

Cite this: *RSC Adv.*, 2019, 9, 4072

Fabrication of three-dimensional polyetherimide bead foams *via* supercritical CO₂/ethanol co-foaming technology

Dong Feng, Li Li * and Qi Wang

The fabrication of light-weight and high-performance polymer foams, especially special engineering plastic foams, with complicate three-dimensional (3D) geometry remains a great challenge worldwide. In this study, microcellular polyetherimide (PEI) bead foams with 3D geometry and high expansion ratio were successfully prepared by using supercritical CO₂ (scCO₂)/ethanol (EtOH) as co-blowing agent. The co-foaming mechanism and the effect of EtOH on foaming properties were studied. The results indicated that the addition of EtOH increased the solubility of co-blowing agent in PEI matrix by promoting the interactions between them, thus broadening the foaming temperature window and significantly increasing the expansion ratio, up to 7.12. The obtained PEI foams with 3D geometry had the cell size of 58.54 μm and cell density of 3.66 × 10⁶ cells per cm³, as well as excellent mechanical strength, e.g., tensile stress of 6.59 MPa and compression stress of 6.87 MPa. This co-foaming technology also has a great potential in fabricating other high-performance polymer foams.

Received 26th November 2018

Accepted 23rd January 2019

DOI: 10.1039/c8ra09706b

rsc.li/rsc-advances

1. Introduction

Polymer foams have been widely used in industry, agriculture, packaging, construction, transportation, *etc.*, because of their light weight, excellent specific strength, heat insulation and cost-saving features over their solid analogs.^{1,2} Currently, the most widely used polymer foams are general plastic foams, mainly including polystyrene (PS), polythene (PE), polypropylene (PP), *et al.*^{3,4} However, due to the poor performance of the matrix, general plastic foams usually show some drawbacks, e.g. low mechanical strength, poor thermal stability and flame retardancy, thus are hardly to meet the requirements of high-tech industries like aerospace and aeronautic,⁵ military and electronic fields.⁶ It is urgent to develop lightweight and high-performance polymer foams, especially special engineering plastic foams, used for high-tech industries.

Polyetherimide (PEI) is an amber and amorphous high-performance engineering thermoplastic with excellent processability, intrinsic flame retardancy and less smoke generation, and can be long-term used at the temperature above 200 °C. However, ascribing to the high processing temperature of PEI (near to 400 °C), it is quite difficult to find a proper chemical blowing agent for PEI due to its degradability or explosibility at so high temperature. Therefore, PEI foams are usually prepared by physical foaming methods using inert gas including carbon dioxide (CO₂) and nitrogen (N₂) as blowing

agent. At present, the main fabrication methods for PEI foams are solid-state intermittent foaming technology, melt extrusion foaming technology and reactive extrusion foaming technology, among which solid-state intermittent foaming technology^{5,7} is the widely used one, and has also been applied in the preparation of some other high-performance polymer foams like polyether-ether-ketone (PEEK),^{8,9} polyphenylene sulfide (PPS)^{10,11} and polysulfone (PSU).^{12,13} For this technology, the polymer resins are first saturated with supercritical CO₂ (scCO₂) for a quite long time at quite low temperature, then foaming by quickly increasing the temperature.¹⁴ Yu, *et al.*¹⁵ adopted this method to get PEI foaming film with the cell size of 180 nm and cell density of 1.58 × 10¹³ cells per cm³. The saturation time for such PEI film with only a thickness of 0.45 mm was more than 20 h at 8 MPa and 40 °C, and the foaming temperature was 170 °C. With the increasing of the thickness of PEI samples used for foaming, the saturation time must be greatly increased. For example, the saturation time for 1.5 mm thickness PEI sample was above 300 h at 5 MPa and room temperature,⁵ and the expansion ratio and cell size of the obtained PEI foaming products were 2.5 and 30 nm to 4 μm, respectively. Limited by so long saturation time of solid-state intermittent foaming technology, the PEI foam products are mostly one-dimensional sheets, resulting in that PEI foams even with micro/nano-cellular structures still have narrow applications.^{16,17} Then, Goel *et al.* proposed a developed solid-state foaming technology,¹⁸ *i.e.* polymer resin was first saturated in scCO₂ at a relatively high temperature, over *T_g* or *T_m* for amorphous or semi-crystalline polymer, followed by rapid pressure quench. In this way, the saturation time could be greatly reduced because

State Key Laboratory of Polymer Materials Engineering, Polymer Research Institute of Sichuan University, Chengdu 610065, China. E-mail: powerlily@scu.edu.cn; Fax: +86-28-85402465; Tel: +86-28-85405133



of the higher diffusion rate of scCO₂ at high temperature, but the solubility of CO₂ was decreased, which was liable to led to the undesired concentration gradient of CO₂ in the polymer matrix, especially the thick polymer samples, making the fabrication of PEI foams with high expansion ratio and uniform cell structure quite difficult.

Accordingly, to prepare PEI and PEI based composite foams with 3D shapes and satisfied cell structure, we developed a mould assisted batch foaming technology by using scCO₂ as blowing agent. In this foaming process, polymer granules were initially filled in a stainless steel mould with many micro nozzles, through which scCO₂ could be injected. After being saturated at high temperature for a relatively short time, the bead foam products were obtained *via* rapidly depressurization of the stainless steel mould. To further improve the cell structure and expansion ratio of our PEI bead foam products, in this paper, ethanol (EtOH) was further introduced as co-blowing agent with scCO₂, and the foaming mechanism and foaming behaviors of PEI with scCO₂/EtOH were investigated. The foaming technology described in this paper was high efficiency and could possibly be applied to fabricate other high-performance thermoplastic foams, like PEEK, PSU and PPS, *etc.*, with 3D geometry and uniform cell structure.

2. Experimental

2.1. Materials

PEI (ULTEM 1000) with T_g of 215 °C was purchased from Sabic, America. CO₂ was supplied by Xuyuan Co., Ltd. (China) with purity higher than 99%. EtOH was supplied by Chengdu Changlian Co., Ltd. (China). All materials were used without further purification.

2.2. Preparation of PEI foamed beads and bead foamed parts

After being dried in a drying oven at 120 °C for 4 h, 50 g PEI granules were placed into a 0.5 L stainless steel vessel, then the vessel was put into a high pressure vessel with certain EtOH (0%, 15%, 30%, 45% and 60%, respectively, with mol ratio to scCO₂) in it at room temperature. The vessel was first flushed with 0.2 MPa CO₂ for about 1 min to exhaust the air inside, and then immediately filled with CO₂ to certain pressure (3, 4, 5 and 6 MPa, respectively), followed by heating to different temperature (175 °C, 183 °C, 191 °C, 199 °C, 207 °C and 215 °C) within 10 min and saturating at this temperature for 90 min. The final saturation pressure was 8 MPa, 10 MPa, 12 MPa and 14 MPa, respectively. After that, the vessel was rapidly depressurized (5 MPa s⁻¹), and PEI foamed beads were obtained. The mol of scCO₂ was calculated using appropriate PVT data from saturation pressure, saturation temperature and vessel volume.^{19,20}

For bead foamed parts, the PEI granules were first filled into a stainless steel mould with micro nozzles. Then, the mould was put into the high pressure vessel. After being saturated at a certain condition, the bead foamed parts were obtained *via* rapidly depressurization (5 MPa s⁻¹) of the high pressure vessel. The foaming process was shown in Fig. 1.

