Supplementary Materials

Glass-Forming Tendency of Molecular Liquids and the Strength of the Intermolecular Attractions

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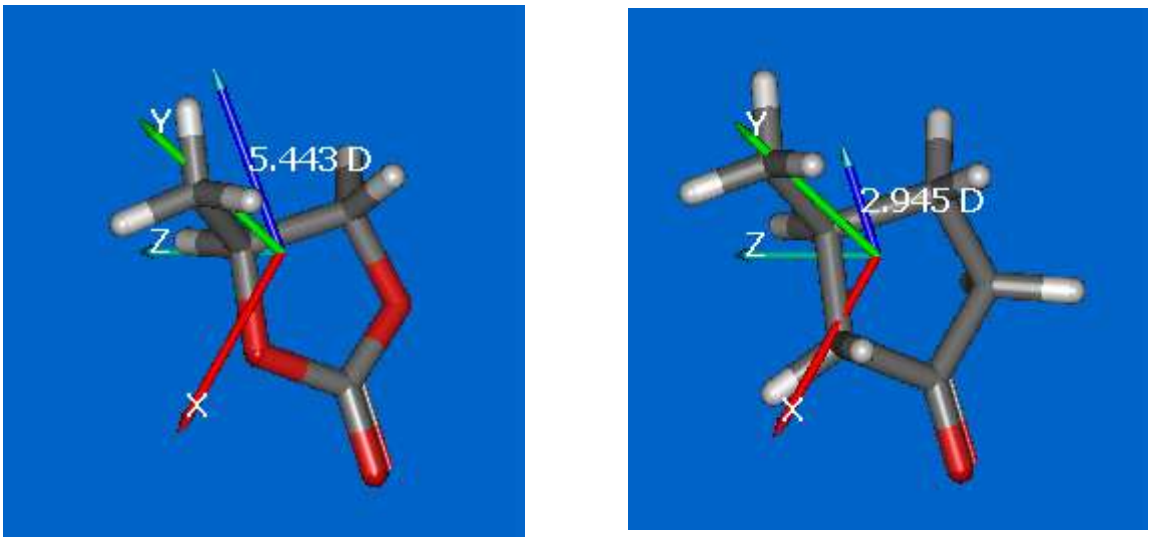
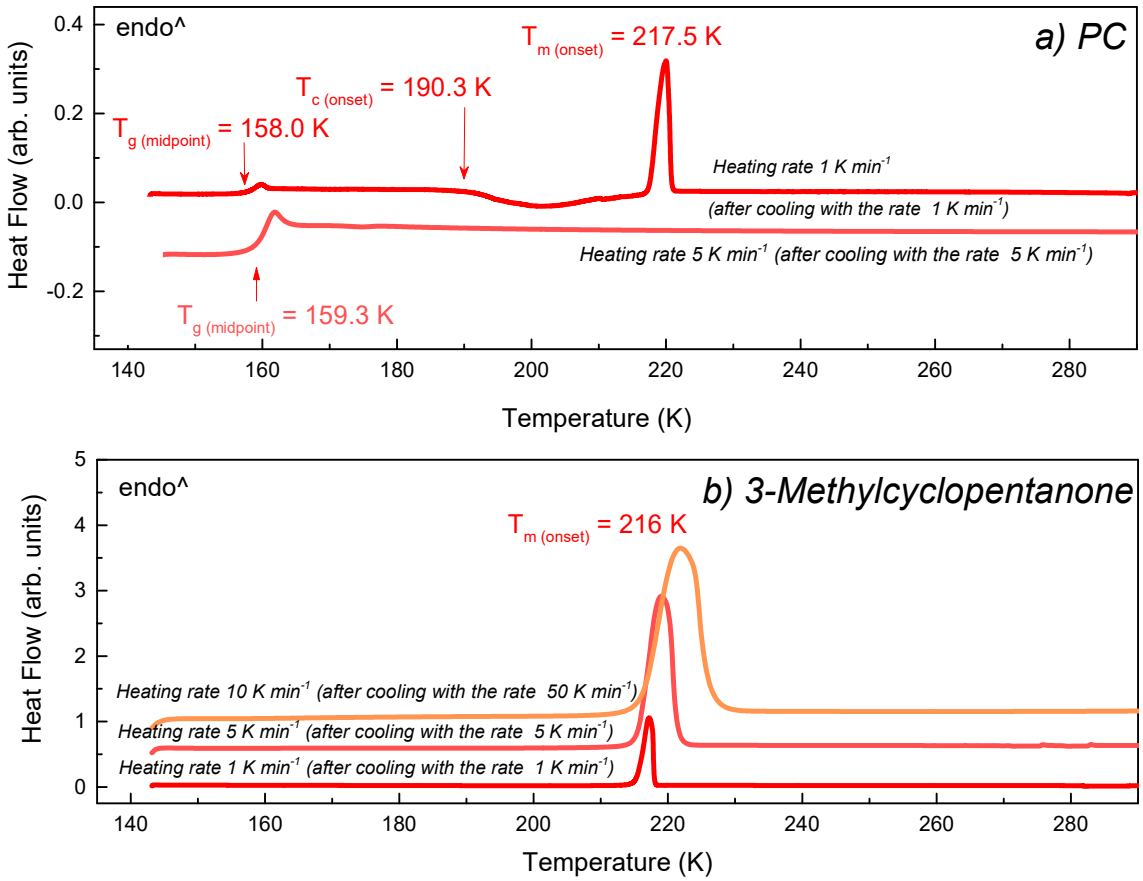
Supplementary Figures

