

Supporting Information

Aqueous solutions of surface-active ionic liquids: remarkable alternative solvents to improve the solubility of triterpenic acids and their extraction from biomass

Emanuelle L. P. de Faria[†], Selesa V. Shabudin[†], Ana Filipa M. Cláudio[†], Mónica Válega[‡], Fernando M. J. Domingues[§], Carmen S. R. Freire[†], Armando J. D. Silvestre[†], Mara G. Freire^{†}*

[†]CICECO – Aveiro Institute of Materials, Chemistry Department, University of Aveiro, University Campus of Santiago, 3810-193 Aveiro, Portugal

[‡]QOPNA – Chemistry Department, University of Aveiro, University Campus of Santiago, 3810-193 Aveiro, Portugal

[§]Chemistry Department, University of Aveiro, University Campus of Santiago, 3810-193 Aveiro, Portugal

*E-mail address: maragfreire@ua.pt

Supporting Information Contents:

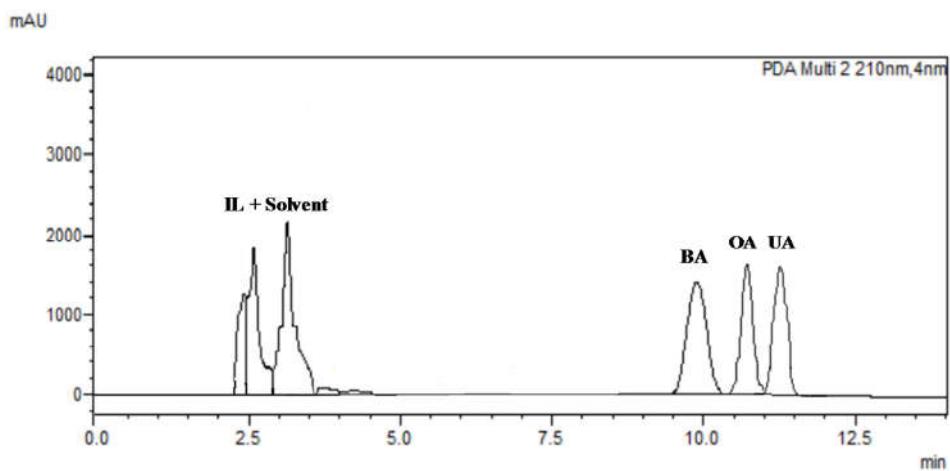
Number of pages: 7

Number of figures: 3

Number of tables: 5

Figures/Tables

(a)



(b)

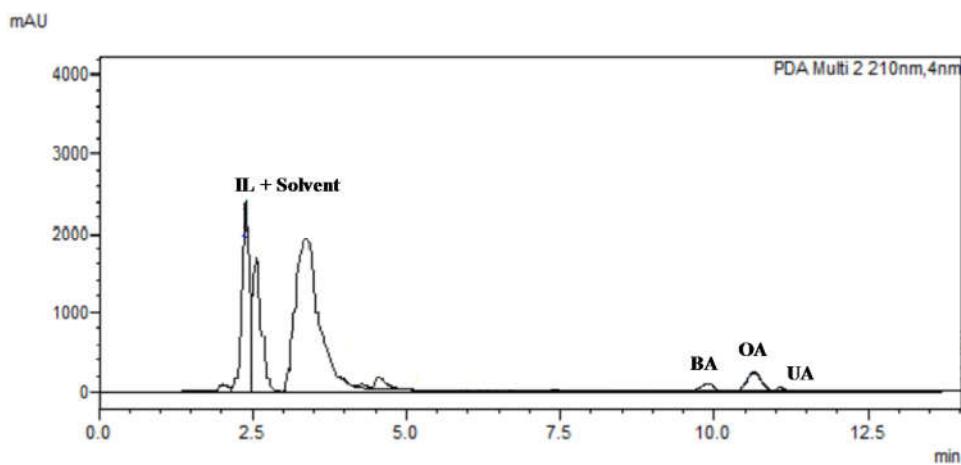


Fig. S1. Illustrative HPLC chromatograms corresponding to (a) standards TTAs, namely, UA, OA and BA, dissolved in an aqueous solutions of $[C_{14}C_1im]Cl$; (b) $[C_{14}C_1im]Cl$ (500 mM) aqueous solution after the extraction from apple peels at 80°C (b).

Table S1. Solubility of ursolic acid in aqueous solutions of ILs at 30°C.