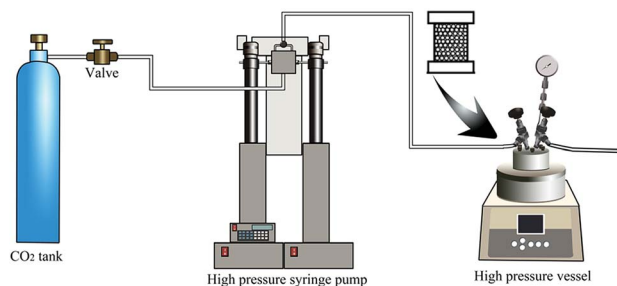


Fig. 1 Schematic illustration of the mould assisted batch foaming process.

2.3. Molecular dynamic (MD) simulation

In order to better understand the interactions between scCO₂/EtOH co-solvent and PEI chains, MD simulation of PEI/CO₂/EtOH systems was studied using Materials Studio 8.0. The simulations were carried out with the COMPASS force field, and PEI molecular chains with 10 repeat units were built due to the limited computing resources. After the optimization of the chain geometry, cubic box containing one PEI chain, 1000 CO₂ and different number of EtOH molecules were constructed in Amorphous Cell tools with a periodic boundary condition, the snapshot of cubic box containing different EtOH molecules was shown in Fig. 2, and the corresponding parameters were shown in Table 1. After annealing treated from 300 K to 500 K at 10 K min⁻¹ ramp rate for 5 times, 100 ps NPT MD simulation was carried out at 190 °C and 12 MPa using Forcite tools. The interaction energy between PEI and scCO₂/EtOH was obtained using the mean value of the last 10 configurations.

2.4. Characterization

2.4.1. Solubility of blowing agents. PEI dried granules were placed in a high pressure vessel (0.5 L) at 190 °C and 12 MPa for different time. After the vessel was cooled by cold water to ambient temperature, the PEI pellets were quickly taken out and immediately weighted on a Mettler AE240 analytical (Labx Co., Ltd., USA) with ±10 µg accuracy to measure the CO₂ and CO₂/EtOH adsorbed. Five measurements were performed for each sample.

2.4.2. Differential scanning calorimetry (DSC). Glass transition behaviors of PEI with different blowing agent were performed on a Q20 DSC apparatus (TA Instruments Co., Ltd., New Castle, USA) in nitrogen atmosphere from 40 °C to 240 °C at the heating rate of 20 °C min⁻¹. To avoid the escape of foaming agent during the testing, the samples were placed into a sealed crucible.

2.4.3. High pressure capillary (HP-capillary). High pressure capillary rheological behaviors of PEI foaming systems were carried out on Rosand RH7D HP-capillary rheometer (Malvern Instrument, UK) with a long die ($L/D = 16$) at 340 °C from 100–2000 s⁻¹. In order to avoid the escape of foaming agent during the testing process, PSU was used to seal both sides of the charging barrel, and PEI/CO₂ or PEI/CO₂/EtOH foaming systems must be immediately transferred into the charging barrel of the

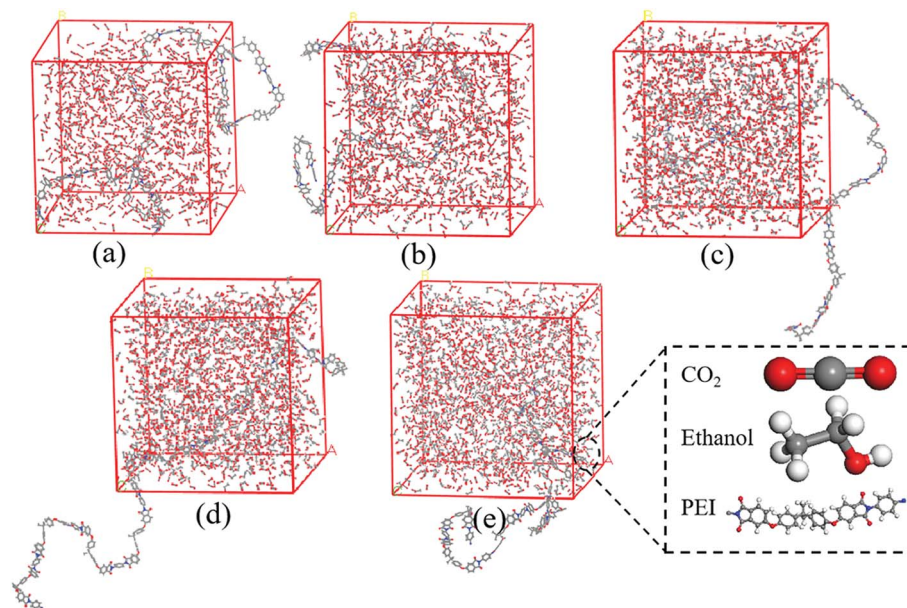


Fig. 2 Snapshot of cubic box containing different molecules. (a) PEI + 1000CO₂; (b) PEI + 1000CO₂ + 150EtOH; (c) PEI + 1000CO₂ + 300EtOH; (d) PEI + 1000CO₂ + 450EtOH; (e) PEI + 1000CO₂ + 600EtOH.

Table 1 Systems of PEI/CO₂/EtOH in MD simulations

Compositions	Number of PEI chain	Number of CO ₂	Number of EtOH
PEI/scCO ₂	1	1000	0
PEI/CO ₂ -15% EtOH	1	1000	150
PEI/CO ₂ -30% EtOH	1	1000	300
PEI/CO ₂ -45% EtOH	1	1000	450
PEI/CO ₂ -60% EtOH	1	1000	600

HP-capillary rheometer once being taken out from the high pressure vessel.

2.4.4. Scanning electron microscope (SEM). Cell structure of PEI foam was analyzed using Inspect F SEM instrument (FEI Co., Ltd., USA) at 0.5 torr and 20 kV. The image processing software Nano Measurer was used to calculate the cell density (N_f , number of cells per unit foamed polymer volume) and the average cell diameter. N_f was determined using eqn (1).

$$N_f = \left[\frac{nM^2}{A} \right]^{3/2} \quad (1)$$

where n is the number of cells in the SEM micrograph, M is the magnification factor, and A is the area of the micrograph (in cm²).

