

SUPPLEMENTARY INFORMATION FOR

High-throughput design and optimization of fast lithium ion conductors by the combination of bond-valence method and density functional theory

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Table S1. The activation energies (E_a) of lithium migration calculated by BV-based method and DFT method for 46 compounds.

| No. | Compounds | E_a by DFT | E_a by BV | Comments |
|-----|--|--------------|-------------|----------------------|
| 1 | LiMnPO ₄ | 0.31[1] | 1.03 | Olivine[2] |
| 2 | LiFePO ₄ | 0.32[1] | 0.93 | Olivine[3] |
| 3 | LiCoPO ₄ | 0.28[1] | 0.90 | Olivine[4] |
| 4 | LiNiPO ₄ | 0.22[1] | 0.91 | Olivine[2] |
| 5 | LiMgPO ₄ | 0.30[5] | 0.88 | Olivine[6] |
| 6 | LiAlSiO ₄ | 0.20[5] | 0.34 | Olivine[7] |
| 7 | LiScSiO ₄ | 0.33[5] | 0.95 | Olivine[9] |
| 8 | LiInSiO ₄ | 0.40[5] | 0.97 | Olivine[10] |
| 9 | LiTmSiO ₄ | 0.43[5] | 0.98 | Olivine[11] |
| 10 | LiAlGeO ₄ | 0.11[5] | 0.56 | Olivine[12] |
| 11 | LiInGeO ₄ | 0.28[5] | 0.86 | Olivine[13] |
| 12 | LiScGeO ₄ | 0.22[5] | 0.84 | Olivine[14] |
| 13 | LiMn ₂ O ₄ | 0.39[15] | 0.61 | Spinel[16] |
| 14 | LiNi _{1/2} Ge _{3/2} O ₄ | 0.47[17] | 0.67 | Spinel[18] |
| 15 | LiNi _{1/2} Mn _{3/2} O ₄ | 0.32[17] | 0.86 | Spinel[19] |
| 16 | LiTi ₂ O ₄ | 0.28[20] | 0.64 | Spinel[21] |
| 17 | Li ₄ Ti ₅ O ₁₂ | 0.30[22] | 0.59 | Spinel[23] |
| 18 | LiCoO ₂ | 0.84[24] | 2.15 | Layer[25] |
| 19 | LiCrO ₂ | 0.80[26] | 1.96 | Layer[27] |
| 20 | Li ₂ MnO ₃ | 0.54[28] | 1.08 | Layer[29] |
| 21 | Li ₂ IrO ₃ | 0.52[30] | 1.12 | Layer[31] |
| 22 | LiTiS ₂ | 0.70[32] | 1.93 | Layer[33] |
| 23 | LiFeSO ₄ OH | 0.20[34] | 0.82 | Layer[35] |
| 24 | LiMnBO ₃ | 0.36[36] | 0.57 | Borate/Carbonate[37] |
| 25 | LiFeBO ₃ | 0.22[36] | 0.58 | Borate/Carbonate[38] |
| 26 | LiCoBO ₃ | 0.37[36] | 0.63 | Borate/Carbonate[39] |
| 27 | LiBO ₂ | 0.50[40] | 0.83 | Borate/Carbonate[41] |
| 28 | Li _{2.6} C _{0.33} B _{0.67} O ₃ | 0.21[42] | 0.47 | Borate/Carbonate[42] |
| 29 | Li ₂ CO ₃ | 0.30[43] | 0.54 | Borate/Carbonate[44] |
| 30 | Li ₃ OCl | 0.37[45] | 0.98 | Antiperovskite[46] |
| 31 | Li ₃ OBr | 0.41[45] | 1.29 | Antiperovskite[46] |
| 32 | β-Li ₃ PO ₄ | 0.39[47] | 0.90 | Phosphate[48] |
| 33 | γ-Li ₃ PO ₄ | 0.29[47] | 0.82 | Phosphate[49] |
| 34 | β-Li ₃ PS ₄ | 0.20[50] | 0.78 | Thiophosphate[51] |
| 35 | Li ₇ P ₃ S ₁₁ | 0.24[52] | 0.82 | Thiophosphate[53] |
| 36 | Li ₂ Fe(SO ₄) ₂ | 0.47[54] | 0.88 | Sulphate[55] |
| 37 | Li ₂ Mn(SO ₄) ₂ | 0.54[54] | 0.92 | Sulphate[56] |
| 38 | Li ₂ Co(SO ₄) ₂ | 0.42[54] | 0.84 | Sulphate[55] |
| 39 | LiAlPO ₄ F | 0.55[57] | 0.88 | Tavorite[58] |

