

SUPPLEMENTAL MATERIAL

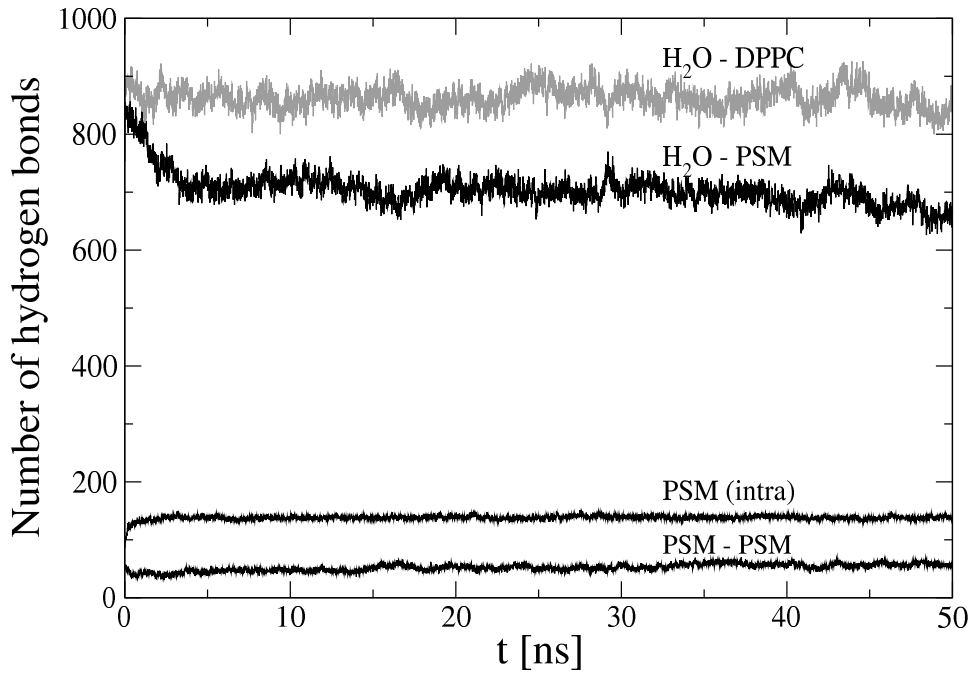


FIG. S-1 Number of different types of hydrogen bonds as a function of time.

TABLE S-I Intermolecular hydrogen bonds: H₂O ··· PSM.

Bond	#N	%	$t_{1/2}$ [ns]
all	711.5 ± 21.1	100.0	–
H ₂ O···O _{P_c}	238.3 ± 7.8	33.5	0.013
H ₂ O···O _{P_d}	233.8 ± 8.0	32.9	0.013
H ₂ O···O _{P_b}	64.6 ± 6.2	9.1	0.009
H ₂ O···H _{NH}	57.7 ± 5.4	8.1	0.185
H ₂ O···O _{OH}	55.4 ± 6.3	7.8	0.010
H ₂ O···O _{PA}	37.9 ± 5.5	5.3	0.009
H ₂ O···H _{OH}	1.4 ± 1.1	0.2	0.008

TABLE S-II Intermolecular hydrogen bonds: H₂O ··· DPPC.

Bond	#N	%	$t_{1/2}$ [ns]
all	862.8 ± 19.8	100.0	–
H ₂ O ··· O _{P_d}	218.4 ± 7.6	25.3	0.009
H ₂ O ··· O _{P_c}	217.8 ± 7.4	25.2	0.009
H ₂ O ··· O _{β2}	207.2 ± 7.8	24.0	0.020
H ₂ O ··· O _{β1}	82.5 ± 7.4	9.6	0.010
H ₂ O ··· O _{P_b}	65.7 ± 6.3	7.6	0.007
H ₂ O ··· O _{α2}	31.5 ± 5.4	3.6	0.009
H ₂ O ··· O _{P_a}	26.1 ± 4.6	3.0	0.006
H ₂ O ··· O _{α1}	13.7 ± 3.6	1.6	0.007