

SUPPORTING INFORMATION

for

**Biomimetic Aerobic C–H Olefination of Cyclic Enaminones at Room Temperature:
Development toward the Synthesis of 1,3,5-Trisubstituted Benzenes**

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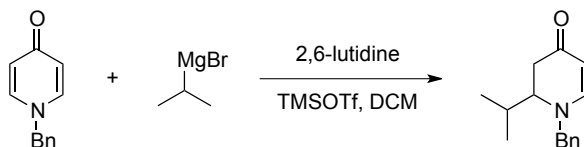
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1. Preparation of the starting materials

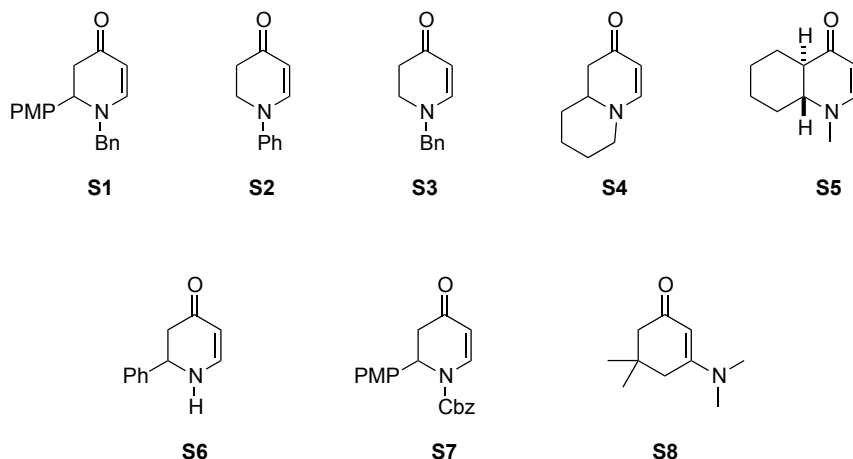
1) Preparation of *N*-benzyl-2-isopropyl-2,3-dihydropyridin-4(1*H*)-one



N-Benzyl-4-pyridone (3.0 g, 16.20 mmol) in dry CH₂Cl₂ (30 mL) was treated with TMSOTf (6.3 mL, 32.39 mmol) at room temperature under a N₂ atmosphere. After the reaction was stirred for 1 h, 2,6-lutidine (3.8 mL, 32.39 mmol) was added, followed by slow addition of a 2.0 M solution of isopropylmagnesium bromide in THF (12.2 mL, 24.29 mmol) with a syringe pump. The reaction was stirred for another 2 h and then quenched with saturated NH₄Cl (15 mL). The mixture was extracted with CH₂Cl₂ (15 mL × 4). The combined organic layers were dried over Na₂SO₄, filtered, and concentrated under reduced pressure. Further purification was carried out by flash column chromatography (3:1 EtOAc/hexanes) to produce 2.1 g (57%) of *N*-benzyl-2-isopropyl-2,3-dihydropyridin-4(1*H*)-one as an orange oil. ¹H NMR (400 MHz, CDCl₃) δ 7.41–7.30 (m, 3H), 7.29–7.24 (m, 2H), 7.12 (d, *J* = 7.4 Hz, 1H), 4.91 (d, *J* = 7.4 Hz, 1H), 4.47 (d, *J* = 15.4 Hz, 1H), 4.39 (d, *J* = 15.4 Hz, 1H), 3.22 (td, *J* = 6.9, 4.0 Hz, 1H), 2.60 (dd, *J* = 16.7, 7.5 Hz, 1H), 2.39 (dd, *J* = 16.7, 3.8 Hz, 1H), 2.28 (dq, *J* = 13.4, 6.7 Hz, 1H), 0.95 (d, *J* = 6.8 Hz, 6H); ¹³C NMR (100 MHz, CDCl₃) δ 191.3, 153.7, 137.0, 129.2, 128.4, 127.4, 97.5, 61.4, 58.8, 36.5, 29.1, 19.8, 18.1; FTIR (Film, cm⁻¹) 3054, 2984, 2876, 1635, 1591, 1496, 1455, 1441, 1421, 1387, 1357, 1344, 1204, 1166, 1094, 1077, 1029, 1014, 1002, 955, 896; HRMS (ESI+) *m/z* calculated for [M+Na]⁺ C₁₅H₁₉NONa: 252.1359, found 252.1354.

