

## Supporting Information

### **CO<sub>2</sub> electrolysis to multi-carbon products in strong acid at ampere-current levels on La-Cu spheres with channels**

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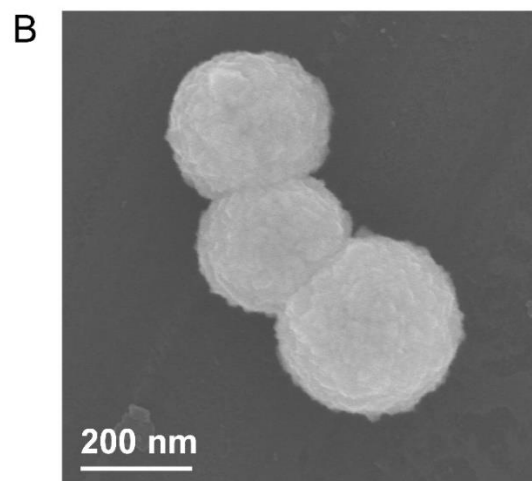
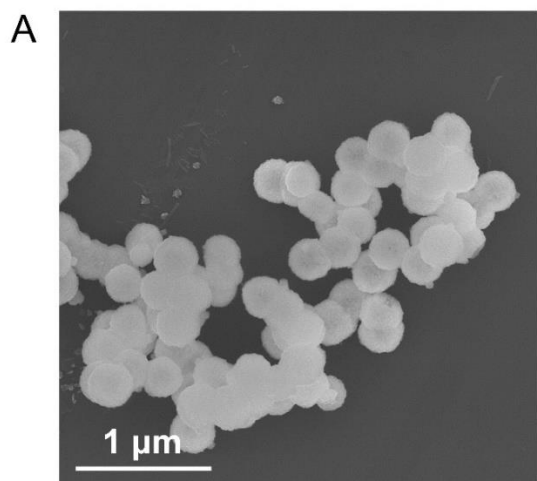
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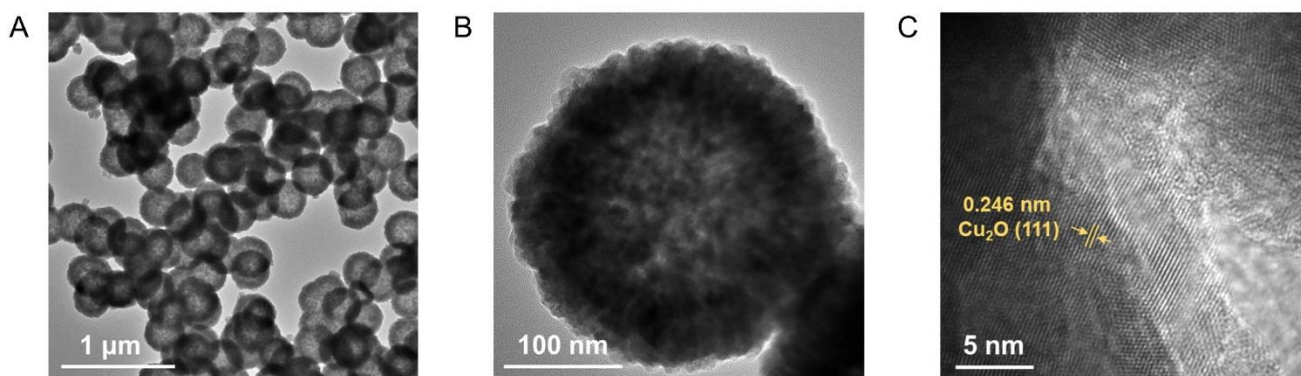
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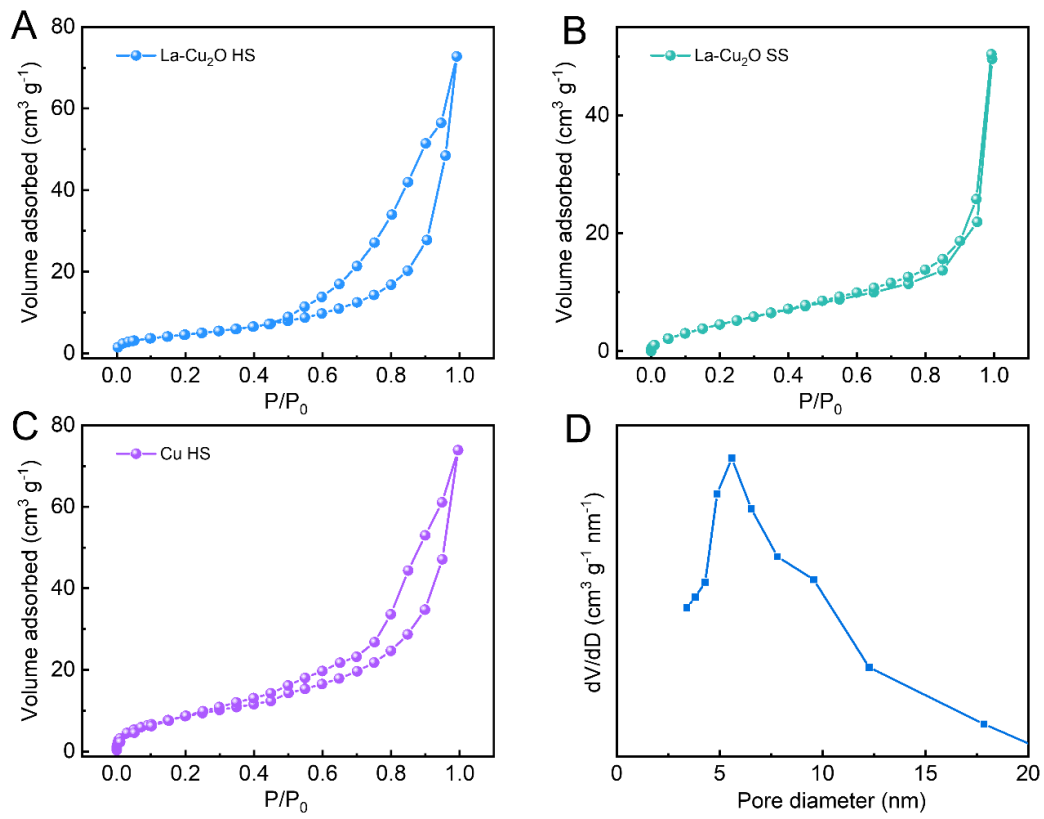
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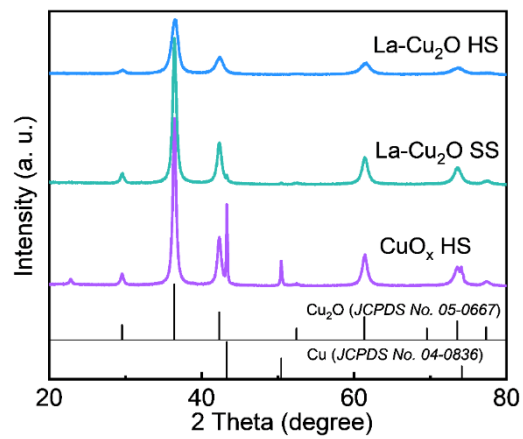
**Supplementary Fig. S1** | SEM images of La-Cu<sub>2</sub>O HS in (A) low and (B) high magnification.



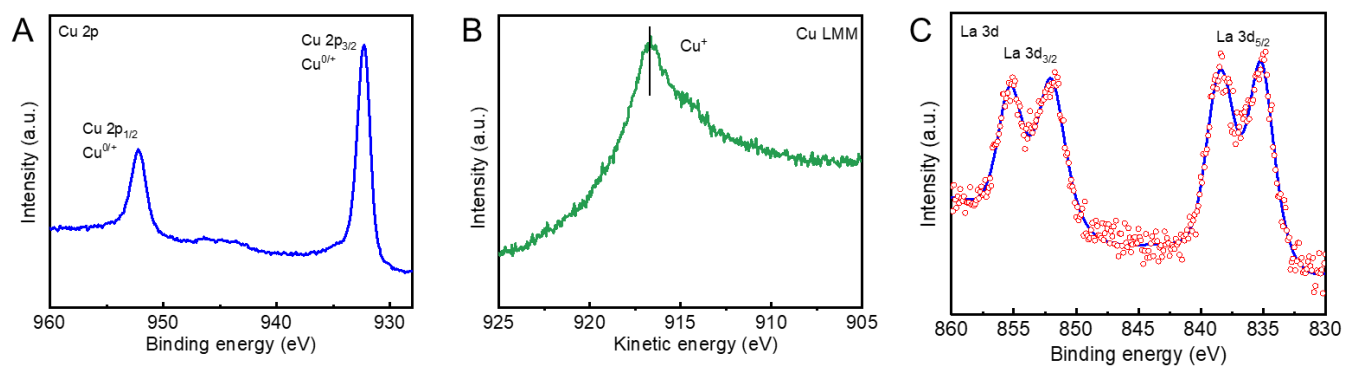
**Supplementary Fig. S2** | (A, B) TEM images and (C) HRTEM image of La-Cu<sub>2</sub>O HS.



**Supplementary Fig. S3** | Nitrogen adsorption-desorption isotherm of (A) La-Cu<sub>2</sub>O HS, (B) La-Cu<sub>2</sub>O SS, and (C) CuO<sub>x</sub> HS. (D) The corresponding BJH pore size distribution of La-Cu<sub>2</sub>O HS.

