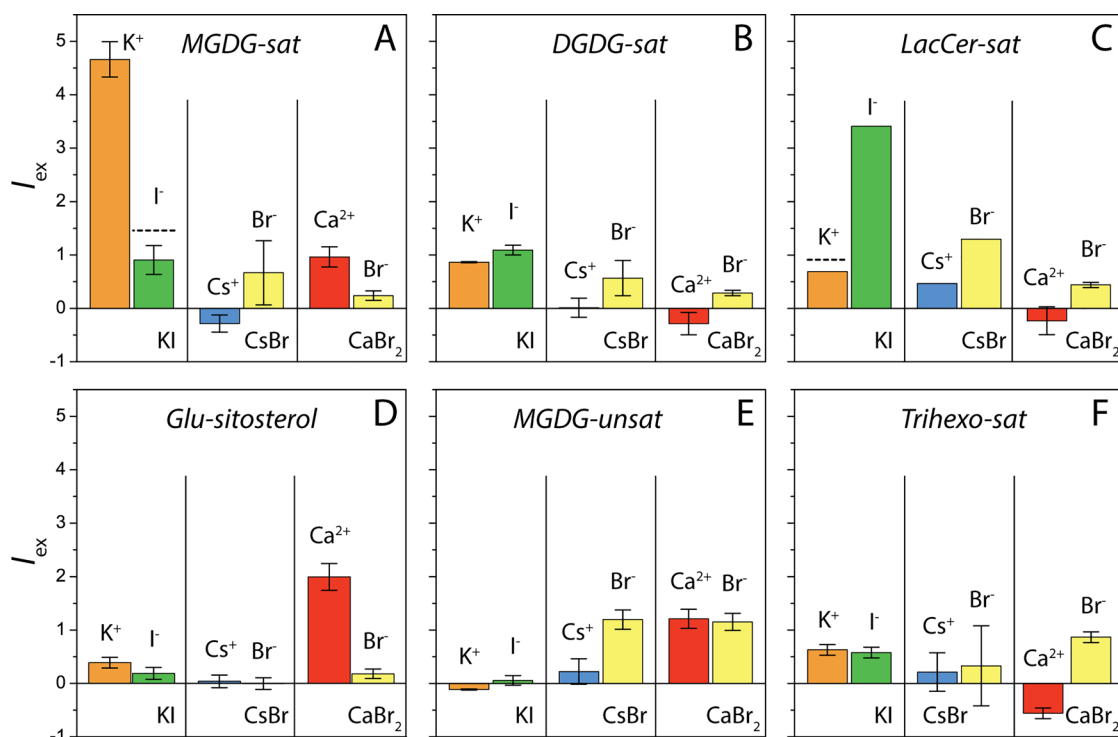


Figure 4. Relative excess fluorescence intensities of K^+ , I^- , Cs^+ , Br^- , and Ca^{2+} near various uncharged glycolipid monolayers on aqueous subphases containing 1 mM KI, CsBr, or CaBr_2 . Data are averages over up to three measurements at $\Pi = 10, 20,$ and 30 mN/m. Error bars represent one standard deviation. They are absent when only one data point was available. Horizontal dashed lines in panels (A) and (C) indicate the intensity level expected from a generic charge-neutralization effect (see the main text).

sitosterol monolayers exhibit stronger response to the subphase nature (Figure S6). Polymorphs of different thermodynamic stability, possibly induced by headgroup hydration/dehydration,³⁷ are observed but again will not be discussed here in detail.

Overall, the investigated electrically neutral glycolipid monolayers can be divided into three classes: (i) highly structured monolayers, characterized by ordered headgroups and ordered alkyl chains (*MGDG-sat*, *LacCer-sat*, and *Glu-sitosterol*), (ii) structured monolayers with ordered alkyl chains but disordered headgroups (*DGDG-sat*, *Trihexo-sat*), and (iii) nonstructured monolayers with no headgroup and no chain order (*MGDG-unsat* and *DGDG-unsat*). Classes (i) and (ii) are distinguishable also by the variation of the tilt angle of the alkyl chains with the surface pressure (Figure 3). The highly structured monolayers in tendency exhibit less variation of the tilt angle due to the rigid headgroup lattice. Those are considered to have a low degree of hydration and to strongly interact, forming a network of hydrogen bonds.³⁷ A more pronounced gradual decrease of the chain tilt angle with increasing lateral pressure is recorded for the structured monolayers characterized by only ordered alkyl chains. This behavior points to a higher flexibility of the monolayer and of the molecular interactions associated with a higher degree of headgroup hydration.³⁸

