# Supporting Information

# Dual-emissive Polymer Dots for Rapid Detection of Fluoride in Pure Water and Biological Systems with Improved Reliability and Accuracy

Qiang Zhao<sup>1,\*</sup>, Chuanqi Zhang<sup>1</sup>, Shujuan Liu<sup>1</sup>, Yahong Liu<sup>1</sup>, Kenneth Yin Zhang<sup>1</sup>, Xiaobo Zhou<sup>1</sup>, Jiayang Jiang<sup>1</sup>, Wenjuan Xu<sup>1</sup>, Tianshe Yang<sup>1</sup>, and Wei Huang<sup>2,\*</sup>

<sup>1</sup>Key Laboratory for Organic Electronics and Information Displays & Institute of Advanced

Materials (IAM), Jiangsu National Synergetic Innovation Center for Advanced Materials

(SICAM), Nanjing University of Posts & Telecommunications, 9 Wenyuan Road, Nanjing

210023, China.

<sup>2</sup>Key Laboratory of Flexible Electronics (KLOFE) & Institute of Advanced Materials (IAM),

Jiangsu National Synergetic Innovation Center for Advanced Materials (SICAM), Nanjing

Tech University (NanjingTech), 30 South Puzhu Road, Nanjing 211816, China.

\*iamqzhao@njupt.edu.cn ; wei-huang@njtech.edu.cn

## 1. Synthesis

The ligands and Ir(III) complexes were synthesized according to the route depicted in Fig. S1.

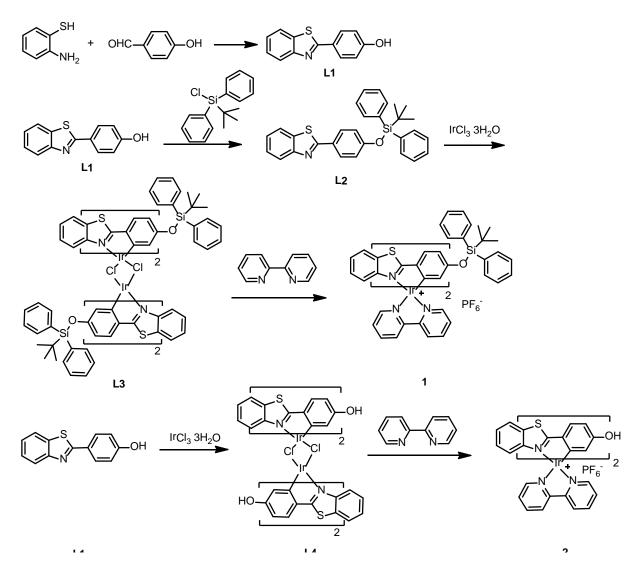


Figure S1. Synthesis of the ligands L1 and L2, and their IrIII complexes 1 and 2.

Synthesis of L1: The N,N-dimethylformamide (5.0 mL) was added to a flask containing 2-aminothiophenol (4.0 g) and 4-hydroxybenzaldehyde (4.12 mL). The mixture was refluxed for 72 h and the temperature of the reaction system was raised to 110  $^{\circ}$ C. Then, the reaction was extracted with water and ethyl acetate. The solvents were evaporated in vacuum and the product was recrystallized from ethanol. Yield: 80%. <sup>1</sup>H NMR (400 MHz, d6-DMSO)  $\delta$ :

10.21 (s, 1H), 8.07 (d, J = 7.9 Hz, 1H), 8.00 - 7.87 (m, 3H), 7.53 - 7.45 (m, 1H), 7.43 - 7.33 (m, 1H), 6.97 - 6.88 (m, 2H).

**Synthesis of L2:** The N,N-dimethylformamide (5.0 mL) was added to a flask containing tert-butyl(chloro)- diphenylsilane (2.0 mL), imidazole (200 mg) and **L1** (1 g). The mixture was refluxed for 24 h under nitrogen. Then the reaction was extracted with water and ethyl acetate. The solvents were evaporated in vacuum and the product was recrystallized from ethanol. Yield: 80%. <sup>1</sup>H NMR (400 MHz, d6-DMSO)  $\delta$ : 6.91 - 6.86 (dd, 2H), 7.51 - 7.39 (m, 8H), 1.05 (s, 9H), 7.71 - 7.66 (m, 4H), 7.96 - 7.79 (d, J = 7.8 Hz, 1H), 8.09 - 8.04 (d, 1H), 7.91 - 7.86 (m, 2H).

