

# **Supplementary information**

## **Unraveling the electronegativity-dominated intermediate adsorption on high-entropy alloy electrocatalysts**

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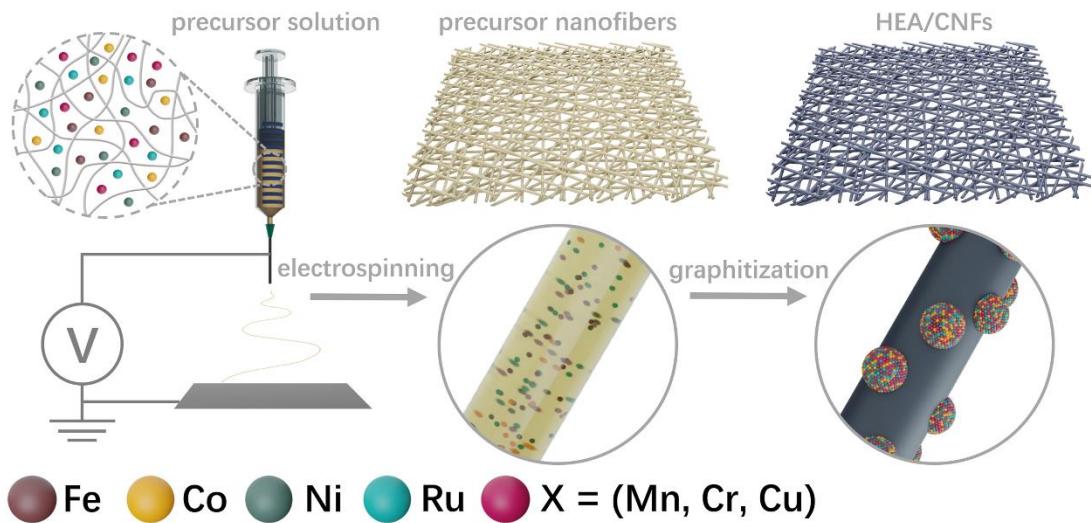
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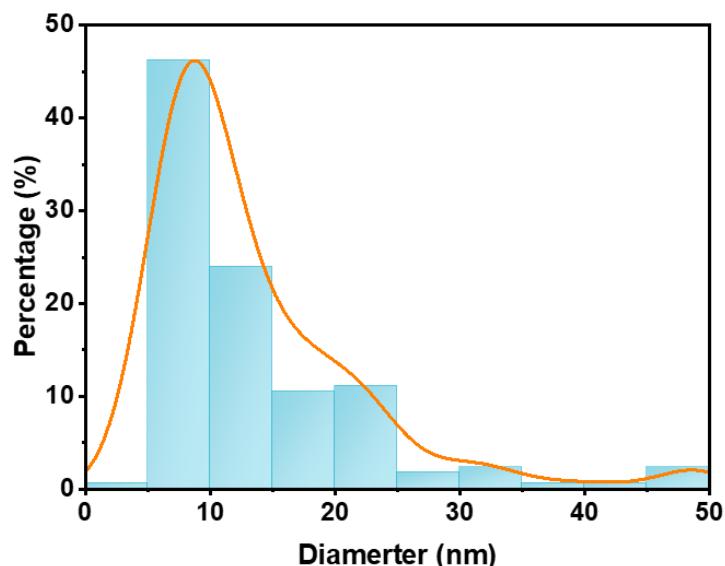
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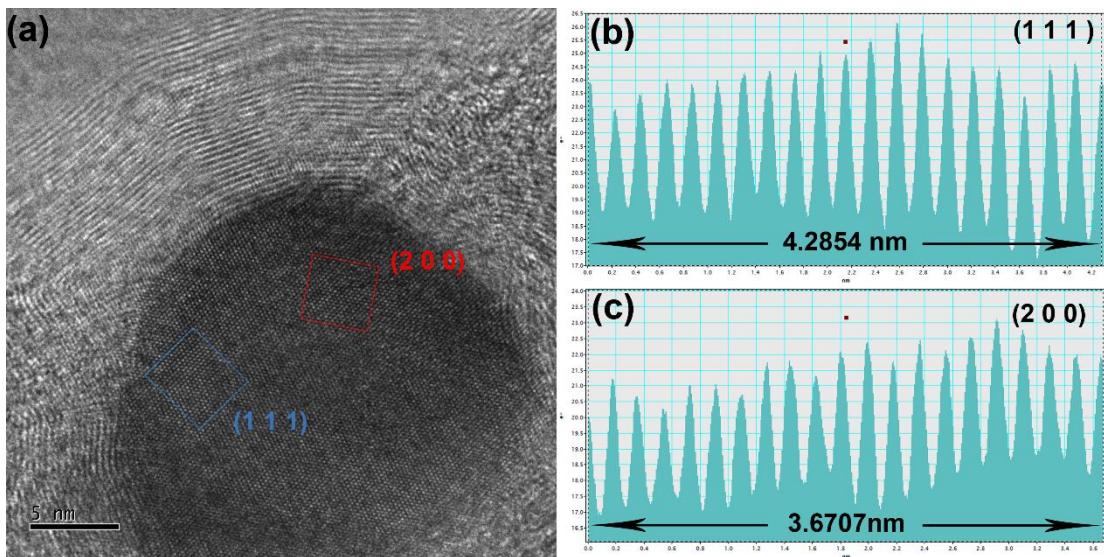
<sup>6</sup> These authors contributed equally: Jiace Hao and Zechao Zhuang. Correspondence and requests for materials should be addressed to H. Z. (email: zhysw@jiangnan.edu.cn)