TRXF experiments were carried out in order to quantify preferential interactions of ions with the glycolipid monolayers in terms of interfacial ion excesses per unit area. The method is highly sensitive to interfacial excesses because the X-ray beam is totally reflected and only illuminates the immediate vicinity of the interface with an evanescent wave.^{39–42} The excess of each ion type is then deduced from the intensity of its element-characteristic X-ray fluorescence. Importantly, in the presence of the 1 mM salt used in the TRXF experiments, the Debye length ($\kappa^{-1} \approx 10$ nm for KI and CsBr, $\kappa^{-1} = 6$ nm for CaBr₂) is comparable to the intensity decay length of the evanescence wave ($\Lambda \approx 7$ nm) for the used combination of incident angle ($\theta_i = 0.11$ or 0.07°) and beam energy ($E = 8.0$ or 15.0 keV). The fluorescence intensity of an ion species preferentially interacting with the surface is therefore higher than that of the corresponding counterion species, which approximately obeys a Gouy–Chapman distribution to achieve charge neutrality on the length scale of κ^{-1} (ref 43) but does not reach a 1:1 ion stoichiometry within the illuminated volume. As a consequence, the measurements allow identification of the preferentially interacting ion species for each salt type.

Figure 4 shows the relative excess fluorescence intensities, I_{ex} , of K⁺, I[−], Cs⁺, Br[−], and Ca²⁺ near monolayers of *MGDG-sat*, *DGDG-sat*, *LacCer-sat*, *Glu-sitosterol*, *MGDG-unsat*, and *Trihexo-sat* on aqueous subphases containing 1 mM KI, CsBr, or CaBr₂. Data are averages over up to three measurements at $\Pi = 10, 20,$ and 30 mN/m. $I_{\text{ex}} = (I - I_0)/I_0$ is the relative change in the measured intensity I with respect to the intensity I_0 expected in the absence of any preferential interactions, i.e., assuming bulk-like ion concentration up to the monolayer surface. I_0 , in turn, is obtained by measuring the intensity from the bare aqueous subphase, I_{bare} , and taking into account the reduction of the illuminated aqueous volume in the presence of the monolayer in the form of a prefactor f , $I_0 = f \cdot I_{\text{bare}}$. This prefactor is given by the electron density profile of the monolayer and robust with respect to minor uncertainties in the characteristics of these profiles (SI). The ion distributions at the bare air/water interface are approximated as constant,

neglecting minor deviations^{39,44} affecting I_{bare} only by few percent.³⁹ As shown in the SI, the employed methodology is consistent with an absolute intensity calibration using charged monolayers.

