

Solvent-Assisted Crystallization of α -Fe₂O₃ Electron Transport Layer for Efficient and Stable Perovskite Solar Cells Featuring Negligible Hysteresis

Akbar Ali Qureshi ^{a, b, c}, Sofia Javed* ^a, Muhammad Aftab Akram ^d, Lukas Schmidt-Mende* ^c, Azhar Fakharuddin* ^c

^a School of Chemical & Materials Engineering, National University of Sciences & Technology, 44000, Islamabad, Pakistan

^b Department of Mechanical Engineering, Bahauddin Zakariya University, 60000, Multan, Pakistan

^c Department of Physics, University of Konstanz, 78464, Konstanz, Germany

^d Department of Materials Science & Engineering, Pak-Austria Fachhochschule, Institute of Applied Sciences & Technology, 22650, Haripur, Pakistan

Correspondence: SJ (sofia.javed@scme.nust.edu.pk), LSM (Lukas.Schmidt-Mende@uni-konstanz.de), AF (Lukas.Schmidt-Mende@uni-konstanz.de).

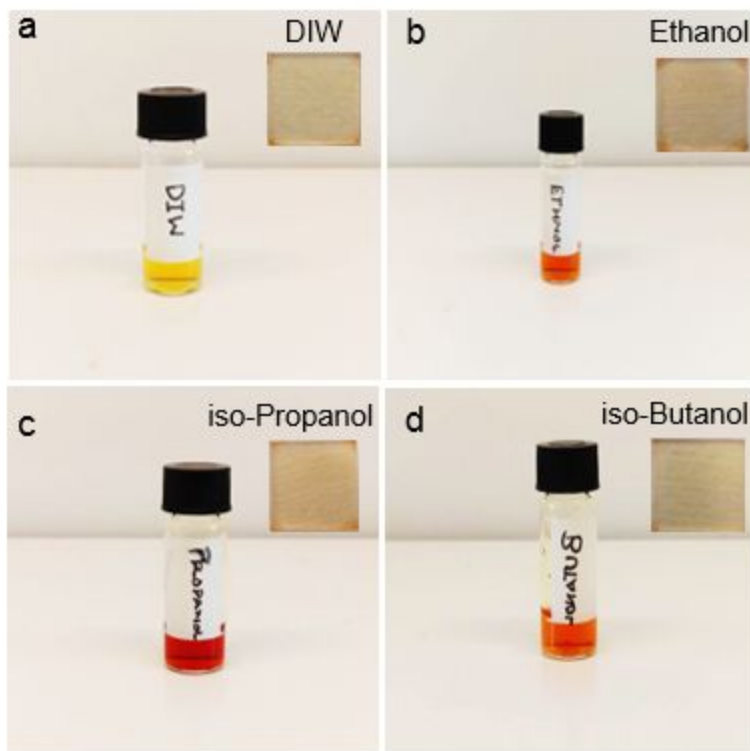


Figure S1. α -Fe₂O₃ dispersion in different solvents: (a) deionized water, (b) ethanol, (c) iso-propanol, and (d) iso-butanol. (Inset shows thin films annealed @ 300 °C for each solvent)

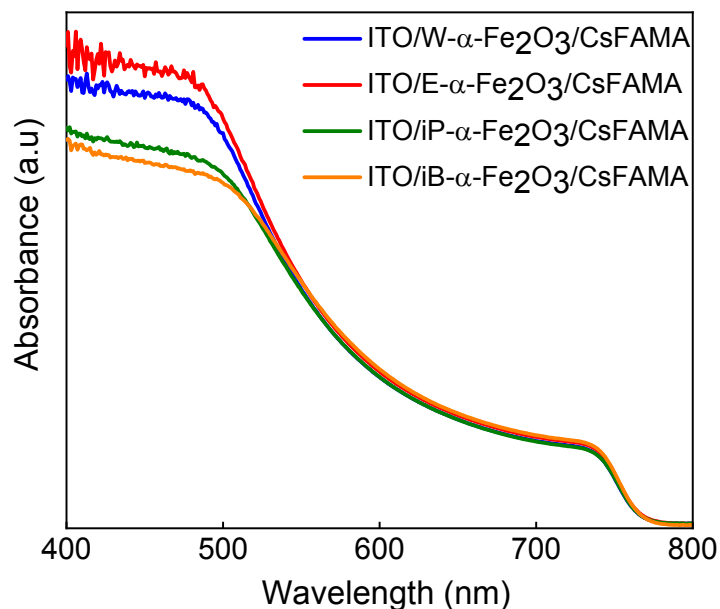


Figure S2. Absorption spectra of CsFAMA films on ITO/ α -Fe₂O₃ ETL with different solvents,

Figure S4 exhibits the absorption spectra of CsFAMA onto W- α -Fe₂O₃, E- α -Fe₂O₃, iP- α -Fe₂O₃, and iB- α -Fe₂O₃, ETLs. The spectra exhibited almost identical absorption over the spectral range of 500 to 800 nm for different α -Fe₂O₃ ETLs. The E- α -Fe₂O₃ and W- α -Fe₂O₃ ETLs manifested significantly enhanced absorption in the spectral range of 400 to 500 nm as compared to other ETLs. The slight difference in shorter wavelengths can be due to parasitical absorption or dense surface coverage.