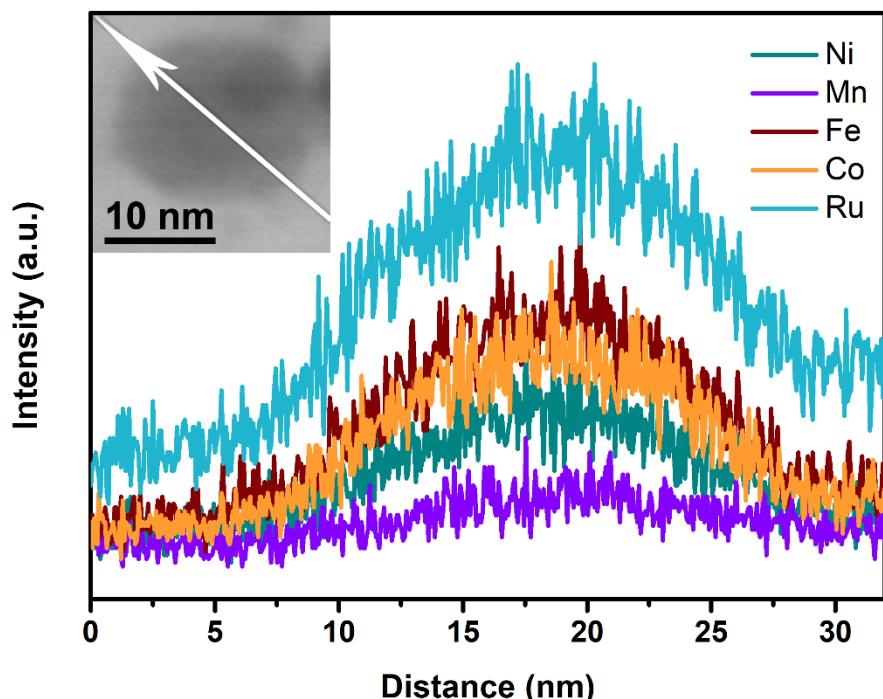
**Supplementary Fig. 1.** Schematic illustration for the synthetic processes of the FeCoNiX Ru/CNFs (X=Mn, Cr and Cu).



