

Supporting Information

**Comprehensive Study of Oxygen Vacancies on
the Catalytic Performance of ZnO for CO/H₂
Activation Using Machine Learning-accelerated
First-principles Simulations**

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1. Computational Methods

1.1 Active Learning for Accelerating Genetic Algorithms (GAs) Global Optimization

Supplemental Note S1. GAs are metaheuristic optimization algorithms inspired by Darwinian evolution. Performing crossover, mutation, and selection operations, the algorithm progresses a population of evolving candidate solutions.

To estimate whether a configuration should be selected and added to the training dataset, two selection procedures are employed. First, the energy uncertainty assessed by the model committee falls between the bounds ($\sigma_{lo.} < \sigma_t < \sigma_{hi.}$), which indicates that the corresponding structure is not well learned and rules out the unphysical configurations. In this strategy, an ensemble of models is concurrently trained with different random seeds that initialize the network parameter and the prediction uncertainty is estimated by the variance of these model outputs. For instance, the mean and variance of the total energy prediction are given by

$$\bar{E} = \frac{1}{M} \sum_1^M E^i \quad (1)$$

$$\sigma^2(E) = \frac{1}{M-1} \sum_1^M |E^i - \bar{E}|^2 \quad (2)$$

where M is the number of models. This estimation can be used for other properties as well, such as forces and stresses. Due to the high nonlinearity of the neural network model and the randomness of the optimizer, the final converged models would be slightly different even with the same inputs. For $\sigma_t < \sigma_{lo.}$, this means that the explored configurations can be predicted with a high accuracy by the Embedded atom neural network(EANN) models. Labelling these configurations does not help to improve the quality of the model but leads to a waste of the computational resources. $\sigma_t > \sigma_{hi}$ illustrates that the simulation sometimes samples unphysical configurations, e.g., overlapping atoms. These configurations may suffer from difficulties in converging self-consistent iterations and are irrelevant to the accuracy of the final model. Therefore, structures corresponding to this case ($\sigma_{lo.} < \sigma_t < \sigma_{hi.}$) are selected as candidates for labelling.

Second, the CUR decomposition method was used to further sift the candidates based on their structural diversity.¹⁻² Here, we use a configuration-averaged Smooth Overlap of Atomic Positions (SOAP) descriptor that compares entire unit cells to one another to quantify the difference between candidate structures. Screening of high-quality structures will greatly improve the quality of potential energy surfaces (PESs), thereby accelerating structure searching effectively.

The workflow has two stop criteria. First, the energy deviation of the explored minima is less than 0.02eV/atom to ensure the accuracy of the PES. Second, the same lowest-energy structure is produced during several consecutive GA optimization processes. Although the EANN model can generate energies and forces that are very close to the density functional theory (DFT) reference model, a small intrinsic error still exists. Therefore, the structures with sufficiently low energies (the top 100 most stable structures for the same composition) were relaxed with DFT. These structures converged very fast within 10-20 ionic steps, indicating well-trained machine learning potentials (MLPs).

Table S1. Energy and force root mean squared error (RMSE) on DFT-labelled structures. The Nframes is the number of the training structures, as distinguished by the chemical formula (system) and the number of atoms per cell (Natoms).

| System | Natoms | Nframes | E_{RMSE} (eV/atom) | F_{RMSE}^o (eV/Å) | F_{RMSE}^{zn} (eV/Å) |
|--------|--------|---------|----------------------|---------------------|------------------------|
| O0Zn14 | 14 | 25 | 0.0263 | nan | 0.0594 |
| O0Zn15 | 15 | 683 | 0.0212 | nan | 0.0945 |
| O0Zn16 | 16 | 4164 | 0.0179 | nan | 0.1045 |
| O0Zn17 | 17 | 8 | 0.0121 | nan | 0.0428 |
| O0Zn19 | 19 | 63 | 0.023 | nan | 0.0895 |
| O0Zn24 | 24 | 229 | 0.0185 | nan | 0.0548 |
| O0Zn28 | 28 | 22 | 0.019 | nan | 0.0529 |
| O0Zn29 | 29 | 18 | 0.0203 | nan | 0.1185 |
| O0Zn30 | 30 | 56 | 0.0257 | nan | 0.1138 |
| O0Zn31 | 31 | 67 | 0.0129 | nan | 0.0612 |
| O0Zn32 | 32 | 87 | 0.017 | nan | 0.0529 |