2) Synthesis of other cyclic enaminones

Other cyclic enaminones were prepared according to the reported procedures as follows (Scheme S1): **S1**,^[1] **S2**,^[2] **S3**,^[3] **S4**,^[4] **S5**,^[4] **S6**,^[2] **S7**,^[5] **S8**.^[6]

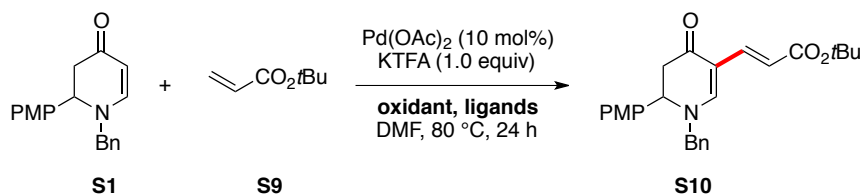


Scheme S1. Preparation of cyclic enaminones.

2. Optimization data for the aerobic dehydrogenative alkenylation

1) Initial trials

Table S1. Initial attempts of aerobic oxidative alkenylation of cyclic enaminones

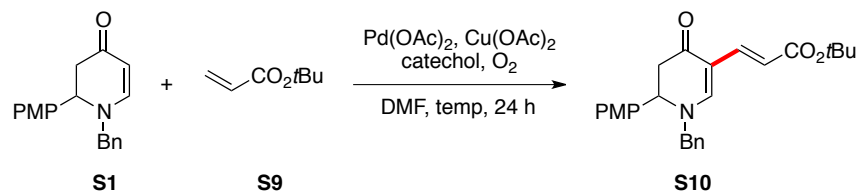


Entry ^a	Cu ^{II} (0.2)	Oxidant ^b	Ligand (0.2)	% Consumption ^c	% Yield ^d
1	–	air	–	50	27
2	–	O ₂	–	71	44
3	Cu(OAc) ₂	O ₂	–	91	41
4	–	O ₂	Boc-β-Ala-OH	85	49
5	–	O ₂	Boc-Phe-OH	76	45

^a Other conditions: **S1** (0.2 M), **S9** (2 equiv), Pd(OAc)₂ (10 mol %), KTFA (1 equiv) in DMF (0.5 mL) at 80 °C through 24 h. ^b The pressure was 1 atm. ^c NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^d NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard.

2) Introduction of catechol

Table S2. Aerobic oxidative alkenylation of cyclic enaminones with catechol

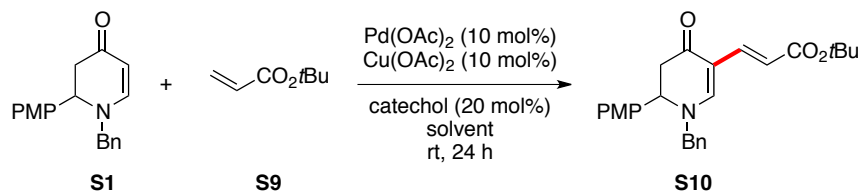


Entry ^a	Pd(OAc) ₂ (mol %)	Cu(OAc) ₂ (mol %)	Catechol (mol %)	Oxidant (1 atm)	Temp (°C)	% Consumption ^b	% Yield ^c
1	5	5	10	O ₂	80	69	40
2	5	5	10	O ₂	rt	49	36
3	5	10	20	O ₂	rt	47	47
4	10	10	20	O ₂	rt	85	78
5	5	10	20	air	rt	49	44
6	10	10	20	air	rt	79	72
7	10	10	10	O ₂	rt	78	72
8	10	20	40	O ₂	rt	77	69

^a Other conditions: **S1** (0.2 M), **S9** (4 equiv) in DMF (0.5 mL) through 24 h. ^b NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^c NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard.