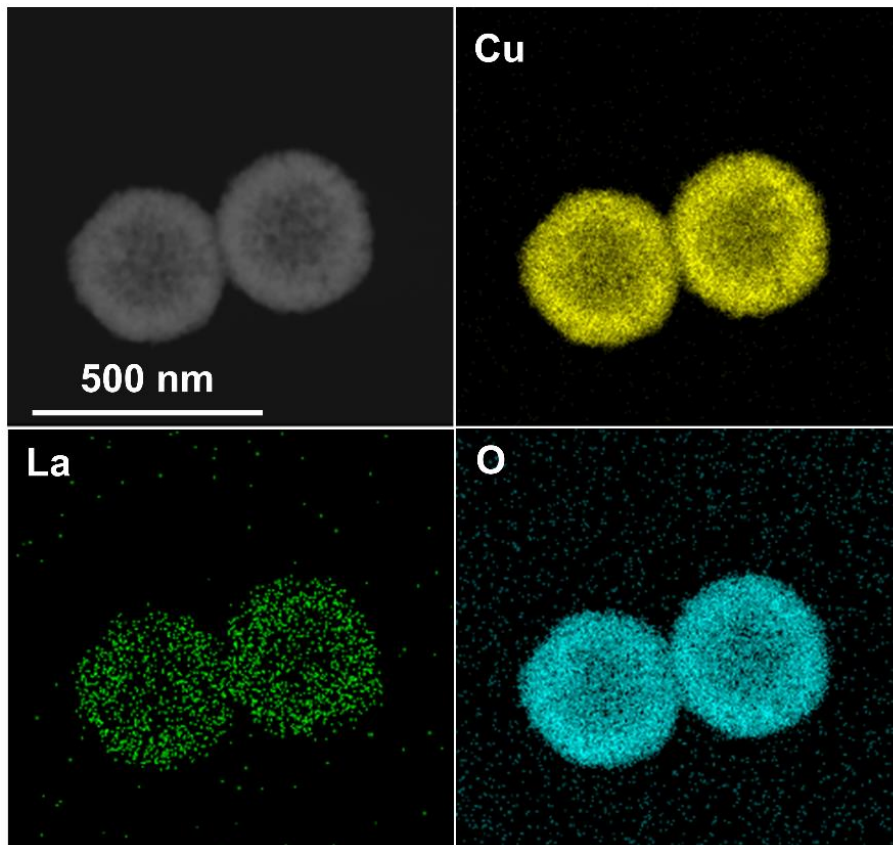


**Supplementary Fig. S4** | XRD patterns of La-Cu<sub>2</sub>O HS, La-Cu<sub>2</sub>O SS, and CuO<sub>x</sub> HS.

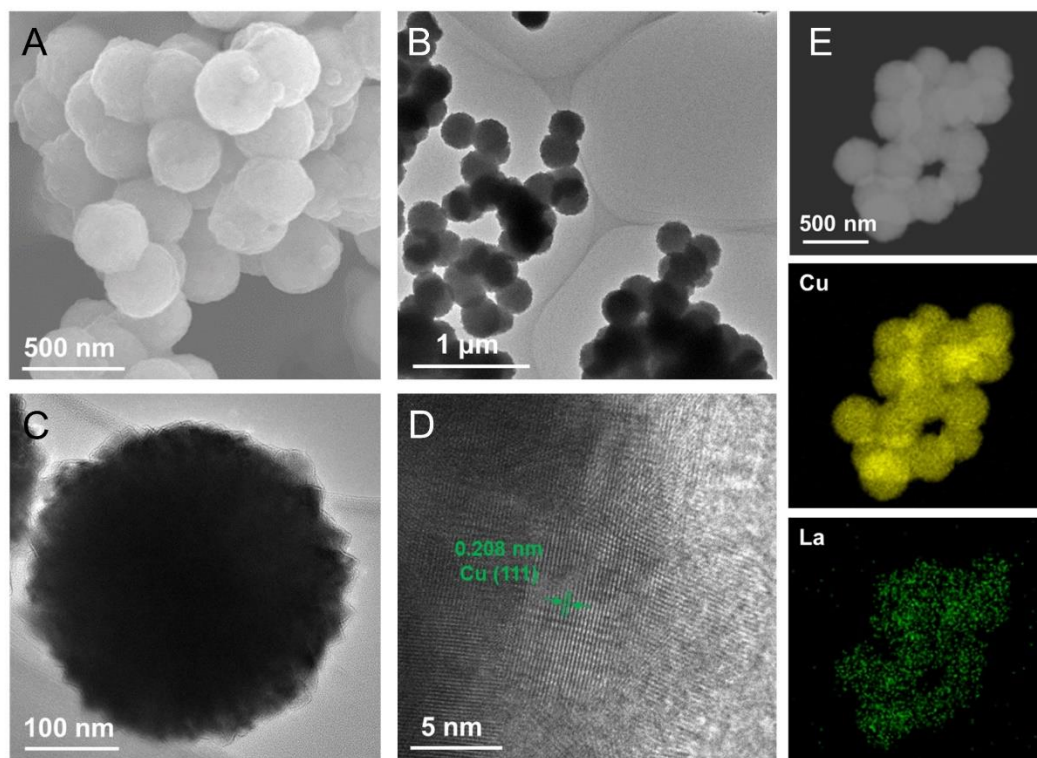


**Supplementary Fig. S5** | (A) Cu 2p XPS profile, (B) Cu Auger LMM spectra, and (C) La 3d XPS profile of La-Cu<sub>2</sub>O

HS.

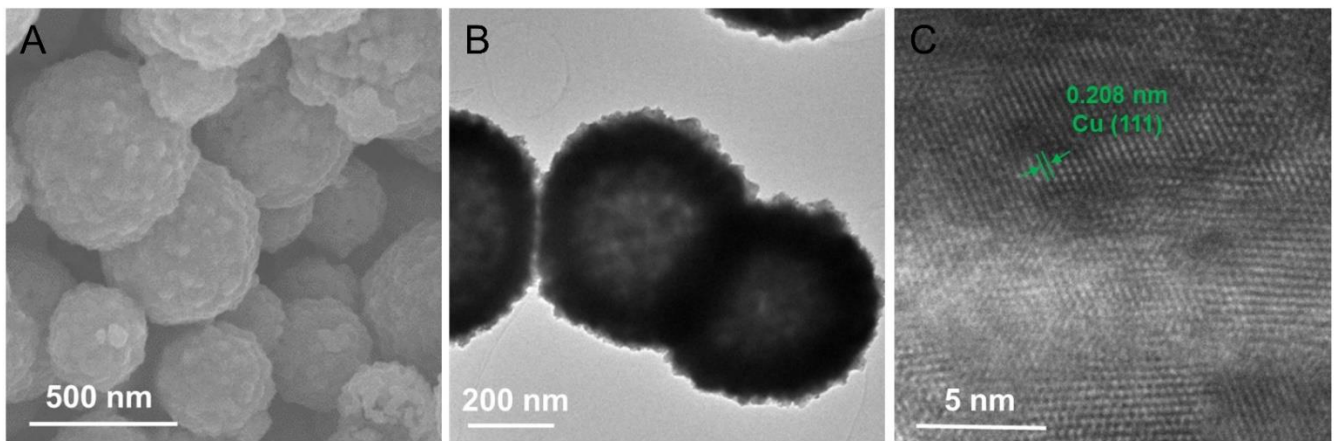


**Supplementary Fig. S6** | EDS mapping of La-Cu<sub>2</sub>O HS. Yellow, green, and blue represent Cu, La, and O elements, respectively. The results indicate that Cu, La, and O elements are uniformly distributed on La-Cu<sub>2</sub>O HS.

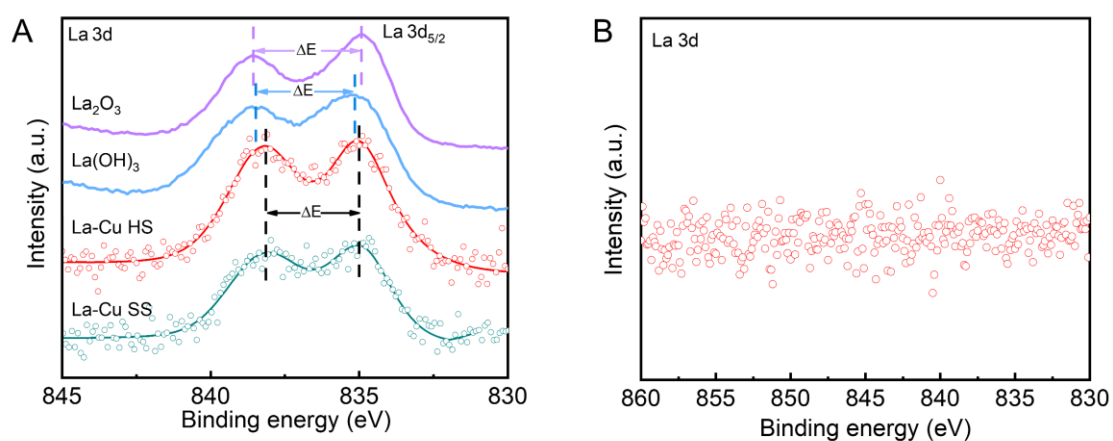


**Supplementary Fig. S7** | (A) SEM, (B, C) TEM, (D) HRTEM images, and (E) EDS mapping of La-Cu SS (yellow and green represent Cu and La elements, respectively).

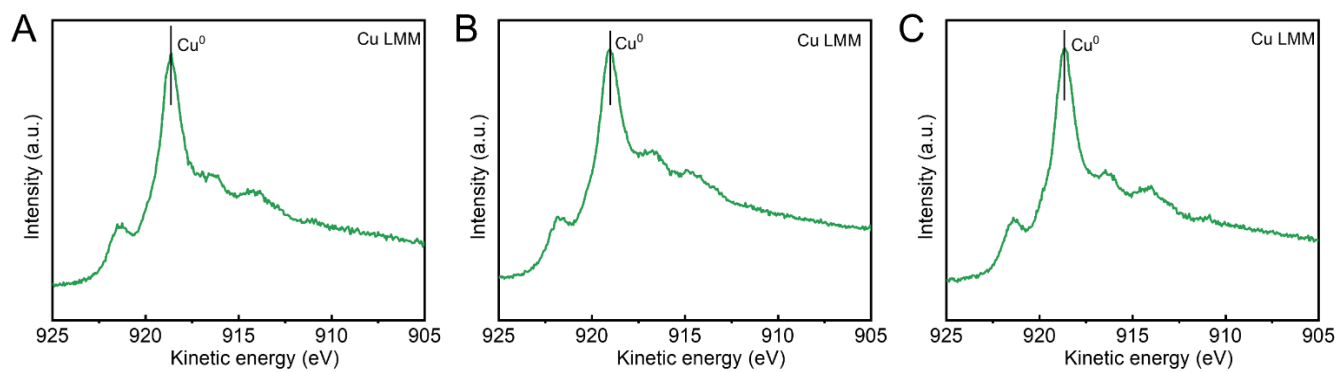




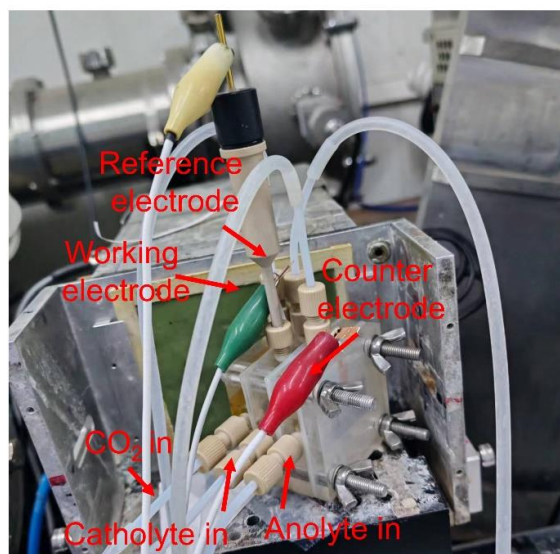
**Supplementary Fig. S8** | (A) SEM, (B) TEM, and (C) HRTEM images of Cu HS.