| | | | | | |
|---------|----|------|--------|--------|--------|
| O10Zn10 | 20 | 1411 | 0.0102 | 0.102 | 0.0963 |
| O10Zn16 | 26 | 2604 | 0.0106 | 0.1175 | 0.13 |
| O10Zn24 | 34 | 334 | 0.0147 | 0.1578 | 0.1323 |
| O11Zn0 | 11 | 790 | 0.039 | 0.3809 | Nan |
| O11Zn15 | 26 | 28 | 0.0149 | 0.1697 | 0.1745 |
| O11Zn16 | 27 | 3017 | 0.0195 | 0.1426 | 0.1435 |
| O12Zn16 | 28 | 3339 | 0.026 | 0.1588 | 0.1545 |
| O13Zn15 | 28 | 231 | 0.0158 | 0.1383 | 0.1441 |
| O13Zn16 | 29 | 3143 | 0.0811 | 0.531 | 0.5044 |
| O14Zn15 | 29 | 53 | 0.0166 | 0.1467 | 0.1414 |
| O14Zn16 | 30 | 2765 | 0.1102 | 0.2619 | 0.2041 |
| O15Zn16 | 31 | 3234 | 0.015 | 0.1929 | 0.1596 |
| O16Zn14 | 30 | 341 | 0.0933 | 0.2363 | 0.1821 |
| O16Zn16 | 32 | 2337 | 0.0223 | 0.3803 | 0.3455 |
| O2Zn17 | 19 | 250 | 0.017 | 0.1463 | 0.1002 |
| O4Zn0 | 4 | 79 | 0.0509 | 0.5566 | Nan |
| O6Zn4 | 10 | 40 | 0.0811 | 0.3265 | 0.2949 |
| O6Zn6 | 12 | 1415 | 0.022 | 0.205 | 0.194 |
| O6Zn8 | 14 | 306 | 0.0266 | 0.1843 | 0.1749 |
| O7Zn16 | 23 | 33 | 0.0212 | 0.15 | 0.1378 |
| O7Zn8 | 15 | 309 | 0.0284 | 0.2006 | 0.1864 |
| O8Zn16 | 24 | 1129 | 0.0067 | 0.0606 | 0.0867 |
| O8Zn8 | 16 | 3588 | 0.0149 | 0.1055 | 0.0862 |
| O9Zn16 | 25 | 1445 | 0.0084 | 0.0871 | 0.1041 |
| O36Zn48 | 84 | 1428 | 0.0043 | 0.0772 | 0.0918 |
| O37Zn48 | 85 | 1254 | 0.029 | 0.099 | 0.1049 |
| O38Zn48 | 86 | 1433 | 0.0048 | 0.0915 | 0.1082 |
| O39Zn48 | 87 | 1452 | 0.0122 | 0.1012 | 0.1136 |
| O40Zn48 | 88 | 1428 | 0.0092 | 0.1156 | 0.1186 |
| O41Zn48 | 89 | 1357 | 0.0125 | 0.1386 | 0.1314 |
| O42Zn48 | 90 | 1414 | 0.0063 | 0.1295 | 0.126 |
| O43Zn48 | 91 | 1440 | 0.0144 | 0.1321 | 0.1287 |
| O44Zn48 | 92 | 1451 | 0.0225 | 0.1981 | 0.1351 |
| O45Zn48 | 93 | 1369 | 0.0178 | 0.1434 | 0.1369 |
| O46Zn48 | 94 | 1339 | 0.0191 | 0.1622 | 0.1319 |

| | | | | | |
|---------|-----|-------|--------|--------|--------|
| O47Zn48 | 95 | 1397 | 0.0215 | 0.167 | 0.1371 |
| O48Zn48 | 96 | 1359 | 0.0193 | 0.176 | 0.1431 |
| ALL | nan | 55764 | 0.0363 | 0.1917 | 0.1673 |