Synthesis of L3: A mixture of 2-ethoxyethanol and water (3:1, v/v) was added to a flask containing  $IrCl_3 \cdot 3H_2O$  (1 mmol) and ligand L2 (2.5 mmol). The mixture was refluxed for 24 h. After cooling down, the red solid precipitate was filtered and washed with ethanol to give crude cyclometalated Ir(III) chloro-bridged dimer.

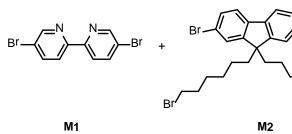
**Synthesis of 1:** The solution of cyclometalated Ir(III) chloro-bridged dimer L3 (200 mg) and 2,2'-bipyridine (64 mmol) in CH2Cl2/MeOH [30 mL, 2:1 (v/v)] was heated to reflux. After 4 h, the red solution was cooled down to room temperature and then a 5-fold excess of potassium hexafluorophosphate was added. The suspension was stirred for 1 h and then was filtered to remove insoluble inorganic salts. The solution was evaporated to dryness under reduced pressure. It was chromatographed by using CH2Cl2/acetone (30:1) to afford red solid in a 60% yield. <sup>1</sup>H NMR (400 MHz, d6-DMSO)  $\delta$ : 8.70 (d, J = 8.3, 2H), 8.22 (dd, 2H), 7.89 (d, J = 8.0, 2H), 7.77 (d, J = 8.3, 2H), 7.66 (dd, J = 7.1, 2H), 7.56 (d, J = 5.1, 2H), 7.59 - 7.32 (dd, 4H), 7.22 - 7.15(m, J = 24.18H), 6.83 (dd, 2H), 6.73 (dd, J = 8.4, 2H), 5.60 (d, J = 2.1, 2H), 5.45 (d, J = 8.4, 2H), 1.0 - 0.71 (s, 18H). <sup>13</sup>C NMR (100 MHz, d6-DMSO,  $\delta$ ):180.08,

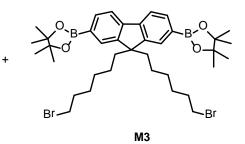
158.22, 156.27, 153.02, 150.25, 141.01, 138.34, 135.18, 134.08, 131.86, 130.77, 129.12, 127.50, 124.70, 115.90, 60.418, 58.49, 31.46, 29.53, 26.21, 22.07, 19.34, 14.32. MS (MALDI-TOF -MS) [m/z]: 1277.634 (4-PF6)<sup>+</sup>.

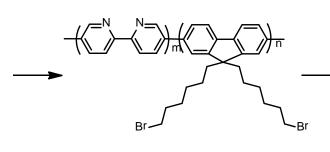
Synthesis of L4: A mixture of 2-ethoxyethanol and water (3:1, v/v) was added to a flask containing IrCl<sub>3</sub>·3H<sub>2</sub>O (1 mmol) and L1 (2.5 mmol). The mixture was refluxed for 24 h. After cooling down, the red solid precipitate was filtered and washed with ethanol to give crude cyclometalated Ir(III) chloro-bridged dimer.

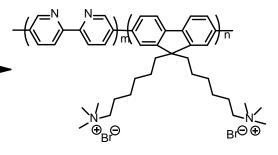
**Synthesis of 2:** The solution of cyclometalated Ir(III) chloro-bridged dimer **L4** (200 mg) and 2,2'-bipyridine (50 mg) in CH2Cl2/CH3OH [30 mL, 2:1 (v/v)] was heated to reflux. After 4 h, the red solution was cooled down to room temperature and then a 10-fold excess of potassium hexafluorophosphate was added. The suspension was stirred for 2 h and then was filtered to remove insoluble inorganic salts. The solution was evaporated to dryness under reduced pressure. It was chromatographed by using CH2Cl2/acetone (20:1) to afford red solid in a 60% yield. <sup>1</sup>H NMR (400 MHz, d6-DMSO)  $\delta$ : 9.85 (s, 2H), 8.80 (d, J = 8.2 Hz, 1H), 8.33 - 8.24 (m, 1H), 8.11 m, J = 6.1 Hz, 2H), 7.84 (d, J = 8.4 Hz, 1H), 7.82 - 7.74 (m, 1H), 7.30 (m, J = 7.7 Hz, 1H), 7.07 (m, J = 7.8 Hz, 1H), 6.53 (dd, J = 8.4, 2.2 Hz, 1H), 5.94 (d, J = 8.4 Hz, 1H), 5.74 (s, 2H).

