Supplementary information:

Nanoscale Graphene Oxide (nGO) as Artificial Receptors: Implications for Biomolecular Interactions and Sensing

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nGO synthesis and characterization.

nGOs were synthesized according to reported procedure¹. All materials were from Sigma Aldrich unless indicated otherwise. Briefly, 0.2 g 20 nm carbon fibers (Catalytic Materials LLC) were added to a mixture of H₂SO₄ (5ml), K₂S₂O (0.15 g) and P₂O₅ (0.15 g) at 80 °C. Mixture was stirred for 5 hours, cooled to room temperature, washed with DI water (200 ml) then filtered (Whatman, Grade 4), dried and transferred to a 50 ml round bottom flask. 25 ml H_2SO_4 was added and the mixture was chilled to 0 °C using an ice bath. KMnO₄ (1 g) was added to *slowly* with stirring and temperature maintained below 10 ^oC. After addition, reaction was allowed to proceed at room temperature over night. Product was transferred in to large flask and chilled to 0° C with stirring. 100 ml DI water was added *slowly*, with temperature maintained below 55 °C using an ice bath. 5 ml of 30% H_2O_2 was then added *slowly*, turning solution yellow. Mixture was then *slowly* titrated with 3.4% HCL to neutralize bases. Resultant nGO mixture was collected and dialyzed against DI water for ten days using MW 7000 dialysis membranes (Sigma Aldrich). After dialysis, mixture was briefly centrifuges at 3000 rcf to remove large aggregates. Remaining nGOs were air dried, weighed and dispersed in water at 4 mg/ml concentration. Weight concentration was confirmed with lypholyzed aliquots.

 Absorbance profiles of nGO were obtained by diluting successive aliquots of concentrated GO (4 mg/ml, 5 μ l) in to 600 μ l of 5 mM phosphate buffer (PB, pH 7.4, Figure S1a). Extinction coefficient was calculated by monitoring changes at $\lambda = 234$ nm. For the 20 nm nGO, extinction coefficient (ε_{234}) of 0.0272 ml μ g⁻¹ cm⁻¹ was observed (Figure S1b). Considering this extinction coefficient, concentrated stock (4 mg/ml, 109 OD) was diluted to 5.9 OD (214 µg/ml) using PB for further use. Post dilution concentration (5.9 OD) was verified absorbance spectroscopy.

Figure S1 Calibration of nGO concentration through optical methods. a. Absorbance profiles of nGO in 5 mM PB, pH 7.42. **b.** Extinction coefficient of nGO was derived through linear fitting of nGO concentration (ug/ml) vs. absorbance at 234 nm. $\Box_{234} = 0.0272$ ml μ g⁻¹ cm⁻¹

 Physical characterizations of nGO were performed using DLS, TEM and AFM. DLS characterizations (diameter and zeta potential) were performed in PB using Malvern Zetasizer Nano, while TEM (Hitachi 8100) and AFM (Bruker Nanoscope III) samples were in water before desiccation. For TEM imaging, 10 ul of dilute nano GO was dropped casted onto lacey carbon grid with formvar backing (Ted Pella). Sample was allowed to dry in vacuum before TEM. Diameters of 30 specimens were measured to create the TEM size distribution. AFM samples were prepared by spin coating at 4000 RPM onto silicon substrates with 500 nm silicon dioxide. Images were acquired in tapping mode, diameters were measured by taking the FWHM of each nGO "island". In all, 24 flakes of nGO were characterized on the AFM. Resultant size distributions from all three methods can be observed in figure S2. It is important to recall that DLS measurements assume 3D spherical colloids, thus the numbers are to be taken more as the tumbling diameter of hydrated nGO in solution, rather than actual size.

Figure S2 Size and charge characterizations of nano GO. a. DLS size and zeta potential distributions of nano GO. DLS measurements differ from AFM/TEM measurements because it considers colloids to be spherical, a false assumption in the case of nGOs. The resultant DLS sizes are therefore better interpreted as tumbling diameter of hydrated nGO in solution. Nonetheless, the monodispersity of nGOs is illustrated **b.** AFM size distributions of nGO in probability plot format (demonstrating Gaussian "normal" distribution) and in bar graph formal (with normal curve fitting). **c.** TEM size distributions of nGO.