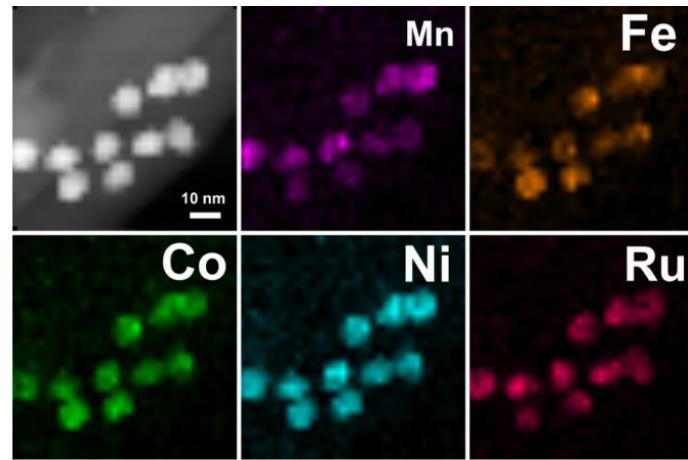
**Supplementary Fig. 2.** Average diameters of the FeCoNiMnRu HEA NPs supported on CNFs.



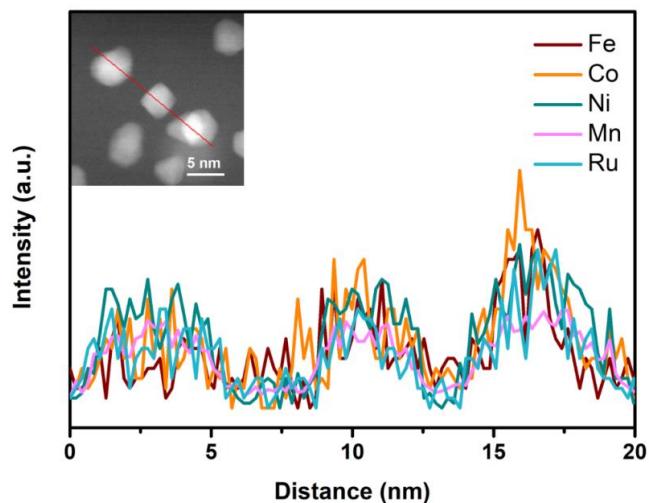
**Supplementary Fig. 3.** (a) HRTEM image of the FeCoNiMnRu HEA NP supported on CNFs. (b, c) Intensity profile recorded from the area indicated by the quadrat box in panel.



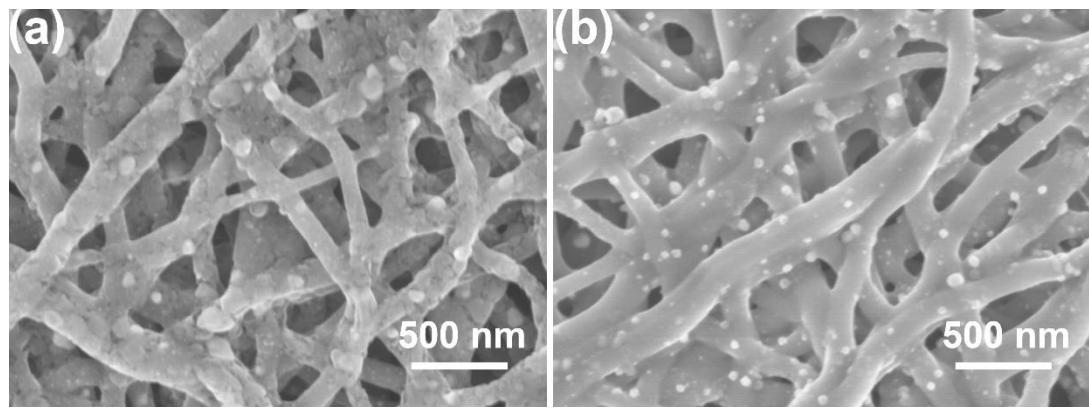
**Supplementary Fig. 4.** Line-scan STEM-EDS spectra of a FeCoNiMnRu NP supported on CNFs.