3) Solvent screen

Table S3. Solvent effect on the aerobic oxidative alkenylation of cyclic enaminones

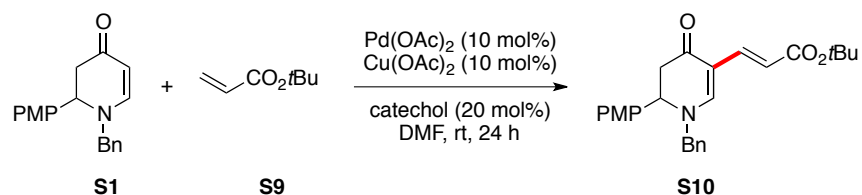


Entry ^a	Solvent	% Consumption ^b	% Yield ^c
1	DMF	85	78
2	MeCN	73	65
3	DMSO	70	62
4	NMP	51	43
5	THF	42	39
6	toluene	65	57

^a Other conditions: **S1** (0.2 M), **S9** (4 equiv), Pd(OAc)₂ (10 mol %), Cu(OAc)₂ (10 mol %), catechol (20 mol %) in DMF (0.5 mL) at rt under O₂ (1 atm) through 24 h. ^b NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^c NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard.

4) Alkene concentration

Table S4. Alkene concentration on the aerobic oxidative alkenylation of cyclic enaminones

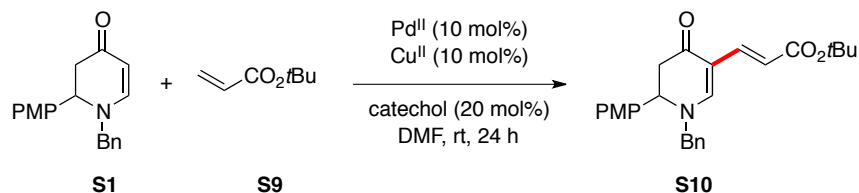


Entry ^a	S9 (equiv)	% Consumption ^b	% Yield ^c
1	1.0	59	57
2	2.0	72	69
3	3.0	69	62
4	4.0	85	78
5	5.0	77	74
6	6.0	82	74
7	8.0	77	77

^a Other conditions: **S1** (0.2 M), Pd(OAc)₂ (10 mol %), Cu(OAc)₂ (10 mol %), catechol (20 mol %) in DMF (0.5 mL) at rt under O₂ (1 atm) through 24 h. ^b NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^c NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard.

5) Catalyst combination

Table S5. Catalyst screen for the aerobic oxidative alkenylation of cyclic enaminones

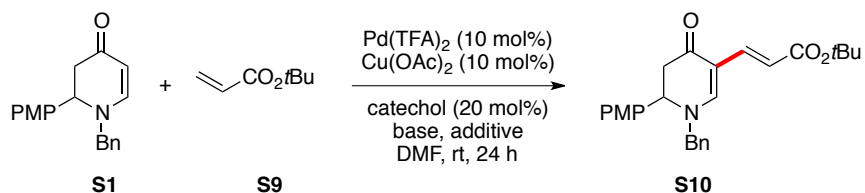


Entry ^a	Pd ^{II} (10 mol %)	Cu ^{II} (10 mol %)	% Consumption ^b	% Yield ^c
1	Pd(OAc) ₂	Cu(OAc) ₂	85	78
2	Pd(TFA) ₂	Cu(TFA) ₂	73	63
3	PdCl ₂	CuCl ₂	53	55
4	Pd(TFA) ₂	Cu(OAc) ₂	90	81
5	Pd(OAc) ₂	Cu(TFA) ₂	86	81
6	PdCl ₂	Cu(OAc) ₂	61	57
7	Pd(OAc) ₂	CuCl ₂	48	37
8	PdCl ₂	Cu(TFA) ₂	43	33
9	Pd(TFA) ₂	CuCl ₂	38	27

^a Other conditions: **S1** (0.2 M), **S9** (4 equiv), catechol (20 mol %) in DMF (0.5 mL) at rt under O₂ (1 atm) through 24 h. ^b NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^c NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard.