Positive values of I_{ex} in Figure 4 correspond to an accumulation of the respective ion species at the respective monolayer, and negative values correspond to a depletion. It is seen that significant ion accumulation occurs only for certain monolayer/salt combinations. Strong accumulation is observed for K⁺ near monolayers of *MGDG-sat*, which implies preferential interactions of K⁺ with the monolayer surface. The moderate excess of the counterion I[−], on the other hand, must be interpreted as a secondary effect of charge neutralization. The horizontal dashed line in Figure 4A indicates the intensity level expected from this secondary effect within a Poisson–Boltzmann model described in the SI. The measured I[−] excess is somewhat below this estimate, suggesting even slightly unfavorable interactions of I[−] with the interface. The magnitude of the K⁺ excess, as deduced from I_{ex} within the model, is $\Gamma_{\text{K}} \approx 0.02$ nm^{−2}, corresponding to an area per adsorbed ion of $A_{\text{K}} = 1/\Gamma_{\text{K}} \approx 50$ nm² or 1 ion per approximately 100 lipids. This excess is roughly 2 orders of magnitude smaller than previously measured ion excesses compensating a certain number of charges per lipid in charged lipid monolayers.^{41,42} Nonetheless, the preferential interaction of K⁺ with the charge-neutral monolayer is significant. At this point, it should be noted that neither charged impurities nor monolayer ionization by X-ray irradiation can be the cause for the observed accumulation because other cations including the divalent Ca²⁺ ions do not accumulate significantly at the same surface under the same conditions (Figure 4A). The same reasoning holds for other monolayer/salt combinations. A similar extent of ion accumulation is found for KI near monolayers of *LacCer-sat* (Figure 4C). Interestingly, the situation is reversed: I[−] instead of K⁺ ions exhibit preferential interactions with this interface, with an ion excess of $\Gamma_{\text{I}} \approx 0.014$ nm^{−2}. The horizontal dashed line again indicates the intensity level expected from a generic charge-neutralization effect. Pronounced ion accumulation is also found for Ca²⁺ near *Glu-sitosterol* monolayers (Figure 4D), with $\Gamma_{\text{Ca}} \approx 0.012$ nm^{−2}. The depth of the ion-adsorbing potential, ΔG , can be estimated from Γ via Boltzmann inversion, $\Delta G = -k_{\text{B}}T \ln(\Gamma/c_0d)$, where k_{B} is the Boltzmann constant and d the width of the ion-adsorbing region. For a reasonable d range ($2 \text{ \AA} < d < 7 \text{ \AA}$), we obtain $\Delta G = -11.2 \pm 2$ kJ/mol (K⁺ at *MGDG-sat*), $\Delta G = -10.3 \pm 2$ kJ/mol (I[−] at *LacCer-sat*), and $\Delta G = -9.7 \pm 2$ kJ/mol (Ca²⁺ at *Glu-sitosterol*). These numbers roughly correspond to 1/3 of the free energy per hydrogen bond in water.⁴⁵

All monolayers exhibiting pronounced preferential interactions with at least one ion species belong to monolayer class (i) featuring headgroup order. On the other hand, monolayers of classes (ii) or (iii), without headgroup order, do not seem to exhibit any clear trends. This notion suggests that the defined structural motifs displayed by headgroup-ordered surfaces are responsible for the pronounced ion selectivity. It further provides a possible explanation for the observation that different headgroup chemistries, which lead to different structural motifs, exhibit selectivity for different ions. The selectivity likely arises due to a match between the hydrogen bond configurations inside of the defined saccharide “pockets” and those of the ions’ hydration shells, which are known to be species-dependent. With that, the observed phenomenon appears to be related to the remarkable ion selectivity of

crown ethers owing to their characteristic polyether cavity.⁴⁶ Indeed, crown ethers such as dibenzo-18-crown-6 (18C6), an 18-atom heterocycle containing 6 oxygen atoms, selectively capture K^+ , while other polyether rings like 15-crown-5 (15C5) or 21-crown-7 (21C7) selectively capture Na^+ (Ca^{2+}) and Cs^+ ions, respectively.^{47,48} Our results demonstrate impressively that uncharged hydrophilic surfaces in general and saccharide surfaces in particular can selectively attract ions, the species being dependent on the structural motifs displayed by the surfaces. At first glance, the determined ion excesses appear to be weak. Note, however, that the bulk ion concentrations in the present study were chosen to be very low. The excess increases with the bulk concentration (albeit underproportionally) and therefore can be expected to reach considerable levels at biologically or technologically relevant concentrations. Selectivity to only one ion species in a salt solution inevitably leads to a charge separation that, in turn, results in electrostatic repulsion between two such surfaces.^{6,49} This notion provides a route to a better understanding of the ion-specificity in the swelling of wood.⁶ In a biological context, the effective surface charge induced by preferential interactions of ions with headgroup-ordered glycolipid microdomains ("lipid rafts")⁵⁰ is suited to attract proteins and to accelerate their binding. Moreover, such preferential ion interactions lead to additional coupling between the lateral and perpendicular equations of state of multilamellar membrane systems.⁵¹ Our results motivate further systematic investigations with the aim to identify correlations between the structural features of the crystalline saccharide surfaces and the preferentially adsorbing ion species. In the longer term, atomistic molecular dynamics simulations appear to be suited to shed additional light on the underlying physical mechanisms.

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.jpcllett.8b03865.

Additional information on methods and results, including fluorescence studies, additional calculations, modeling of the electron density profiles, prediction of ion excesses, GIXD results, chemical structure representations, Langmuir compression isotherms, and peak indexing, (PDF)

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Notes

The authors declare no competing financial interest.

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