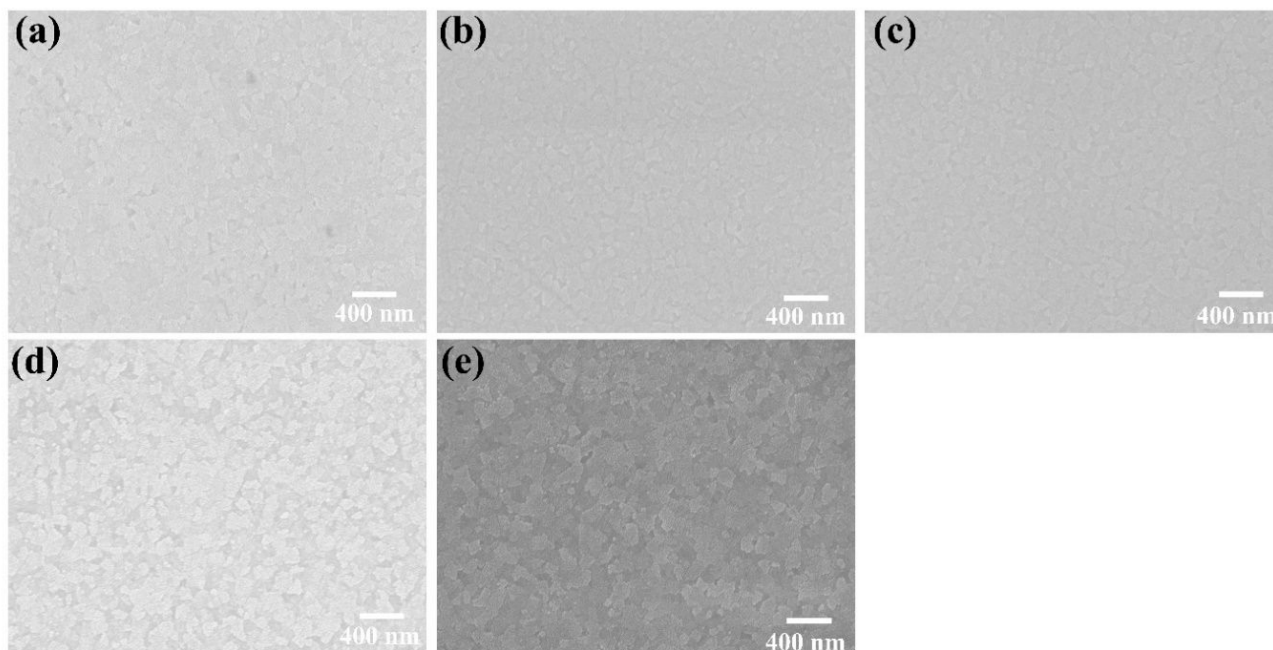


Figure S3. Top view SEM image of ITO/ α -Fe₂O₃ thin films in different solvents: (a) W- α -Fe₂O₃, (b) E- α -Fe₂O₃, (c) iP- α -Fe₂O₃, (d) iB- α -Fe₂O₃, and (e) pristine ITO.

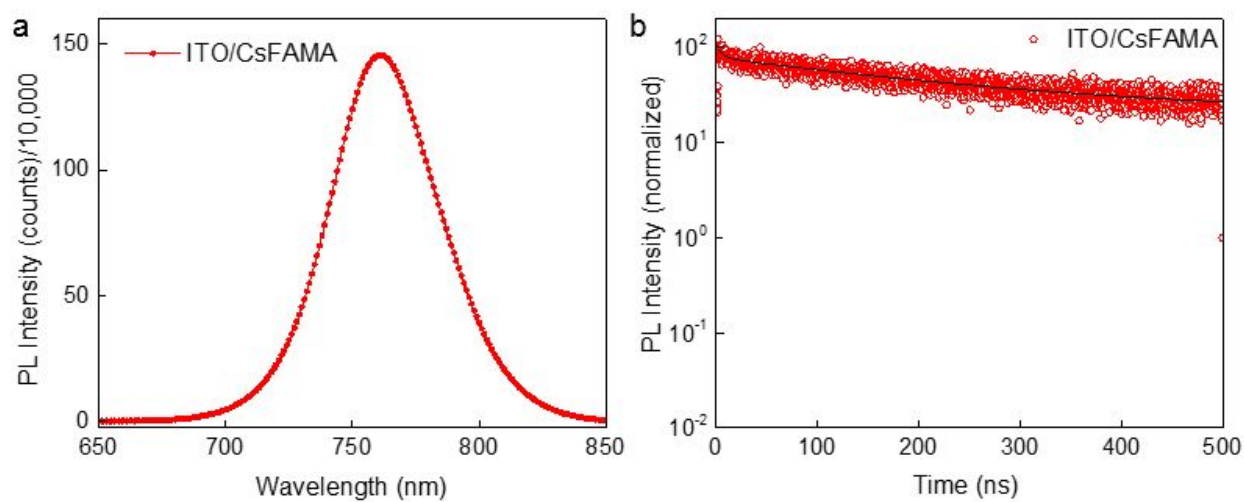


Figure S4. (a) Steady-state PL spectra of CsFAMA film on ITO substrate, and (b) time-resolved PL transients of the same.