6) Additional optimization

Table S6. Further study on the aerobic oxidative alkenylation of cyclic enaminones



Entry ^a	S1 (M)	Base (0.2 equiv)	Additive	% Consumption ^b	% Yield ^c
1	0.2	–	–	90	81
2	0.2	pyridine	–	20	16
3	0.2	Et ₃ N	–	52	51
4	0.2	Na ₂ CO ₃	–	72	69
5	0.2	NaHCO ₃	–	77	73
6	0.4	–	–	100	86
7	0.1	–	–	58	43

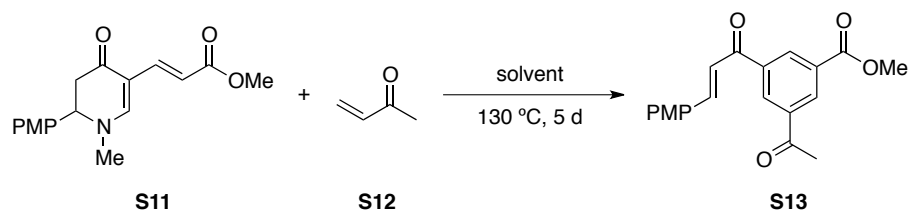
8	0.05	–	–	77	68
9	0.03	–	–	61	52
10	0.2	–	4Å MS	100	91 (89^d)

^a Other conditions: **S9** (4 equiv), Pd(TFA)₂ (10 mol %), Cu(OAc)₂ (10 mol %), catechol (20 mol %) in DMF (0.5 mL) at rt under O₂ (1 atm) through 24 h. ^b NMR % conversion of the starting enaminone with Ph₃SiMe (1.0 equiv) as the internal standard. ^c NMR % yield with Ph₃SiMe (1.0 equiv) as the internal standard. ^d Isolated yield.

3. Condition optimization for the Diels-Alder tandem reaction

1) Solvent screen

Table S7. Solvent effect on the Diels-Alder tandem reaction

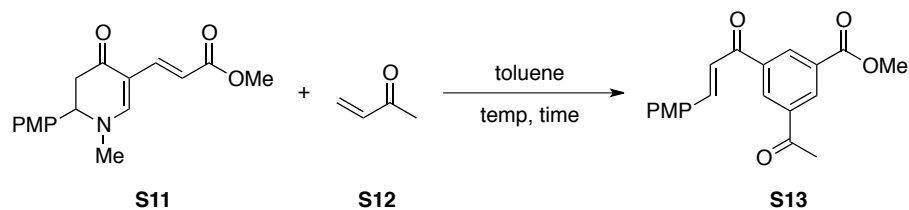


Entry ^a	Solvent (1 mL)	% Yield
1	toluene	44
2	DCE	45
3	MeCN	36
4	dioxane	35
5	DMAc	28
6	THF	0

^a Conditions: Enaminone **S11** (0.07 M), alkene **S12** (2 equiv) at 130 °C for 5 d. ¹H NMR yield based on 1.0 equiv of Ph₃SiMe as the internal standard.

2) Reaction time and temperature

Table S8. Reaction time and temperature effect on the Diels-Alder tandem reaction



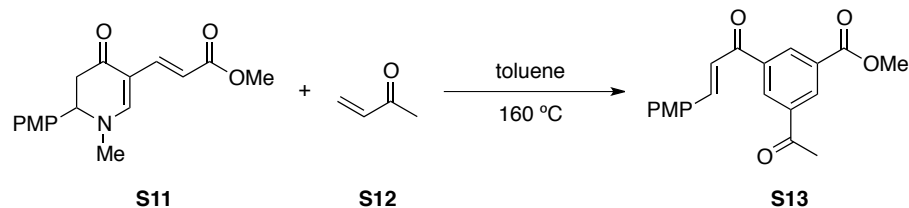
Entry ^a	Time (h)	Temp (°C)	% Yield
1 ^b	0.5	180	17
2 ^b	2	180	44
3 ^b	4	180	48
4 ^b	8	180	51

5 ^b	10	180	52
6 ^b	2	200	46
7 ^b	4	200	19
8	24	150	31
9	24	160	56
10	24	170	40
11	24	180	45
12	15	180	43

^a Conditions: Enaminone **S11** (0.07 M), alkene **S12** (4 equiv) in toluene (1 mL). ¹H NMR yield based on 1.0 equiv of Ph₃SiMe as the internal standard. ^b Under microwave.