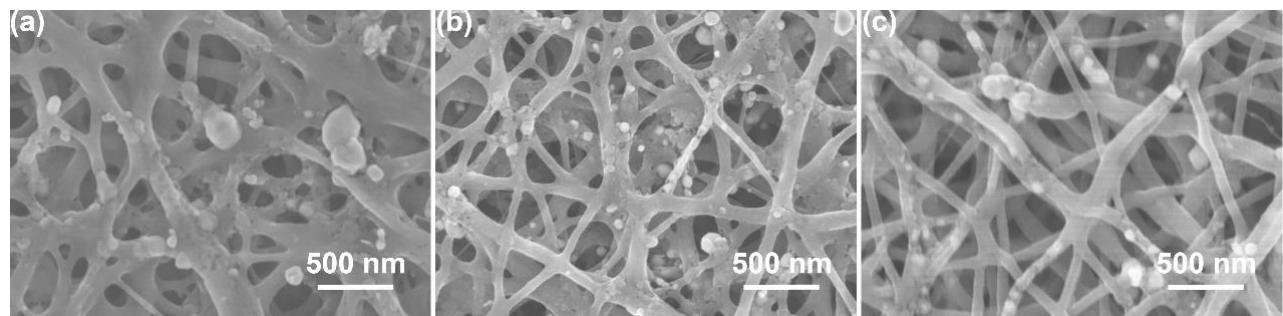
**Supplementary Fig. 5.** STEM-EDS mapping of the HEA NPs supported on CNFs.



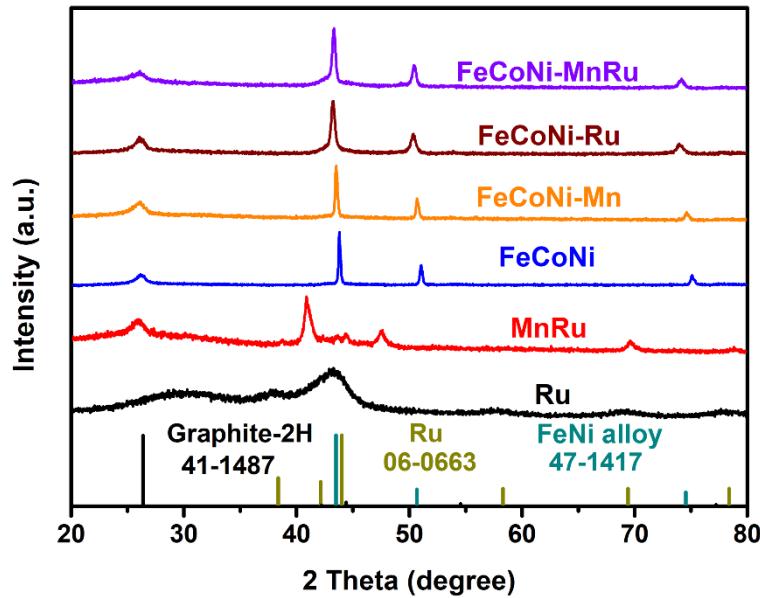
**Supplementary Fig. 6.** Line-scan STEM-EDX spectra of the HEA NPs supported on CNFs.



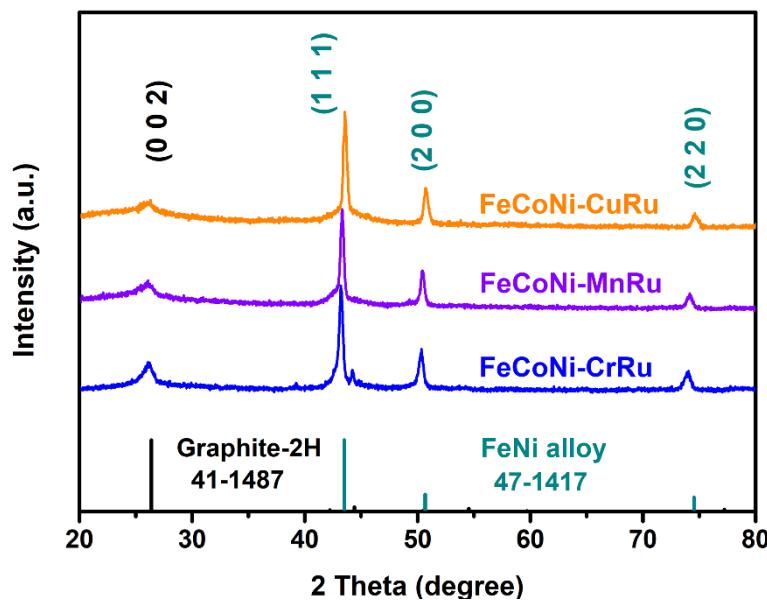
**Supplementary Fig. 7.** FE-SEM images of the (a) FeCoNiCrRu/CNFs and (b) FeCoNiCuRu/CNFs.