Conventional-GO synthesis and characterization.

 Conventional GOs were prepared using a modified Hummer's method using graphite powder (Bay carbon, SP-1). In a typical reaction, 2 g of graphite, 1 g of NaNO₃, and 46 mL of H_2SO_4 were stirred together in an ice bath. Next, 6 g of $KMnO₄$ was slowly added maintaining the solution temperature \leq 20 °C. Once mixed, the solution is transferred to a 35 °C water bath and stirred for about 1 h. To that green paste, 40 mL of water was added very slowly and reaction mixture was transferred to a preheated 90 °C water bath. The reaction mixture was stirred for 60 min. Finally, 200 mL of water was added, followed by the slow addition of 6 mL of H_2O_2 (30%). The warm solution was cooled down and washed three times with 200 ml of 10% HCL and 200 ml of water, and the filtered. The filtered cake was then dispersed in water by mechanical agitation overnight. Large aggregates were then removed by lowspeed centrifugation at 500 rpm. The final product was the dialysed for two weeks with twice daily water changes.

 Extinction coefficient of conventional GOs was characterized in 5 mM PB buffer. Extinction coefficient was calculated by monitoring changes at $\lambda = 234$ nm. Absorbance titration showed extinction coefficient (ϵ_{234}) of 0.0360 ml μ g⁻¹ cm⁻¹ was observed (Figure S3). Size distribution were characterized by DLS and SEM (Figure S4).

Figure S3 Calibration of conventional GO concentration through optical methods. a. Absorbance profiles of conventional GO in 5 mM PB, pH 7.42. **b.** Extinction coefficient of nGO was derived through linear fitting of nGO concentration (ug/ml) vs. absorbance at 234 nm. $\Box_{234} = 0.0360$ ml μ g⁻¹ cm⁻¹

Figure S4 Size and charge characterizations of conventional GO. a. DLS size distributions of conventional GO. Due to the polydisperse nature of conventional GOs, results fluctuate greately runrun. Further DLS characterization was thus not informative.. **b.** SEM size distributions of nGO in probability plot format and in bar graph formal. The non linear nature of the probability plot demonstrates the polydisperse and non-Gaussian nature of conventionally synthesized GO.

Absorbance characterization of nGO and conventional GO:Reporter molecule complexes.

Absorbance titrations were performed to characterize changes in *fluorescent reporters* with complexation in nGO. Experiments were carried out in PB, and monitored using the HP 8452 UV/VIS spectrophotometer. To obtain appropriate absorbance profile resolution, 400 µl of PB were first spiked with concentrated fluorophores thus that $\lambda_{\text{max}} \sim 0.2$ prior to GO addition. After recording profiles of uncomplexed fluorophore, aliquots of 2 mg/ml nGO were added sequentially, with a 2 min delay in

between to allow for complexation. Absorbance profiles were recorded after each GO addition (Figure S5 and S6). In all cases, overall volume of solution changed by <4% with GO addition

Figure S5. Absorbance profiles of various fluorophore reporters titrated with nGO.

Figure S6. Absorbance profiles of various fluorophore reporters titrated with conventional GO.

Raman Characterization of conventional and nGO

 Raman spectroscopy was performed to characterize changes in nGO as a consequence of complexation with reporter molecules. Samples for Raman spectroscopy were prepared by quenching the fluorescence of respective fluorophore molecules using nGO, as monitored through a UV-lamp. 5 µl of fully quenched specimens were dried on to Si substrates with 300 µm of CVD deposited SiO₂. After desiccating in vacuum for 6 hrs, samples were immediately characterized on a Nanophoton Raman-11 confocal Raman microscope (excitation wavelength 532nm)(figure S7).

Figure S7. **(a)** Typical Raman D and G band for the conventional-GO and different fluorophore complexes. **(b)** Effect of G band frequency and width upon complexation with different fluorophore.