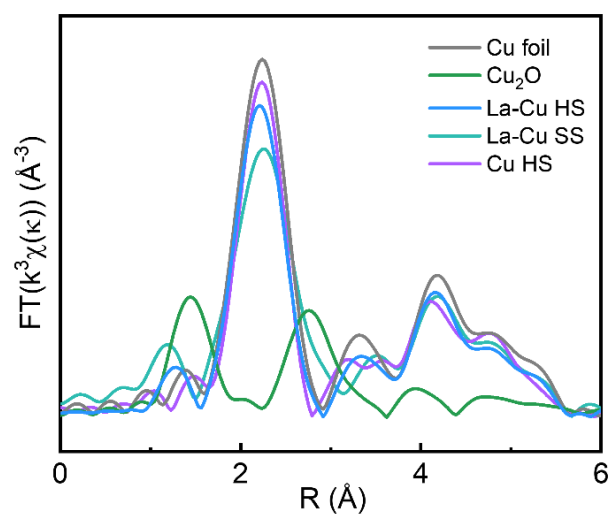
**Supplementary Fig. S9** | La 3d XPS profile of (A) La-Cu HS, La-Cu SS,  $\text{La}_2\text{O}_3$ ,  $\text{La}(\text{OH})_3$  and (B) Cu HS.



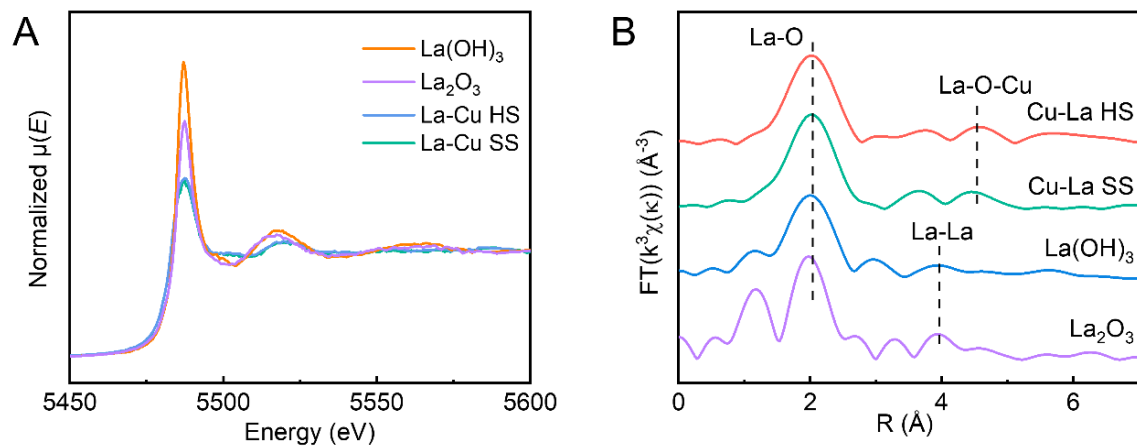
**Supplementary Fig. S10** | Cu Auger LMM spectra of (A) La-Cu HS, (B) La-Cu SS, and (C) Cu HS.



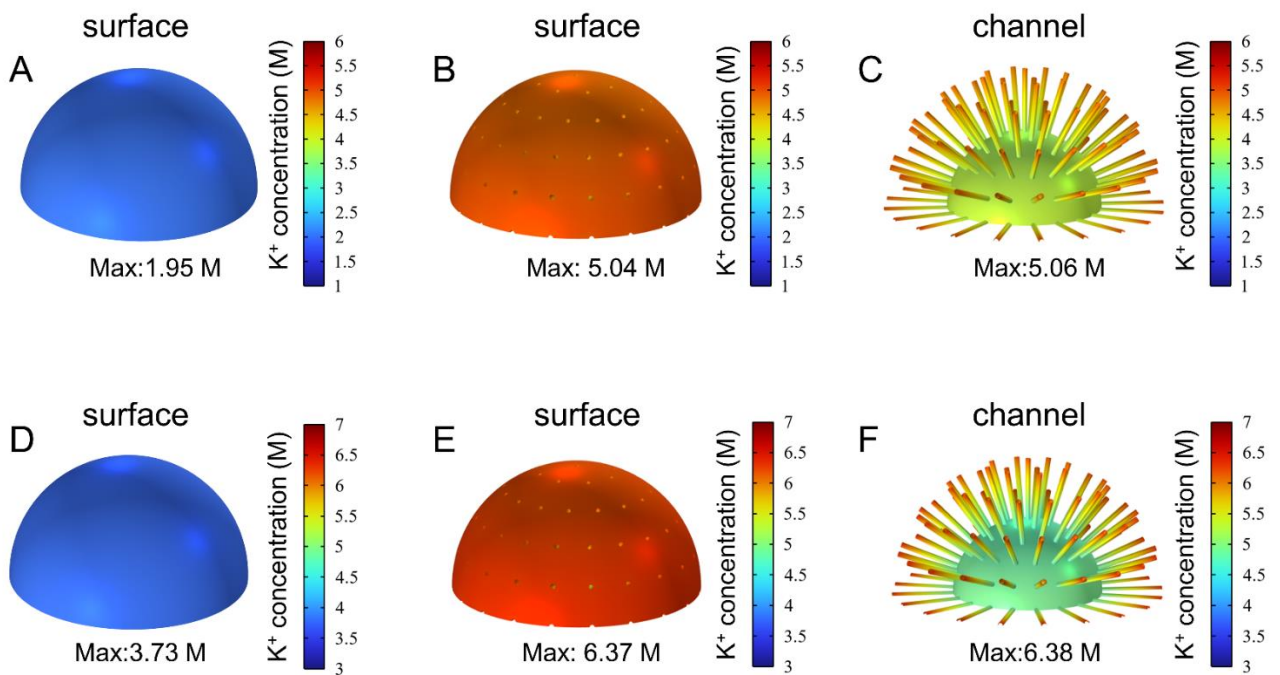
**Supplementary Fig. S11** | The optical photo of flow cell used for *in situ* XAS experiment.



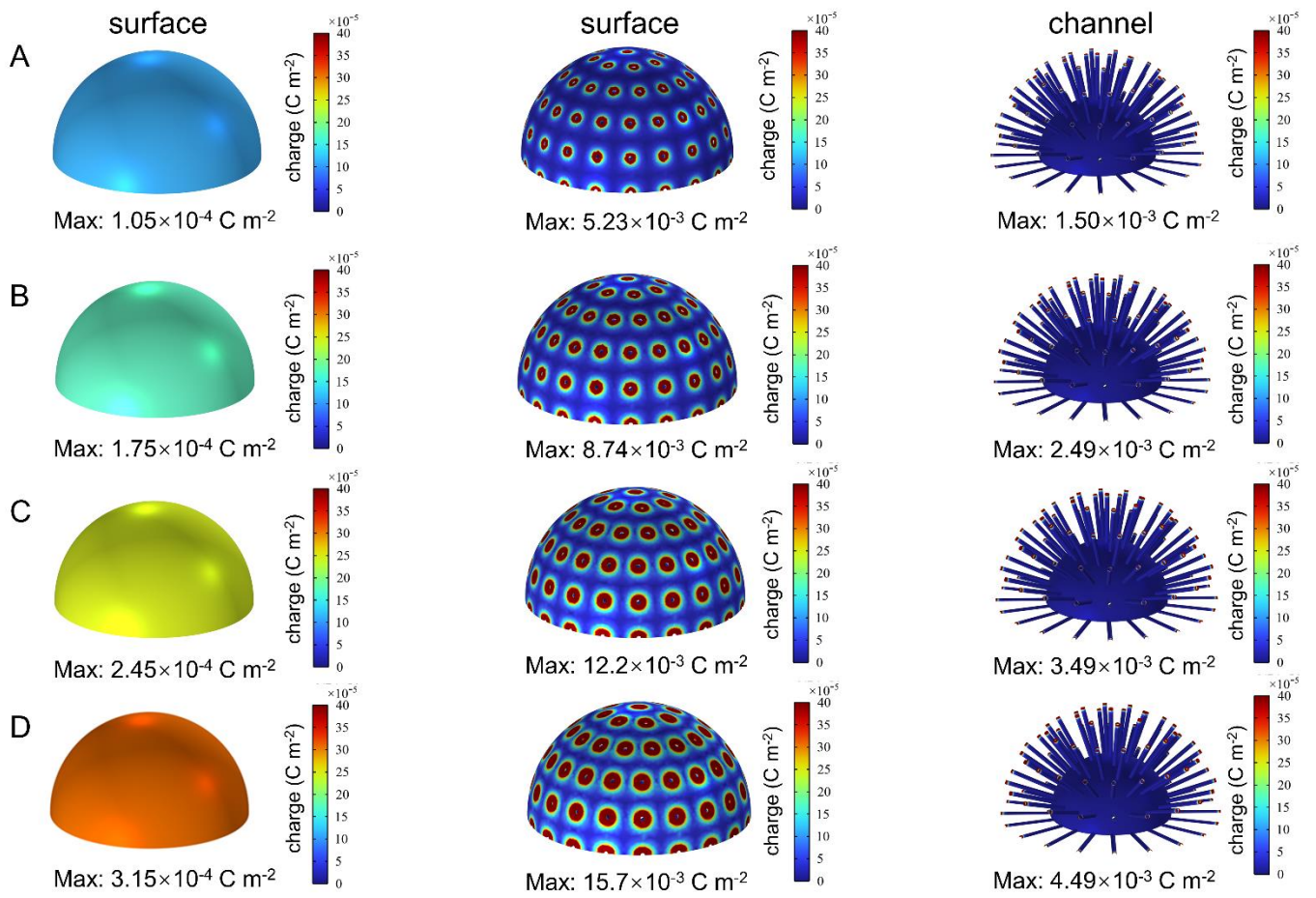
**Supplementary Fig. S12** | Fourier transform extended X-ray absorption fine structure spectra of La-Cu HS, La-Cu SS, and Cu HS.



**Supplementary Fig. S13** | (A) La L<sub>3</sub>-edge XANES spectra and (B) FT EXAFS spectra of La-Cu HS, La-Cu SS, La(OH)<sub>3</sub> and La<sub>2</sub>O<sub>3</sub>.

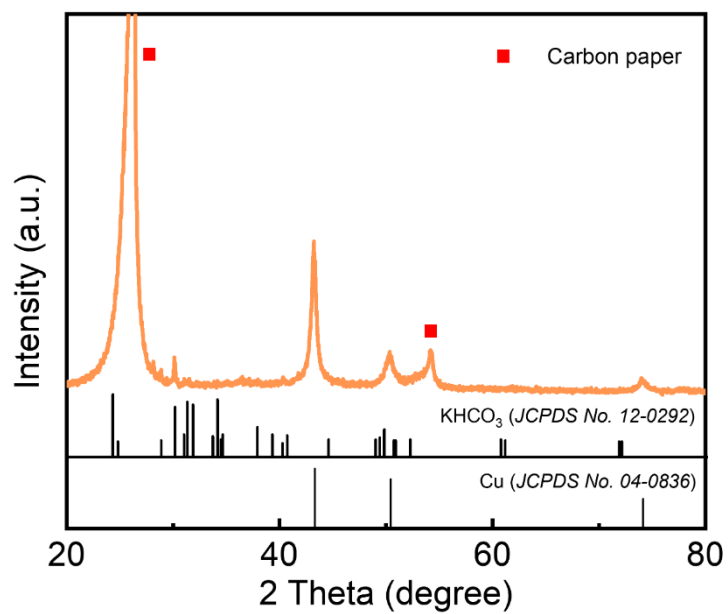


**Supplementary Fig. S14** | The  $K^+$  concentration distribution over surface of (A) solid sphere, (B) hollow sphere with channels, and within channels of (C) hollow sphere at -300 mA. The  $K^+$  concentration distribution over surface of (D) solid sphere, (E) hollow sphere with channels, and within channels of (F) hollow sphere at -500 mA.