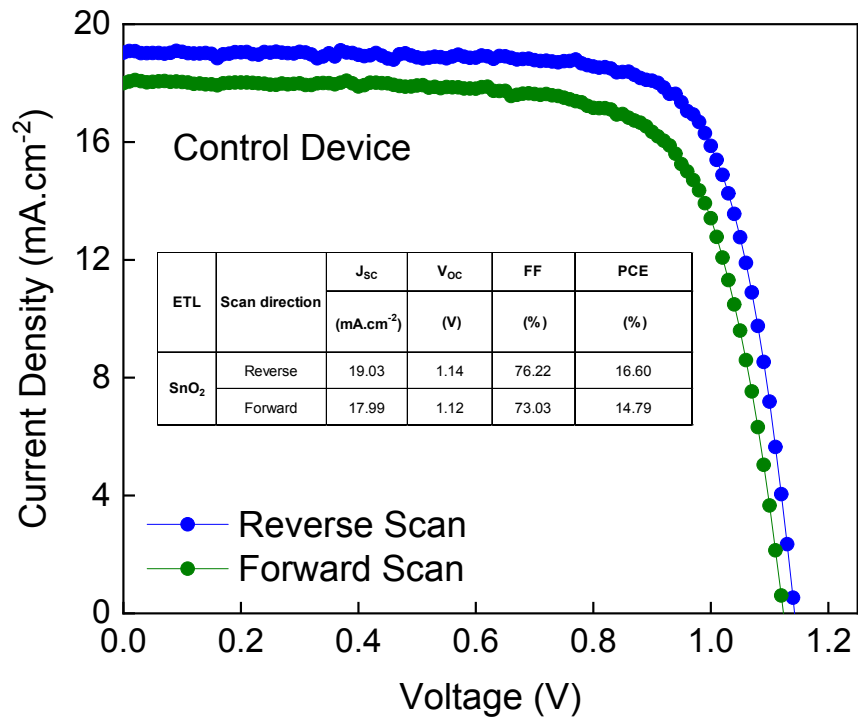


Figure S2. Light J-V curves of a control device with SnO₂ ETL.

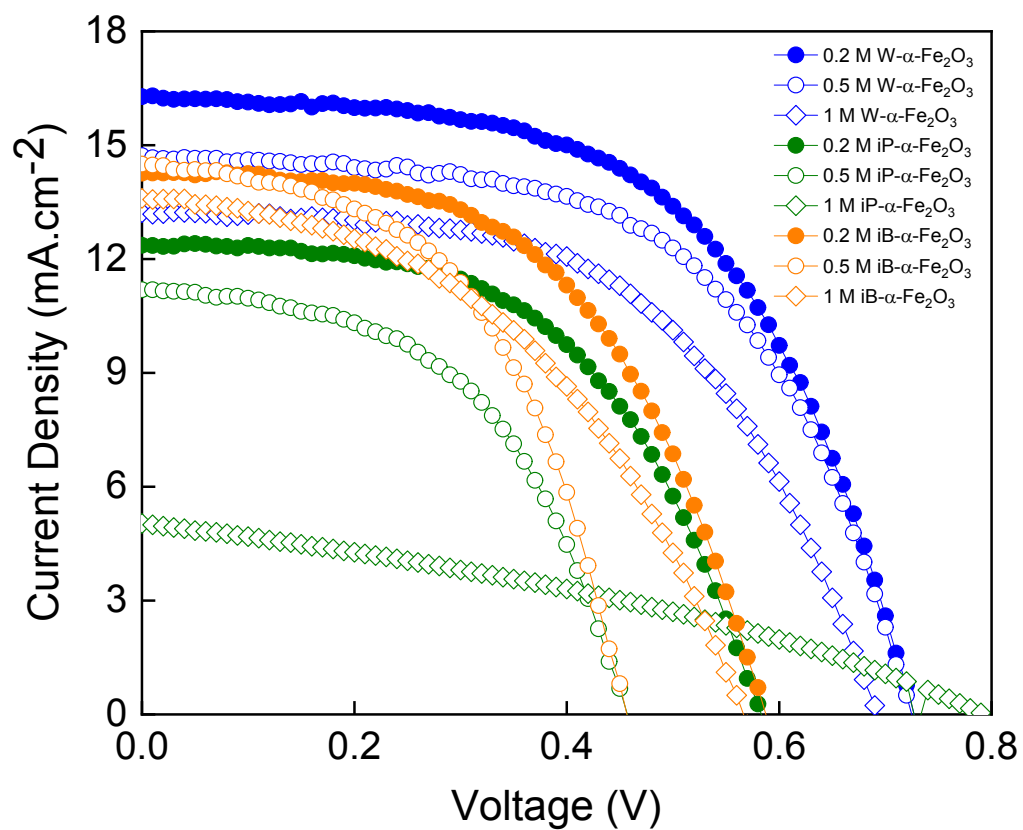


Figure S6. Light J-V curves of perovskite solar cells with α -Fe₂O₃ ETLs in different solvents and concentrations.

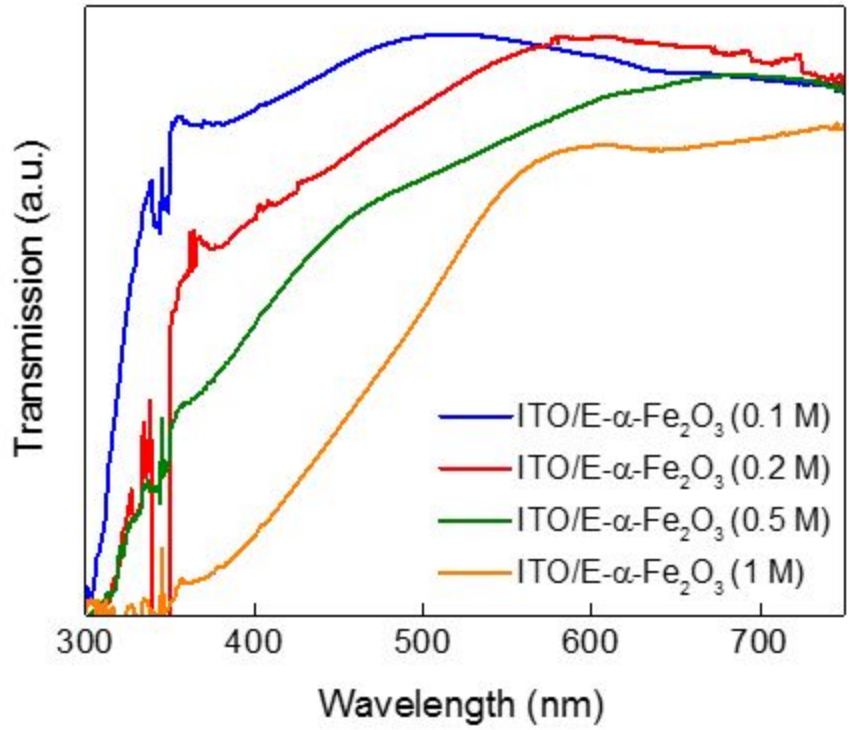


Figure S7. Transmission spectra of ITO/E- α -Fe₂O₃ thin films with different concentrations.