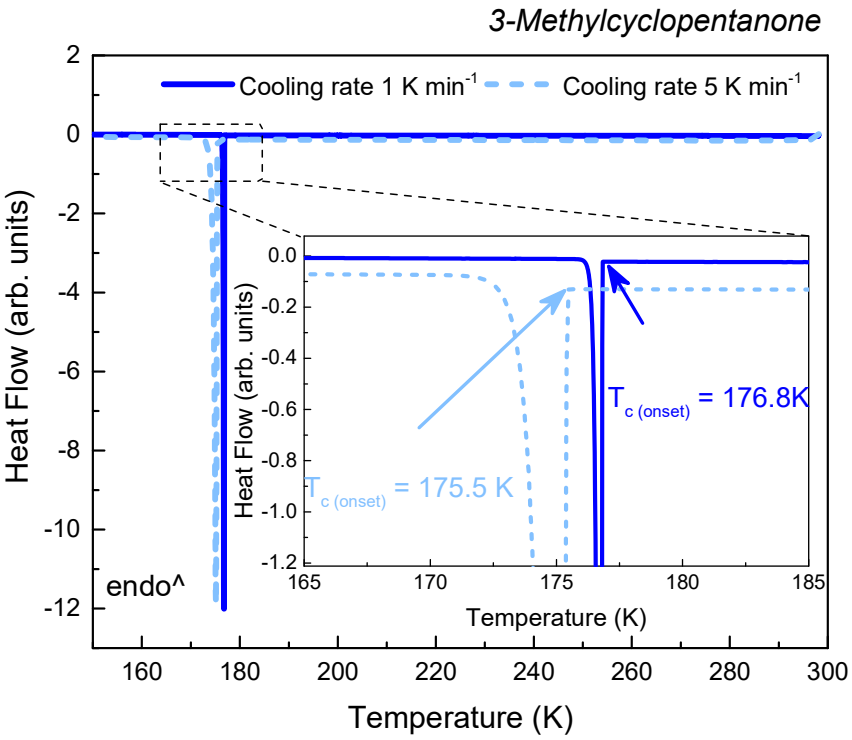
(a) (b)