The CPEs and its procedure were synthesized according to the route depicted in Fig. S2.



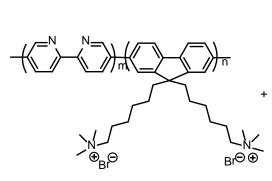


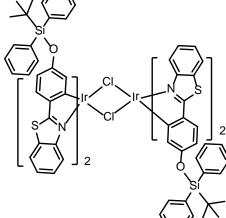




Pr1

Pr2





Br

Br

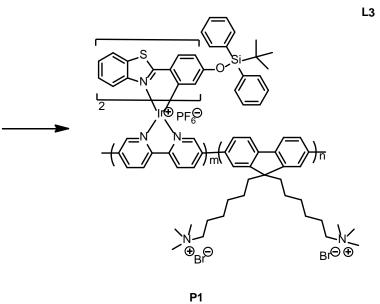


Figure S2. The synthetic routes of the CPEs.

Synthesis of **M1**, **M2** and **M3**: The monomers **M1**, **M2** and **M3** were synthesized according to the previous reports.<sup>1-3</sup>

The detail of synthetic of **P1** was stated in the methods part of the article.

## 2. Absorption spectra

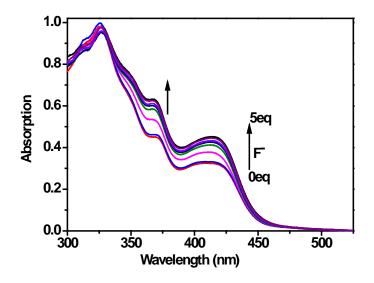


Figure S3. The response of absorption spectra of complex 1 to different concentrations of F- (with the addition of F- from 0 to 5-fold equivalents). The concentration of complex 1 was  $1.26 \times 10-4$  mol/L. The test was carried out in tetrahydrofuran.

#### 3. Mass spectra

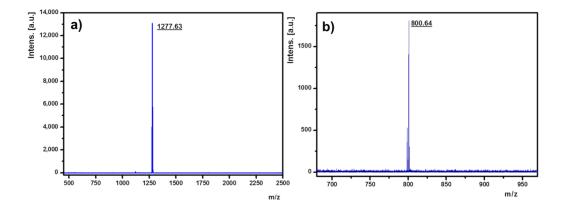
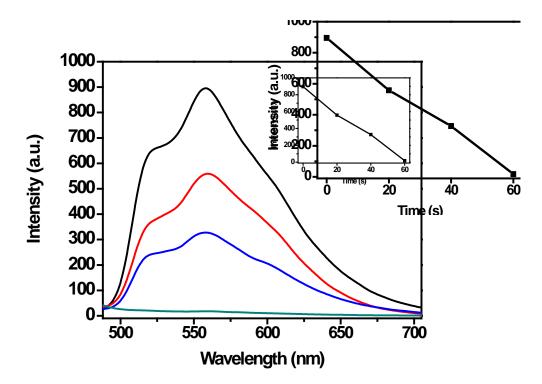


Figure S4. The mass spectra of compound 1 before (a) and after (b) addition of F.

#### 4. Time-dependent PL spectra

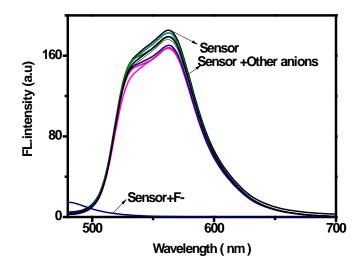
The time-dependent PL spectrum of compound 1 with saturation concentrations of  $F^-$  was shown in Fig. S5. It can be seen that there is almost no time-dependent effects in the PL sensing process, and the emission intensity reaches to its saturation value within 60 seconds due to the rapid deprotection reaction.