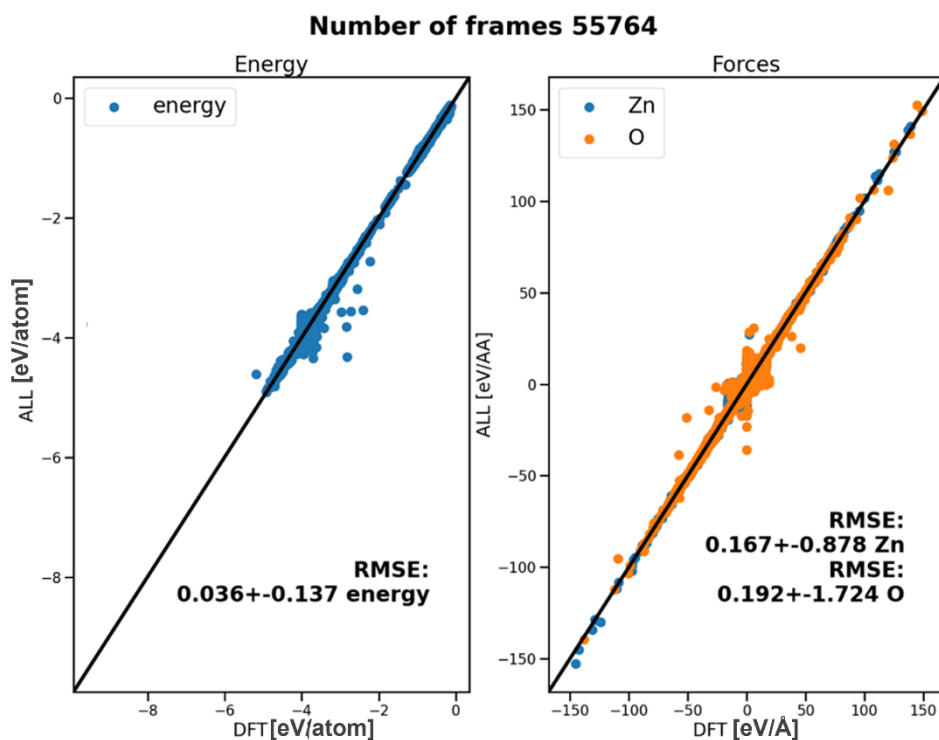


Figure S1. Energy and force RMSE on DFT-labelled structures. The horizontal and vertical axes of Figure S1a are the energies per atom (eV/atom) calculated from DFT and EANNP, respectively. The fitting error (RMSE) of the energy predicted by the EANNP is 0.036 eV/atom. The RMSE in forces is 0.192 eV/Å for O and 0.167 eV/Å for Zn. In the process of preparing the training data, some unstable structures with bond distances being relatively largely different from the optimized ones for training the potential.

1.2 Oxygen Chemical Potential (μ_O) Calculation

Table S2. Calculated energy, the thermal correction to Gibbs Free energy and Gibbs free energy (eV).

| | E_{DFT} | ZPE | $U(T)$ | $TS(T)$ | $k_B T \ln\left(\frac{P}{P_0}\right)$ | $\mu(T, P)$ |
|------------------------------------|------------------|------|--------|---------|---------------------------------------|-------------|
| H ₂ O(673 K, 0.01 Kpa) | -14.22 | 0.58 | 0.25 | 1.53 | -0.53 | -15.45 |
| H ₂ (673 K, 1.5 Mpa) | -6.76 | 0.27 | 0.21 | 1.09 | 0.16 | -7.20 |
| CO ₂ (673K, 0.0765 Mpa) | -22.99 | 0.32 | 0.28 | 1.75 | -0.02 | -24.16 |
| CO(673 K, 1 Mpa) | -14.80 | 0.14 | 0.21 | 1.56 | 0.13 | -15.88 |

Supplemental Note S2. For H₂-assisted OV generation, the pressure of H₂O is set to 0.01 Kpa according to the experiment.³⁻⁴ For CO-assisted OV generation, the equilibrium pressure of CO₂ is evaluated at 673 K, 0.0765 Mpa, according to the experimental results (CO conversion at ~17% and CO₂ selectivity at ~45%).⁵ Based on the above calculations for μ_O under the experimental conditions (-8.25 eV and -8.28 eV when the surface is equilibrium with H₂ and CO, respectively). The μ_O under the standard condition (298.15K, 0.1 Mpa) is -5.15eV. Therefore, the $\Delta\mu_O = -3.10$ eV and -3.13 eV when the surface is equilibrium with H₂ and CO, respectively.