3) Stoichiometry

Table S9. Effect of reagent stoichiometry on the Diels-Alder tandem reaction

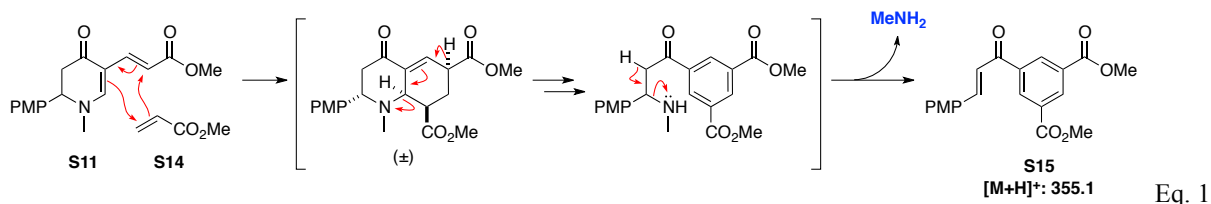


Entry ^a	S12 (equiv)	Time (h)	% Yield
1	1.1	24	27
2	2.0	24	42
3	4.0	24	56
4	8.0	24	63 (66^b)
5	15.0	24	59
6	30.0	24	61
7	2.0	48	43
8 ^c	2.0	24	17

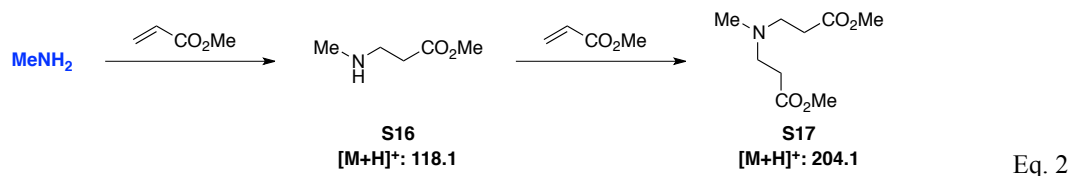
^a Other conditions: Enaminone **S11** (0.07 M) in toluene (1 mL) at 160 °C. ¹H NMR yield based on 1.0 equiv of Ph₃SiMe as the internal standard. ^b Isolated yield. ^c Under 1 atm O₂.

4. Mechanistic study of the Diels-Alder tandem reaction

We proposed that methylamine was released during retro-Michael fragmentation in the Diels-Alder tandem sequence of synthesizing 1,3,5-trisubstituted benzenes. In order to support our hypothesis, diene **S11** (2.5 mg, 8.3 μmol) was treated with dienophile **S14** (6.0 μL, 66.4 μmol) in toluene (0.5 mL) at 160 °C for 15 h (Eq. 1). The crude mixture from the tandem reaction was initially subjected to a GC-MS analysis to detect methylamine. However, no methylamine was found.



In light of the excess amount of dienophile in presence (*i.e.* 8.0 equiv of acrylate **S14**), we postulated that the release of a stoichiometric amount of methylamine could be readily trapped by acrylate **S14** to yield aza-Michael adducts **S16** and/or **S17** in Eq. 2.