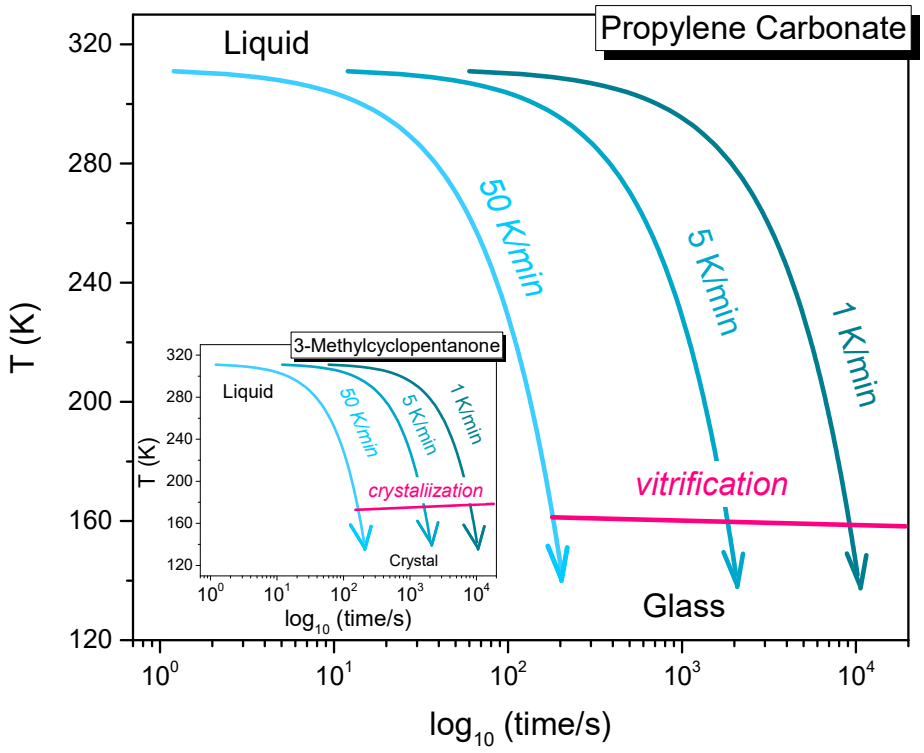
Supplementary Figure 1. Orientation of the dipole moment. Optimized geometry of (a) PC and (b) 3-methyl-cyclopentanone. Blue arrows indicate orientation of dipole moment with respect to principal inertial axes.



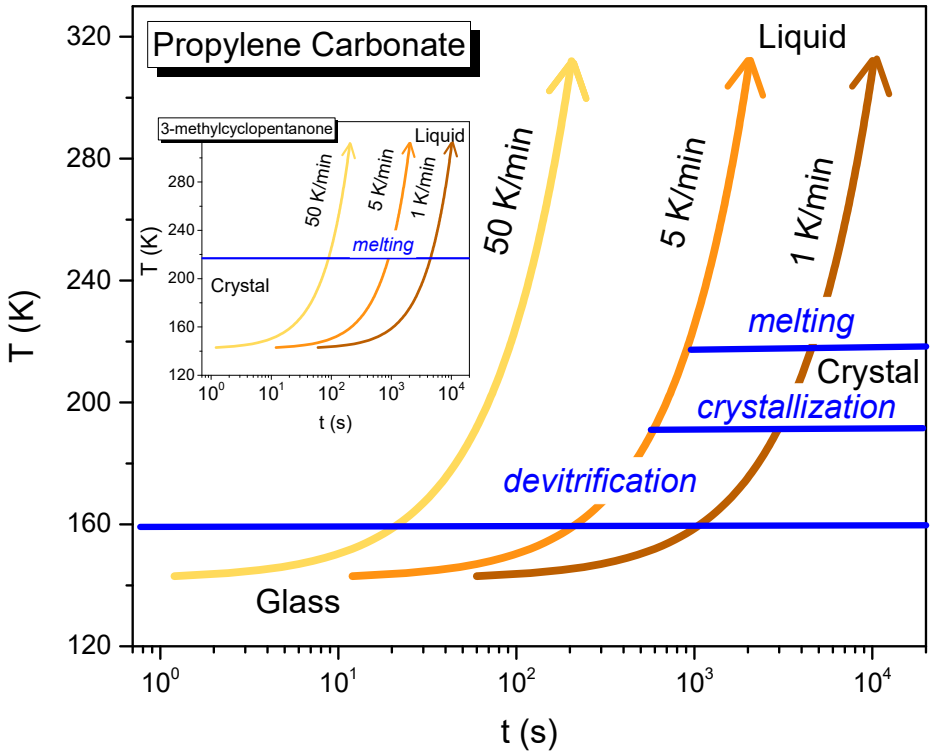
Supplementary Figure 2. Results of the calorimetric measurements. DSC thermograms for (a) PC and (b) 3-Methylocyclopentanone recorded on heating with different rates. No glass transition event was detected in DSC of 3-methylcyclopentanone.