2. Stability of the ZnO Surfaces

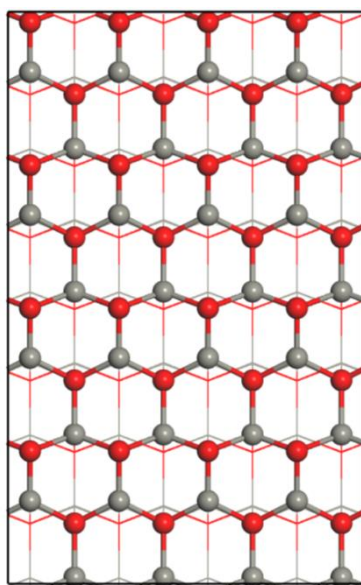


Figure S2. Top view of p(4x4) ZnO surface.

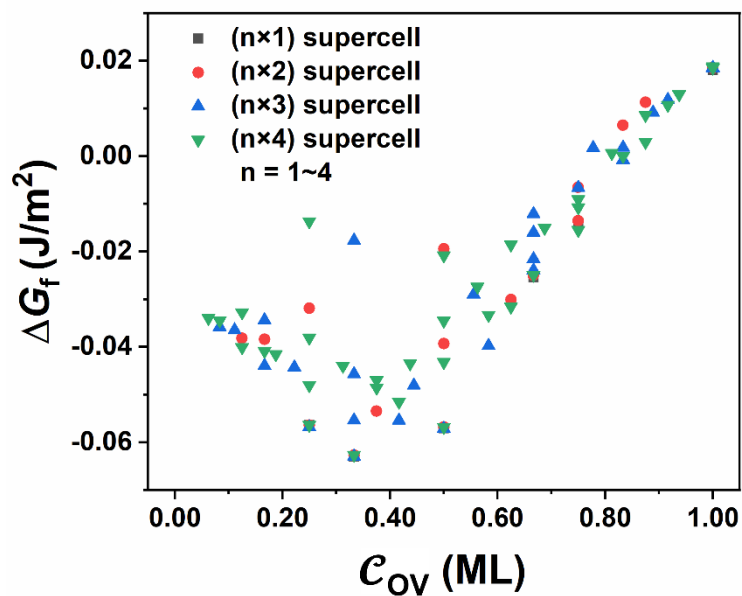


Figure S3. OV formation energies of ZnO (10 $\bar{1}$ 0) at different C_{OV} referring to the perfect ZnO(10 $\bar{1}$ 0) surface under the CO reduction atmosphere (673K, 1Mpa).

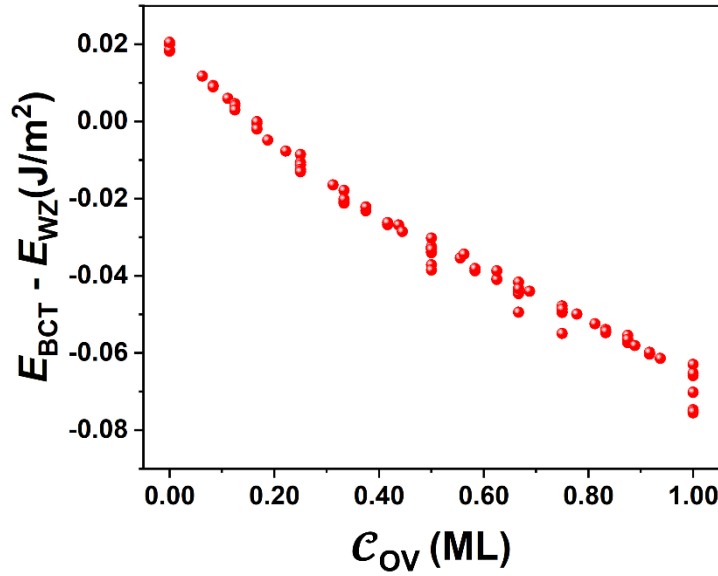


Figure S4. The energy difference between the BCT surface and WZ surface as a function of OV concentrations.