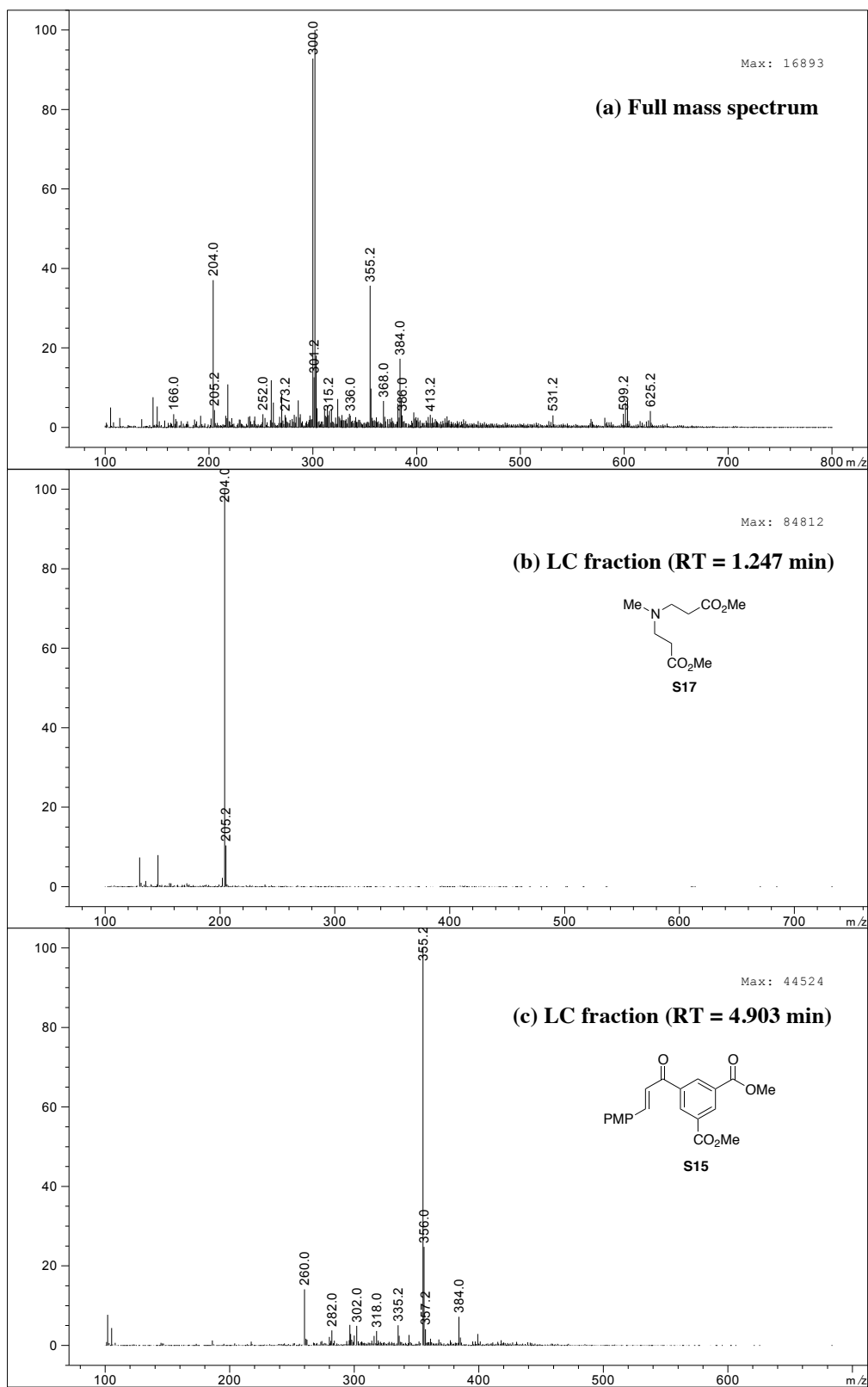


Hence, an LC-MS analysis was conducted on an Agilent LC-MSD ChemStation (Figure S1). An initial full MS scan of the crude mixture confirmed the product peak at m/z 355.2 ($[\mathbf{S15}+\mathbf{H}]^+$, Figure S1-a). In addition, another prominent peak at m/z 204.0 ($[\mathbf{S17}+\mathbf{H}]^+$) was also detected corresponding to the aza-Michael adduct **S17** (Figure S1-a). No peak from the possible adduct **S16** was observed. To our delight, the LC-MS separation reconfirmed the two distinct peaks from **S17** (Figure S1-b, RT = 1.247 min) and **S15** (Figure S1-c, RT = 4.903 min) respectively, which provides strong evidence for our proposed mechanism. In fact, the fast scavenging of byproduct methylamine may be one driving force for the tandem reaction to proceed smoothly.

5. References

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Figure S1. LC-MS analysis of a Diels-Alder tandem reaction between **S11** and **S14**



6. ¹H NMR and ¹³C NMR spectra for new compounds