| | | | | |
|----|---|----------|------|------------------|
| 40 | LiMgSO ₄ F | 0.20[57] | 0.89 | Tavorite[59] |
| 41 | LiFeSO ₄ F | 0.44[60] | 1.02 | Tavorite[61] |
| 42 | FeSO ₄ F | 0.21[62] | 0.60 | Tavorite[62] |
| 43 | LiTi ₂ (PO ₄) ₃ | 0.25[63] | 0.68 | NASICON[64] |
| 44 | Li ₂ NiO ₂ | 0.45[65] | 0.60 | Orthorhombic[66] |
| 45 | Li ₂ O | 0.30[67] | 0.63 | Cubic[68] |
| 46 | LiF | 0.73[69] | 1.07 | Cubic[70] |

References

1. Kutteh, R. & Avdeev, M. Initial Assessment of an Empirical Potential as a Portable Tool for Rapid Investigation of Li⁺ Diffusion in Li⁺-Battery Cathode Materials. *J. Phys. Chem. C* **118**, 11203-11214 (2014).
2. García-Moreno, O. *et al.* Influence of the Structure on the Electrochemical Performance of Lithium Transition Metal Phosphates as Cathodic Materials in Rechargeable Lithium Batteries: A New High-Pressure Form of LiMPO₄ (M = Fe and Ni). *Chem. Mater.* **13**, 1570-1576 (2001).
3. Rouse, G., Rodriguez-Carvajal, J., Patoux, S. & Masquelier, C. Magnetic Structures of the Triphylite LiFePO₄ and of Its Delithiated Form FePO₄. *Chem. Mater.* **15**, 4082-4090 (2003).
4. Kubel, F. Crystal structure of lithium cobalt double orthophosphate, LiCoPO₄. *Z. Kristallogr.* **209**, 755-755 (1994).
5. Jalem, R., Nakayama, M., & Kasuga, T. An efficient rule-based screening approach for discovering fast lithium ion conductors using density functional theory and artificial neural networks. *J. Mater. Chem. A* **2**, 720-734 (2014).
6. Hanic, F., Handlovic, M., Burdova, K. & Majling, J. Crystal structure of lithium magnesium phosphate, LiMgPO₄: Crystal chemistry of the olivine-type compounds. *J. Cryst. Spectrosc. Res.* **12**, 99-127 (1982).
7. Ceriani, C., Fois, E. & Gamba, A. The role of extra-framework cations on the structure of dehydrated Li-ABW. A computer simulation study. *Microporous Mesoporous Mater.* **57**, 73-81 (2003).
8. Blasse, G. & Bril, A. Structure and Eu³⁺-fluorescence of lithium and sodium lanthanide silicates and germinates. *J. Inorg. Nucl. Chem.* **29**, 2231-2241 (1967).
9. Steele, I. M., Pluth, J. J. & Ito, J. Crystal structure of synthetic LiScSiO₄ olivine and comparison with isotypic Mg₂SiO₄. *Zeitschrift für Kristallographie - Crystalline Materials* **147**, 119-127 (1978).
10. Redhammer, G. J. & Roth, G. LiInSiO₄: a new monovalent-trivalent olivine. *Acta Crystallographica C* **59**, 38-40 (2003).
11. Kono, Y., Uematsu, X. & Sato, M. Structural analysis and electrical property of Li-Ln-Si oxide systems. *Kidorui* **24**, 148-149 (1994).
12. Tripathi, A., Kim, S. J., Johnson, G. M. & Parise, J. B. Synthesis and single-crystal structure of a lithium aluminogermanate with the zeolite ABW topology. *Microporous Mesoporous Mater.* **34**, 273-279 (2000).
13. Touboul, M. & Toledano, P. Structure du germanate d'indium et de lithium. *Acta*