Fluorescent Titration of nGO and Conventional GO.

Excitation and emission wavelengths used for reporter molecules are listed below:

 The fluorescent titration curves were fitted by first considering the equilibrium binding constant between the fluorophores and the quenching sites on GOs.

$$
K_s = \frac{[F - Q]}{[F][Q]}
$$

Where [F] and [Q] are free fluorophore and quenching site concentrations respectively, and [F - Q] is the concentration of the complex. With consideration that $[F_0] = [F] + [F - Q]$, where $[F_0]$ is total fluorophore concentration, and the number of binding sites can be described by $[Q] = n*[GO]$, the following relationship can be derived.

$$
\frac{[F]}{[F_o]} = (1 + n * K_s [GO])^{-1}
$$

 Because of this equation is similar to an inverted Stern-Volmer quenching equation, we can consider Stern-Volmer quenching constant $K_{s-v} = n*K_s$. The Stern-Volmer quenching constant therefore encompasses the quantity of fluorophores bound by individual nGOs, which can be useful for describing overall binding. All final curved were therefore fitted to the Stern-Volmer equation(figure S8 and S9). In these cases, static quenching is indicated by large K_{S-V} constants.²

$$
\frac{[F]}{[F_o]} = (1 + K_{S-V} [GO])^{-1}
$$

 In regression analysis, the separation of variables relating to GO and fluorophore concentration allows K_{S-V} to take on units dictated by [GO]. It is therefore possible to use regression to fittings by considering [GO] in g/ml or in molar concentrations. For ease of comparison, GO concentrations were converted from g/ml to molar concentrations using the density reported by Dikin et al $(1.8 \text{ g/ml})^3$ and the size of nano GO as measured by AFM and TEM. For nGO, a molecular weight of $6.75*10^9$ g/mol was used. Due to their polydispersity, an useful estimation of molecular weights could not be obtained for the conventional GOs.

Figure S8. nGO fluorescence-quenching curves of the five fluorophores used in the initial nose array. Concentrations of fluorophores used are as followed: eGFP (9.4 nM), PY (522.9 nM), R6G (41.8 nM), AO (2.9 nM), RB (331 nM). Stern-Volmer quenching constants were determined by non-linear regression.

Figure S9. Conventional-GO fluorescence-quenching curves of the five fluorophores used in the initial nose array. Concentrations of fluorophores used are as followed: eGFP (9.4 nM), PY (522.9 nM), R6G (41.8 nM), AO (2.9 nM), RB (331 nM). Stern-Volmer quenching constants were determined by non-linear regression.

Protein	MW	pl	ε_{280} (M ⁻¹ cm ⁻¹)
	(kDa)		
Bovine serum albumin (BSA)	66.3	4.8	46860
β -galactosidase (β -Gal)	540	46	1128600
Hemoglobin (Hem)	64.5	68	125000
Histone ⁴⁴	21.5	11	3840
Lipase (Lip)	58	56	54350
Lysozyme (Lys)	14.4	11	38000
Myoglobin (Mayo)	17	72	13940
Ribonuclease A (Rib-A)	13.7	94	10000

Table S1. The eight analyte proteins used in sensing, and their properties of interest.

Sensing data classification

A Brief overview of Linear Discriminant Analysis (LDA)

Differential sensing utilizes many receptors that exhibit selective but nonspecific binding toward analytes. The cross-reactive nature of these receptors enables them to react to different degrees with a multitude of different analytes. They are thusly termed "differential", rather than "specific". Since these receptors are not specific for any individual analyte, their collective response in an array is unique to a particular analyte being studied. This creates a signal pattern in the array that is unique and specific for each analyte, and can be used as a "memory" or "training matrix" for identification of this analyte.⁵ This approach is similar to how a mammalian nose/tongue classifies and identifies unknown odors/test.

