

Supporting Information

Implantable Tin Porphyrin-PEG Hydrogels with pH-responsive Fluorescence

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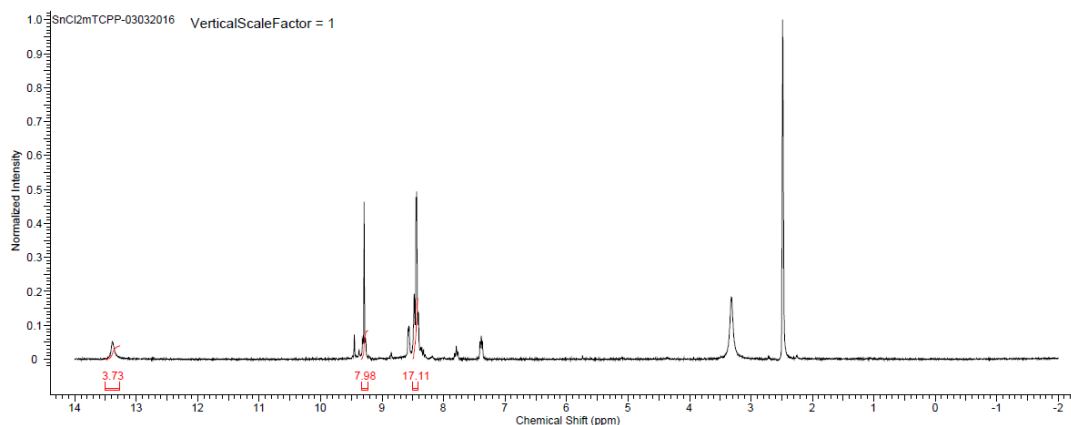


Fig. S1 NMR of SnCl₂mTCCP hydrogels. The peak at 3.33 ppm is water in (CD₃)₂SO [1].

[1] G.R. Fulmer, A.J.M. Miller, N.H. Sherden, H.E. Gottlieb, A. Nudelman, B.M. Stoltz, J.E. Bercaw, K.I. Goldberg, NMR Chemical Shifts of Trace Impurities: Common Laboratory Solvents, Organics, and Gases in Deuterated Solvents Relevant to the Organometallic Chemist, *Organometallics* 29(9) (2010) 2176-2179.

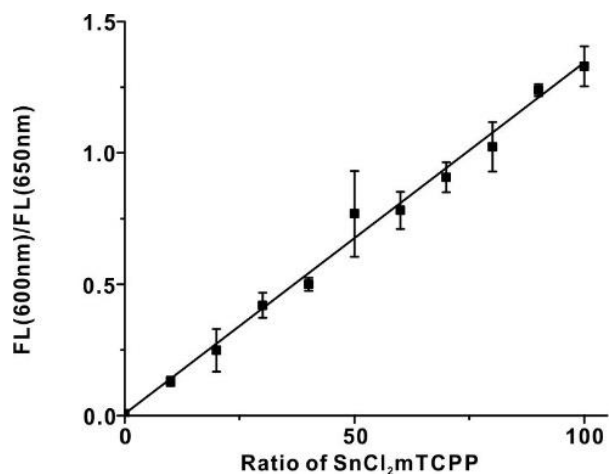


Fig. S2 Standard curve of SnCl₂mTCPP concentration ratio and fluorescent intensities ratio of 600 nm and 650 nm, $R^2=0.99838$. Data show mean +/- std. dev. for n=3 separate polymerization reactions and purifications.

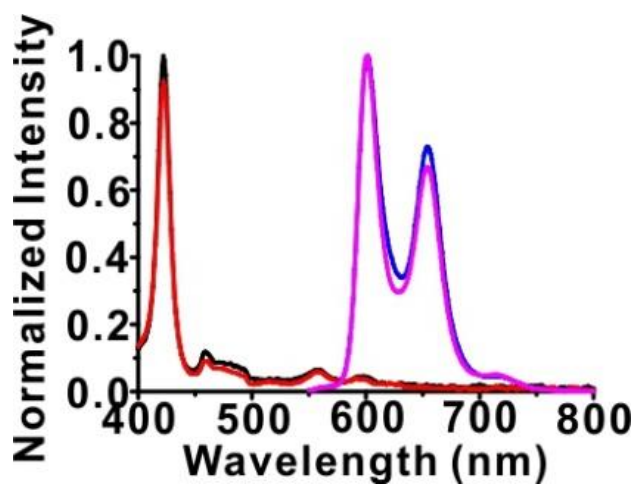


Fig S3. Absorbance spectra of synthesized SnCl₂mTCPP (black) and the post chelated one (red), and their emission spectra (blue for synthesized sample and purple for post chelated one).

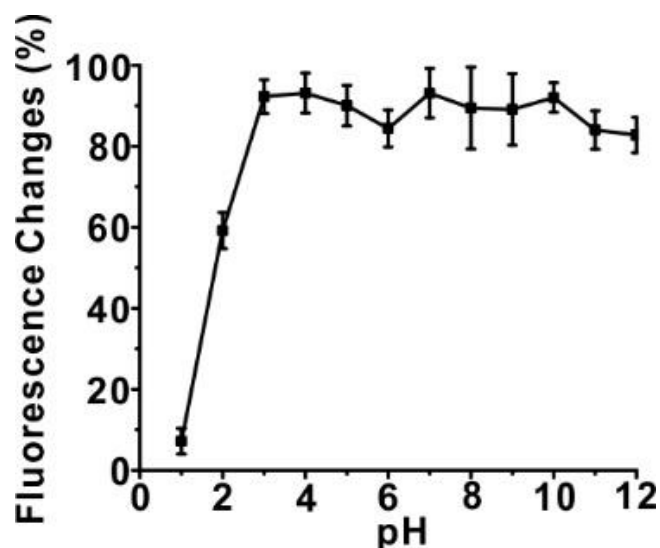


Fig S4. pH sensitivity of 2H-mTCPP hydrogels in pH=1-12 100mM sodium phosphate. Data show mean +/- std. dev. for n=3 separate polymerization reactions and purifications.

Table S1. Lifetime values and corresponding fitted parameters of tin hydrogel

pH	A ₁	τ ₁ (ns)	A ₂	τ ₂ (ns)	<τ> (ns)
1	0.02±0.005	2.2±0.3	0.98±0.03	0.8±0.02	0.9±0.05
2	0.16±0.02	1.7±0.1	0.84±0.04	0.9±0.03	1.1±0.09
3	0.48±0.008	1.3±0.01	0.52±0.02	0.5±0.01	1.1±0.02
4	0.62±0.05	1.4±0.06	0.38±0.06	0.7±0.1	1.2±0.1
5	0.46±0.03	1.5±0.05	0.54±0.04	0.7±0.05	1.2±0.09
6	0.58±0.04	1.4±0.06	0.42±0.06	0.6±0.08	1.2±0.1
7	0.73±0.03	1.6±0.02	0.27±0.03	0.7±0.07	1.5±0.06
8	0.97±0.02	1.5±0.02	0.03±0.007	3.1±0.6	1.6±0.08
9	0.97±0.02	1.6±0.02	0.03±0.009	2.8±0.6	1.6±0.09
10	0.75±0.02	1.7±0.02	0.25±0.02	1.0±0.06	1.6±0.05
11	0.37±0.009	1.8±0.03	0.63±0.01	1.3±0.02	1.6±0.04
12	1±0*	1.6±0.01*			1.6±0.1*

* mono-exponential decay