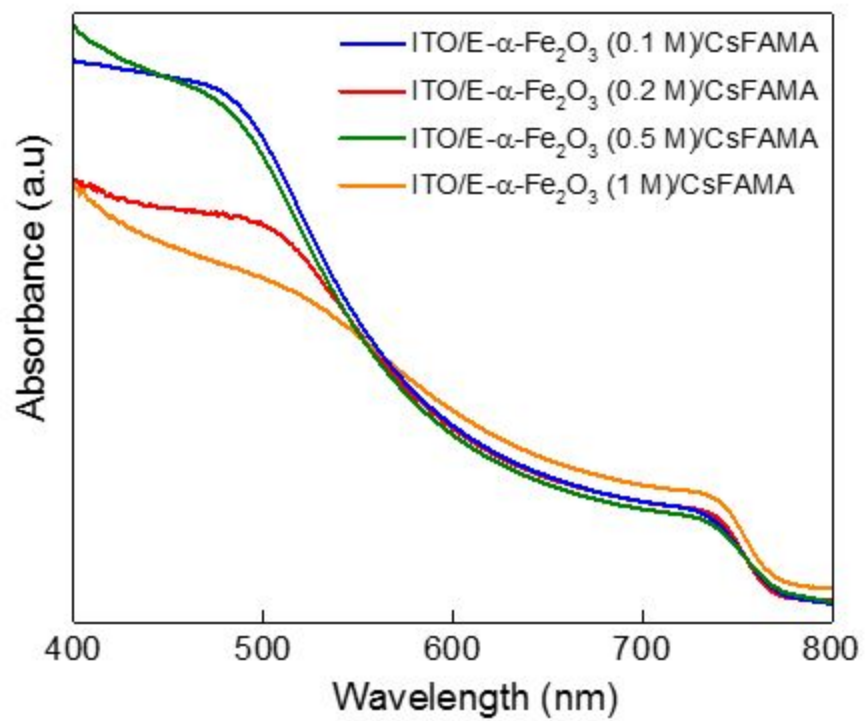


Figure S8. Absorption spectra of ITO/E- α -Fe₂O₃ thin films with different concentrations.