Supplemental Note S3. To take the unit cells of different sizes and symmetries into consideration, we have explored the surfaces with a number of OVs using unit cells of $p(1 \times 1)$, $p(2 \times 1)$, $p(3 \times 1)$, $p(4 \times 1)$, $p(1 \times 2)$, $p(2 \times 2)$, $p(3 \times 2)$, $p(4 \times 2)$, $p(1 \times 3)$, $p(2 \times 3)$, $p(3 \times 3)$, $p(4 \times 3)$, $p(1 \times 4)$, $p(2 \times 4)$, $p(3 \times 4)$, $p(4 \times 4)$. Firstly, the largest unit cell, $p(4 \times 4)$ ZnO surface with 256 numbers of atoms, is large enough, containing 4 O atoms in each row and column on the top layer, respectively, as shown in **Figure S2**. Secondly, they cover a range of surfaces with different symmetries. These warrant that the data contain general features of the ZnO surface. To ensure the energy convergency of models of various sizes, high density K-points were used. The KSPACING was set to 0.167 \AA^{-1} . In order to compare the OV formation energies from different unit sizes, they are converted to the unit of J m^{-2} . As shown in **Figure S3**, the surface with 0.33 ML OVs is the most likely surface under the experimental conditions. And with the increasing concentration of OVs, the energy difference between the BCT surface and WZ surface increase in **Figure S4**.

3. Geometric and Electronic Structures of the ZnO Surfaces

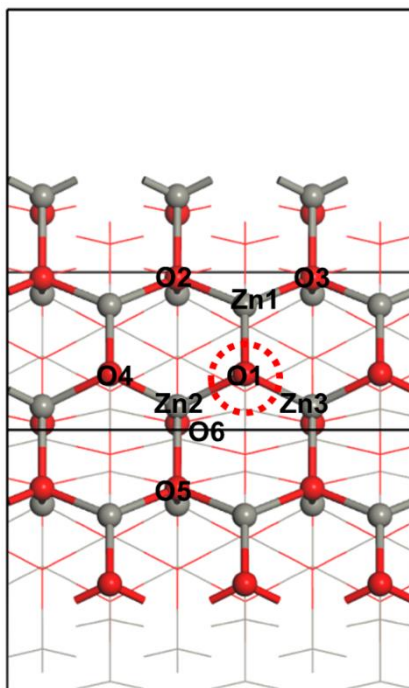


Figure S5. Structure of the optimized BCT surface. Key bond lengths and Zn-Zn pair distances before and after introducing an OV defect at the O1 site are listed in **Table S3**.

Table S3. Computed key atomic distances (\AA) for the defect-free BCT and WZ before and after introducing an OV. The labels of atoms are shown in **Figure S5**.

| | BCT | BCT with 0.17 ML OVs | Δ BCT | WZ | WZ with 0.17 ML OVs | Δ WZ |
|---------|------|-------------------------|--------------|------|------------------------|-------------|
| Zn1-Zn2 | 2.93 | 2.50 | -0.43 | 2.97 | 2.57 | -0.40 |
| Zn1-Zn3 | 2.94 | 2.51 | -0.43 | 2.97 | 2.57 | -0.40 |
| Zn2-Zn3 | 3.25 | 2.95 | -0.31 | 3.25 | 3.03 | -0.22 |
| Zn2-O1 | 1.90 | nan | nan | 1.92 | nan | nan |
| Zn2-O4 | 1.90 | 1.94 | 0.03 | 1.92 | 1.94 | 0.02 |
| Zn2-O5 | 2.00 | 2.00 | 0.00 | 1.99 | 1.99 | 0.00 |
| Zn2-O6 | 2.08 | 2.07 | -0.01 | 2.07 | 2.06 | -0.01 |
| Zn1-O1 | 1.84 | nan | nan | 1.85 | nan | Nan |
| Zn1-O2 | 1.90 | 1.95 | 0.05 | 1.92 | 1.99 | 0.07 |
| Zn1-O3 | 1.90 | 1.95 | 0.05 | 1.92 | 1.98 | 0.06 |