- Crystallographica C* **43**, 2004-2006 (1987).
14. Genkina, E. A., Timofeeva, V. A. & Bykov, A. B. Structure of the Olivine-type Li, Sc-germanate LiScGeO_4 . *J. Struct. Chem.* **27**, 167-168 (1986).
 15. Nakayama, M., Kaneko, M. & Wakihara, M. First-principles study of lithium ion migration in lithium transition metal oxides with spinel structure. *Phys. Chem. Chem. Phys.* **14**, 13963-13970 (2012).
 16. Berg, H., Thomas, J. O., Wen, L. & Farrington, G. C. A neutron diffraction study of Ni substituted LiMn_2O_4 . *Solid State Ionics* **112**, 165-168 (1998).
 17. Nakayama, M., Jalem, R. & Kasuga, T. Electronic structure of spinel-type $\text{LiNi}_{1/2}\text{Ge}_{3/2}\text{O}_4$ and $\text{LiNi}_{1/2}\text{Mn}_{3/2}\text{O}_4$ as positive electrodes for rechargeable Li-ion batteries studied by first-principles density functional theory. *Solid State Ionics* **262**, 74-76 (2014).
 18. Hirota, K., Ohtani, M., Mochida, N. & Ohtsuka, A. Formation and structure of spinel solid solution in $\text{Li}_2\text{O} - \text{MO} - \text{GeO}_2$ (M=Zn,Co,Ni) system. *J. Ceram. Soc. Jpn.* **96**,92-96 (1988).
 19. Gryffroy, D., Vandenberghe, R.E., Legrand, E. A Neutron Diffraction Study of some Spinel Compounds Containing Octahedral Ni and Mn at a 1:3 Ratio. *Mater. Sci. Forum* **79**, 785-790 (1991).
 20. Bhattacharya, J. & Van der Ven, A. Phase stability and nondilute Li diffusion in spinel $\text{Li}_{1+x}\text{Ti}_2\text{O}_4$. *Phys. Rev. B* **81**, 104304 (2010).
 21. Cava, J. R., Murphy, D. W. & Zahurak, S. M. The crystal structures of the lithium-inserted metal oxides $\text{Li}_{0.5}\text{TiO}_2$ anatase, LiTi_2O_4 spinel, and $\text{Li}_2\text{Ti}_2\text{O}_4$. *J. Solid State Chem.* **53**, 64-75 (1984).
 22. Ziebarth, B., Klinsmann, M., Eckl, T. & Elsässer, C. Lithium diffusion in the spinel phase $\text{Li}_4\text{Ti}_5\text{O}_{12}$ and in the rocksalt phase $\text{Li}_7\text{Ti}_5\text{O}_{12}$ of lithium titanate from first principles. *Phys. Rev. B* **89**, 174301(2014).
 23. Deschanvres, A., Raveau, B. & Sekkal, Z. Mise en évidence et étude cristallographique d'une nouvelle solution solide de type spinelle $\text{Li}_{1+x}\text{Ti}_{2-x}\text{O}_4$ $0 \leq x \leq 0.33$. *Mater. Res. Bull.* **6**, 699-704 (1971).
 24. Van der Ven, A., Ceder, G., Asta, M. & Tepsch, P. D. First-principles theory of ionic diffusion with nondilute carriers. *Phys. Rev. B*, **64**, 184307 (2001).
 25. Holzapfel, M., Haak, C. & Ott, A. Lithium-Ion Conductors of the System $\text{LiCo}_{1-x}\text{Fe}_x\text{O}_2$, Preparation and Structural Investigation. *J. Solid State Chem.* **156**, 470-479 (2001).
 26. Lyu, Y. C. *et al.* Atomic insight into electrochemical inactivity of lithium chromate (LiCrO_2): Irreversible migration of chromium into lithium layers in surface regions. *J. Power Sources* **273**, 1218-1225 (2015).
 27. Lu, Z.-H. & Dahn, J. R. Structure and Electrochemistry of Layered $\text{Li}[\text{Cr}_x\text{Li}_{(1/3-x/3)}\text{Mn}_{(2/3-2x/3)}]\text{O}_2$. *J. Electrochem. Soc.* **149**, A1454-1459 (2002).
 28. Xiao, R. J., Li, H. & Chen, L. Q. Density Functional Investigation on Li_2MnO_3 . *Chem. Mater.* **24**, 4242-4251 (2012).
 29. Strobel, P. & Lambert-Andron, B. Crystallographic and magnetic structure of Li_2MnO_3 . *J. Solid State Chem.* **75**, 90-98 (1988).
 30. Chen, Y. C., Huo, M., Song, L. J. & Sun, Z. L. Electrical and lithium ion dynamics in Li_2IrO_3 from density functional theory study. *RSC Adv.* **4**, 42462-42466 (2014).
 31. O'Malley, M. J., Verweij, H. & Woodward, P. M. Structure and properties of ordered Li_2IrO_3 and Li_2PtO_3 . *J. Solid State Chem.* **181**, 1803-1809 (2008).
 32. Van der Ven, A., Thomas, J. C. & Xu, Q. C. Nondilute diffusion from first principles: Li diffusion