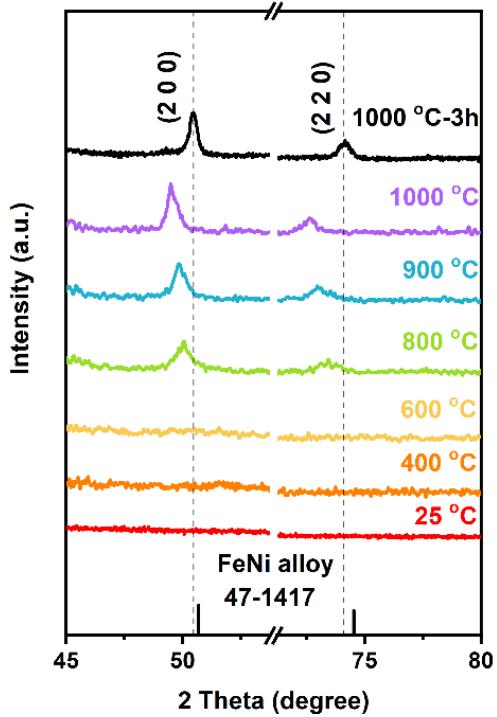
**Supplementary Fig. 8.** FE-SEM images of the as-prepared (a) FeCoNi/CNFs, (b) FeCoNiMn/CNFs, and (c) FeCoNiRu/CNFs.



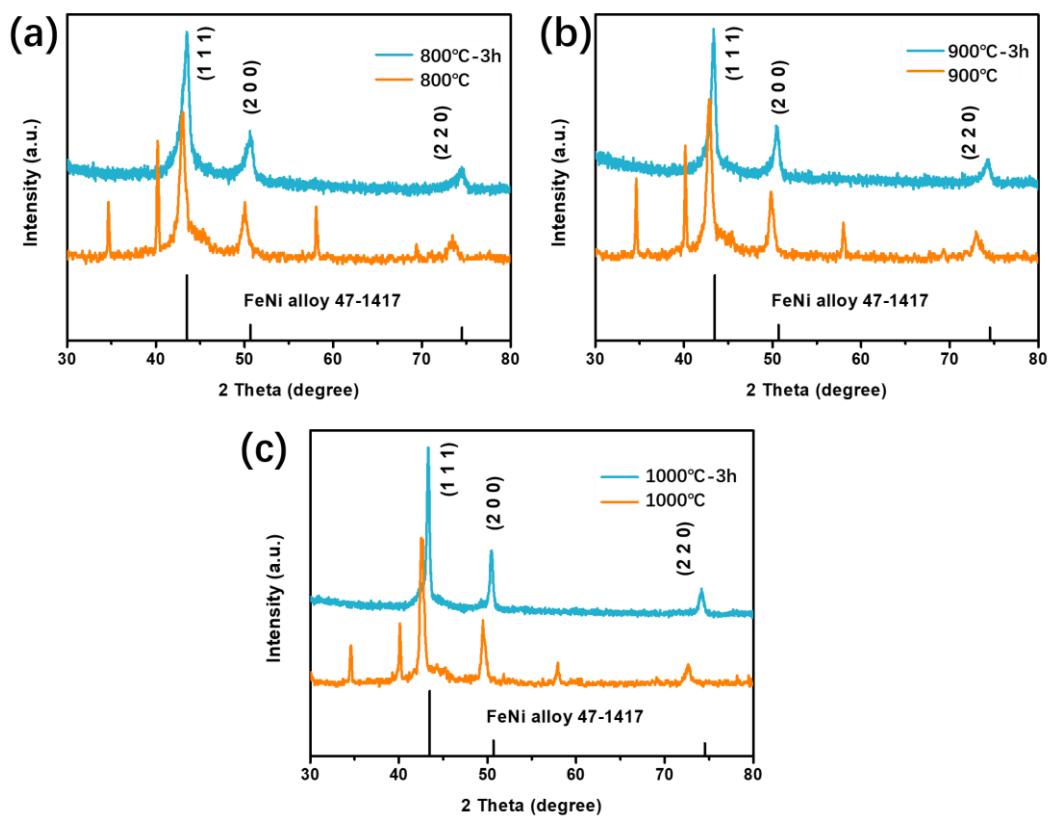
**Supplementary Fig. 9.** XRD patterns of FeCoNi/CNFs, FeCoNiMn/CNFs, and FeCoNiRu/CNFs.