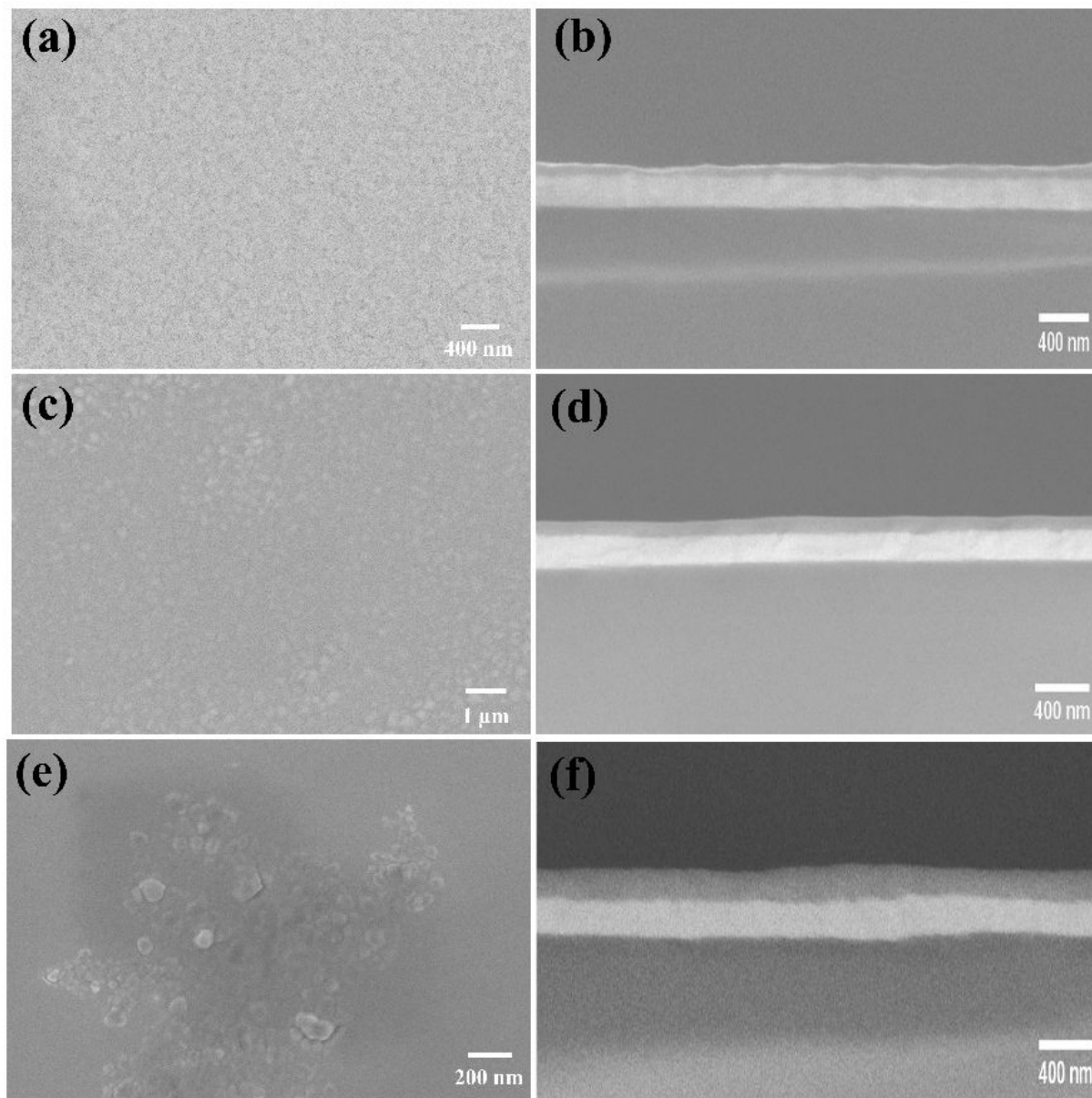


Figure S9. Top view SEM images of E- α -Fe₂O₃ compact layers deposited on ITO with concentrations: (a) 0.2 M, (c) 0.5 M, and (e) 1 M, Cross-sectional SEM images E- α -Fe₂O₃ compact layers deposited on ITO with concentrations: (b) 0.2 M, (d) 0.5 M, and (f) 1 M.

In order to confirm the reproducibility of our results, we fabricated at least 12 devices for each molar concentration and measured their photovoltaic performance (Figure S10). The statistical analysis showed a narrow distribution of the photovoltaic parameters as well as the highest performance for 0.2 M E- α -Fe₂O₃ ETL confirming the reproducibility and reliability of our results.

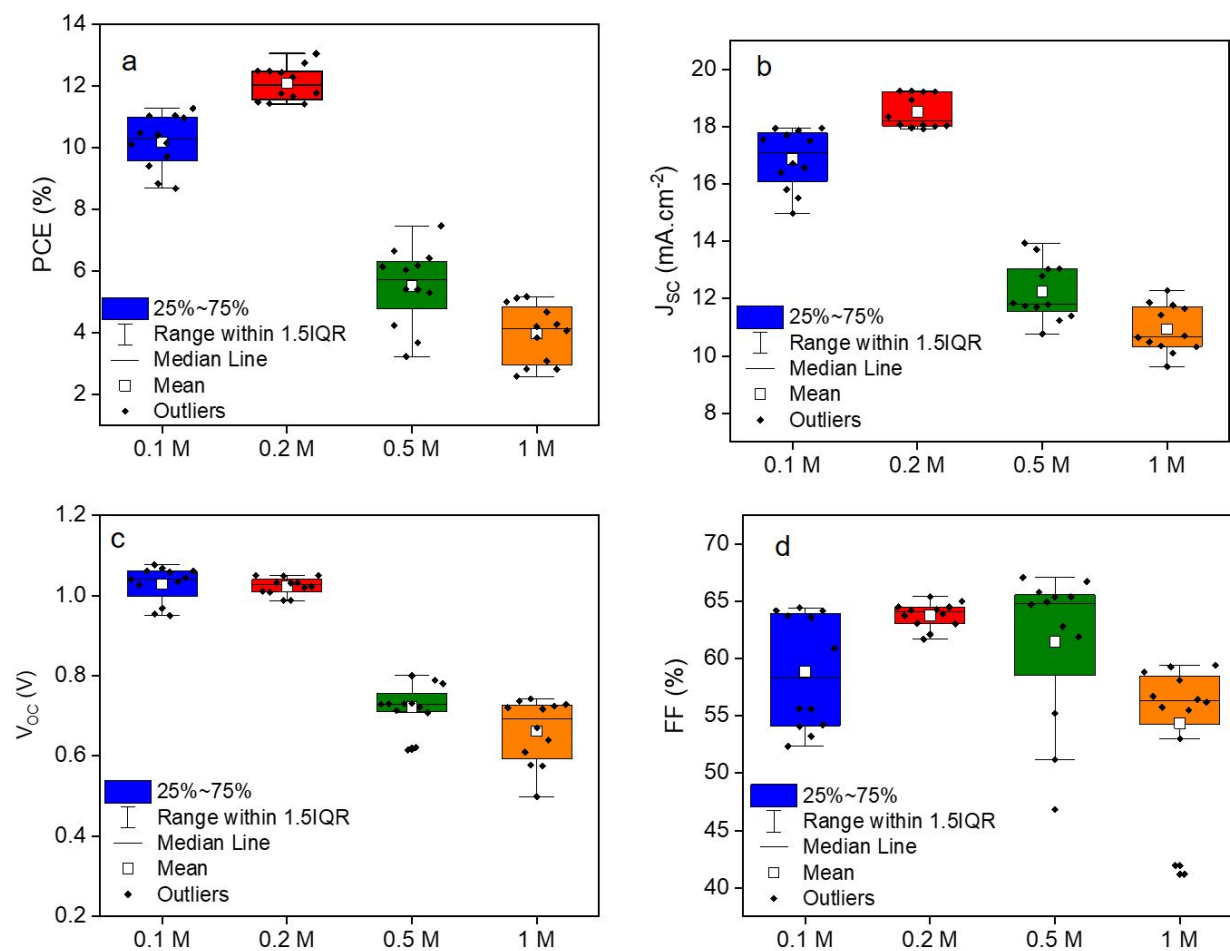


Figure S3. Statistical analysis of the photovoltaic parameters of the PSCs made using different molar concentrations of the α -Fe₂O₃ ETL in ethanol): (a) PCE, (b) J_{sc}, (c) V_{oc}, and (d) FF.