- in Li_xTiS_2 . *Phys. Rev. B* **78**, 104306 (2008).
33. Dahn, J. R., McKinnon, W. R., Haering, R. R., Buyers, W. J. L. & Powell, B. M. Structure determination of Li_xTiS_2 by neutron diffraction. *Can. J. Phys.* **58**, 207-213 (1980).
 34. Eames, C., Clark, J. M., Rouse, G., Tarascon, J.M. & Islam, M. S. Lithium Migration Pathways and van der Waals Effects in the LiFeSO_4OH Battery Material. *Chem. Mater.* **26**, 3672-3678 (2014).
 35. Subban, C. V. *et al.* Preparation, Structure, and Electrochemistry of Layered Polyanionic Hydroxysulfates: LiMSO_4OH (M = Fe, Co, Mn) Electrodes for Li-Ion Batteries. *J. Am. Chem. Soc.* **135**, 3653-3661 (2013).
 36. Seo, D.H. *et al.* First-principles study on lithium metal borate cathodes for lithium rechargeable batteries. *Phys. Rev. B* **83**, 205127 (2011).
 37. Bondareva, O. S., Simonov, M. A., Egorov-Tismenko, Y. K. & Belov, N. V. Crystal structure of $\text{LiZn}[\text{BO}_3]$ and $\text{LiMn}[\text{BO}_3]$. *Kristallografiya* **23**, 487-490 (1978).
 38. Legagneur, V. *et al.* LiMBO_3 (M=Mn, Fe, Co): synthesis, crystal structure and lithium deinsertion/insertion properties. *Solid State Ionics* **139**, 37-46 (2001).
 39. Piffard, Y., Rangan, K. K., An, Y., Guyomard, D. & Tournoux, M. Cobalt Lithium Orthoborate, LiCoBO_3 . *Acta Crystallographica C* **54**, 1561-1563 (1998).
 40. Islam, M. M., Bredow, T. & Heitjans, P. Formation and Mobility of Li Point Defects in LiBO_2 : A First-Principles Investigation. *J. Phys. Chem. C* **115**, 12343-12349 (2011).
 41. Kirfel, A., Will, G. & Stewart, R. F. The chemical bonding in lithium metaborate, LiBO_2 . Charge densities and electrostatic properties. *Acta Crystallographica B* **39**, 175-185 (1983).
 42. Meng, F. H. *et al.* (in prepare)
 43. Shi, S. Q., Qi, Y., Li, H. & Hector, L. G. Defect Thermodynamics and Diffusion Mechanisms in Li_2CO_3 and Implications for the Solid Electrolyte Interphase in Li-Ion Batteries. *J. Phys. Chem. C* **117**, 8579-8593 (2013).
 44. Idemoto, Y., Ricardson, J. W. Jr., Koura, N., Kohara, S. & Loong, C.-K. Crystal structure of $(\text{Li}_x\text{K}_{1-x})_2\text{CO}_3$ ($x = 0, 0.43, 0.5, 0.62, 1$) by neutron powder diffraction analysis. *J. Phys. Chem. Solids* **59**, 363-376 (1998).
 45. Zhang, Y., Zhao, Y. S. & Chen, C. F. Ab initio study of the stabilities of and mechanism of superionic transport in lithium-rich antiperovskites. *Phys. Rev. B* **87**, 134303 (2013).
 46. Zhao, Y. S. & Daemen, L. L. Superionic Conductivity in Lithium-Rich Anti-Perovskites. *J. Am. Chem. Soc.* **134**, 15042-15047 (2012).
 47. Du, Y. A. & Holzwarth, N. A. W. Mechanisms of Li^+ diffusion in crystalline γ - and β - Li_3PO_4 electrolytes from first principles. *Phys. Rev. B* **76**, 174302 (2007).
 48. Keffer, C., Mighell, A. D., Mauer, F., Swanson, H. & Block, S. Crystal structure of twinned low-temperature lithium phosphate. *Inorganic Chemistry* **6**, 119-125 (1967).
 49. Yakubovich, O. V. & Urosova, V. S. Electron distribution in lithiophosphatite: Crystallochemical features of orthophosphates with hexagonal close packing. *Kristallografiya* **42**, 301-308 (1997).
 50. Lepley, N. D., Holzwarth, N. A. W. & Du, Y. A. Structures, Li^+ mobilities, and interfacial properties of solid electrolytes Li_3PS_4 and Li_3PO_4 from first principles. *Phys. Rev. B*, **88**, 104103 (2013).
 51. Homma, K. *et al.* Crystal structure and phase transitions of the lithium ionic conductor Li_3PS_4 . *Solid State Ionics* **182**, 53-58 (2011).
 52. Xiong, K., Longo, R. C., Santosh, K. C., Wang, W. & Cho, K. Behavior of Li defects in solid