2.4.5. Expansion ratio. The expansion ratio ϕ of the foam was calculated using eqn (2).²¹

$$\phi = \frac{\rho}{\rho_f} \quad (2)$$

where ρ and ρ_f are the densities of unfoamed and foamed PEI, respectively, and can be tested on a MDMDY-350 automatic density instrument (Meidi analytical Co., Ltd., China).

2.4.6. Mechanical properties. Mechanical properties including tensile strength and compression strength of PEI

foamed parts were tested at ambient temperature on an Instron 5567 Tensile & Compression Tester (Instron Co., Ltd., United States). The tensile strength was conducted at a speed of 5 mm min⁻¹ with a 60 mm × 10 mm × 5 mm rectangular sample according to GB/T 9641-1988/ISO 1926-1979, while the compression strength testing of a cylinder sample with diameter of 30 mm and height of 10 mm was conducted at a cross-head speed of 2 mm min⁻¹ according to GB/T 8813-2008/ISO 844: 2004.

3. Results and discussion

3.1. Solubility of CO₂/EtOH

The solubility of blowing agent in a polymer matrix is an important factor that affects the cell structure of the final foam. The solubility of CO₂ in PEI at 190 °C and 12 MPa for different saturation time was investigated, as shown in Fig. 3a. The CO₂ concentration was defined as the mass ratio of CO₂ to PEI matrix. Apparently, with the increasing of saturation time, CO₂ concentration in PEI increased and reached the equilibrium after 90 min, *i.e.* 21.5 mg g⁻¹ PEI. With the further addition of EtOH into scCO₂ system, the solubility parameter as well as the polarity of scCO₂ could be significantly improved ascribing to the hydrogen-bond interactions between CO₂, EtOH and polymer matrix.^{27,28} This was beneficial to the increase of the diffusivity of CO₂ and would lower the equilibrium saturation time of the co-blowing agent to less than 90 min. However, to better compare the solubility of CO₂/EtOH co-blowing agent in PEI and only CO₂ in PEI, the saturation time for the system with EtOH was still fixed to 90 min. It could be clearly seen from Fig. 3b that CO₂/EtOH co-blowing agent presented greatly increased solubility in PEI matrix compared with CO₂. For example, with the addition of 60% mol ratio EtOH, the solubility of co-blowing agent increased to 86.2 mg g⁻¹ PEI, almost 4

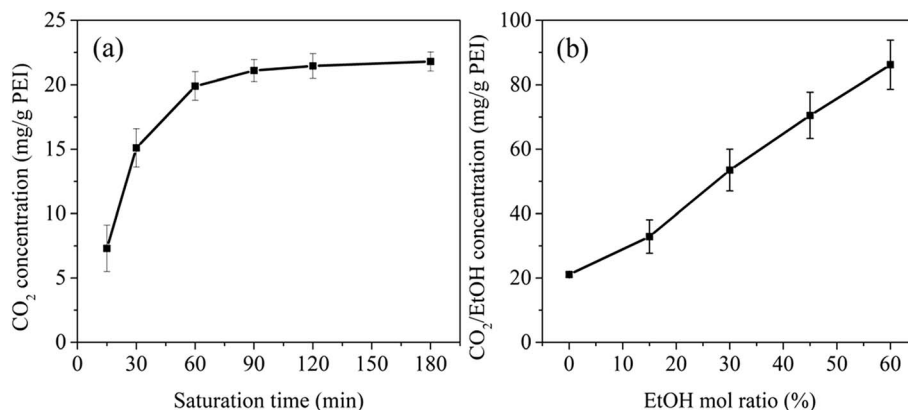


Fig. 3 CO₂ concentration vs. saturation time (a) and CO₂/EtOH equilibrium concentration vs. EtOH mol ratio (b).

times than that of CO₂. Zhao *et al.*²² had reported that the hydroxyl groups of ethanol could connect with the ether group of polymer. So, here, EtOH might interact with both CO₂ and PEI, thus increasing the plasticization effect of blowing agent on PEI and resulting in the increase of the flexibility of PEI molecular chains as well as the greater solubility of CO₂/EtOH mixture in PEI matrix. The increased content of blowing agent would help to obtaining PEI 3D foams with high expansion ratio.

3.2. Plasticization effect of blowing agent

To further illustrate the plasticization effect of blowing agent on PEI molecules, the DSC heating curves (Fig. 4a) and the high pressure capillary curves (Fig. 4b) of PEI with CO₂ and CO₂/EtOH were analyzed. It could be found from Fig. 4a that a significant decrease of T_g appeared with the dissolving of CO₂, *i.e.* from 215 °C to about 190 °C. CO₂ is a Lewis acid, the carbon atom lacking of electrons in CO₂ molecule can act as electron acceptor, so has special interactions with the groups like phenyl group and carbonyl group in polymer chains.^{23,24} In this way, CO₂ could act as a small molecular plasticizer for PEI and magnified the distances of PEI molecules, augmenting their free volumes and weakening their attractions, thus increasing

their mobility and reducing the T_g of PEI. At the same CO₂ concentration, the addition of EtOH further decreased the T_g of the foaming system, and the more the EtOH content, the more the decrease of T_g . For example, the addition of 60% mol ratio EtOH led to the decrease of T_g from about 190 °C to 130 °C, strongly confirming the improved plasticization effect of CO₂/EtOH on PEI. This was beneficial to decrease the foaming temperature, *i.e.* reducing the foaming energy.

The plasticization effect of blowing agent also influenced the melt viscosity of the foaming system. As shown in Fig. 4b, compared with neat PEI, the shear viscosity of PEI/blowing agent system apparently reduced. Especially when 60% mol ratio of EtOH was introduced, the shear viscosity of the system was almost insensitive to shear rate, *i.e.* changed little with the increase of shear rate.

Overall, it would be a best choice to overcome the problem existed in the present PEI foaming technology by using scCO₂/EtOH as co-blowing agent.

3.3. Foaming behavior of PEI/scCO₂/EtOH system

3.3.1. Effect of EtOH concentration. Up to now, the maximum expansion ratio reported for PEI foam by using scCO₂ or scN₂ as blowing agent was no more than 5 and the foaming

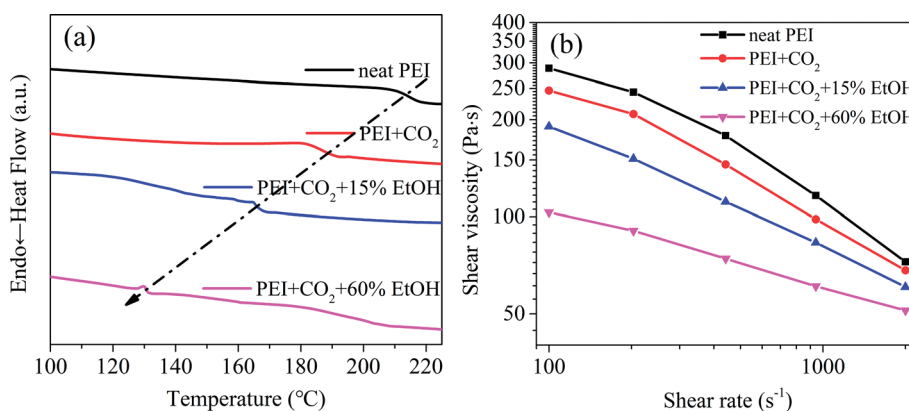


Fig. 4 DSC curves (a) and high pressure capillary curves (b) of different PEI/blowing agent systems.