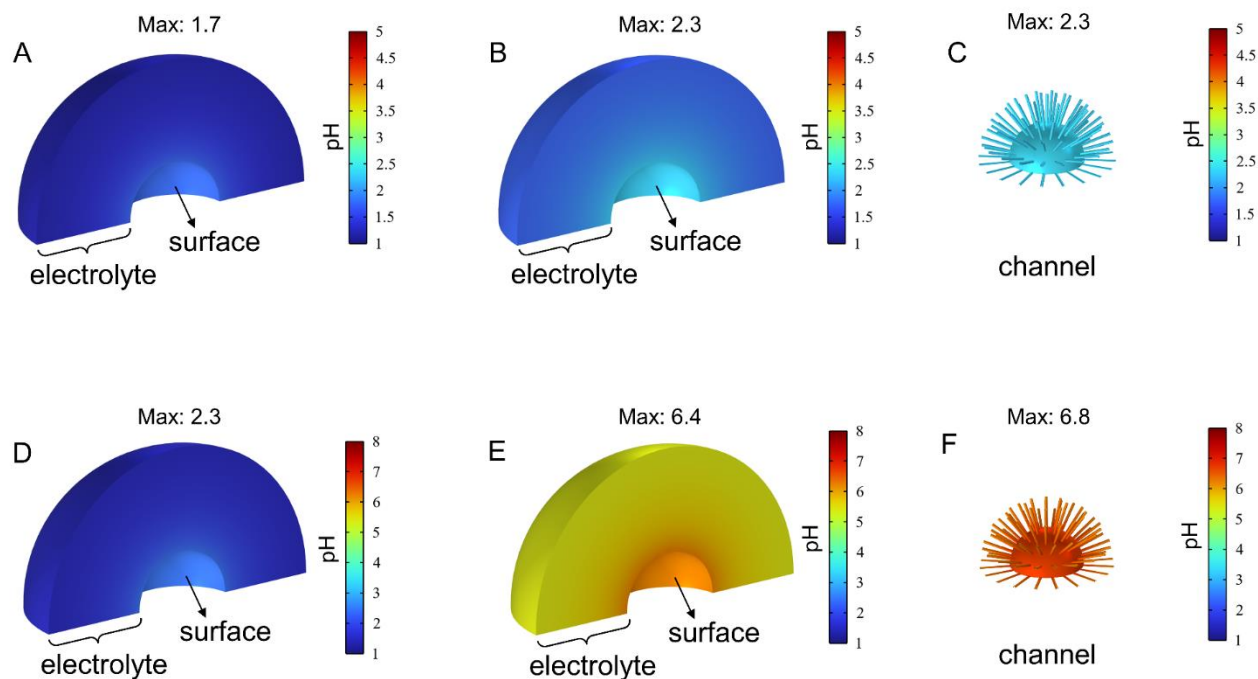


**Supplementary Fig. S15** | The negative charge distribution over solid sphere, hollow sphere with channels, and within channels of hollow sphere at (A) -300 mA, (B) -500 mA, (C) -700 mA, and (D) -900 mA.

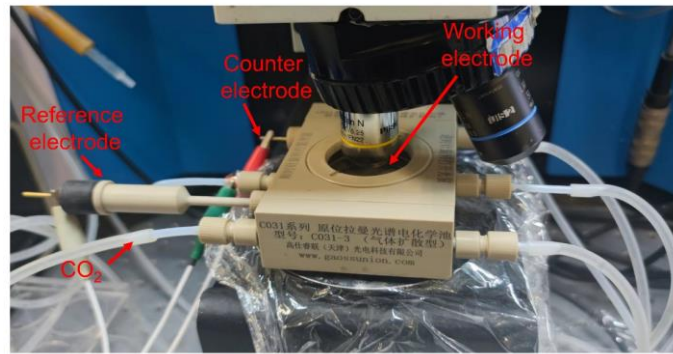




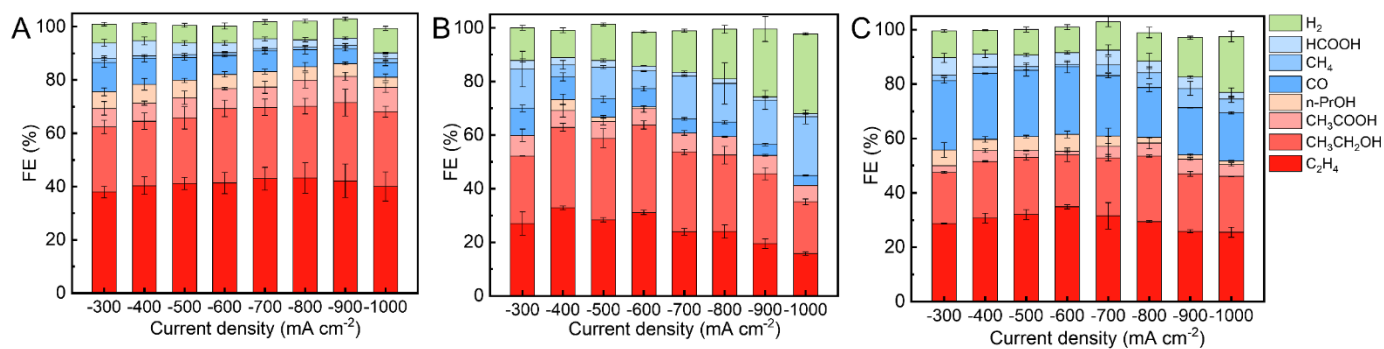
**Supplementary Fig. S16** | The XRD patterns of the electrode loaded with La-Cu HS after  $\text{CO}_2$  reduction.



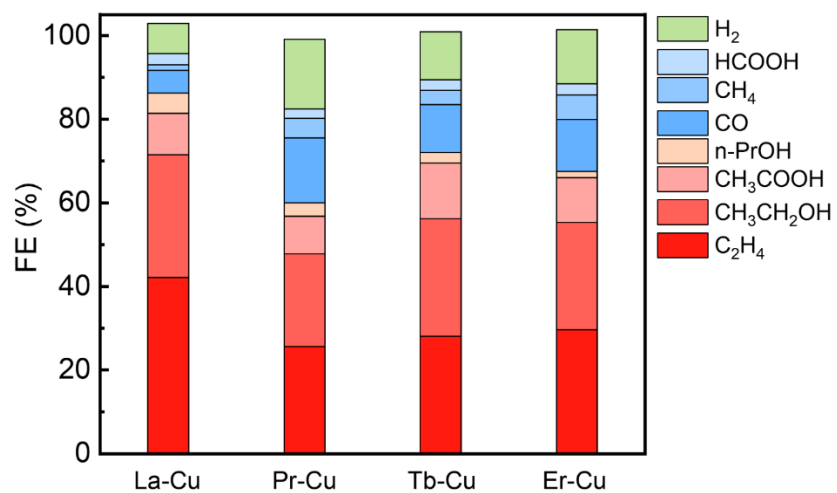
**Supplementary Fig. S17** | The pH distribution in electrolyte near surface of (A) solid sphere, (B) hollow sphere with channels, and within (C) channel of hollow sphere at -300 mA. The pH distribution in electrolyte near surface of (D) solid sphere, (E) hollow sphere with channels, and within (F) channels of hollow sphere at -500 mA.



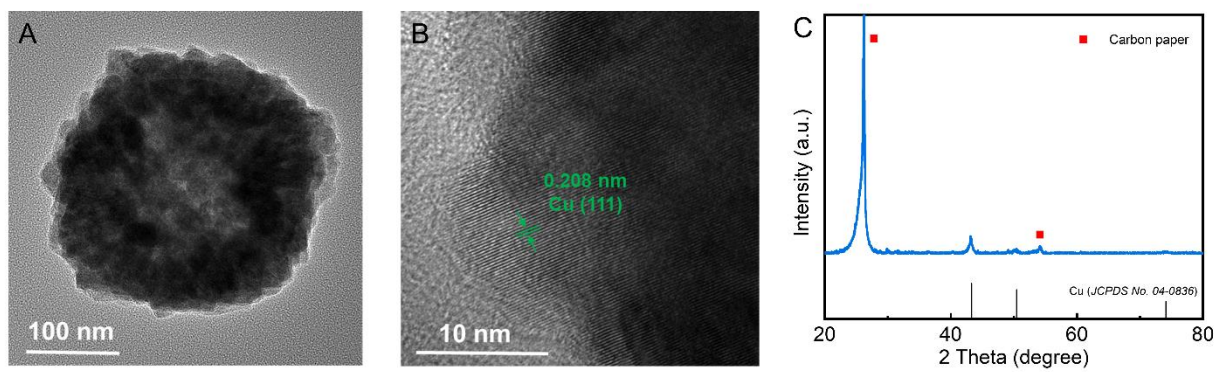
**Supplementary Fig. S18** | The photograph of cell used for *in situ* SERS spectroscopy.



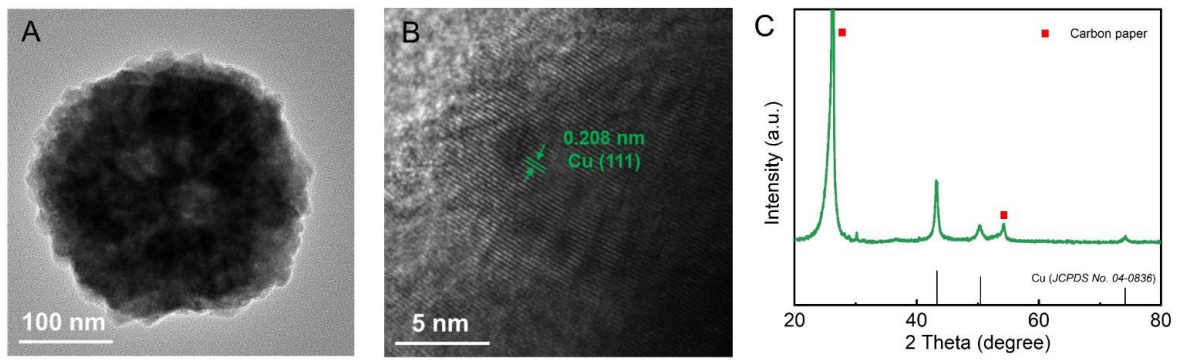
**Supplementary Fig. S19** | The product distribution over (A) La-Cu HS, (B) La-Cu SS, and (C) Cu HS under different current densities. Values are means and error bars indicate s.d. ( $n = 3$  replicates).