Ionic Liquid/Concentration	[Ursolic Acid] $\pm \sigma$ / (g.L ⁻¹)					
	50 mM	150 mM	250 mM	500 mM	750 mM	1000 mM
[C ₈ C ₁ im]Cl	-	0.004 \pm 0.001	0.028 \pm 0.001	0.091 \pm 0.003	0.280 \pm 0.016	0.435 \pm 0.007
[C ₁₀ C ₁ im]Cl	0.007 \pm 0.000	0.192 \pm 0.001	0.461 \pm 0.009	1.008 \pm 0.439	1.474 \pm 0.048	1.612 \pm 0.014
[C ₁₂ C ₁ im]Cl	0.059 \pm 0.001	0.412 \pm 0.009	0.699 \pm 0.045	1.173 \pm 0.068	1.544 \pm 0.028	1.714 \pm 0.052
[C ₁₄ C ₁ im]Cl	0.321 \pm 0.007	0.544 \pm 0.023	0.868 \pm 0.008	1.656 \pm 0.038	1.919 \pm 0.005	2.852 \pm 0.009
[C ₁₆ C ₁ im]Cl	0.227 \pm 0.005	0.892 \pm 0.010	1.404 \pm 0.036	2.481 \pm 0.057	0.815 \pm 0.110	ND
[C ₁₈ C ₁ im]Cl	0.419 \pm 0.019	1.180 \pm 0.018	ND	ND	ND	ND
[C ₄ C ₁ im][C ₈ H ₁₇ SO ₄]	0.179 \pm 0.007	1.931 \pm 0.001	3.130 \pm 0.001	1.849 \pm 0.005	1.549 \pm 0.005	1.343 \pm 0.050
[P _{44,14}]Cl	0.235 \pm 0.003	0.299 \pm 0.003	0.348 \pm 0.003	0.370 \pm 0.004	0.488 \pm 0.006	0.813 \pm 0.029

*ND – Not determined.

Table S2. Solubility of ursolic acid in aqueous solutions of conventional surfactants at 30°C.

Conventional Surfactant/Concentration	[Ursolic Acid] $\pm \sigma$ / (g.L ⁻¹)					
	50 mM	150 mM	250 mM	500 mM	750 mM	1000 mM
SDS	0.325 \pm 0.001	1.350 \pm 0.044	ND	ND	ND	ND
CTAB	0.078 \pm 0.001	0.259 \pm 0.002	0.980 \pm 0.036	1.607 \pm 0.009	ND	ND
SDBS	1.240 \pm 0.077	2.092 \pm 0.066	2.480 \pm 0.045	ND	ND	ND

*ND – Not determined.

Table S3. CMC values of the studied surface-active ILs at 30°C.

[Ionic Liquids]	CMC (mM) $\pm \sigma$ (measured)	Literature
[C ₈ C ₁ im]Cl	233.0 \pm 1.4	238.0 ¹
[C ₁₀ C ₁ im]Cl	58.7 \pm 0.8	57.2 ² ; 55.0 ³
[C ₁₂ C ₁ im]Cl	15.2 \pm 0.9	15.1 ³ ; 13.5 ⁴
[C ₁₄ C ₁ im]Cl	3.9 \pm 0.1	3.8 ⁵ ; 3.6 ⁴
[C ₁₆ C ₁ im]Cl	0.9 \pm 0.2	0.9 ^{3,4} / 1.1 ³
[C ₁₈ C ₁ im]Cl	0.3 \pm 0.03	0.4 ² ; 0.4 ³
[C ₄ C ₁ im][C ₈ H ₁₇ SO ₄]	30.3 \pm 0.7	31.0 ⁶

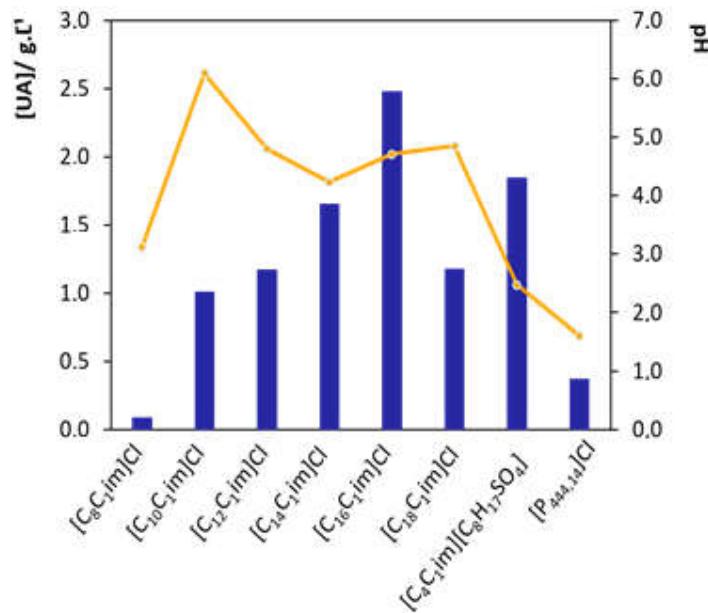


Fig. S2. pH () and UA solubility () in ILs aqueous solutions at 30°C.