As the array size increases, the data generated becomes gradually more complex. The accurate interpretation of this data can thus be challenging. To solve this dilemma, chemometric tools can be deployed. The most popular chemometric analysis tools have been linear discriminant analysis (LDA), principal component analysis (PCA), and hierarchical cluster analysis (HCA).^{6, 7}

In LDA, the "memory" or "training matrix" is generated using "known samples." The discriminant functions are then obtained to create to maximize the separation between analyte classes while minimizing the separation within each analyte class. This approach is often called the supervised technique, since the classification of the analytes is known before the analysis. The confidence of the separation in LDA plots are checked using jack-knife analysis. The jackknife method removes a single individual data point from the set and recreates the mathematical functions in the absence of the omitted point. That data point is then treated as a pseudounknown, and is re-classified in accordance with a new set of discriminant functions derived in its absence. This process is then repeated for every data point. The accuracy of the reassignment represents the classification ability of the model. Once a training matrix is complete, true unknowns can be used as inputs and classified according to the LDA plots created from the training matrix. By calculating the Mahalanobis distance, that is, the proximity to known group centers, each unknown sample can then be identified.

	nGO				Conventional GO					
PROT EIN	eGFP	PY	R6G	AO	RB	eGFP	PY	R ₆ G	AO	RB
β -Gal	36541.3	-4517.4	4668.9	5160.4	-144.2	2302.0	-4537.5	7255.0	3392.5	-954.5
β -Gal	43045.3	-3737.4	6992.9	2522.4	-1146.2	3849.0	-3738.5	8626.0	3697.5	6114.5
β -Gal	32189.3	-3983.4	6964.9	2409.4	-3149.2	4041.0	-4048.5	11223.0	4440.5	-51.5
β -Gal	44752.3	-3050.4	8595.9	3221.4	-871.2	3954.0	-4247.5	8958.0	4301.5	5244.5
β -Gal	42237.3	-3259.4	8536.9	2518.4	1796.8	3988.0	-4627.5	10769.0	5892.5	9202.5
BSA	5067.3	1632.6	-3197.1	-181.6	-4465.2	1322.5	-1676.5	5273.0	-363.0	-996.0
BSA	3540.3	1378.6	-1228.1	-496.6	-2685.2	1187.5	-503.5	5400.0	-90.0	661.0
BSA	3962.3	2256.6	-1489.1	-246.6	-3357.2	1130.5	-277.5	5732.0	-73.0	873.0
BSA	5238.3	3549.6	-713.1	-237.6	-2967.2	979.5	-176.5	6807.0	-79.0	207.0
BSA	5044.3	1043.6	-1909.1	-598.6	-1799.2	1005.5	-1212.5	4946.0	-254.0	605.0
Hem	20222.6	3636.8	6037.4	-1720.1	3214.5	1683.5	-2084.5	7462.0	-1116.5	4584.0
Hem	21346.6	4962.8	5546.4	-1843.1	3466.5	1776.5	-1989.5	8206.0	-1053.5	-1513.0
Hem	18338.6	4673.8	6784.4	-1779.1	2173.5	2021.5	-2163.5	5579.0	-1009.5	2946.0
Hem	18582.6	4086.8	7999.4	-1857.1	-1978.5	1933.5	-1730.5	6529.0	-1163.5	4369.0
Hem	18418.6	3167.8	7655.4	-2038.1	-919.5	1823.5	-1191.5	7486.0	-1019.5	2545.0
His	27673.3	15869.6	20397.9	3287.4	-2903.2	1565.0	2907.0	3841.5	462.5	-1127.0
His	31552.3	16073.6	27220.9	3073.4	-3446.2	1141.0	3454.0	3083.5	461.5	-4015.0
His	26198.3	15670.6	28247.9	2311.4	-2520.2	1518.0	4290.0	1671.5	577.5	-4129.0
His	29381.3	11259.6	25047.9	1049.4	-2683.2	1587.0	4592.0	3576.5	609.5	-1195.0
His	32668.3	18596.6	30154.9	2757.4	-1309.2	1702.0	4761.0	5126.5	445.5	-2380.0
Lip	1251.6	-2226.3	13037.4	-2439.1	2193.5	601.5	-2777.5	1571.5	-380.0	-1558.5
Lip	1918.6	-1022.3	13681.4	-2320.1	3880.5	613.5	-2550.5	-648.5	-363.0	-3299.5
Lip	1817.6	-1159.3	15665.4	-1973.1	5730.5	432.5	-1447.5	1016.5	-349.0	-1475.5
Lip	1491.6	-1041.3	15416.4	-1917.1	1092.5	580.5	-2773.5	1613.5	-338.0	-3276.5
Lip	2029.6	-1755.3	15101.4	-2161.1	2364.5	481.5	-2101.5	1185.5	-290.0	-4465.5
Lys	5039.1	5776.4	13037.4	1561.7	-2364.0	345.0	7702.0	16988.0	879.5	-1792.0
Lys	5611.1	3280.4	17929.4	1128.7	-2603.0	647.0	5701.0	14649.0	546.5	-3091.0
Lys	5708.1	5447.4	13681.4	1317.7	-4222.0	456.0	5200.0	15345.0	463.5	-2210.0
Lys	6059.1	6003.4	15416.4	1267.7	-2788.0	429.0	5280.0	10449.0	546.5	-2474.0
Lys	8485.1	5852.4	15101.4	1860.7	-2231.0	455.0	4432.0	16075.0	400.5	-2800.0
Myo	7637.6	920.8	11903.4	-2254.1	4197.5	1667.5	886.0	10232.0	-436.5	-342.5
Myo	6175.6	1407.8	12499.4	-1721.1	4788.5	1131.5	1206.0	11095.0	-320.5	-2613.5
Myo	9219.6	1860.8	10602.4	-1612.1	4153.5	1315.5	899.0	8456.0	-484.5	-2380.5
Myo	7108.6	1795.8	13288.4	-1770.1	4129.5	1796.5	2599.0	9336.0	-352.5	-1380.5
Myo	7903.6	1348.8	12938.4	-1825.1	2504.5	1388.5	2255.0	10302.0	-368.5	-327.5
RibA	38510.6	14011.8	1806.4	1522.9	-333.5	2055.5	5725.0	17285.0	749.5	581.5
RibA	43226.6	15482.8	121.4	1816.9	816.5	1764.5	4776.0	16086.0	738.5	-1748.5
RibA	42664.6	15721.8	1333.4	1867.9	3911.5	1798.5	5859.0	12906.0	688.5	-1516.5
RibA	45363.6	16680.8	-528.6	2421.9	3642.5	1760.5	5275.0	13378.0	884.5	-1426.5
RibA	40963.6	15940.8	219.4	1947.9	2551.5	1786.5	6713.0	18272.0	791.5	-1534.5