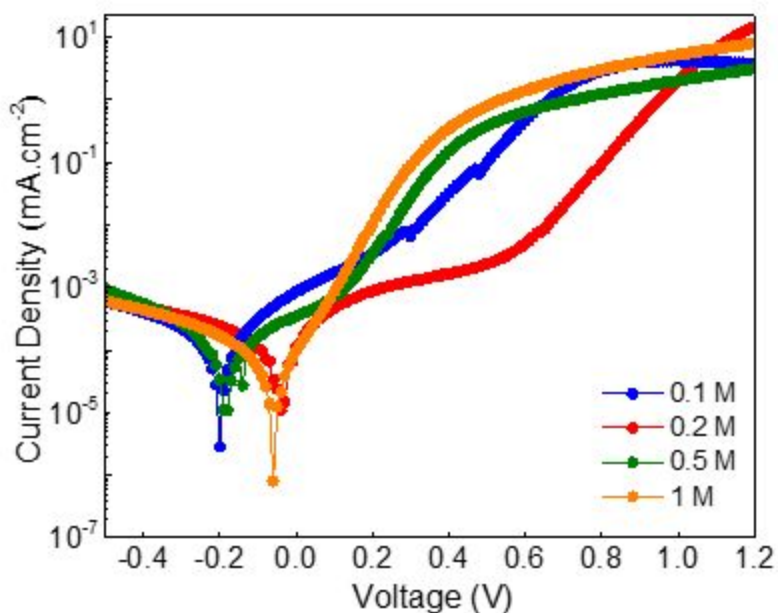


Figure S4. Dark J-V curves of perovskite solar cells with E- α -Fe₂O₃ ETL in different concentrations.

Figure S12a depicts the top view SEM image of CsFAMA perovskite deposited over-optimized 0.2 M E- α -Fe₂O₃ ETL. The dense and crack-free surface coverage of the perovskite film was observed for the 0.2 M E- α -Fe₂O₃ ETL. S12b and Figure S13 show the top view AFM image of CsFAMA perovskite deposited over-optimized 0.2 M E- α -Fe₂O₃ ETL. The large size grains are evident from AFM images with clearly visible grain boundaries. The longitudinal fusion of grain boundaries and vertical growth of the perovskite as in Figure S12c confirmed the better electronic extraction and transport from perovskite to ETL interface. Figure S12d depicts the cross-sectional SEM image of a 0.2 M E- α -Fe₂O₃ ETL-based perovskite solar cell. We further examined the morphology of perovskite for each concentration of E- α -Fe₂O₃ ETL (Figure S14). The crack-free dense morphology was observed for each ETL affirming the superiority and reproducibility of triple cation perovskite.

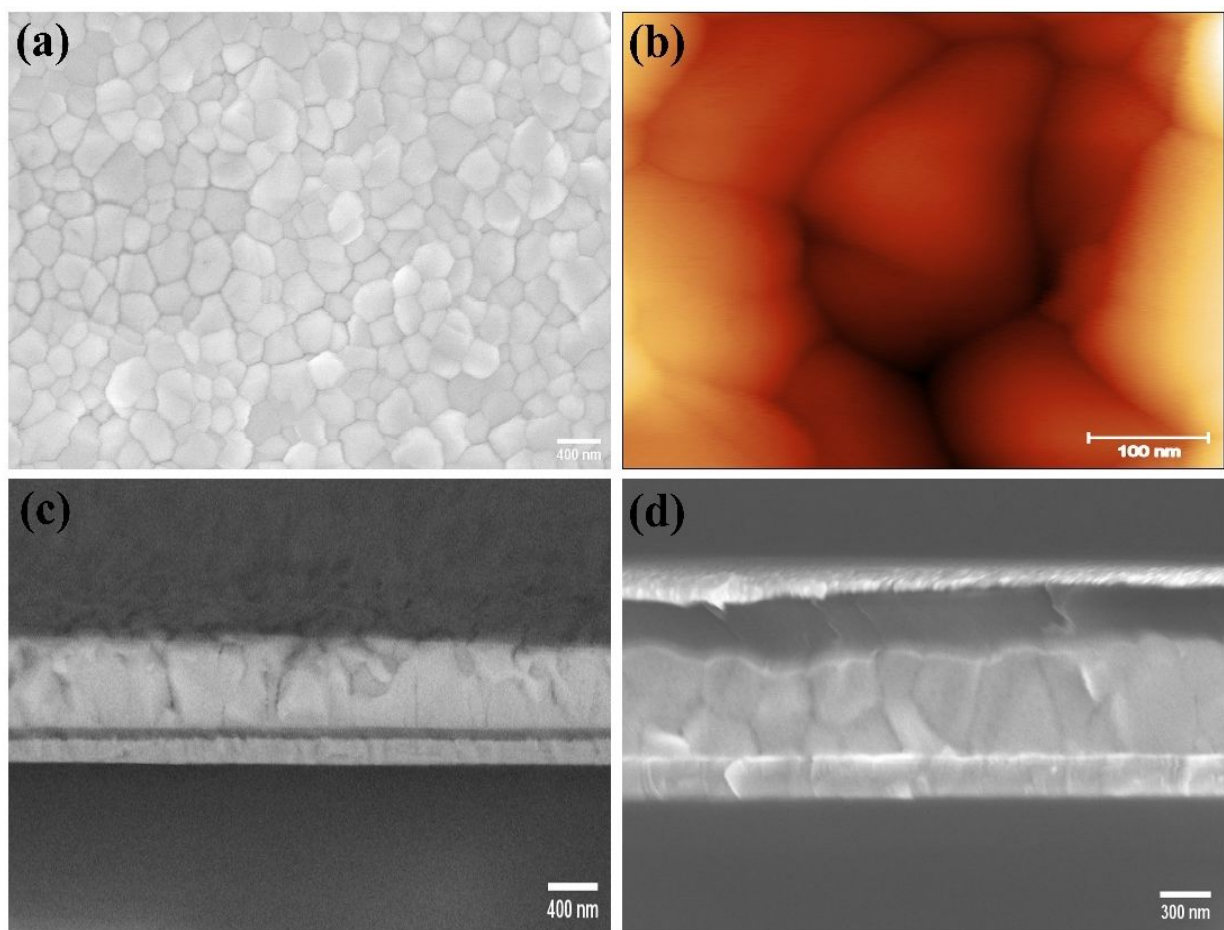


Figure S5. (a) Top view SEM image of CsFAMA perovskite thin films deposited on ITO/E- α -Fe₂O₃ ETL, (b) Top view AFM image of CsFAMA perovskite thin films deposited on ITO/E- α -Fe₂O₃ ETL, (c) Cross-sectional SEM image of CsFAMA perovskite thin films deposited on ITO/E- α -Fe₂O₃ ETL (0.5 M), and (d) cross-sectional SEM image of a full device with optimized solvent and optimized concentration