4. Catalytic Activities of the Body-centered-tetragonal (BCT) and Wurtzite (WZ) Surfaces with Various Concentrations of OVs

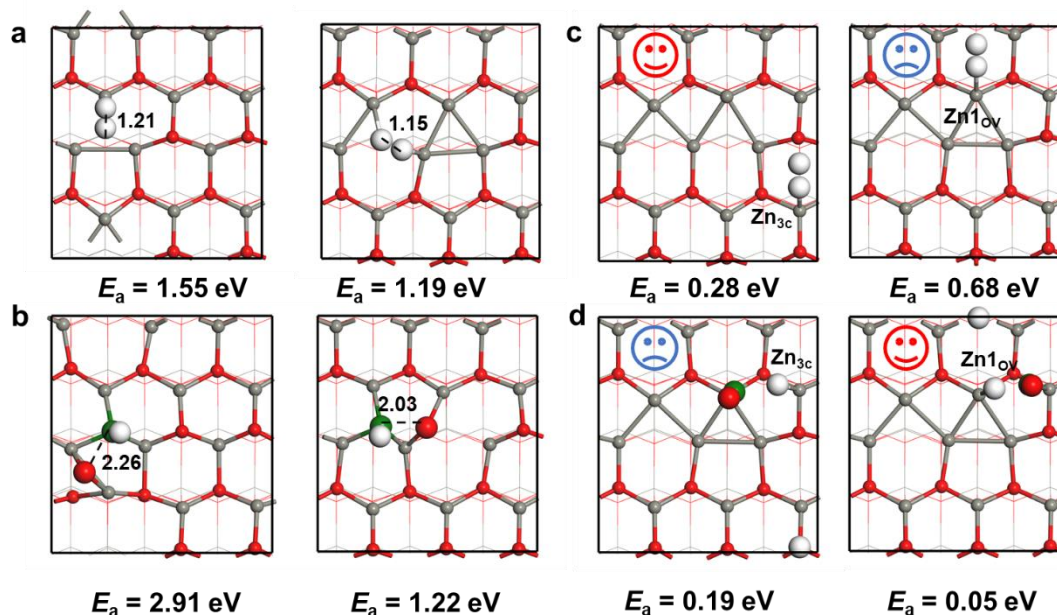


Figure S6. The configurations of the H_2 homolytic dissociation (a) and $CH-O$ dissociation (b) on model I and model III. (c) The configurations of the heterolysis I on the $Zn_{3c}\text{-}O_{3c}$ and $Zn_{1OV}\text{-}O_{3c}$ sites. (d) The configurations of the CO hydrogenation hetero_H_I on $Zn_{1OV}\text{-}Zn_{3c}$ and $Zn_{3c}\text{-}Zn_{1OV}$ sites.

Table S4. Calculated the adsorption energy of H_2 and CO on BCT and WZ surface with various concentrations of OVs.

| | BCT surface | | WZ surface | |
|---------|-------------|------------|------------|------------|
| | CO (eV) | H_2 (eV) | CO (eV) | H_2 (eV) |
| 0.00 ML | -0.27 | -0.05 | -0.38 | -0.13 |
| 0.33 ML | -0.23 | -0.04 | -0.31 | -0.08 |
| 0.67 ML | -0.21 | -0.04 | -0.27 | -0.06 |
| 1.00 ML | -0.20 | -0.01 | -0.33 | -0.06 |

Table S5. Calculated energy barriers and enthalpy changes of H₂ dissociation on BCT surface with various concentrations of OVs.