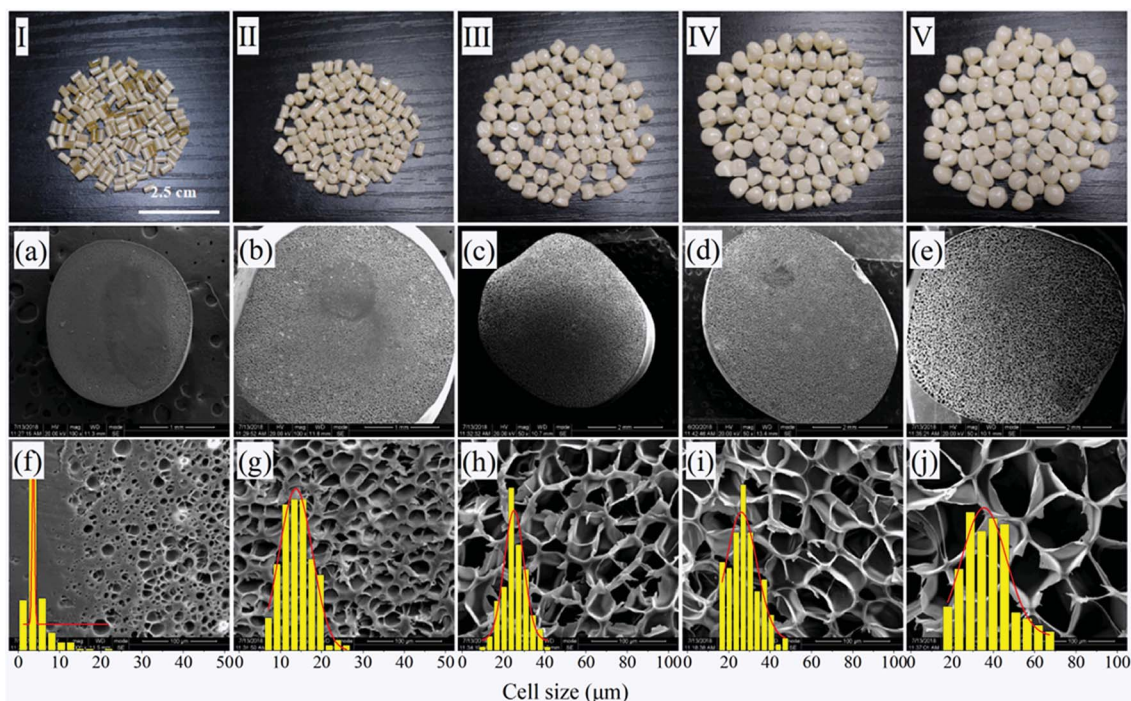


Fig. 5 Macrographs (I–V) and SEM micrographs of PEI foamed beads (a–j) obtained at 190 °C and 12 MPa using CO₂/EtOH as blowing agent. EtOH concentration: (a and f) 0%; (b and g) 15%; (c and h) 30%; (d and i) 45%; (e and j) 60% mol ratio.

temperature and pressure were quite high,^{25,26} which are hardly to be accepted by industries. In this paper, EtOH was adopted as co-blowing agent with scCO₂, and ascribing to its plasticization effect on PEI molecules with scCO₂, the foaming temperature was largely decreased to about 190 °C. Fig. 5 showed the photos and SEM micrographs of PEI foamed beads using scCO₂/EtOH as blowing agent, and the corresponding expansion ratio, cell size and cell density of the beads with different EtOH mol ratio were depicted in Fig. 6. Obviously, PEI granules only using scCO₂ as blowing agent showed partially foaming behavior, while with the addition of EtOH, all the formed beads exhibited smooth surface with fine closed cell structure, as well as significant increases in both expansion ratio and cell size. When EtOH mol ratio increased from 15% to 60%, the expansion ratio and cell size of the foamed beads increased from 1.53 to 6.49 and 14.15 to 37.39 μm, respectively, but the cell density decreased, from 1.15×10^8 cell per cm³ to 1.02×10^7 cell per

cm³. As studied previously, EtOH had obvious plasticization effect on PEI, so could further decrease the surface tension of PEI and reduce the resistance of cell growth, leading to the enlarged cell size and the increased porosity of the foams.²⁷

To give a deep understanding on the increase of expansion ratio with the addition of EtOH, the interactions between CO₂/EtOH co-blowing agent and PEI molecular chains were studied with the help of MD simulations, as shown in Fig. 2. The interaction energy between CO₂/EtOH and a PEI molecular chain, which could quantitatively predict their compatibility, was evaluated from the energy of each system, as expressed in eqn (3).²⁷

$$E_{\text{inter}} = E_{\text{PCE}} - (E_{\text{PEI}} + E_{\text{CE}}) \quad (3)$$

where E_{inter} refers to interaction energy, E_{PCE} represents the total energy of a PEI chain and CO₂/EtOH blowing agent. E_{PEI}

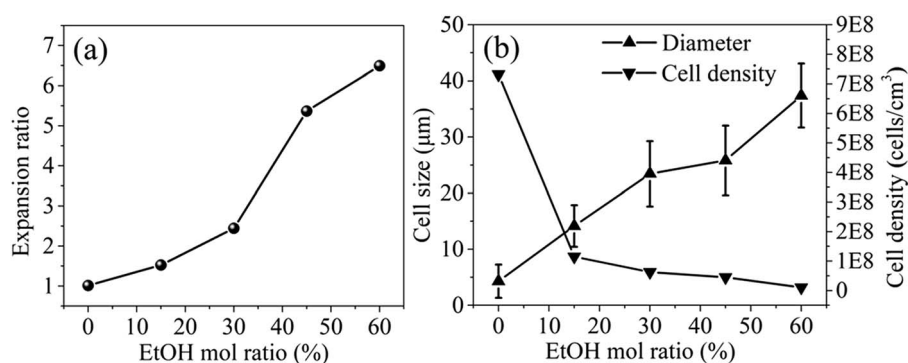


Fig. 6 Expansion ratio (a), cell size and cell density (b) of PEI foamed beads obtained at 190 °C and 12 MPa with different EtOH mol ratio.