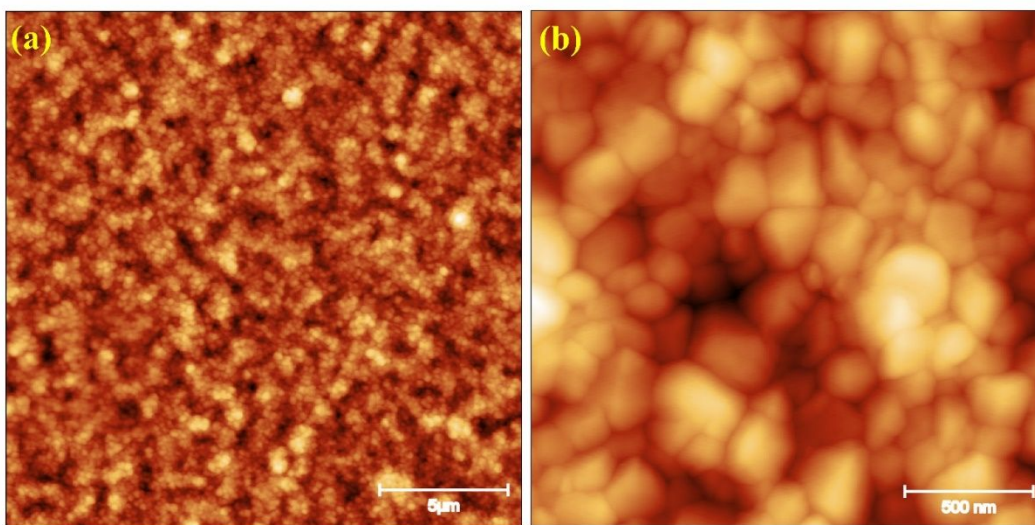


Figure S6. Top view AFM images of CsFAMA perovskite thin films deposited on ITO/E- α -Fe₂O₃ ETL in different resolutions, (a) 5 μ m, and (b) 500 nm.

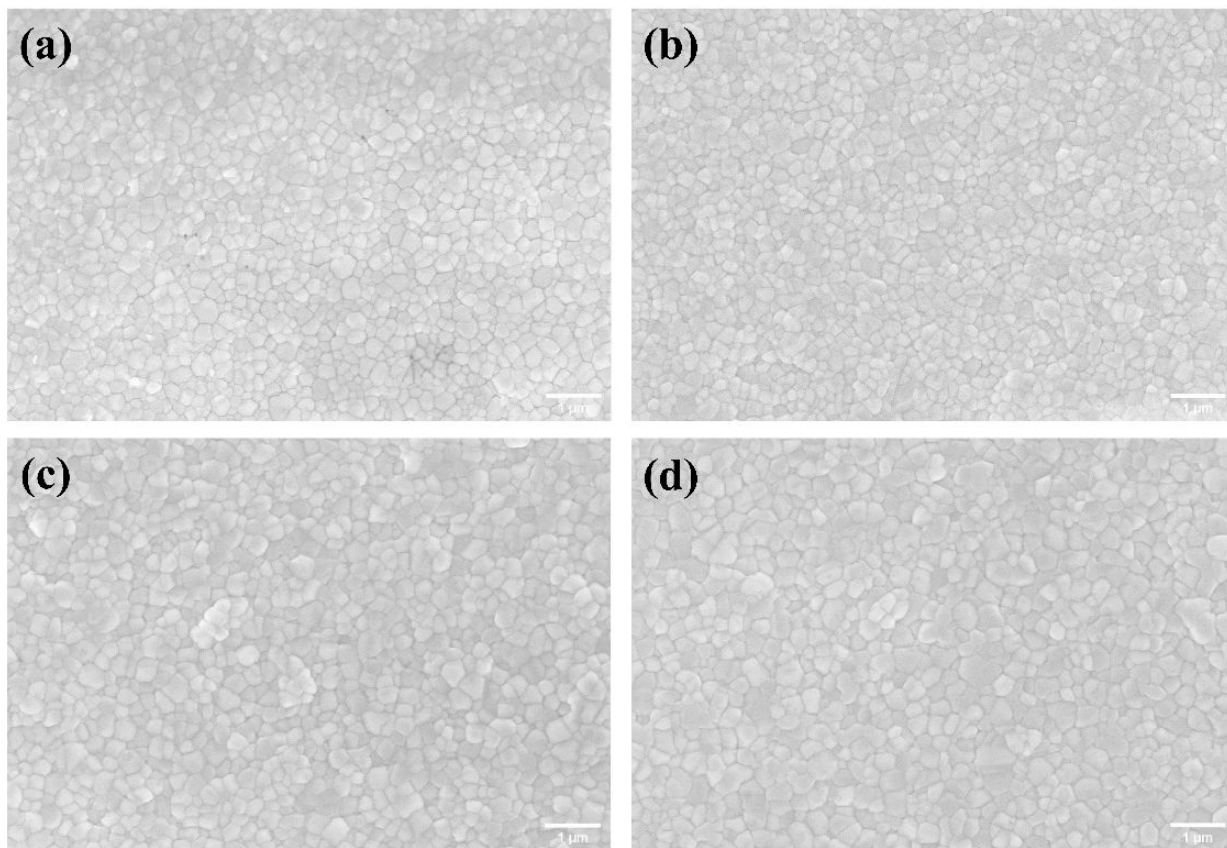


Figure S7. Top view SEM images of CsFAMA perovskite deposited on E- α -Fe₂O₃ ETLs in different concentrations, (a) 0.1 M, (b) 0.2 M, (c) 0.5 M, and (d) 1 M.

