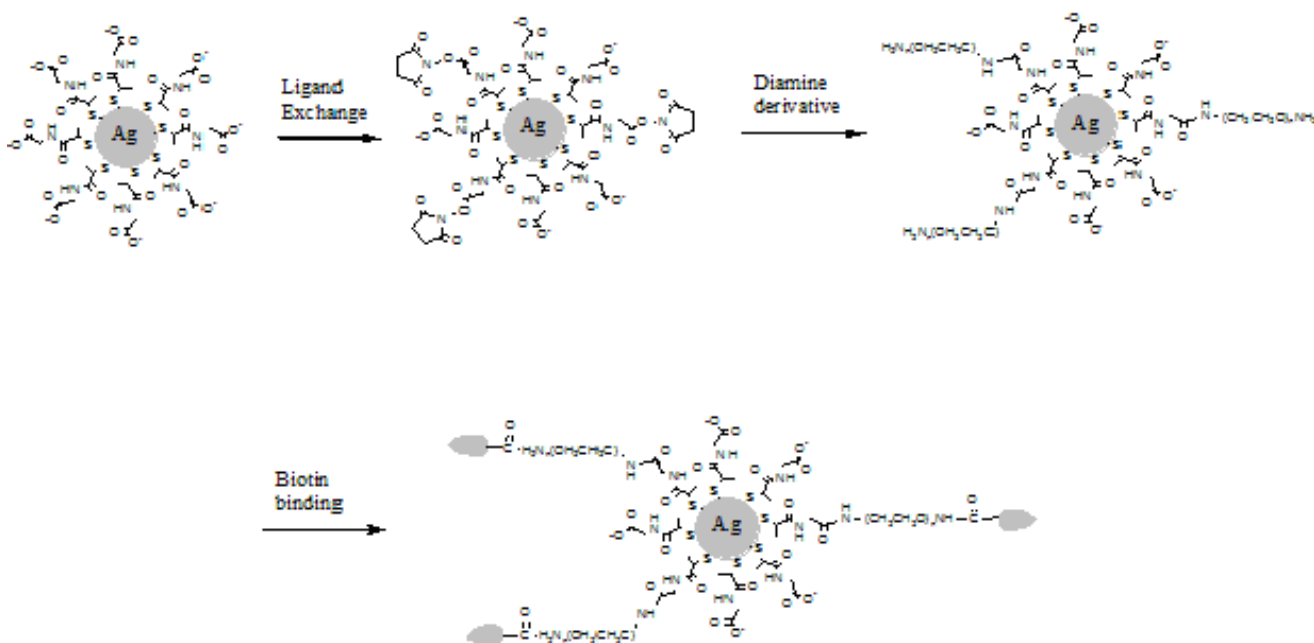


## Supplemental Information

### Silver-Enhanced Fluorescence Emission of Single Quantum Dot Nanocomposites

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#### Synthesis and Characterization of Biotin-coated Silver Nanoparticles

All reagents and spectroscopic grade solvents were used as received from Sigma-Aldrich. Cy5-labeled avidin was commercially from Molecule Probe. RC dialysis membrane (MWCO 50,000) was purchased from Spectrum Laboratories, Inc. Nanopure water ( $>18.0 \text{ M}\Omega\text{cm}^{-1}$ ) purified using Millipore Milli-Q gradient system, was used in all experiments. (2-mercapto-propionylamino) acetic acid 2,5-dioxo-pyrrolidin-1-ylester was synthesized as our previous report [1].

### **Preparation and terminal succinimidylation of tiopronin-coated silver nanoparticle.**

Tiopronin-coated silver nanoparticles were prepared using a modified Brust reaction with a mole ratio of tiopronin / silver nitrate = 1 / 6 in methanol using excess amount of sodium borohydride as reducing agent [2]. These silver particles were succinimidylated via ligand exchange [2,3]. (2-mercapto-propionylamino) acetic acid 2,5-dioxo-pyrrolidin-1-ylester ( $2 \times 10^{-6}$  M) and silver particle ( $2 \times 10^{-7}$  M) were co-dissolved in a mixing solvent of water / ethanol (v/v = 1/1) and stirred for 72 h at room temperature. The ligands displacements were expected to occur on the metal cores in a mole ratio of 1:1 [4,5]. Unbound compounds were removed by centrifuging at 6,000 rpm for 30 min. The residuals were dispersed in water and then further purified by dialysis against water (MWCO 50,000).

