

Supplementary Information

(6 pages, 1 figure, 2 tables)

Energy use and carbon footprints differ dramatically for diverse wastewater-derived carbonaceous substrates: An integrated exploration of biokinetics and life-cycle assessment

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Notation

X_S	Slowly biodegradable substrates
X_{Sads}	Adsorbed substrates
S_O	Dissolved oxygen
S_S	Readily biodegradable organic substrates
X_I	Inert particulate organic material
X_H	Active heterotrophic organisms
X_{STO}	Internal storage product of heterotrophic organisms
$Y_{H,S}$	Yield coefficient for direct growth
$Y_{H,STO}$	Yield coefficient for growth on stored products
Y_{STO}	Yield of stored products on S_S
k_{ads}	Substrate adsorption rate
q_m	Saturation constant for absorption
k_H	Hydrolysis rate constant
f_{STO}	Ratio of substrate utilized for storage
q_{max}	Maximum substrate uptake rate
$u_{H,STO}$	Maximum growth rate on stored products
K_O	Half saturation constant for S_O
K_S	Half saturation constant for substrate S_S
K_{STO}	Half saturation constant for storage substance
K_X	Half hydrolysis saturation constant
b_H	Endogenous respiration rate of heterotrophs
b_{STO}	Endogenous respiration rate of storage products
f_i	Fraction of inert particulate substance

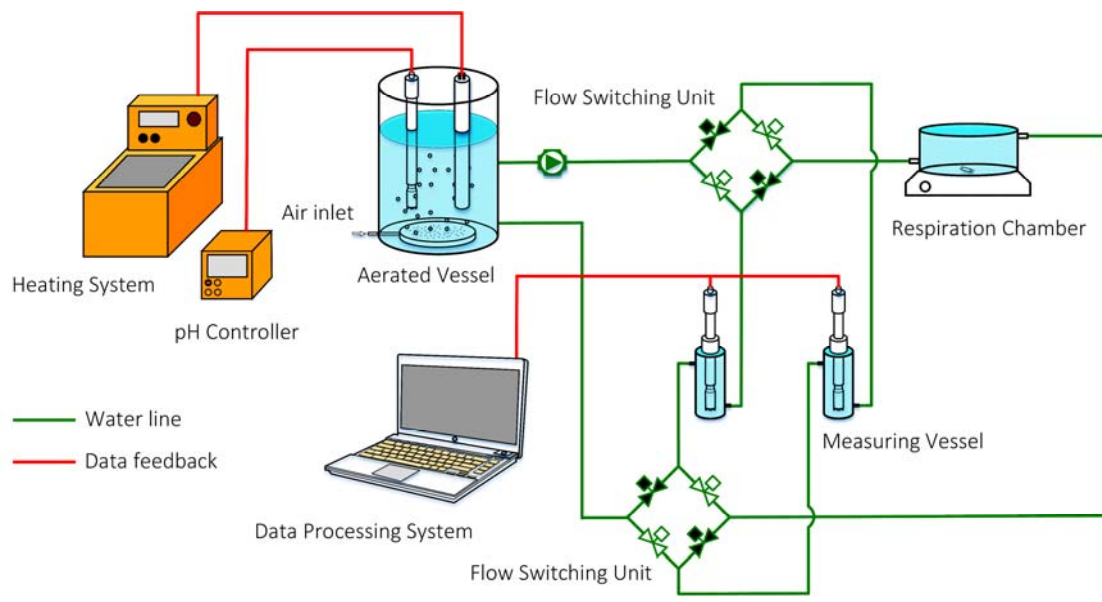


Figure S1 A schematic illustration of the proposed activated sludge respirometer.

Table S1 Kinetic and stoichiometric parameters for different substrates.

Parameters	Substrates				Source
	HAc	HPr	SolS	BSA	
$Y_{H,S}$ (mg COD/mg COD)	0.57	0.56	0.63	0.55	this study
$Y_{H,STO}$ (mg COD/mg COD)	0.65	0.63	0.65	0.60	this study
Y_{STO} (mg COD/mg COD)	0.81	0.76	0.92	0.74	this study
k_{ads} (h ⁻¹)	-	-	15.0	15.0	ref 1
q_m (mg COD/mg COD)	-	-	0.071	0.098	this study
k_H (d ⁻¹)	-	-	0.91	0.85	this study
f_{STO} (d ⁻¹)	0.67	0.45	0.72	0.45	this study
q_{max} (d ⁻¹)	4.4	1.8	-	-	this study
$\mu_{H,STO}$ (d ⁻¹)	1.1	1.3	1.2	1.1	this study
K_O (mg/L)	0.2	0.2	0.2	0.2	ref 2
K_S (mg COD/L)	0.7	0.7	-	-	ref 2
K_{STO} (mg COD/mg COD)	0.5	0.4	0.4	0.5	this study
K_X (mg COD/mg COD)	-	-	0.05	0.01	this study
b_H (d ⁻¹)	0.2	0.2	0.2	0.2	ref 3
b_{STO} (d ⁻¹)	0.2	0.2	0.2	0.2	ref 3
f_i (-)	0.2	0.2	0.2	0.2	ref 3

Table S2 Matrix of the proposed kinetic model.

Process	X_S	X_{Sads}	S_o	S_s	X_I	X_H	X_{STO}	Kinetics
Adsorption	-1	1						$k_{ads} \cdot \frac{q_m - (X_{sads}/X_H)}{q_m} \cdot X_S$
Hydrolysis		-1		1				$k_H \cdot \frac{(X_{sads}/X_H)}{K_X + (X_{sads}/X_H)} \cdot X_H$
Growth on S_s			$1-1/Y_{H,S}$	$-1/Y_{H,S}$		1		$f_{STO} \cdot q_{max} \cdot Y_{STO} M_O \cdot M_S \cdot X_H$
Storage of S_s			$1-1/Y_{STO}$	$-1/Y_{STO}$			1	$(1 - f_{STO}) \cdot q_{max} \cdot Y_{H,S} \cdot M_O \cdot M_S \cdot X_H$
Growth on X_{STO}			$1-1/Y_{H,STO}$			1	$-1/Y_{H,STO}$	$\mu_{H,STO} \cdot M_O \cdot \frac{(X_{STO}/X_H)}{K_{STO} + (X_{STO}/X_H)} \cdot \frac{K_S}{S_S + K_S} \cdot X_H$
Endogenous respiration			f_i-1		f_i	-1		$b_H \cdot M_O \cdot X_H$
Endogenous respiration of X_{STO}			-1				-1	$b_{STO} \cdot M_O \cdot X_{STO}$

M stands for a Monod kinetic function e.g. $M_S = S_S / (K_S + S_S)$.

References

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