**Figure S5.** The time-dependent PL spectrum of compound 1 with saturation concentrations of F.

#### 5. Ion selective response

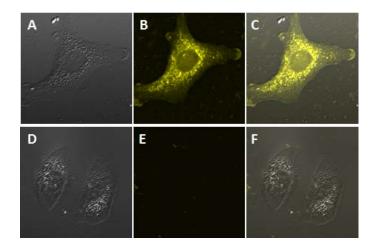
Upon screening of a large variety of common anions, such as  $CO_3^{2-}$ ,  $HSO_4^{-}$ ,  $NO_2^{-}$ ,  $NO_3^{-}$ ,  $CH_3COO^{-}$ ,  $HSO_3^{--}$ ,  $HCO_3^{--}$ ,  $CI^{--}$ ,  $Br^{--}$ , and  $F^{--}$ , probe **1** only exhibited significantly decreased emission (at 565 nm) in the presence of  $F^{--}$ . Other anions did not trigger any significant enhancement or disturb the response toward  $F^{--}$ .



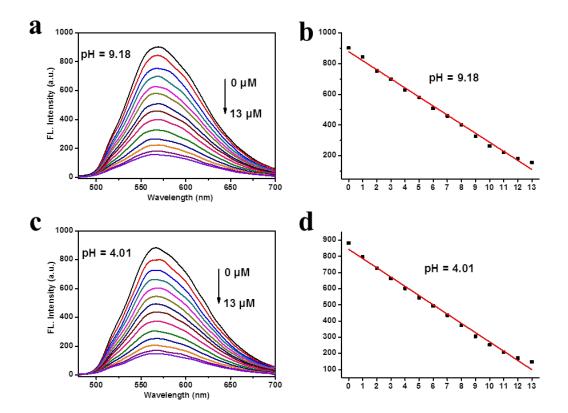
**Figure S6.** PL spectra ( $\lambda_{ex} = 370 \text{ nm}$ ) of complex **1** (20 µM) in the presence of 40 equivalents of various anions (CO<sub>3</sub><sup>2-</sup>, HSO<sub>4</sub><sup>-</sup>, NO<sub>2</sub><sup>-</sup>, NO<sub>3</sub><sup>-</sup>, CH<sub>3</sub>COO<sup>-</sup>, HSO<sub>3</sub><sup>-</sup>, HCO<sub>3</sub><sup>-</sup>, Cl<sup>-</sup>, Br<sup>-</sup>, and F<sup>-</sup>) in THF.

#### 6. Confocal images

To demonstrate the potential of the probe **1** in cellular sensing and imaging, the confocal images of living KB cells after incubation with probe **1** in PBS buffer for 60 min at 37 °C. Then the experimental group was further incubated with  $F^{-}(20 \ \mu\text{M})$  for 30 min at 37 °C, and the control group was directly incubated for 90 min. We used confocal luminescence imaging to demonstrate the feasibility of the probe for F<sup>-</sup> detection. The emission collected at 500-600 nm (B and E). As shown in the Fig. S5, after incubated with F, the lumienscence was significantly decreased, indicating that the probe exhibits good performances for F<sup>-</sup> sensing and imaging.



**Figure S7.** Confocal images of KB cell in the presence of the probe before (A-C) and after (D-F) incubated with F<sup>-</sup>. Excitation wavelength was 405 nm; The A and D were bright-field, B and E was collected at 500-600 nm, and C and F were the overlay images. All the images share the same bar of 20  $\mu$ m.



#### 7. Sensing performance of 1 and P1 under different pH values

**Figure S8.** (a) PL spectra of complex **1** in sodium tetraborate solution (ST, pH=9.18) with addition of F<sup>-</sup> from 0 to 13  $\mu$ M upon excitation at 365 nm. (b)  $I_{600}/I_{420}$  as a function of [F<sup>-</sup>] and its trendline in ST buffer solution. (c) PL spectra of complex **1** in potassium hydrogen phthalate solution (PHP, pH=4.00) with addition of F<sup>-</sup> from 0 to 13  $\mu$ M upon excitation at 365 nm. (d)  $I_{600}/I_{420}$  as a function of [F<sup>-</sup>] and its trendline in PHP buffer solution.