Table 2 Interaction energy between a single PEI chain and CO₂-EtOH co-blowing agent obtained from MD simulations

Compositions	E_{PCE} (kJ mol ⁻¹)	E_{PEI} (kJ mol ⁻¹)	E_{CE} (kJ mol ⁻¹)	E_{inter} (kJ mol ⁻¹)
PEI/CO ₂	7753.67	787.61	6420.47	545.59
PEI/CO ₂ -15% EtOH	7188.69	780.04	6244.68	163.97
PEI/CO ₂ -30% EtOH	12 101.28	6998.86	5667.01	-564.58
PEI/CO ₂ -45% EtOH	4302.21	842.68	4678.33	-1218.80
PEI/CO ₂ -60% EtOH	2327.62	587.90	3042.50	-1302.78

and E_{CE} are the individual energy of a PEI chain and CO₂/EtOH blowing agent. The values of E_{PEC} , E_{PEI} and E_{CE} can be obtained from MD simulations and are the average values of the related properties in the final 10 configurations.

As shown in Table 2, under the same foaming conditions, with the increase of EtOH mol ratio, E_{inter} gradually decreased, even to a negative value when EtOH mol ratio was above 30%. Positive E_{inter} indicates the endothermic reaction. For PEI/CO₂ system, E_{inter} was 545.59 kJ mol⁻¹, confirming the relatively low adsorption capacity of PEI molecules on CO₂. That was why partially foaming behavior happened for this system. While with the addition of EtOH (15% mol ratio), E_{inter} decreased from 545.59 kJ mol⁻¹ to 163.97 kJ mol⁻¹, meaning that the compatibility between PEI and the blowing agent enhanced attributing to the synergistic effect of EtOH on both PEI and CO₂. In this way, CO₂/EtOH co-blowing agent could better exert its plasticization effect and foaming efficiency, making the expansion ratio of PEI foamed beads slightly increased to 1.53. Further increasing EtOH mol ratio from 30% to 60%, E_{inter} values became negative, *i.e.* from -564.58 kJ mol⁻¹ to -1302.78 kJ mol⁻¹, indicating the occurrence of the exothermic reaction during the foaming process, possibly caused by the increase of the bonding energy between PEI and scCO₂/EtOH solvent, which helped to enhance the solubility of scCO₂/EtOH in PEI matrix. As a result, the expansion ratio of PEI foamed beads greatly increased, from 2.44 to 6.49.²⁸ The variation trend

of E_{inter} was in good agreement with the rising tendency of expansion ratio from thermodynamical view, suggesting that MS simulations might be an alternative way to predicting foaming behaviors.

Considering the comprehensive cell structure of PEI foamed beads, 45% mol ratio of EtOH was chose in the letter analyses.

3.3.2. Effect of foaming temperature. The effects of foaming temperature on cell structures were investigated at the foaming pressure of 12 MPa and EtOH mol ratio of 45%. The photos and SEM micrographs of PEI foamed beads were shown in Fig. 7, and the corresponding expansion ratio, cell size and cell density at different foaming temperature were depicted in Fig. 8. As reported by Miller *et al.*, the minimum foaming temperature was equivalent to the effective T_g of a polymer/solvent system.^{5,7} The introduction of EtOH enhanced the solubility of scCO₂/EtOH in PEI matrix, thus improving the movement of PEI chains and reducing the T_g of the foaming system. In this way, the minimum foaming temperature of system could be reduced to about 175–215 °C, much lower than those reported in present ref. 5 and 9. It could be seen from Fig. 7 and 8 that with the increase of the foaming temperature, the cell size and expansion ratio of PEI foamed beads increased, while the cell density decreased, due to the accelerated movement of PEI molecular chains with the temperature, which weakened the resistance of the matrix to the bubble growth. For example, when temperature increased from 175 °C to 191 °C,

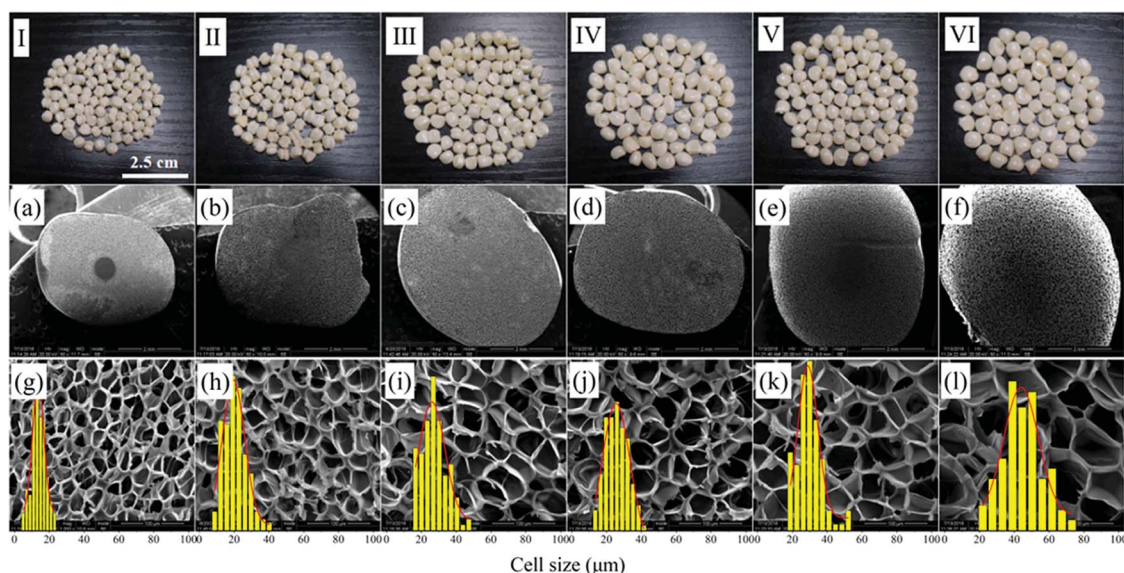


Fig. 7 Macrographs (I–VI) and SEM micrographs of PEI foamed beads (a–l) obtained at different foaming temperature. (a and g) 175 °C; (b and h) 183 °C; (c and i) 191 °C; (d and j) 199 °C; (e and k) 207 °C; (f and l) 215 °C.