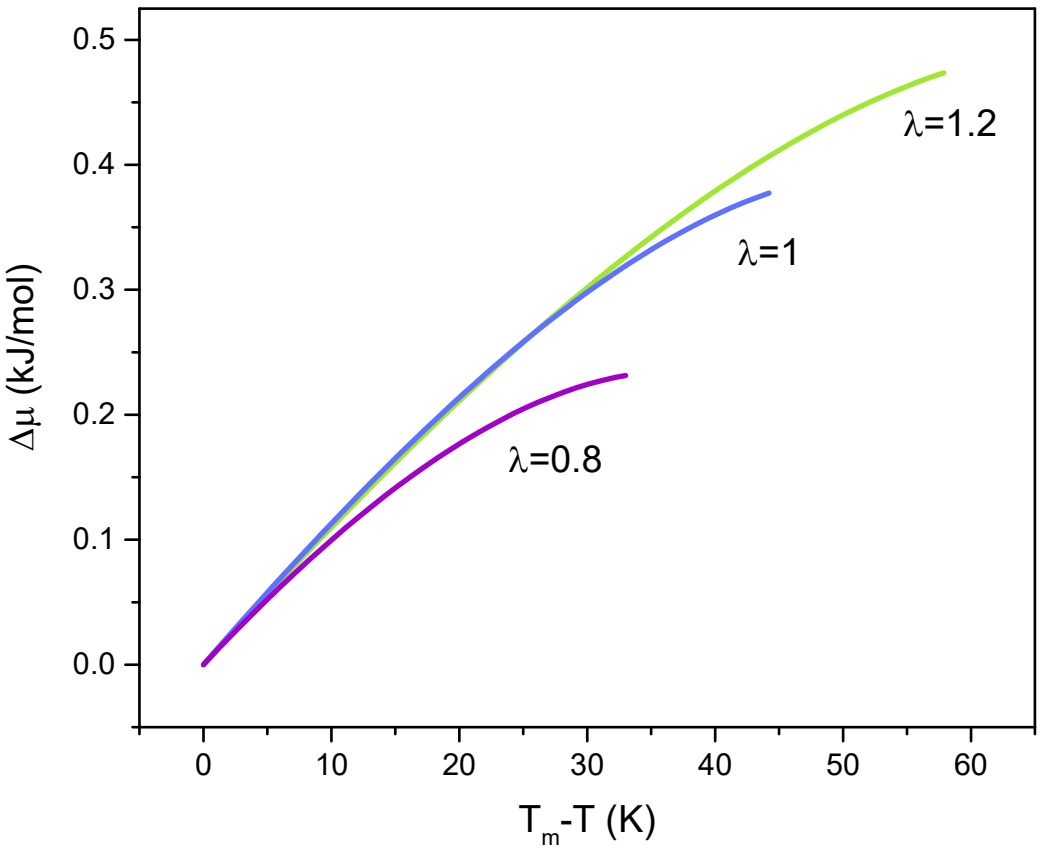
Supplementary Figure 3. Results of the calorimetric measurements. DSC thermograms recorded for 3-Methylcyclopentanone on cooling with different rates.