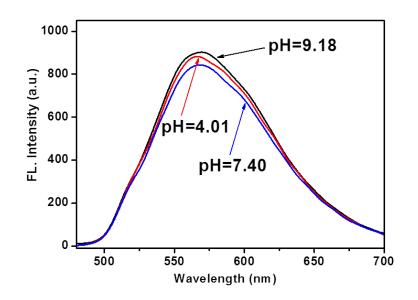
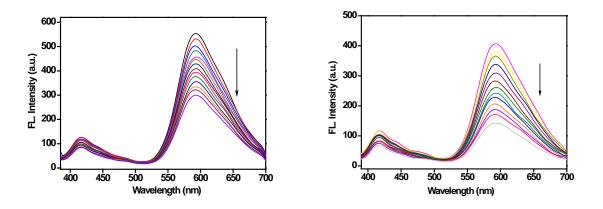
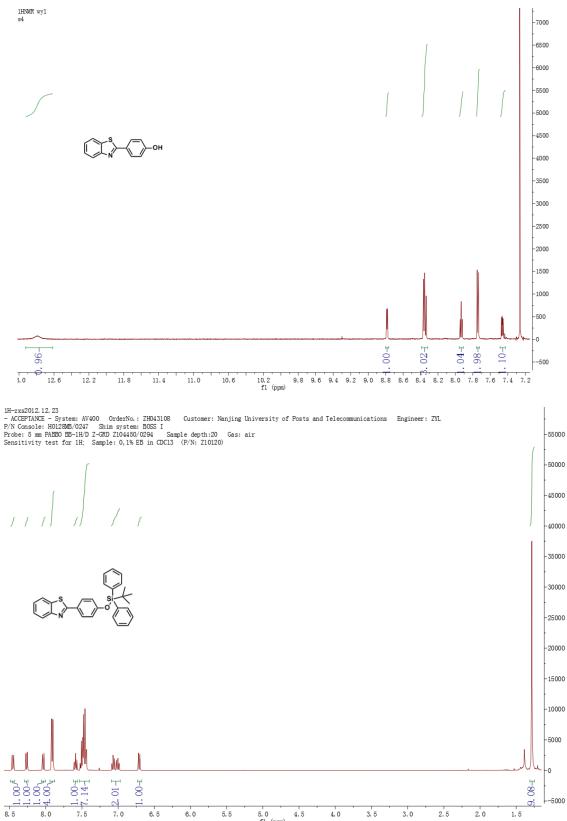


Figure S9. PL spectra of complex 1 in different pH buffer soulution without the addition of F



**Figure S10.** PL spectra of **P1** in aqueous solution under pH = 4.01 (left) or 9.18 (right) with addition of F<sup>-</sup> from 0 to 13 µM upon excitation at 365 nm.

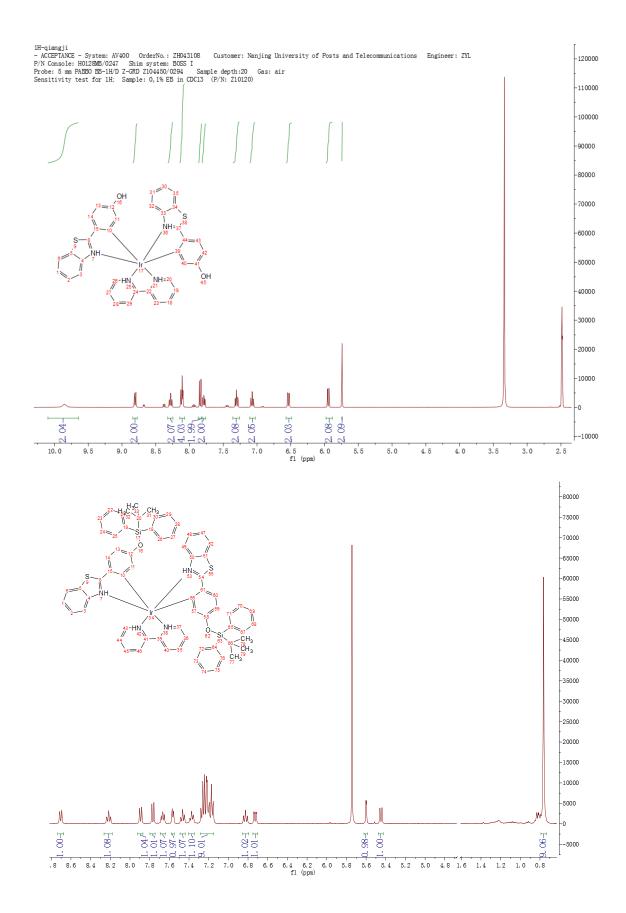
# 8. $^{1}$ H NMR



5.0 f1 (ppm) 4.5

4.0

3.5 3.0



# 9. References

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