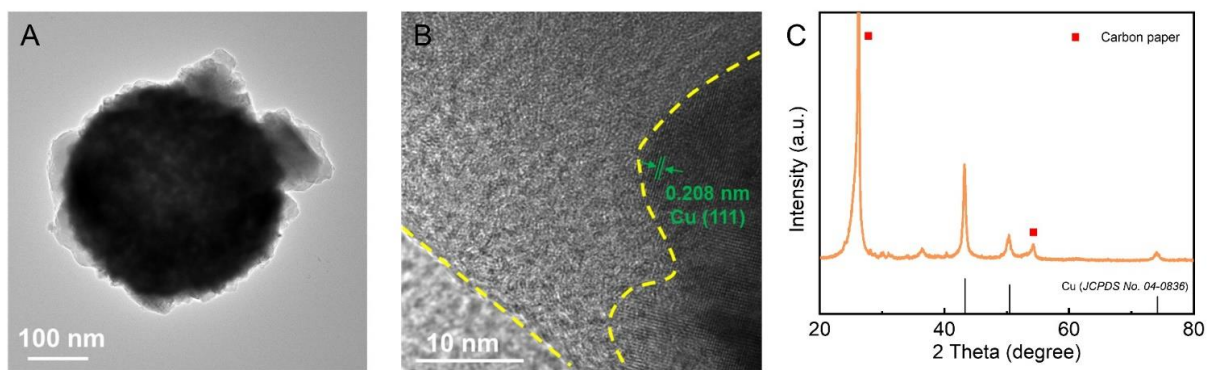
**Supplementary Fig. S20** | The products distribution of La-Cu, Pr-Cu, Tb-Cu and Er-Cu at  $-900 \text{ mA cm}^{-2}$ .



**Supplementary Fig. S21** | (A) TEM image, (B) HRTEM image, and (C) XRD pattern of La-Cu HS with La/Cu ratio of 0.1.

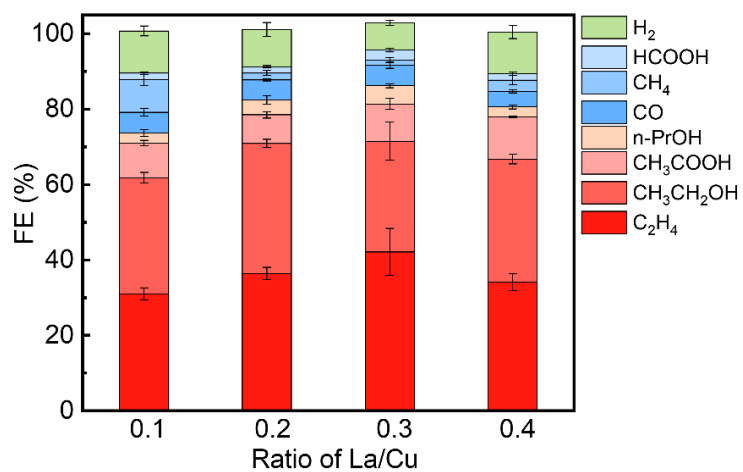


**Supplementary Fig. S22** | (A) TEM image, (B) HRTEM image, and (C) XRD pattern of La-Cu HS with La/Cu ratio of 0.2.

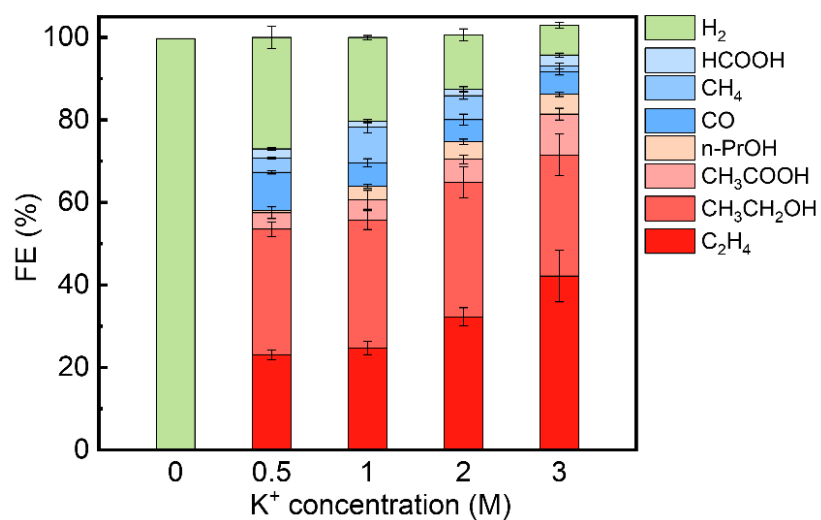


**Supplementary Fig. S23** | (A) TEM image, (B) HRTEM image, and (C) XRD pattern of La-Cu HS with La/Cu ratio of 0.4.

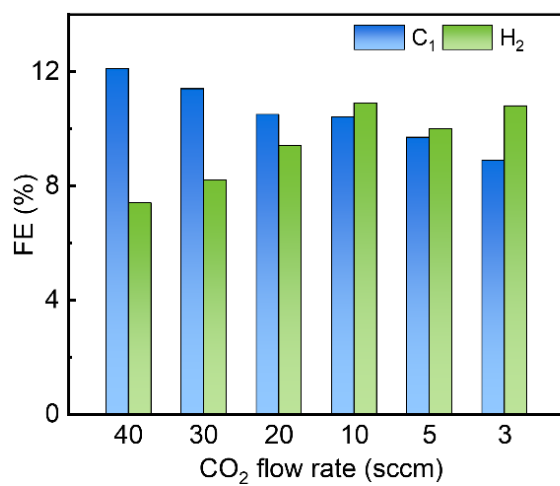




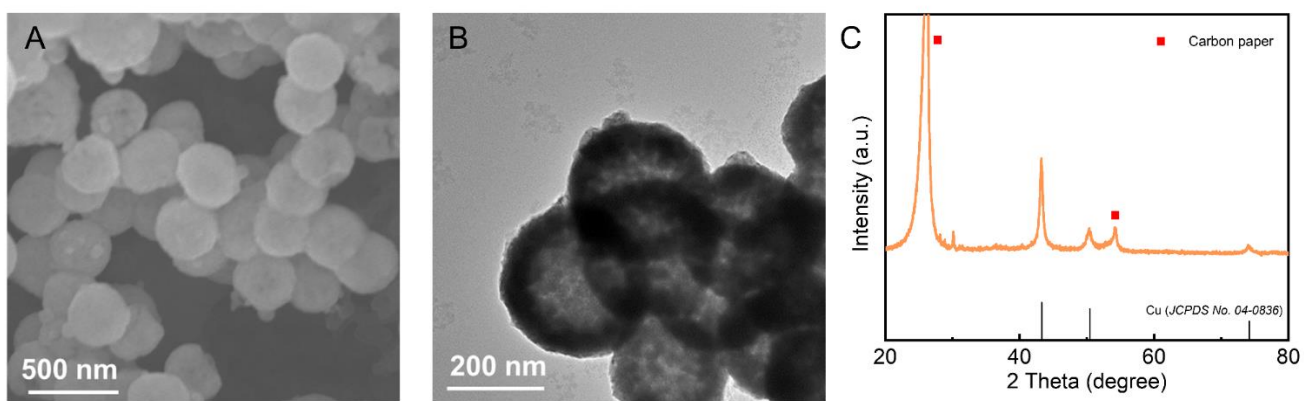
**Supplementary Fig. S24** | Product distribution of La-Cu HS with various La/Cu ratio at  $-900 \text{ mA cm}^{-2}$ . Values are means and error bars indicate s.d. ( $n = 3$  replicates).



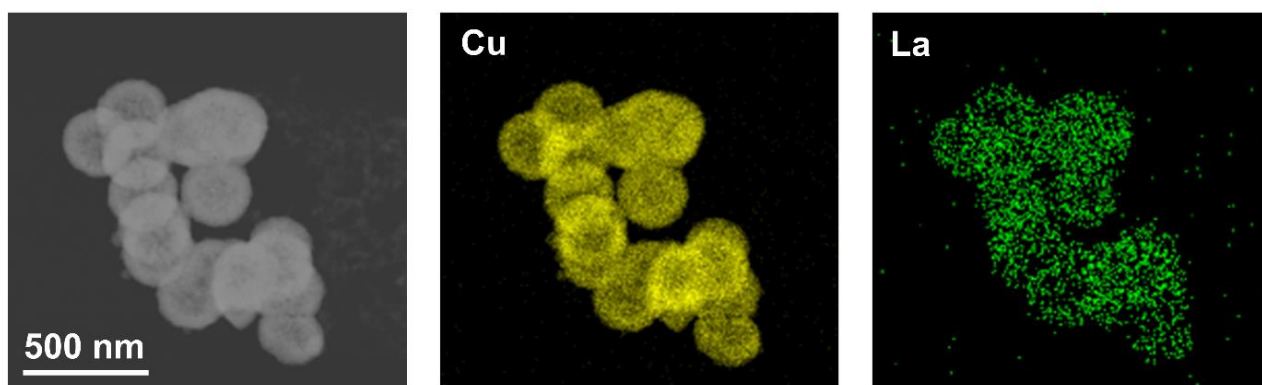
**Supplementary Fig. S25** | The products FE of La-Cu HS in 0.05 M H<sub>2</sub>SO<sub>4</sub> aqueous solution containing different KCl concentrations at -900 mA cm<sup>-2</sup>. Values are means and error bars indicate s.d. (*n* = 3 replicates).