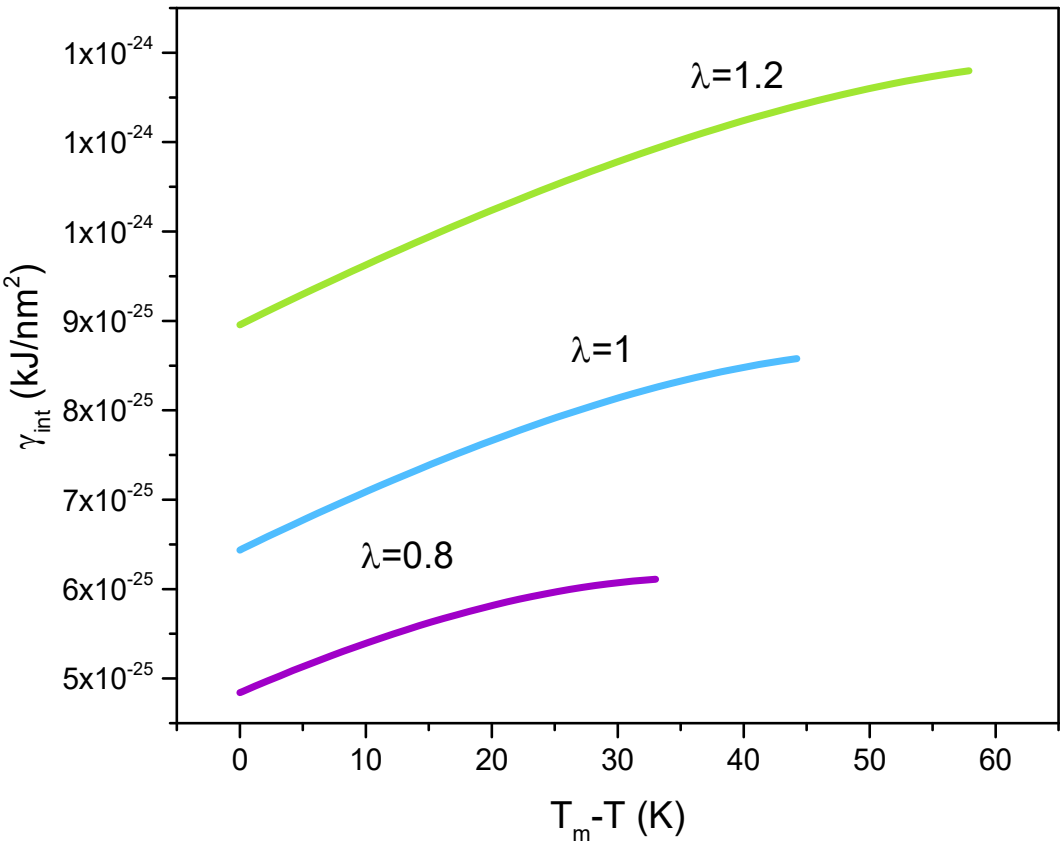
Supplementary Figure 4. CCT curve for PC and 3-methylcyclopentanone (inset).



Supplementary Figure 5. CHT curve for PC and 3-methylcyclopentanone (inset).



Supplementary Figure 6. Temperature evolution of the thermodynamic driving force towards crystallization. Possible temperature behavior of $\Delta\mu$ for LJ system with varying strength of the intermolecular attractive forces, predicted using Eq. 3.



Supplementary Figure 7. Temperature evolution of the specific surface energy. Possible temperature behavior of γ_{int} for LJ system with varying strength of the intermolecular attractive forces, predicted using Eq. 4.