- electrolyte lithium thiophosphate $\text{Li}_7\text{P}_3\text{S}_{11}$: A first principles study. *Comput. Mater. Sci.* **90**, 44-49 (2014).
53. Yamane, H. *et al.* Crystal structure of a superionic conductor, $\text{Li}_7\text{P}_3\text{S}_{11}$. *Solid State Ionics* **178**, 1163-1167 (2007).
 54. Clark, J. M. *et al.* High voltage sulphate cathodes $\text{Li}_2\text{M}(\text{SO}_4)_2$ (M = Fe, Mn, Co): atomic-scale studies of lithium diffusion, surfaces and voltage trends. *J. Mater. Chem. A* **2**, 7446-7453 (2014).
 55. Reynaud, M. *et al.* $\text{Li}_2\text{Fe}(\text{SO}_4)_2$ as a 3.83 V positive electrode material. *Electrochem. Commun.* **21**, 77-80 (2012).
 56. Reynaud, M., Rousse, G., Chotard, J. N., Rodriguez-Carvajal, J. & Tarascon, J. M. Marinite $\text{Li}_2\text{M}(\text{SO}_4)_2$ (M = Co, Fe, Mn) and $\text{Li}_1\text{Fe}(\text{SO}_4)_2$: Model Compounds for Super-Super-Exchange Magnetic Interactions. *Inorg. Chem.* **52**, 10456-10466 (2013).
 57. Jalem, R., Nakayama, M. & Kasuga, T. Lithium ion conduction in tavorite-type LiMXO_4F (M–X: AlP, MgS) candidate solid electrolyte materials. *Solid State Ionics* **262**, 589-592 (2014).
 58. Simonov, V. I. & Belov, N. V. The determination of the structure of amblygonite by the Minimum-Function Method. *Kristallografiya* **3**, 429-438 (1958).
 59. Sebastian, L., Gopalakrishnan, J. & Piffard, Y. Synthesis, crystal structure and lithium ion conductivity of LiMgFSO_4 . *J. Mater. Chem.* **12**, 374-377 (2002).
 60. Tripathi, R., Gardiner, G. R., Islam, M. S. & Nazar, L. F. Alkali-ion Conduction Paths in LiFeSO_4F and NaFeSO_4F Tavorite-Type Cathode Materials. *Chem. Mater.* **23**, 2278-2284 (2011).
 61. Melot, B. C. *et al.* Magnetic Structure and Properties of the Li-Ion Battery Materials FeSO_4F and LiFeSO_4F . *Chem. Mater.* **23**, 2922-2930 (2011).
 62. Mueller, T., Hautier, G., Jain, A. & Ceder, G. Evaluation of Tavorite-Structured Cathode Materials for Lithium-Ion Batteries Using High-Throughput Computing. *Chem. Mater.* **23**, 3854-3862 (2011).
 63. Lu, X. *et al.* (in prepare)
 64. Aatiq, A., Menetrier, M., Croguennec, L., Suard, E. & Delmas, C. On the structure of $\text{Li}_3\text{Ti}_2(\text{PO}_4)_3$. *J. Mater. Chem.* **12**, 2971-2978 (2002).
 65. Kang, K., Morgan, D. & Ceder, G. First principles study of Li diffusion in I- Li_2NiO_2 structure. *Phys. Rev. B* **79**, 014305 (2009).
 66. Rieck, H. & Hoppe, R. Ein neues Oxonicolat: Li_2NiO_2 . *Zeitschrift fuer Anorganische und Allgemeine Chemie* **392**, 193-196 (1972).
 67. Islam, M. M., Bredow, T. & Minot, C. Theoretical Analysis of Structural, Energetic, Electronic, and Defect Properties of Li_2O . *J. Phys. Chem. B* **110**, 9413-9420 (2006).
 68. Farley, T.W.D., Hayes, W., Hull, S., Hutchings, M. T. & Vrtis, M. Investigation of thermally induced Li^+ ion disorder in Li_2O using neutron diffraction. *J. Phys.: Condens. Matter* **3**, 4761-4781 (1991).
 69. Chen, Y. C., Ouyang, C. Y., Song, L. J. & Sun, Z. L. Electrical and Lithium Ion Dynamics in Three Main Components of Solid Electrolyte Interphase from Density Functional Theory Study. *J. Phys. Chem. C* **115**, 7044-7049 (2011).
 70. Streltsov, V. A., Tsirelson, V. G., Ozerov, R. P. & Golovanov, O. A. Electron and Thermal Parameters of Ions in CaF_2 . Results of the Application of the Regularized Least Squares Method. *Kristallografiya* **33**, 90-97 (1988).