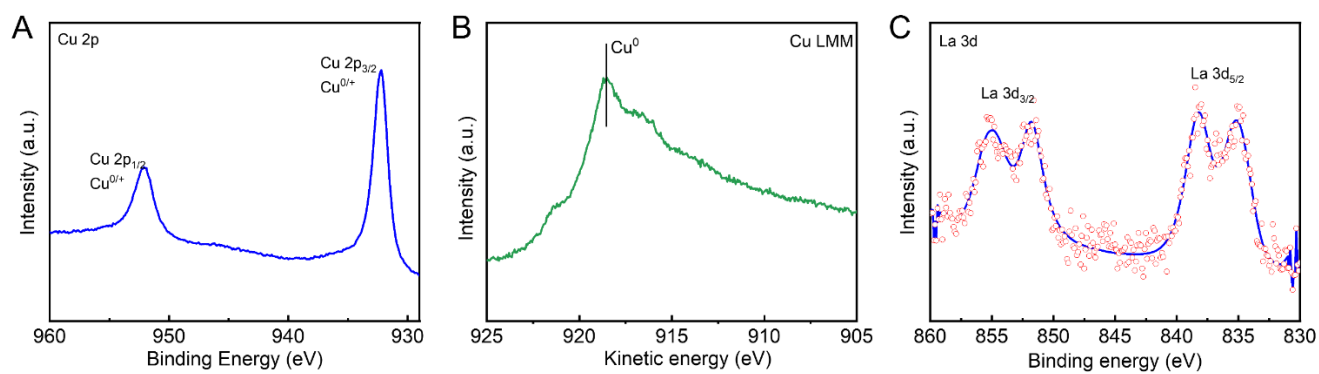
**Supplementary Fig. S26** | C<sub>1</sub> products and H<sub>2</sub> FE of La-Cu HS under different CO<sub>2</sub> gas flow rate at -900 mA cm<sup>-2</sup>



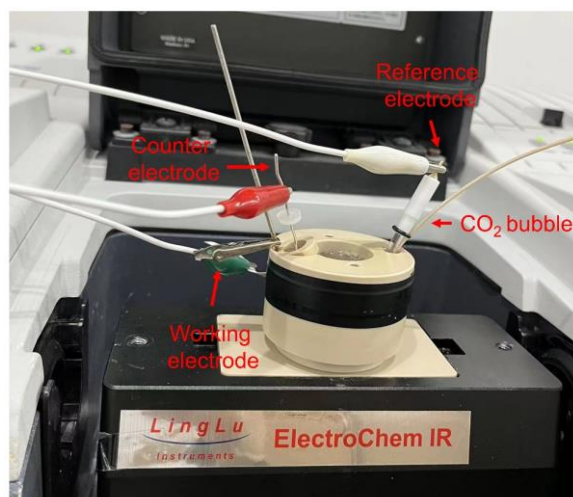
**Supplementary Fig. S27** | (A) SEM image, (B) TEM image, and (C) XRD pattern of La-Cu HS after 40 h CO<sub>2</sub>RR.



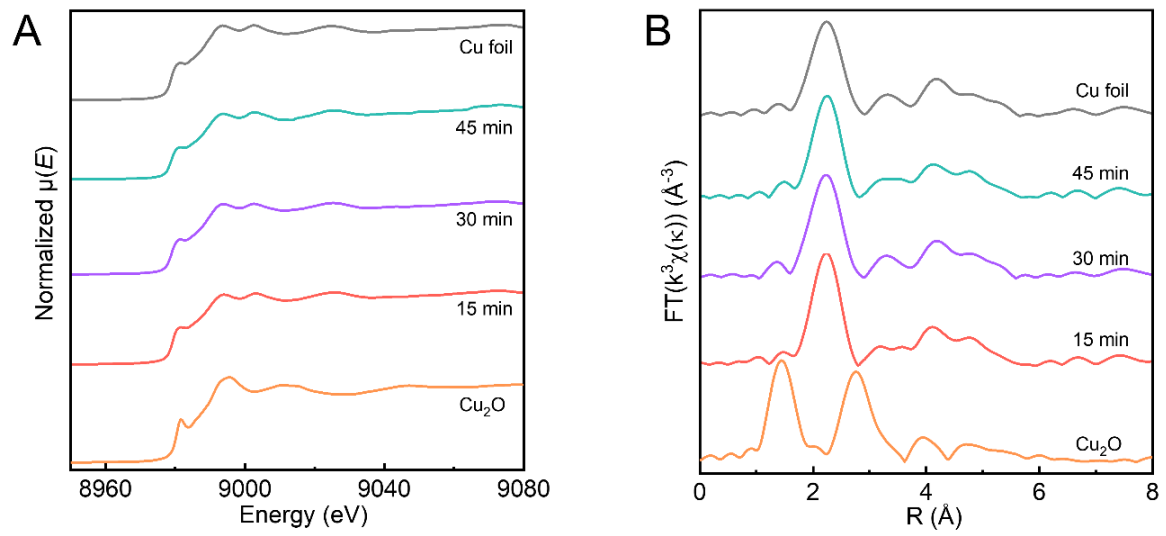
**Supplementary Fig. S28** | EDS mapping of La-Cu HS after 40 h CO<sub>2</sub>RR. Yellow and green represent Cu and La elements, respectively. The results indicate that Cu and La elements are uniformly distributed on La-Cu HS.



**Supplementary Fig. S29** | (A) Cu 2p<sub>3/2</sub> XPS spectrum, (B) Cu Auger LMM spectrum, and (C) La 3d XPS profile of La-Cu HS after 40 h CO<sub>2</sub>RR.

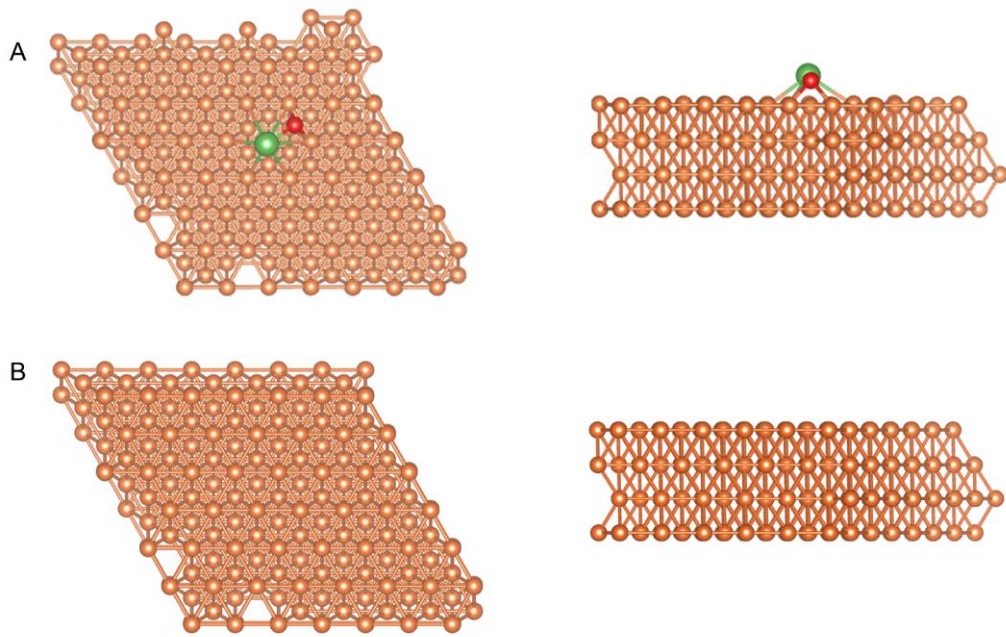


**Supplementary Fig. S30** | The photograph and scheme of cell used for *in situ* ATR-SEIRAS spectroscopy.

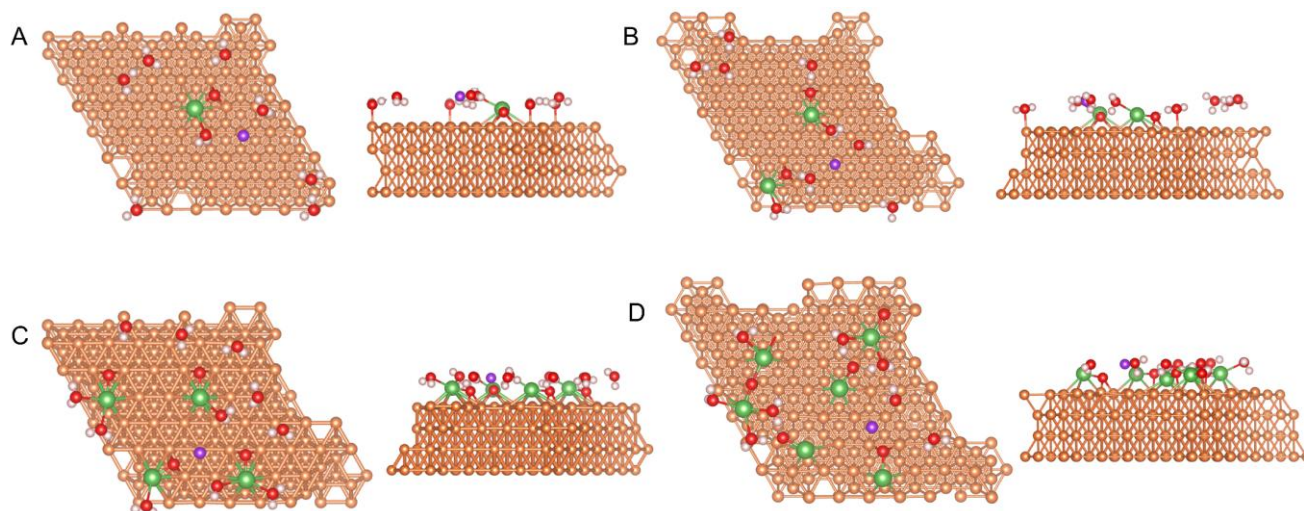


**Supplementary Fig. S31** | (A) *In-situ* Cu K-edge XANES spectra and (B) *in-situ* FT EXAFS spectra of La-Cu HS at different reaction time.