Table S2. Training matrix data from 100 nM sensing experiments. Data from the nGO array and conventional GO array are shown in tandem.

(a)

Nano-GO array: correlation of fluorescence response to LDA scores

 (b) Conventional-GO array: correlation of fluorescence response to LDA scores

 (c) Nano-GO array: correlation of fluorescence response to analyte protein properties

Figure S10. (a) Correlation of fluorescent response from the nGO array in presence of 100 nM analyte proteins, and their LDA scores. (b) Correlation of fluorescent response from the conventionalGO array in presence of 100 nM analyte proteins, and their LDA scores. (c) Correlation between the fluorescent responses of the nGO array and analyte protein properties (100 nM). Here, PY response appears strongly correlated to protein PI. For Correlations with MW, all correlations appears dominated by β-Gal's

significantly higher mass. With β-gal removed, remaining correlations are not suggestive of data interdependence. (d) Correlation between the fluorescent responses of the conventional-GO array and analyte protein properties (100 nM). Similarly to the nGO array, PY response appears to correlate strongly to PI.

Table S3. Results of 100 nM unknown sample identification using the nGO array.

Table S4. Results of 100 nM unknown sample identification using the conventional GO array.

T**able S5.** Training matrix data from 10 nM sensing experiments.

Figure S11. Protein discrimination at 10 nM using only three dimensions (R6G, PY and eGFP) demonstrated comparable results to separation using five dimensions. This further validated the redundancy of the discarded dimensions, RB and AO.

Table S6. Unknown identification at 10 nM using nGO array.				
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