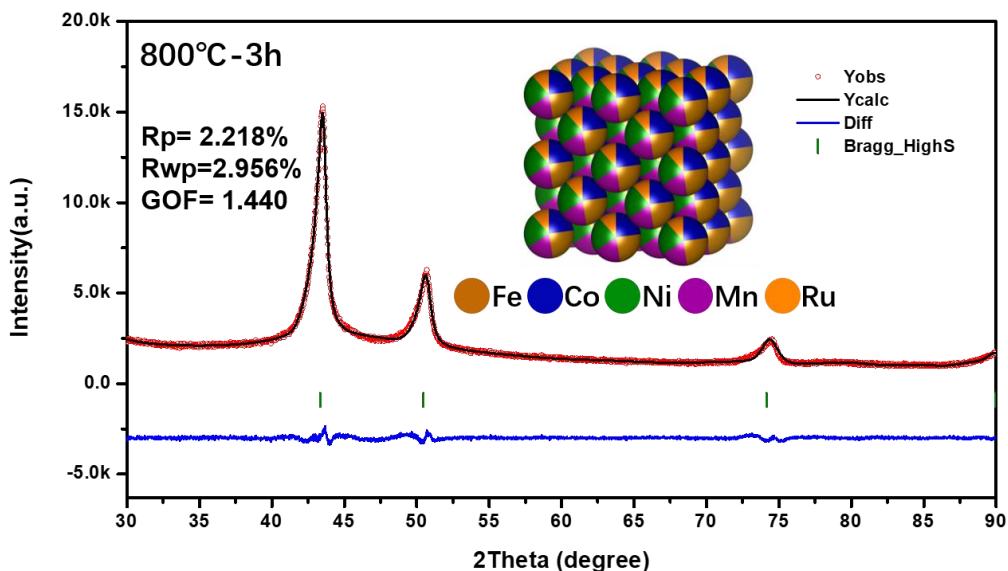
**Supplementary Fig. 10.** XRD patterns of FeCoNiMnRu/CNFs, FeCoNiCrRu/CNFs, and FeCoNiCuRu/CNFs.