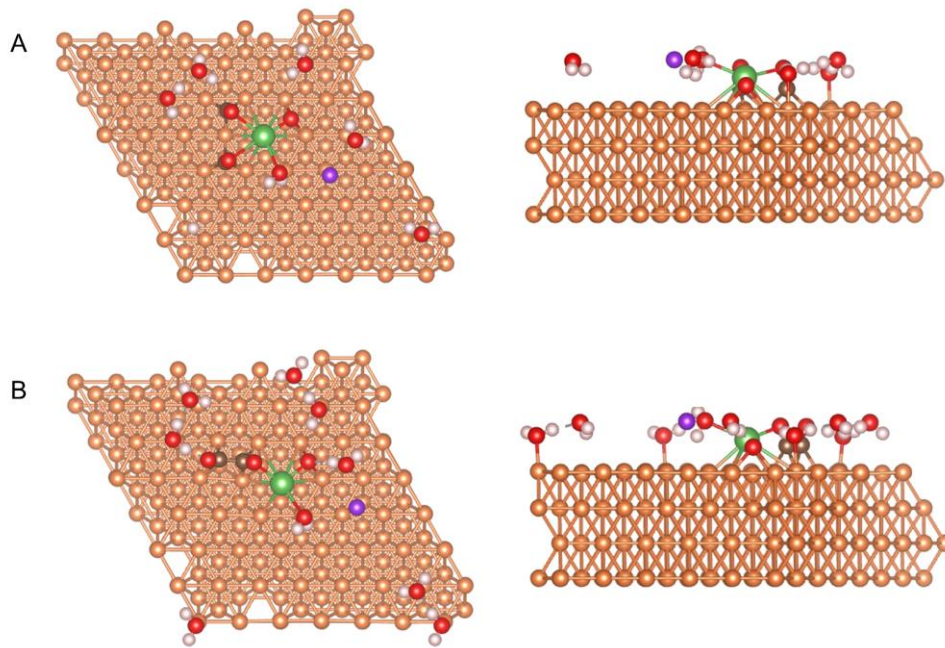




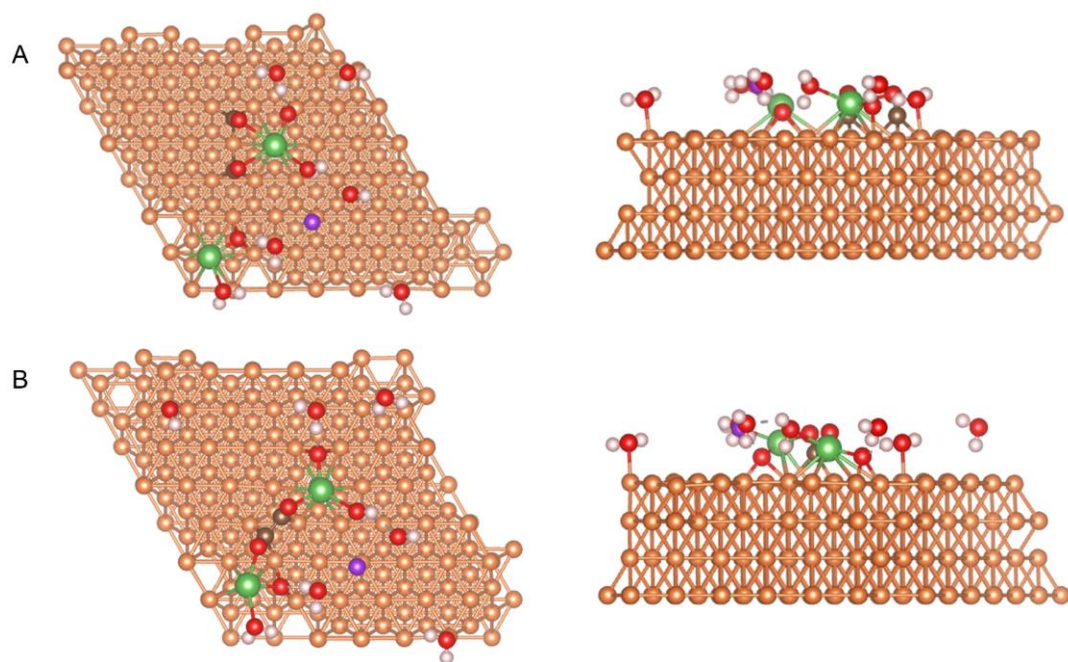
**Supplementary Fig. S32** | Top and side view of (A) La-O-Cu (111) and (B) Cu (111) models.



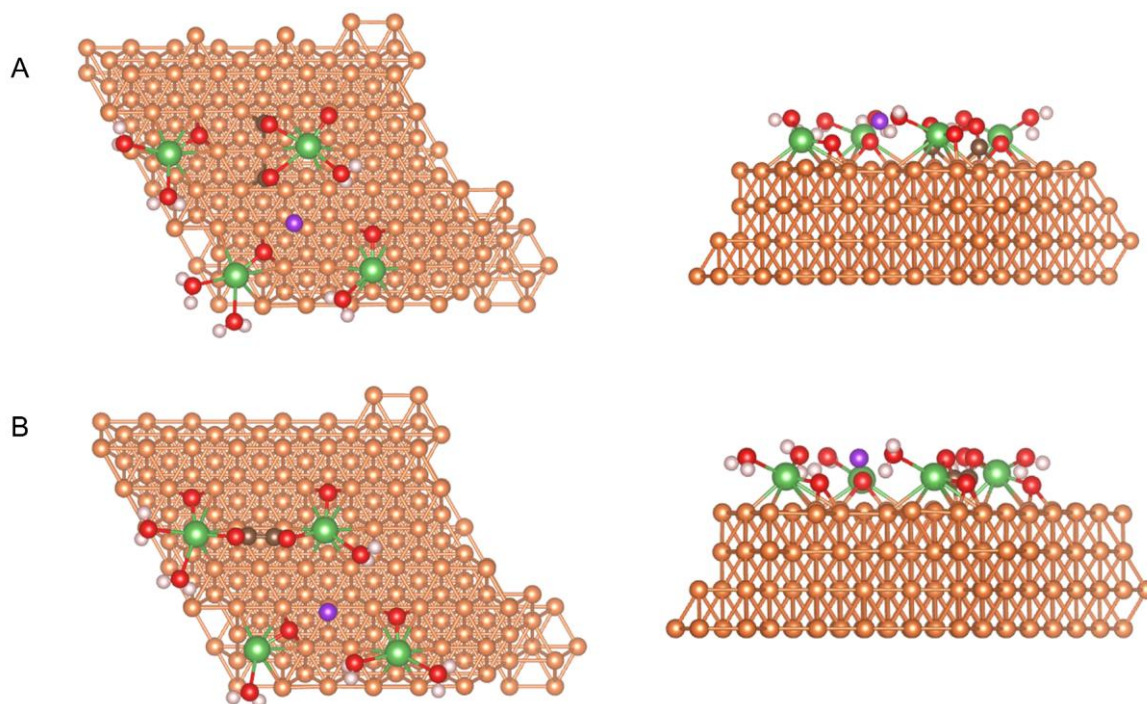
**Supplementary Fig. S33** | Top and side view of (A) 1La-Cu, (B) 2La-Cu, (C) 4La-Cu and (D) 6La-Cu models with H<sub>2</sub>O and K<sup>+</sup> species on surface. The bronze, green, red, white and purple balls represent Cu, La, O, H and K atoms, respectively.



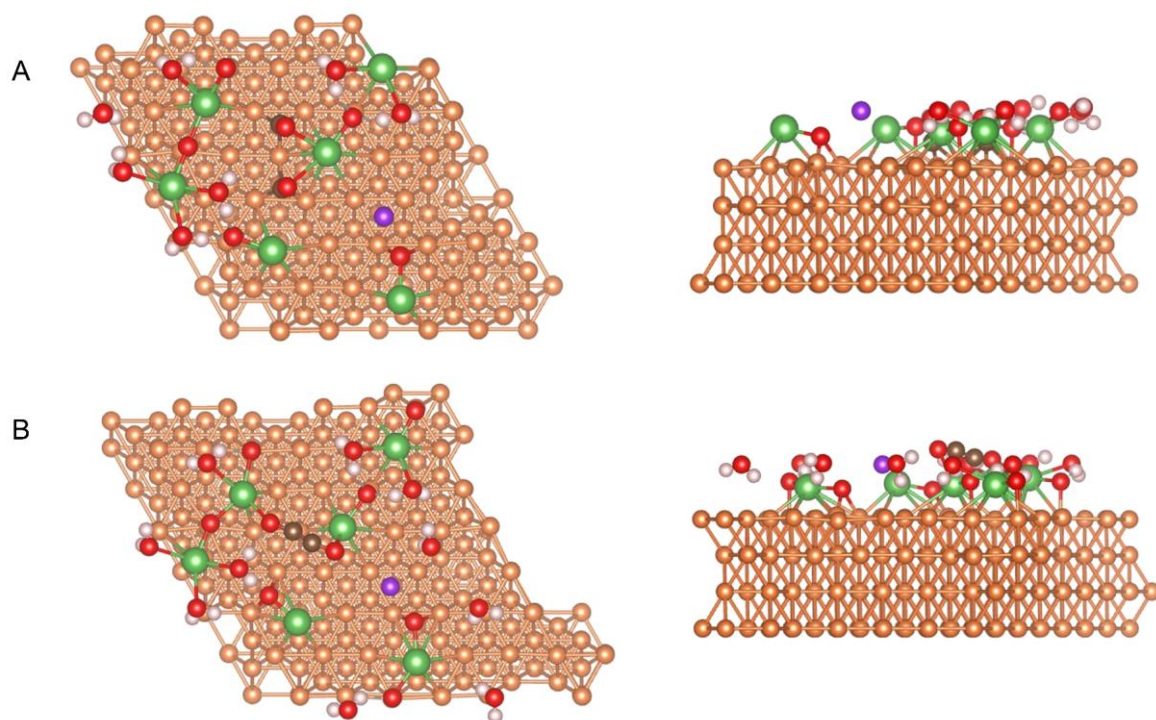
**Supplementary Fig. S34** | Top and side view of (A)  $2^*CO$  and (B)  $O^*CCO$  over 1La-Cu model with  $H_2O$  and  $K^+$  species on surface. The bronze, green, red, black, white and purple balls represent Cu, La, O, C, H and K atoms, respectively.



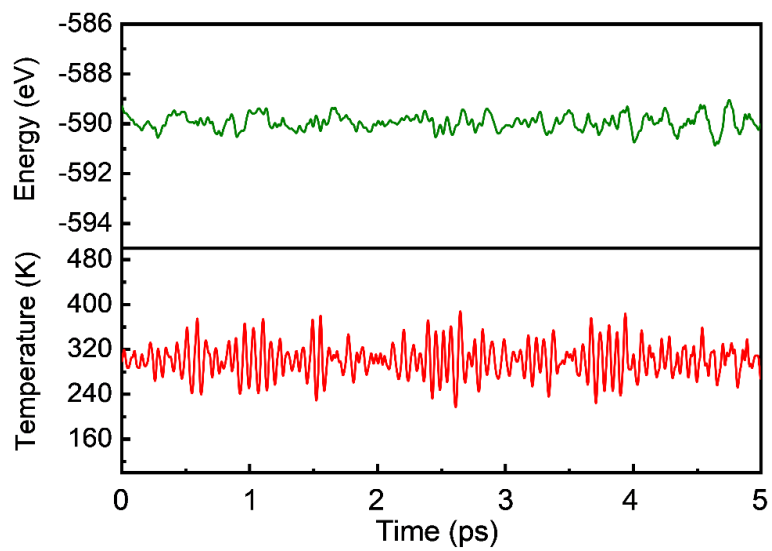
**Supplementary Fig. S35** | Top and side view of (A) 2\*CO and (B) O\*CCO over 2La-Cu model with H<sub>2</sub>O and K<sup>+</sup> species on surface. The bronze, green, red, black, white and purple balls represent Cu, La, O, C, H and K atoms, respectively.