| | Heterolysis I | | Heterolysis II | | Homolysis | |
|---------|---------------|-----------------|----------------|-----------------|------------|-----------------|
| | E_a (eV) | ΔH (eV) | E_a (eV) | ΔH (eV) | E_a (eV) | ΔH (eV) |
| 0.00 ML | 0.29 | 0.06 | 0.58 | -0.20 | | |
| 0.33 ML | 0.36 | 0.07 | 0.59 | -0.17 | 1.55 | -0.03 |
| 0.67 ML | 0.33 | 0.05 | 0.58 | -0.16 | 1.10 | -0.15 |
| 1.00 ML | | | | | 0.95 | -0.09 |

Table S6. Calculated energy barriers and enthalpy changes of H₂ dissociation on WZ surface with various concentrations of OVs.

| | Heterolysis I | | Heterolysis II | | Homolysis | |
|---------|---------------|-----------------|----------------|-----------------|------------|-----------------|
| | E_a (eV) | ΔH (eV) | E_a (eV) | ΔH (eV) | E_a (eV) | ΔH (eV) |
| 0.00 ML | 0.19 | -0.21 | 0.54 | -0.45 | | |
| 0.33 ML | 0.21 | -0.15 | 0.53 | -0.38 | 1.39 | -0.33 |
| 0.67 ML | 0.24 | -0.11 | 0.51 | -0.32 | 1.07 | -0.39 |
| 1.00 ML | | | | | 0.92 | -0.35 |

Table S7. Calculated energy barriers of CO hydrogenation on BCT and WZ surface with various concentrations of OVs.

| | BCT surface | | | | WZ surface | | | |
|---------|---------------------------|------------------------------|---------------------------|------------------------------|---------------------------|------------------------------|---------------------------|------------------------------|
| | Hetero | Hetero | Homo | Homo | Hetero | Hetero | Homo | Homo |
| | \underline{H}_I (eV) | \underline{H}_{II} (eV) | \underline{H}_I (eV) | \underline{H}_{II} (eV) | \underline{H}_I (eV) | \underline{H}_{II} (eV) | \underline{H}_I (eV) | \underline{H}_{II} (eV) |
| 0.00 ML | 0.19 | 0.71 | | | 0.19 | 0.79 | | |
| 0.33 ML | 0.18 | 0.62 | 0.30 | 0.86 | 0.14 | 0.68 | 0.28 | 0.70 |
| 0.67 ML | 0.20 | 0.71 | 0.13 | 0.53 | 0.16 | 0.72 | 0.12 | 0.56 |
| 1.00 ML | | | 0.15 | 0.49 | | | 0.17 | 0.39 |

Table S8. Calculated energy barriers of CO breakage on BCT and WZ surface with various concentrations of OVs.

| | BCT surface | | WZ surface | |
|---------|-----------------------|------------------------|-----------------------|------------------------|
| | C-O dissociation (eV) | CH-O dissociation (eV) | C-O dissociation (eV) | CH-O dissociation (eV) |
| 0.33 ML | 5.31 | 2.91 | 4.42 | 2.47 |
| 0.67 ML | 3.06 | 1.05 | 2.71 | 0.81 |
| 1.00 ML | 2.64 | 0.90 | 2.44 | 0.71 |

Supplemental Note S4. We have also investigated the catalytic activities of the WZ surface with various concentrations of OVs, including H₂/CO adsorption, H₂ dissociation, CO hydrogenation and C-O dissociation.

4.1 H₂/CO Adsorption

As shown in **Table S4**, CO shows stronger chemical adsorption than H₂. With the increasing of the OVs, the adsorption of CO on the WZ surface decreases, which is in line with the report of the BCT surface in the main text. It is worth noting that the WZ surface has stronger adsorption capacity for CO and H₂ than BCT surface because of the higher d band center of the WZ surface (-5.77 eV vs. -6.03 eV).

4.2 H₂ Dissociation

As shown in **Table S6**, with the increase of OV concentration, the reactivity of hetero-cleavage of H₂ at the intrinsic site hardly changed, but the reactivity of homo-cleavage of H₂ at the newly generated OV site increased gradually. Similarly, the WZ surface show better H₂ dissociation activity than the BCT surface.