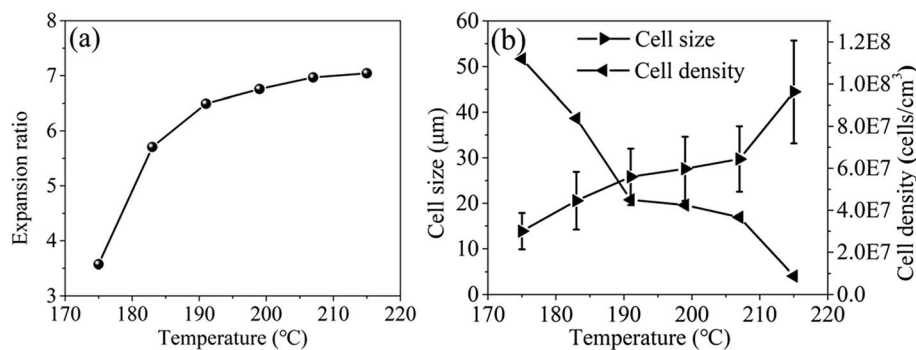


Fig. 8 Expansion ratio (a), cell size and cell density (b) of PEI foamed beads obtained at different temperatures with co-blowing agents.

the expansion ratio and cell size significantly increased from 3.57 and 13.87 μm to 6.50 and 25.57 μm, respectively. However, with the further increase of the temperature, up to 215 °C, the expansion ratio only had a slight increase, from 6.50 to 7.04, indicating that PEI matrix now was insensitive to the foaming temperature from 191 °C to 215 °C, thus a wide foaming temperature window could be obtained for PEI by using EtOH as co-blowing agent with scCO₂.

3.3.3. Effect of saturation pressure. The effect of saturation pressure was investigated at 5 MPa s⁻¹ depressurization rate, 190 °C foaming temperature and 45% EtOH mol ratio. After being saturated at 190 °C for 1.5 h, the inner saturation pressures were 8 MPa, 10 MPa, 12 MPa and 14 MPa, respectively.

The optical images and the cell morphologies of PEI foamed beads were shown in Fig. 9, and the corresponding expansion ratio, cell size and cell density at different saturated pressure were depicted in Fig. 10. The results showed that PEI foamed beads with polished surfaces and fine cells were obtained at each saturation pressure. With the increase of saturation pressure, the cell size of the foamed beads decreased, from 66.54 μm to 11.36 μm, while the cell density increased, from 2.27 × 10⁶ to 1.69 × 10⁸ cells per cm³. It was well known that the higher the saturation pressure, the more the blowing agent will be dissolved in a polymer matrix. According to the homogeneous nucleation theory, higher blowing agent concentration means greater degree of supersaturation and lower energy barrier to

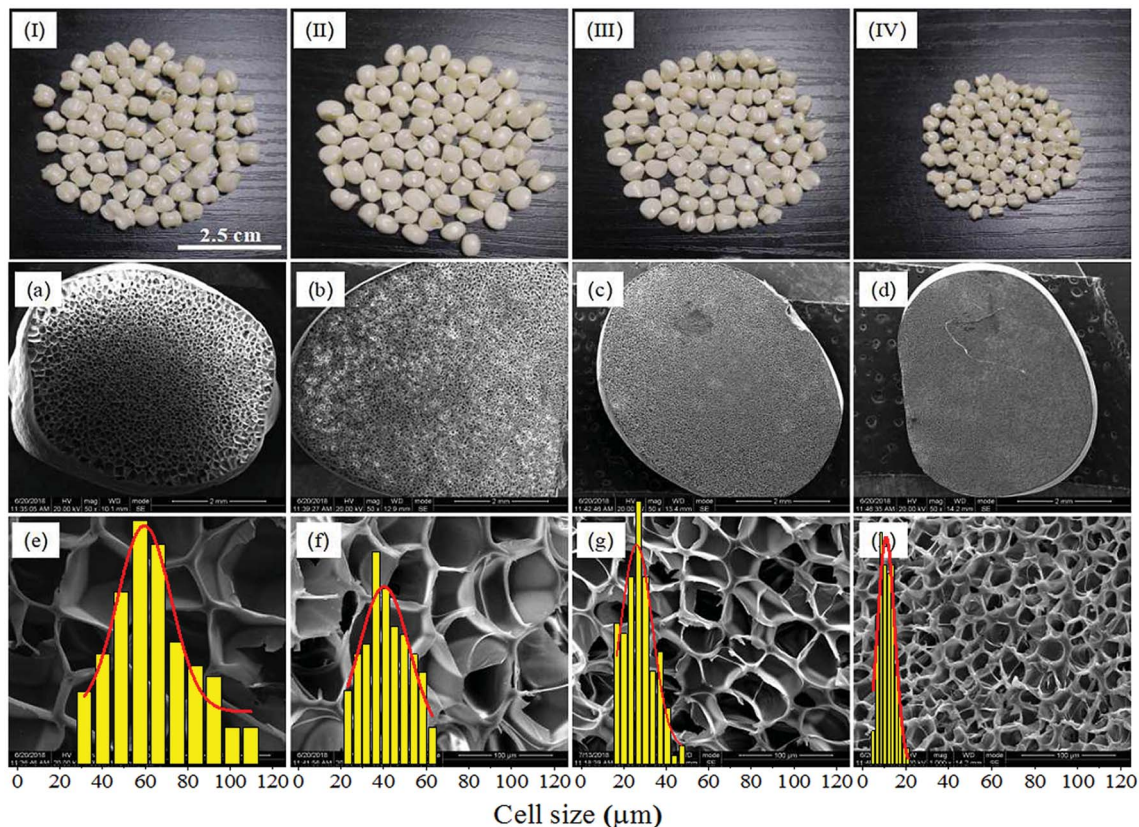


Fig. 9 Macrographs (I–IV) and SEM micrographs of PEI foamed beads (a–h) obtained at different saturation pressure. (a and e) 8 MPa; (b and f) 10 MPa; (c and g) 12 MPa; (d and h) 14 MPa.

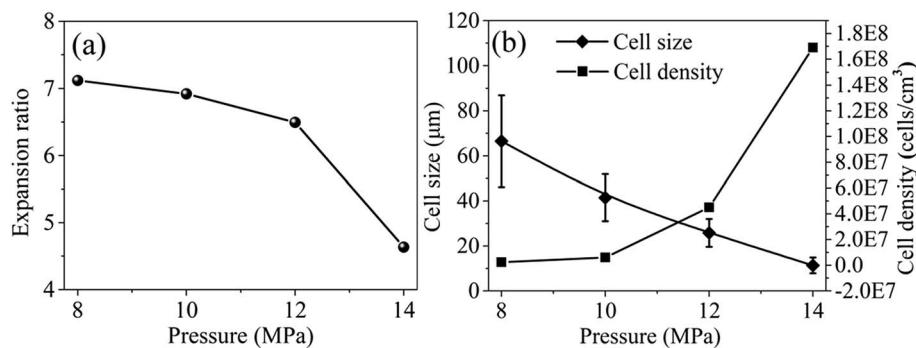


Fig. 10 Expansion ratio (a), cell size and cell density (b) of PEI foamed beads obtained at different saturation pressure.

nucleation, so more cells per unit volume will be generated and higher cell density in the foamed product will be achieved.²⁹ The error bar in Fig. 10b reflected the distribution of the cell size, which became narrow with the increase of saturation pressure, indicating that higher saturation pressure was helpful to improve the uniformity of cells because of the increased cell density over higher supersaturation. When more cell nuclei formed and simultaneously grew in a limited space, the differences in cell growth should be decreased, undoubtedly resulting in more uniform and smaller cell size.