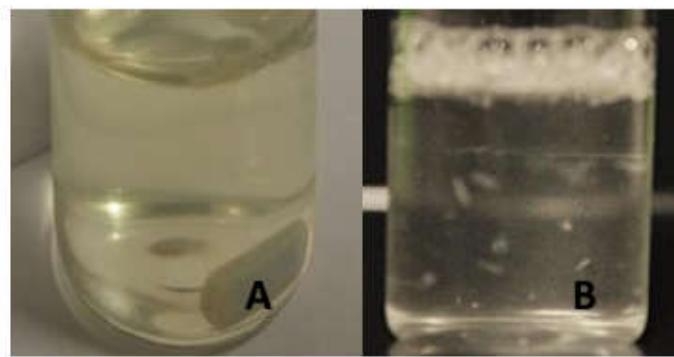


Fig. S3. Demonstration of the ability to use water as an anti-solvent in the precipitation of UA from an aqueous solution of $[C_4C_1im][C_8H_{17}SO_4]$. A: before water addition; B: after water addition.

Table S4. Extraction yields (EY) of TTAs (BA (betulinic), OA (oleanolic) and UA (ursolic) acids) from green apple peels with $[C_{14}C_1im]Cl$ at different temperatures and fixed conditions ($[IL] = 500\text{mM}$, $t = 60\text{ min}$, $R_{S:L} = 1:10$).

$[C_{14}C_1im]Cl$	Triterpenic Acids	T / °C	EY $\pm \sigma$ / wt.%
500 mM	BA	25	ND
		50	0.071 ± 0.001
		80	0.898 ± 0.001
		90	0.105 ± 0.002
500 mM	OA	25	ND
		50	0.437 ± 0.002
		80	1.638 ± 0.001
		90	1.561 ± 0.001
500 mM	UA	25	ND
		50	0.042 ± 0.001
		80	0.081 ± 0.002
		90	0.072 ± 0.002

*ND – Not determined.

Table S5. Extraction yields of TTAs (BA, OA and UA) extracted from green apple peels with volatile organic solvents at fixed conditions ($T = 80^{\circ}\text{C}$ under reflux, $t = 60$ min, $R_{\text{S:L}} = 1:10$).

Solvent	Triterpenic Acids	$EY \pm \sigma / \text{wt.\%}$
Acetone	BA	0.824 ± 0.003
	OA	0.450 ± 0.006
	UA	0.090 ± 0.002
Chloroform	BA	1.393 ± 0.005
	OA	1.049 ± 0.001
	UA	0.032 ± 0.002

References

1. Łuczak, J., Hupka, J., Thöming, J., Jungnickel, C., Self-organization of imidazolium ionic liquids in aqueous solution, *Colloid Surfaces A*, **2008**, 329, 125-133.
2. Łuczak, J., Jungnickel, C., Joskowska, M., Thöming, J., Hupka, J., Thermodynamics of micellization of imidazolium ionic liquids in aqueous solutions, *J. Colloid Interface Sci.*, **2009**, 336, 111-116.
3. Smirnova, N. A., Safonova, E. A., Micellization in solutions of ionic liquids, *Colloid J.*, **2012**, 74, 254-265.
4. El Seoud, O. A., Pires, P. A. R., Abdel-Moghy, T., Bastos, E. L., Synthesis and micellar properties of surface-active ionic liquids: 1-Alkyl-3-methylimidazolium chlorides, *J. Colloid Interface Sci.*, **2007**, 313, 296-304.
5. Bai, G., Lopes, A., Bastos, M., Thermodynamics of micellization of alkylimidazolium surfactants in aqueous solution, *J. Chem. Thermodyn.*, **2008**, 40, 1509-1516.
6. Miskolczy, Z., Sebök-Nagy, K., Biczók, L., Göktürk, S., Aggregation and micelle formation of ionic liquids in aqueous solution, *Chem Phys Lett.*, **2004**, 400, 296-300.