4.3 CO Hydrogenation

As shown in **Table S7**, with the increase of OV concentration, the hydrogenation reactivity with H from heterolysis at the intrinsic reactive sites(O_{3c} and Zn_{3c}) is hardly affected while that from homolysis at the newly generated OV reactive site increases. Compared to the BCT surface, better CO hydrogenation reactivity can be seen except for the H sitting on the oxygen site.

4.4 C-O Activation

As shown in **Table S8**, the C-O bond is more easily broken as the concentration of OVs increases. A better C-O bond activation activity can be seen on the WZ surface.

Overall, as the concentration of OVs increases, a similar behavior trend can be seen on the BCT and WZ surface. That is, the activity occurring on the intrinsic site changes little while the activity occurring on the OV gradually increases. The metastable WZ surface shows better catalytic activity compared to the BCT surface except for the CO hydrogenation reactivity with the hydrogen adsorbed on the oxygen.

Optimized XYZ position for the transition states of H₂ heterolysis I on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

1.0000000000000000
9.7516200000000008 0.0000000000000000 0.0000000000000000
0.0000000000000000 10.4606600000000007 0.0000000000000000
0.0000000000000000 0.0000000000000000 24.3835499999999996

H O Zn

2 46 48

Direct

0.5794563691056167 0.4072243412552114 0.4283745722326627
0.5803411322933714 0.3257721769429525 0.4483508748627591
0.2534539999999978 0.2088630000000009 0.0123030000000028
0.5867870000000012 0.2088630000000009 0.0123030000000028
0.9201199999999973 0.2088630000000009 0.0123030000000028
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0.4201199999999972 0.4588630000000009 0.0507869999999997
0.7534530000000004 0.4588630000000009 0.0507869999999997
1.0867859999999965 -0.0411369999999992 0.0507869999999997
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0.4208134764279651 0.2115195462046242 0.1287308552986575
0.7534547804362631 0.2115039897017647 0.1289450269139011
1.0857266100692891 0.7116739086357894 0.1291582807417617

0.4210128864851237 0.7116588378383538 0.1289482468694068
0.7532811174434554 0.7116234028995488 0.1291915977560263
0.5871111436142508 0.4619982790967059 0.1652254083064379
0.2537269999967597 0.4613986733292224 0.1637990076148176
0.9198248807479761 0.4618134794603986 0.1654351576404245
0.9196039554254717 -0.0383203333811766 0.1656691257483322
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0.5876733428343976 0.7104378215349283 0.2435501022680302
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0.4204873636575015 0.4624617624287893 0.2787811645466054
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0.2536158129392454 0.2721949310747948 0.1645375146157197
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0.9192432019690279 0.7724320985588575 0.1660316783460270
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0.2536039620794369 0.1542520680443110 0.3737040970029016
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0.5820288378969917 0.6835535272770403 0.3820884312891855
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Optimized XYZ position for the transition states of H₂ heterolysis II on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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0.0000000000000000 0.0000000000000000 24.3835499999999996

H O Zn

2 46 48

Direct

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Optimized XYZ position for the transition states of H₂ homolysis on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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0.0000000000000000 0.0000000000000000 24.3835499999999996

H O Zn

2 46 48

Direct

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Optimized XYZ position for the transition states of hetero_H_I on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

1.0000000000000000
9.7516200000000008 0.0000000000000000 0.0000000000000000
0.0000000000000000 10.4606600000000007 0.0000000000000000
0.0000000000000000 0.0000000000000000 24.3835499999999996

H C O Zn
2 1 47 48

Direct

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Optimized XYZ position for the transition states of hetero_H_II on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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H C O Zn
1 1 47 48

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Optimized XYZ position for the transition states of homo_H_I on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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H C O Zn
2 1 47 48

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Optimized XYZ position for the transition states of homo_H_II on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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H C O Zn
1 1 47 48

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Optimized XYZ position for the transition states of C-O breakage on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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C O Zn

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Optimized XYZ position for the transition states of CH-O breakage on ZnO BCT surface with 0.33 ML OVs

AutoCreatByScript

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H C O Zn
1 1 47 48

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