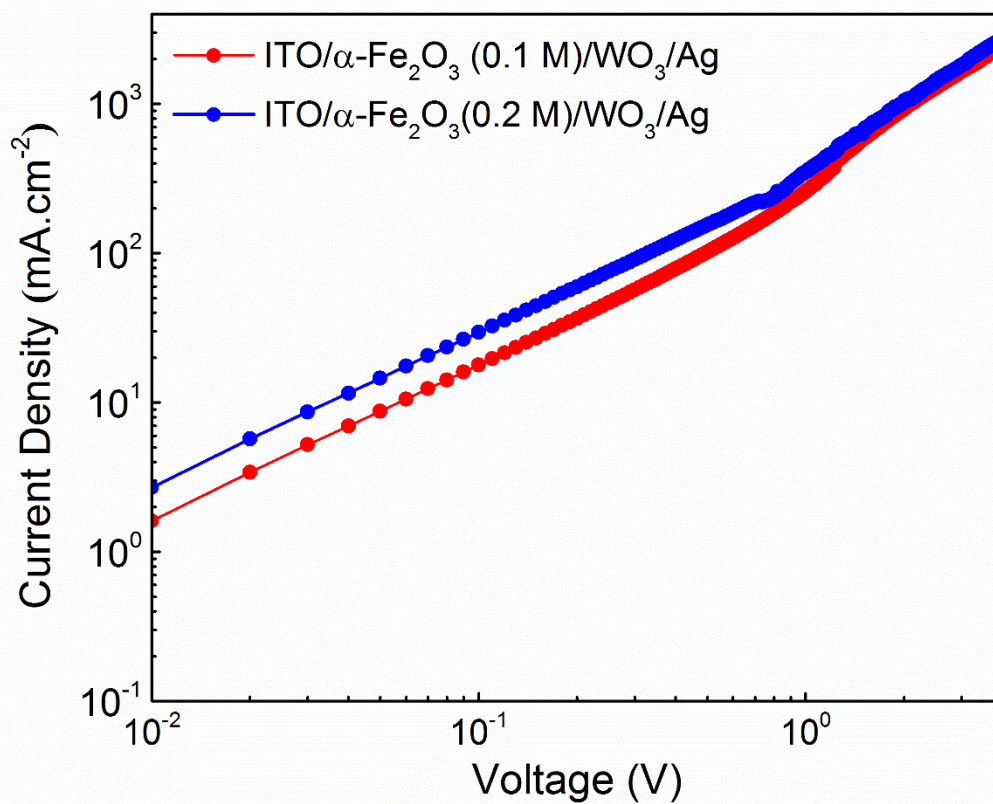


Figure S8. SCLC measurements of the electron-only device with 0.1 M and 0.2 M E- α -Fe₂O₃ ETLs in configuration (ITO/ α -Fe₂O₃/WO₃/Ag).

Table S1 Fast and slow components for the PL decay and their corresponding proportions with ITO/CsFAMA

Description	τ_1 (ns)	A ₁ (%)	τ_2 (ns)	A ₂ (%)	τ_{ave} (ns)
ITO/CsFAMA	6.39	31.28	241.33	68.72	238.52

Table S2 Fast and slow components for the PL decay and their corresponding proportions with ITO/ α -Fe₂O₃/CsFAMA in different solvents

Description	τ_1 (ns)	A ₁ (%)	τ_2 (ns)	A ₂ (%)	τ_{ave} (ns)
W- α -Fe ₂ O ₃	3.40	69.29	22.81	30.71	17.93
E- α -Fe ₂ O ₃	4.82	70.76	31.97	29.24	24.71
iP- α -Fe ₂ O ₃	7.58	52.95	92.17	47.05	85.00
iB- α -Fe ₂ O ₃	2.22	62.72	21.44	37.28	18.59

Table S3 Photovoltaic parameters of PSCs using α -Fe₂O₃ made from different solvents and different concentrations

ETL	Concentration	J _{sc} (mA.cm ⁻²)	V _{oc} (V)	FF (%)	PCE (%)
W- α -Fe ₂ O ₃	0.2 M	16.27	0.72	56.67	6.64
	0.5 M	14.73	0.72	57.63	6.11
	1 M	13.22	0.69	55.62	5.07
iP- α -Fe ₂ O ₃	0.2 M	12.35	0.58	54.08	3.87
	0.5 M	11.19	0.45	51.67	2.60

	1 M	5.03	0.79	34.16	1.36
	0.2 M	14.26	0.58	54.25	4.49
iB- α -Fe ₂ O ₃	0.5 M	14.50	0.45	51.58	3.37
	1 M	13.60	0.56	46.14	3.51

Table S4 Fast and slow components for the PL decay and their corresponding proportions with ITO/E- α -Fe₂O₃/CsFAMA with different concentrations.

Description	Concentration	τ_1 (ns)	A ₁ (%)	τ_2 (ns)	A ₂ (%)	τ_{ave} (ns)
ITO/E-α-Fe₂O₃/CsFAMA	0.1 M	5.05	70.51	32.42	29.49	24.99
	0.2 M	2.55	90.87	18.99	9.13	9.59
	0.5 M	6.12	60.72	65.37	39.28	57.88
	1 M	2.87	55.17	109.29	44.83	105.96