It should be noticed that with the increase of saturation pressure, the expansion ratio of the foamed beads continuously decrease, from 7.12 to 4.63, and the decrease extent enlarged at the saturation pressure above 14 MPa ($P_{\text{initial}} = 6$ MPa) (Fig. 10a). As discussed in previous section, the hydrogen atom (H) in -OH group of EtOH could form strong hydrogen bonds with both scCO_2 and PEI,³⁰ thus improving the solubility of co-blowing agent in PEI matrix. Under the co-plasticization and co-foaming effects of scCO_2 and EtOH on PEI, PEI foamed beads with improved cell structure and high expansion ratio (7.12) were obtained. However, the increase of the saturation pressure was realized by injecting more scCO_2 and EtOH into the high pressure vessel, when other foaming conditions were fixed. The increased blowing agent in system would surely enhance the nucleation ability of the foaming system, leading to the formation of smaller cells, but also imposed greater plasticization effect on PEI molecular chains, which possibly weakened the strength of the PEI cell walls and made them unable to keep so many gases in the cells after pressure release. As a result, the expansion ratio of the foamed beads reduced. This effect might be quite obvious when saturation pressure was above 14 MPa, so the expansion ratio decreased greatly.

3.4. Macromorphology and mechanical properties

Through above analyses and to simultaneously obtaining high expansion ratio and cell density while maintaining cell size as small as possible, the PEI bead foaming products were fabricate at the foaming conditions of 190 °C, 8 MPa and 45% EtOH using a stainless steel mould by applying 5 MPa s^{-1} of depressurization rate. Fig. 11 showed the PEI foamed samples that just fully filled in the stainless mold. It could be seen that the PEI raw

granules were amber-like, while the foam part with high expansion ratio looked white.

Fig. 12 showed the tensile and compression behaviors of the PEI foamed samples. Airex®R82 PEI, a commercial PEI foamed product, with the tensile stress of 2.2 MPa and the compression stress of 1.4 MPa was provided by Airex company in Switzerland. Obviously, PEI bead foaming samples showed higher mechanical properties (tensile stress: 6.59 MPa; compression stress: 6.87 MPa) than those of the commercial Airex®R82 PEI, ascribing to the creation of more uniform microcellular cells in our PEI matrix, with cell size of 58.54 μm and cell density of 3.66×10^6 cells per cm^3 (Fig. 13). This confirmed again that the bead foaming technology proposed in this study was a promising approach for the fabrication of the special engineering thermoplastic foams with complicated 3D shape and high performances.

The inter-bead bonding and the intra-bead bonding force determined the mechanical properties of the final parts.³¹ The bead-to-bead sintering behaviors, including the surface, cut surface and fracture surface, of the PEI bead foaming samples were investigated using FE-SEM, as shown in Fig. 13. It could be seen that the foamed samples presented a good surface quality, and no inter-bead bonding boundary was observed on the cut surface and fracture surface of the foamed samples, implying the strong local sintering quality among the foamed PEI beads.³² The fracture surface after the tensile test again showed that the breaking of PEI bead foams almost occurred intra the bead, further identifying the strong inter-bead sintering.³³

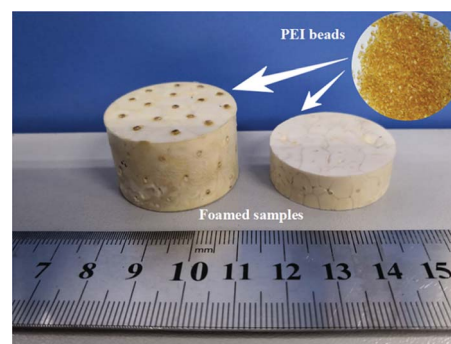


Fig. 11 Photograph of PEI granules and PEI bead foaming samples.

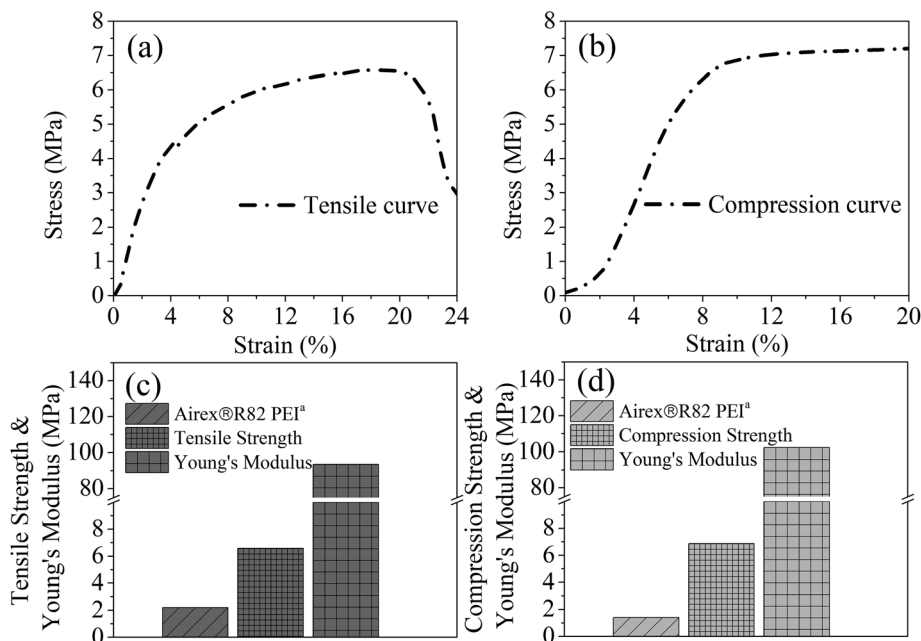


Fig. 12 Tensile (a) and compression (b) stress–strain curves, and the corresponding tensile (c) and compression (d) properties of the PEI bead foaming samples.

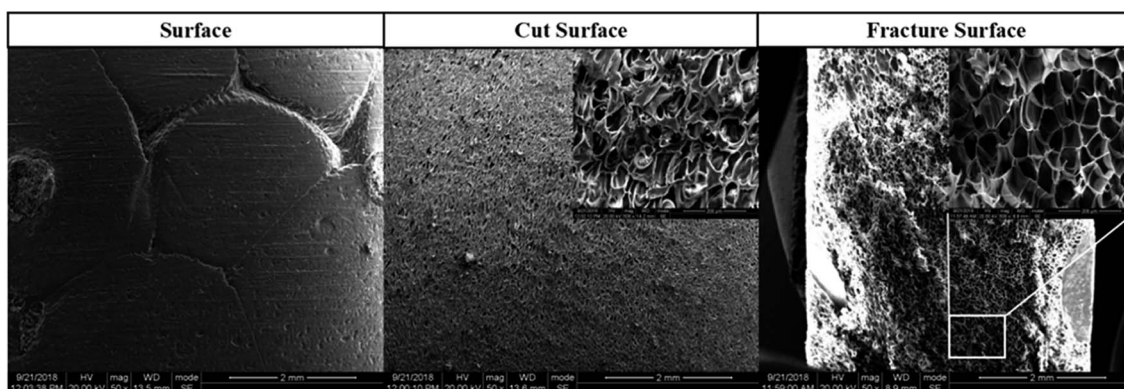


Fig. 13 SEM micrographs of the surface, cut surface and fracture surface of PEI bead foams parts.