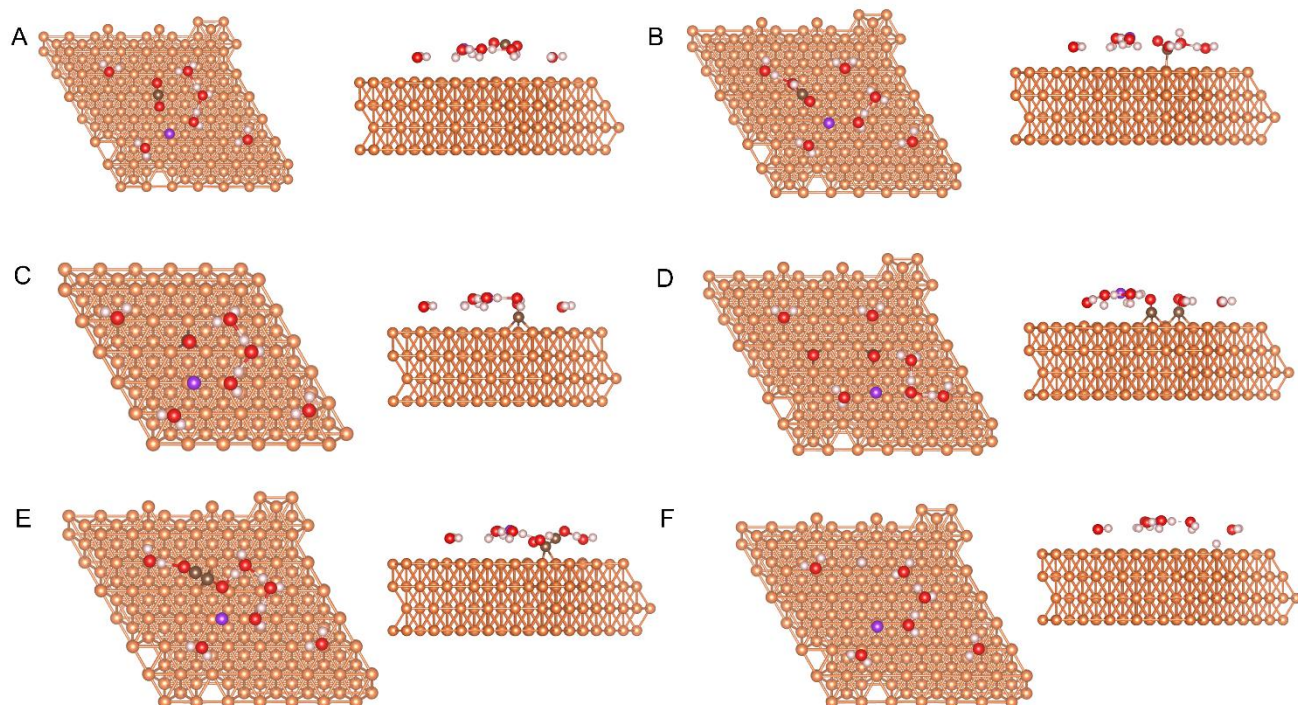
**Supplementary Fig. S36** | Top and side view of (A) 2\*CO and (B) O\*CCO over 4La-Cu model with H<sub>2</sub>O and K<sup>+</sup> species on surface. The bronze, green, red, black, white and purple balls represent Cu, La, O, C, H and K atoms, respectively.



**Supplementary Fig. S37** | Top and side view of (A)  $2^*CO$  and (B)  $O^*CCO$  over 6La-Cu model with  $H_2O$  and  $K^+$  species on surface. The bronze, green, red, black, white and purple balls represent Cu, La, O, C, H and K atoms, respectively.

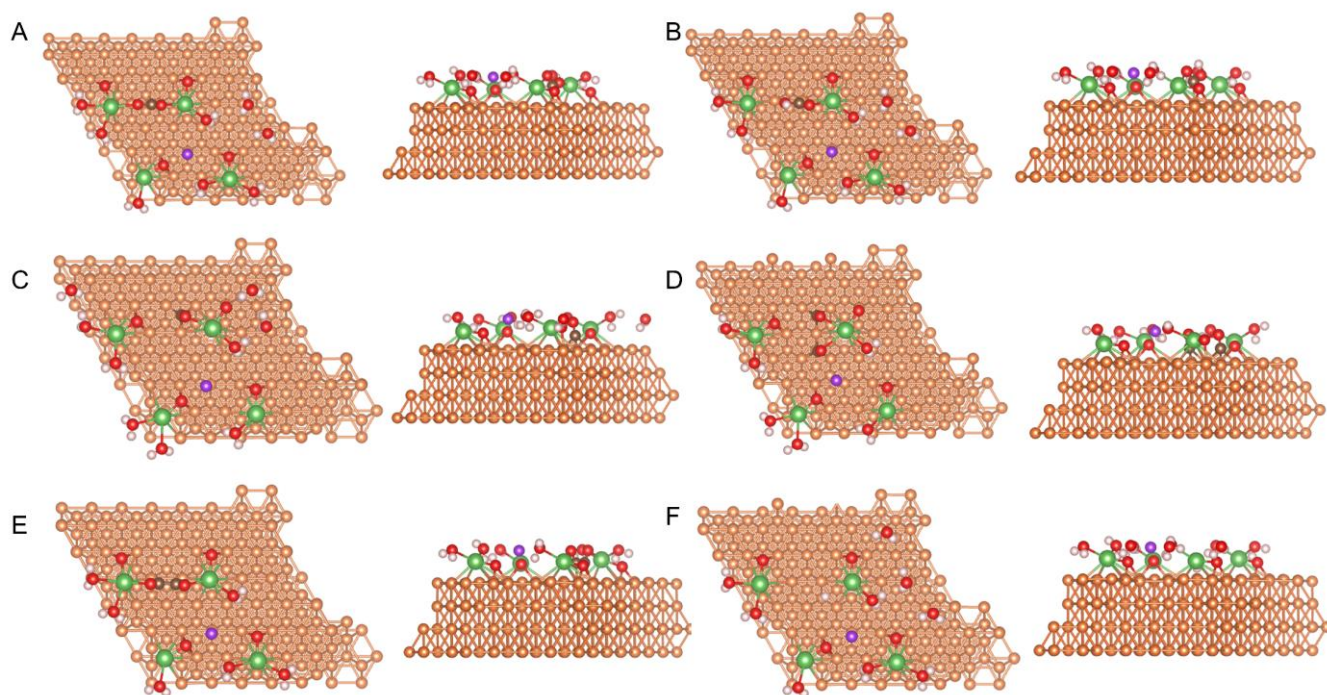


**Supplementary Fig. S38** | Energy and temperature variations during the AIMD simulation for 5 ps.

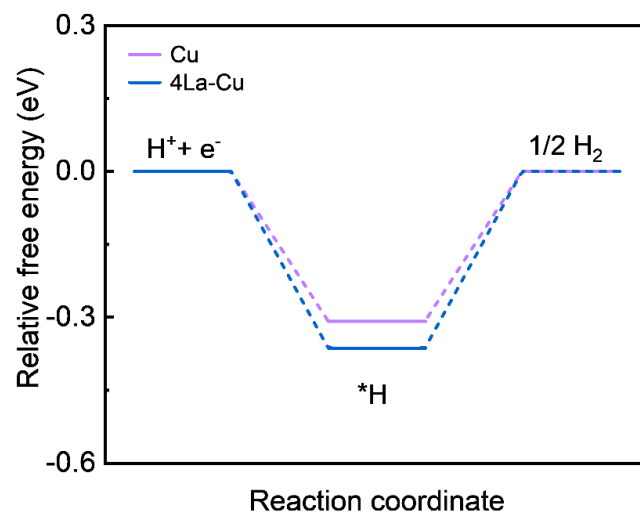


**Supplementary Fig. S39** | Top and side view of (A)  $\text{CO}_2$ , (B)  $\text{*COOH}$ , (C)  $\text{*CO}$ , (D)  $2\text{*CO}$ , (E)  $\text{O*CCO}$  and (F)  $\text{*H}$  over Cu model with  $\text{H}_2\text{O}$  and  $\text{K}^+$  species on surface. The bronze, black, red, white and purple balls represent Cu, C, O, H and K atoms, respectively.





**Supplementary Fig. S40** | Top and side view of (A) CO<sub>2</sub>, (B) \*COOH, (C) \*CO, (D) 2\*CO, (E) O\*CCO and (F) \*H over 4La-Cu model with H<sub>2</sub>O and K<sup>+</sup> species on surface. The bronze, green, black, red, white and purple balls represent Cu, La, C, O, H and K atoms, respectively.



**Supplementary Fig. S41** | Calculated free energy diagram for HER over Cu and 4La-Cu.

**Supplementary Table S1** | Comparison of CO<sub>2</sub>RR-to-C<sub>2+</sub> products over various reported electrocatalysts in acidic electrolyte.

Catalyst	pH	C <sub>2+</sub> FE (%)	<i>j</i> <sub>C<sub>2+</sub></sub> (mA cm <sup>-2</sup> )	Stability (h)	Ref.
La-Cu HS	<1	86.2	775.8	40	This work
ER-CuNS	<1	84	557	30	1
CAL-modified Cu	<1	48	576	10	2
Pd-Cu	2	89	500	4.5	3
Cu/C	<1	~37	~200	~4	4
COF:PFSA-modified PTFE-Cu	<1	75	150	20	5
Cu/PTFE	-	55	~55	9	6
EC-Cu	<1	90	~180	10	7
Carbon/Cu/PTFE	2	64.5*	~200	-	8
Cu <sub>0.9</sub> Zn <sub>0.1</sub>	4	70	210	~35	9
CG-medium	<1	~80	~80	160	10
Cu/15 mM [tolyl-pyr]	~1	45	~1.1	6	11

\* Represents ethylene Faradaic efficiency

## References

1. Ma, Z. et al. CO<sub>2</sub> electroreduction to multicarbon products in strongly acidic electrolyte via synergistically modulating the local microenvironment. *Nat. Commun.* **13**, 7596 (2022).
2. Huang, J. E. et al. CO<sub>2</sub> electrolysis to multicarbon products in strong acid. *Science* **372**, 1074-1078 (2021).
3. Xie, Y. et al. High carbon utilization in CO<sub>2</sub> reduction to multi-carbon products in acidic media. *Nat. Catal.* **5**, 564-570 (2022).
4. Gu, J. et al. Modulating electric field distribution by alkali cations for CO<sub>2</sub> electroreduction in strongly acidic medium. *Nat. Catal.* **5**, 268-276 (2022).
5. Zhao, Y. et al. Conversion of CO<sub>2</sub> to multicarbon products in strong acid by controlling the catalyst microenvironment. *Nature Synthesis* **2**, 403-412 (2023).
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7. Cao, Y. et al. Surface hydroxide promotes CO<sub>2</sub> electrolysis to ethylene in acidic conditions. *Nat. Commun.* **14**, 2387 (2023).
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9. Zhang, J. et al. Accelerating electrochemical CO<sub>2</sub> reduction to multi-carbon products via asymmetric intermediate binding at confined nanointerfaces. *Nat. Commun.* **14**, 1298 (2023).
10. Fan, M. et al. Cationic-group-functionalized electrocatalysts enable stable acidic CO<sub>2</sub> electrolysis. *Nat. Catal.* **6**, 763-772 (2023).
11. Nie, W., Heim, G. P., Watkins, N. B., Agapie, T. & Peters, J. C. Organic additive-derived films on Cu electrodes promote electrochemical CO<sub>2</sub> reduction to C<sub>2+</sub> products under strongly acidic conditions. *Angew. Chem. Int. Ed.* **62**, e202216102 (2023).