Supporting Information

for

Carbon dioxide hydrogenation to aromatic hydrocarbons by using an iron/iron oxide nanocatalyst

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Additional experimental data

Reaction kinetics



Figure S1: Consumption of CO_2 at 400 °C. A: CO_2 peak as a function of reaction time, as recorded by GC–MS. B: The reaction is a heterogeneous first order kinetics, as the plot ln(signal intensity) vs t indicates.



Additional GC–MS chromatograms and mass spectra

Figure S2: GC–MS chromatogram showing the major reaction intermediates that are formed at 400 °C.



Figure S3: Mass spectrum of carbon suboxide (C_3O_2) .



Figure S4: Mass spectrum of 3-oxoacrylaldehyde (C₃H₂O₂).



Figure S5: GC–MS chromatogram showing the major reaction products that are formed at 500 °C.



Figure S6: Mass spectrum of propionaldehyde (C₃H₅O).



Figure S7: Mass spectrum of benzene (C₆H₆).



Figure S8: Mass spectrum of toluene (C₇H₈).



Figure S9: Mass spectrum of *meta*-xylene (C_8H_{10}).



Figure S10: Mass spectrum of *para*-xylene (C_8H_{10}).



Figure S11: Mass spectrum of *ortho*-xylene (C₈H₁₀).



Figure S12: Mass spectrum of mesitylene (C₉H₁₂).

Additional XRD Spectra



Figure S13: X-ray diffractogram of the Fe/Fe₃O₄-catalyst after 5 catalytic runs (2 h each).