4. Conclusions

PEI bead foams with high expansion ratio and 3D shape were prepared by using supercritical CO_2 (scCO_2)/ethanol (EtOH) as co-blowing agent based on our developed mould assisted batch foaming technology. The results showed that the addition of EtOH could significantly increase the interactions between blowing agent with PEI chains, thus helping to enhance the compatibility and improving the solubility of CO_2 /EtOH mixture in PEI. In this way, CO_2 /EtOH mixture could efficiently exert their plasticization effect on PEI molecular chains, decreasing T_g of PEI and making PEI foamed at relatively lower temperature. The increased CO_2 /EtOH in PEI matrix also promoted more nuclei formed, leading to the better foaming of PEI as well as the increased expansion ratio of the final foamed beads, up to 7.12. At proper foaming conditions and with appropriate CO_2 /EtOH content, PEI foamed parts with 3D geometry, cell size

of $58.54 \mu\text{m}$ and cell density of 3.66×10^6 cells per cm^3 were obtained. These PEI foamed parts presented excellent inter-bead sintering behavior, thus having good mechanical properties (6.59 MPa tensile strength and 6.87 MPa compression strength) and a great potential in the applications of some high-tech industries like aerospace, special ship *et al.*, where 3D structure foam materials with light weight, high strength and heat resistance properties are urgently required.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

This work was support by National Natural Science Foundation of China (5171001071) and the Program of Innovative Research

Team for Young Scientists of Sichuan Province (2016TD0010). We would like to thank the Analytical & Testing Center of Sichuan University for providing Materials Studio 8.0 and we would be grateful to Daichuan Ma for his help of computational simulation.

References

- 1 C. B. Park, *Polymeric Foams: Science and Technology*, CRC Press, 2006.
- 2 C. Okolieocha, D. Raps, K. Subramaniam and V. Altstädt, *Eur. Polym. J.*, 2015, **73**, 500–519.
- 3 M. Altan, *Thermoplastic Foams: Processing, Manufacturing, and Characterization*, 2018.
- 4 G. Wang, X. Chen, P. Liu and S. Bai, *J. Appl. Polym. Sci.*, 2016, **134**, 44356.
- 5 D. Miller and V. Kumar, *Polymer*, 2011, **52**, 2910–2919.
- 6 L. Cafiero, S. Iannace and L. Sorrentino, *Eur. Polym. J.*, 2016, **78**, 116–128.
- 7 D. Miller, P. Chatchaisucha and V. Kumar, *Polymer*, 2009, **50**, 5576–5584.
- 8 Q. Yang, G. Zhang, Z. Ma, J. Li and X. Fan, *J. Appl. Polym. Sci.*, 2015, **132**, 42576.
- 9 T. Nemoto, J. Takagi and M. Ohshima, *Polym. Eng. Sci.*, 2010, **50**, 2408–2416.
- 10 M. Itoh and A. Kabumoto, *Furukawa Rev.*, 2005, **28**, 32–38.
- 11 Z. Ma, G. Zhang, Q. Yang, X. Shi, J. Li and X. Fan, *Polym. Compos.*, 2016, **37**, 2527–2540.
- 12 Z. Li, Y. Jia and S. Bai, *RSC Adv.*, 2018, **8**, 2880–2886.
- 13 H. Sun and J. E. Mark, *J. Appl. Polym. Sci.*, 2002, **86**, 1692–1701.
- 14 S. Costeux, *J. Appl. Polym. Sci.*, 2015, **131**, 1–16.
- 15 H. Yu, Y. Lei, X. Yu, X. Wang, T. Liu and S. Luo, *J. Appl. Polym. Sci.*, 2015, **132**, 42325.
- 16 H. Abbasi, M. Antunes and J. I. Velasco, *Eur. Polym. J.*, 2015, **69**, 273–283.
- 17 C. Zhou, N. Vaccaro, S. S. Sundarram and W. Li, *J. Cell. Plast.*, 2012, **48**, 239–255.
- 18 S. K. Goel and E. J. Beckman, *Polym. Eng. Sci.*, 1994, **34**, 1137–1147.
- 19 I. Tsvintzelis, E. Pavlidou and C. Panayiotou, *J. Supercrit. Fluids*, 2007, **42**, 265–272.
- 20 S. Angus, B. Armstrong and K. M. de reuck, *International Thermodynamic Tables of the Fluid State : Carbon Dioxide*, Oxford, 1985.
- 21 E. Rezvanpanah, S. R. Ghaffarian Anbaran and E. Di Maio, *Carbon*, 2017, **125**, 32–38.
- 22 N. Zhao, C. Zhu, L. Howe Mark, C. B. Park and Q. Li, *J. Appl. Polym. Sci.*, 2015, **132**, 11962–11972.
- 23 D. Kajiya, M. Imanishi and K. I. Saitow, *J. Phys. Chem. B*, 2016, **120**, 785–792.
- 24 S. G. Kazarian, M. F. Vincent, F. V. Bright, C. L. Liotta and C. A. Eckert, *J. Am. Chem. Soc.*, 1996, **118**, 1729–1736.
- 25 L. Sorrentino, M. Aurilia and S. Iannace, *Adv. Polym. Technol.*, 2011, **30**, 234–243.
- 26 W. Zhai, W. Feng, J. Ling and W. Zheng, *Ind. Eng. Chem. Res.*, 2012, **51**, 12827–12834.
- 27 D. D. Hu, Y. Gu, T. Liu and L. Zhao, *J. Supercrit. Fluids*, 2018, **140**, 21–31.
- 28 D. Hu, S. Sun, P. Yuan, L. Zhao and T. Liu, *J. Phys. Chem. B*, 2017, **119**, 3194–3204.
- 29 E. Di Maio and E. Kiran, *J. Supercrit. Fluids*, 2018, **134**, 157–166.
- 30 D. Hu, *CIESC J.*, 2018, **69**, 555–562.
- 31 E. K. Lee, PhD thesis, School of Graduate Studies, 2010.
- 32 M. Nofar, A. Ameli and C. B. Park, *Mater. Des.*, 2015, **83**, 413–421.
- 33 W. Zhai, Y.-W. Kim, D. W. Jung and C. B. Park, *Ind. Eng. Chem. Res.*, 2011, **50**, 5523–5531.