Electronic Supplementary Information for

Facile preparation of hierarchically porous carbons from metal-organic gels and their application in energy storage

Wei Xia, Bin Qiu, Dingguo Xia & Ruqiang Zou*

* rzou@pku.edu.cn



Figure S1. PXRD patterns of MOA, MOX and the simulated one from crystallographic information file of the MIL-100(Al) crystal.



Figure S2. Photograph of a super light MOA monolith standing on a piece of note paper.



Figure S3. Images of the MOA: (a) Digital photograph, (b) SEM and c) TEM.



Figure S4. View of MOX samples: (a) Digital photograph, (b) SEM and (c) TEM.



Figure S5. TGA curves of MOX and MOA.

Sample	C (wt %)	H (wt %)	N (wt %)	O (wt %)
MOX-C	93.4	0.8	0.4	5.4
MOA-C	91.3	2.1	0.5	6.1

Table S1 Elemental analysis of the carbon products.



Figure S6. PXRD patterns of the MOX-C and MOA-C.



Figure S7. Raman spectra of MOX-C and MOA-C.



Figure S8. TEM image of MOX-C.



Figure S9. TEM image of MOA-C.



Figure S10. N₂ sorption isotherms at 77K and pore size distributions (inset) based on the QSDFT model of MOX derived carbon before KOH treatment. The BET surface area and pore volume are $2131 \text{ m}^2\text{g}^{-1}$ and $1.63 \text{ cm}^3\text{g}^{-1}$, respectively.



Figure S11. N₂ sorption isotherms at 77K and pore size distributions (inset) based on the QSDFT model of KOH treated MOA-C sample. The BET surface area and pore volume are 2523 m²g⁻¹ and 2.17 cm³g⁻¹, respectively. After KOH activation, the BET surface area of the sample was enhanced (1820 m²g⁻¹ for MOA-C). However, the pore volume dramatically dropped down (3.22 cm³g⁻¹ for MOA-C) due to the collapse of the macropores.



Figure S12. PXRD patterns of a mixture of MOA-C and sulfur (i) before heating and (ii) after heating at 300 °C.



Figure S13. N₂ sorption isotherms at 77K and pore size distributions (inset) based on the QSDFT model of MOA-C before and after sulfur impregnation.



Figure S14. TGA curves of bare sulfur, MOX-C/S and MOA-C/S composites.