

Table S1. Supplementary data on solubility of salts and buffers

Substance (molecular mass)	Solubility	Temp (°C)	Molarity	Reference
NaCl (58.44)	358 g/L.	20	6.13	EC Joint Research Centre Institute for Health and Consumer Protection. http://esis.jrc.ec.europa.eu/
KCl (74.55)	344 g/L	20	4.61	CRC Handbook of Chemistry and Physics 1991-92, 72 nd ed. DR Lide (ed).
CaCl₂ (110.98)	745 g/L	20	6.71	CRC Handbook of Chemistry and Physics 1991-92, 72 nd ed DR Lide (ed).
Na₂SO₄ (142.04)	190g/ 1000g	20	1.29	Wells, RC. 1923. Sodium sulphate: its sources and uses. Bulletin 717, United States Geological Survey Department of the Interior. Government Printing Office, Washington.
Na₂HPO₄ (141.96)		20	0.36	By interpolation of solubility data at other temperatures (Table S2)
Tris (121.14)	~500 g/L	25	4.13	http://en.wikipedia.org/wiki/Tris
HEPES (238.3)		0	2.25	E. Good E, Winget GD, Winter W, Connolly TN, Izawa S, Singh RMM. 1966. Hydrogen Ion Buffers for Biological Research. <i>Biochemistry</i> 5 : 467–477. doi:10.1021/bi00866a011. PMID 5942950.
MOPS (209.3)	980 g/L	20	4.7	http://www.biospectra.us/p-14-mops-free-acid.aspx
Dimethyl glutarate (160.2)	43g/L	20	0.27	U.S. Environmental Protection Agency. Supporting Documents for Risk-Based Prioritization. 3/18/2008. http://www.epa.gov/hpvis/rbp/Dibasic%20esters.Web.SupportDocs.031808.pdf .

Table S2 Solubility of disodium phosphate at different temperatures*

Temp (C)	Solubility (M)
0	0.12
17	0.31
20	0.54
25	0.85
25	0.44
34	2.44
40	3.88
50	5.62
80	6.57

* CRC Handbook of Chemistry and Physics 1991-92. 72nd ed. DR Lide (ed).