### **Biotinylation of silver particles.**

The succinimidylated ligand on the metal particle was lengthened and aminated by two step surface reactions. The succinimidylated metal particles ( $1 \times 10^{-7}$  M) were co-dispersed in water with an excess amount of poly bis(ethylene glycol) (3-aminopropyl) (MW 1,500,  $5 \times 10^{-6}$  M) [3]. The solution was stirred for 2 h, and then a drop ammonium was added dropwise to block non-reacted terminal succinimidyl esters. The solution was removed by centrifugation at 6,000 rpm. The residue was washed with water, and then dispersed in the same volume of water / methanol mixture in v/v = 1/1. The metal particles ( $1 \times 10^{-7}$  M) were biotinylated by co-dissolving an excess amount of biotin N-succinimidyl ester ( $2 \times 10^{-6}$  M). The solution was stirred for 4 h. The suspension was removed by centrifugation at 6,000 rpm. The residue was washed by methanol and water respectively and then dispersed in 10mM PBS buffer solution at pH = 7.2. The metal particles were purified by dialysis against 10mM.

### **Conjugation of biotinylated metal particles and Streptavidin functionalized QDs.**

Streptavidin QD was dissolved in 10 mM PBS buffer solutions (pH = 7.2) and the concentration was controlled to be  $1 \times 10^{-7}$  M. The biotin-metal particles and QDs were mixed at a fixed molar ratio (1:5) to incubate in buffer solution under stirring and the conjugation occurred at room temperature for at least 8 h [6]. The product was collected after centrifuging. Excess QDs were removed by ethanol precipitation.

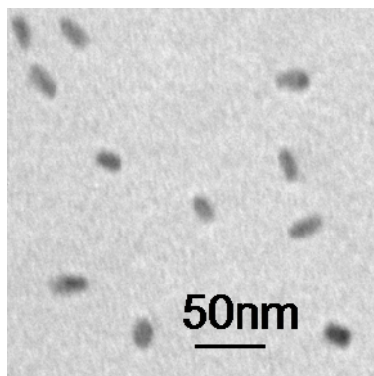
**TEM measurements.** The size distribution of metal core was analyzed with Scion Image Beta Release 2 counting at least 50 particles.

**Reference.**

1. Zhang, J.; Roll, D.; Geddes, C. D.; Lakowicz, J. R. *J. Phys. Chem. B.* **2004**, *108*, 12210.
2. Brust, M.; Walker, M.; Bethell, D.; Schiffrin, D. J.; Whyman, R. *J. Chem. Soc., Chem. Commun.* **1994**, 801.
3. Zhang, J.; Fu, Y.; Lakowicz, J. R. *J. Phys. Chem. C* 2007, *111*, 50.
4. Zhang, J.; Fu, Y.; Chowdhury, M. H.; Lakowicz, J. R. *J. Phys. Chem. C* 2007, *111*, 11784.
5. Templeton, A. C.; Wuelfing, W. P.; Murray, R. W. *Acc. Chem. Res.* 2000, *33*, 27.
6. Ingram, R. S.; Hostetler, M. J.; Murray, R. W. *J. Am. Chem. Soc.* 1997, *119*, 9175.

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**A TEM image of individual Ag nanoparticles**



### Single molecule experiment

Single-molecule measurements were performed using a confocal microscopy system (MicroTime 200, Picoquant, Germany) with an excitation line at 470 nm. Images were recorded by raster scanning (in a bidirectional fashion) the sample over the focused spot of the incident laser with a pixel integration of 0.6 ms. The excitation power into the microscope was maintained less than 0.1  $\mu$ W. Time-dependent fluorescence data were collected with a dwell time of 50 ms. The fluorescence lifetimes of single molecules were measured by time-correlated single photon counting (TCSPC) with the TimeHarp 200 PCI-board (PicoQuant). The data was stored in the time-tagged-time-resolved (TTTR) mode, which allows recording every detected photon with its individual timing information. In combination with a pulsed diode laser, Instrument Response Function (IRF) widths of about 300 ps FWHM can be obtained, which permits the recording of sub-nanosecond fluorescence lifetimes, extendable to less than 100ps with reconvolution. Lifetimes were estimated by fitting to a  $\chi^2$  value of less than 1.2 and with a residuals trace that was fully symmetrical about the zero axis.

### Single-molecule lifetime fitting for intensity decays curves shown in Figure 3

$$\tau_{av} = \frac{\sum_i Ampl_i \cdot Lifet_i^2}{\sum_i Ampl_i Lifet_i}$$

The average lifetime  $\tau_{av}$  shown in the experiment is the amplitude-weighted averaged lifetime calculated from the fit result (Equation 1). In this expression, the values  $Ampl_i$  represent amplitudes of the components, the values  $Lifet_i$  are the decay times.