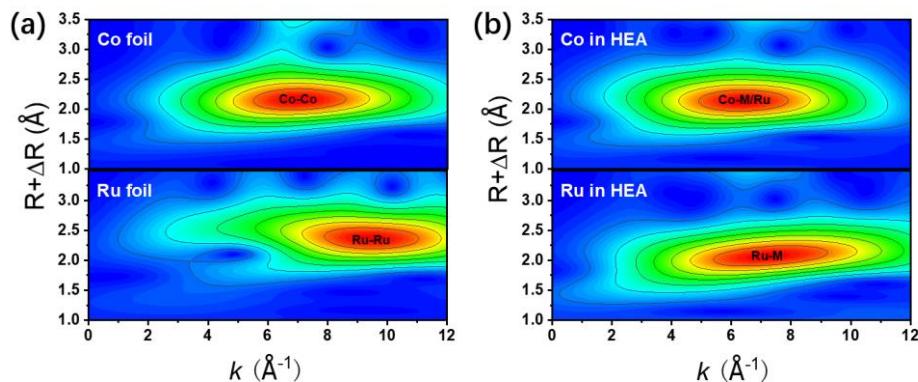
**Supplementary Fig. 11.** The enlarged *in situ* XRD patterns of FeCoNiMnRu/CNFs showing the peak shift of (200) and (220) planes.



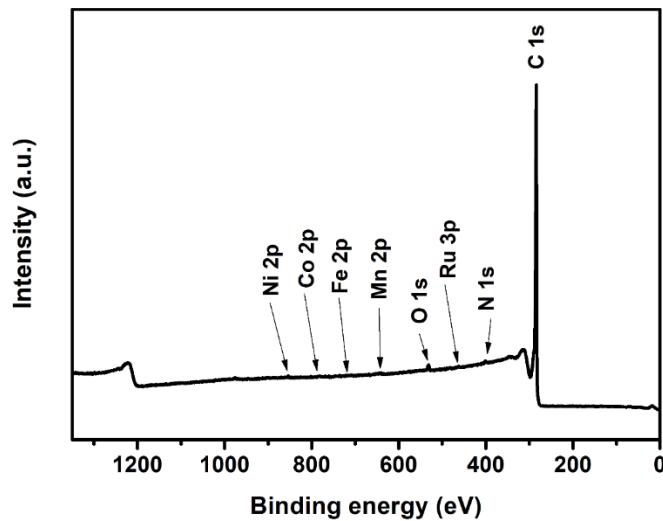
**Supplementary Fig. 12.** XRD patterns of FeCoNiMnRu/CNFs obtained after prolonged heat treatment (upper) and during *in situ* characterization (lower) at (a) 800 °C, (b) 900 °C, and 1000 °C.



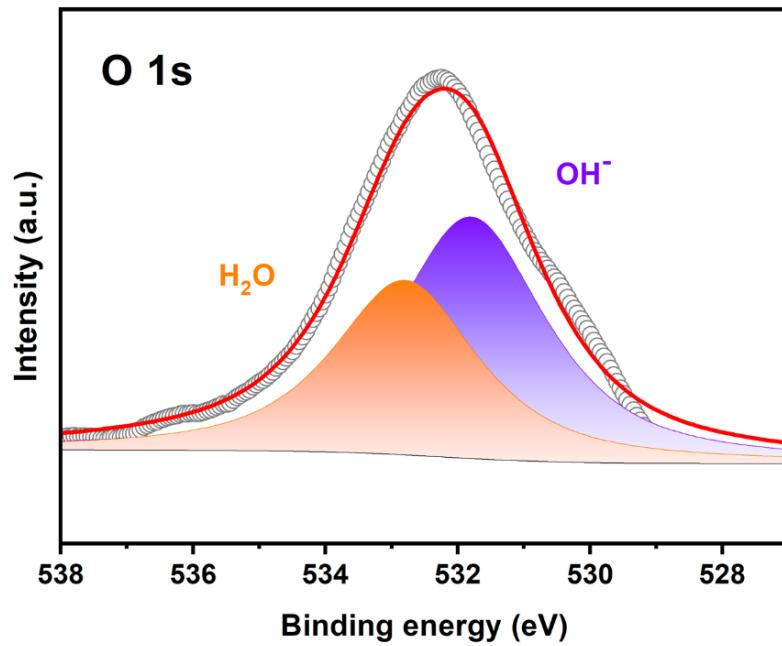
**Supplementary Fig. 13.** XRD pattern of FeCoNiMnRu/CNFs (800 °C-3 h) with a scan rate of 0.5°/min and the corresponding XRD Rietveld refinement.



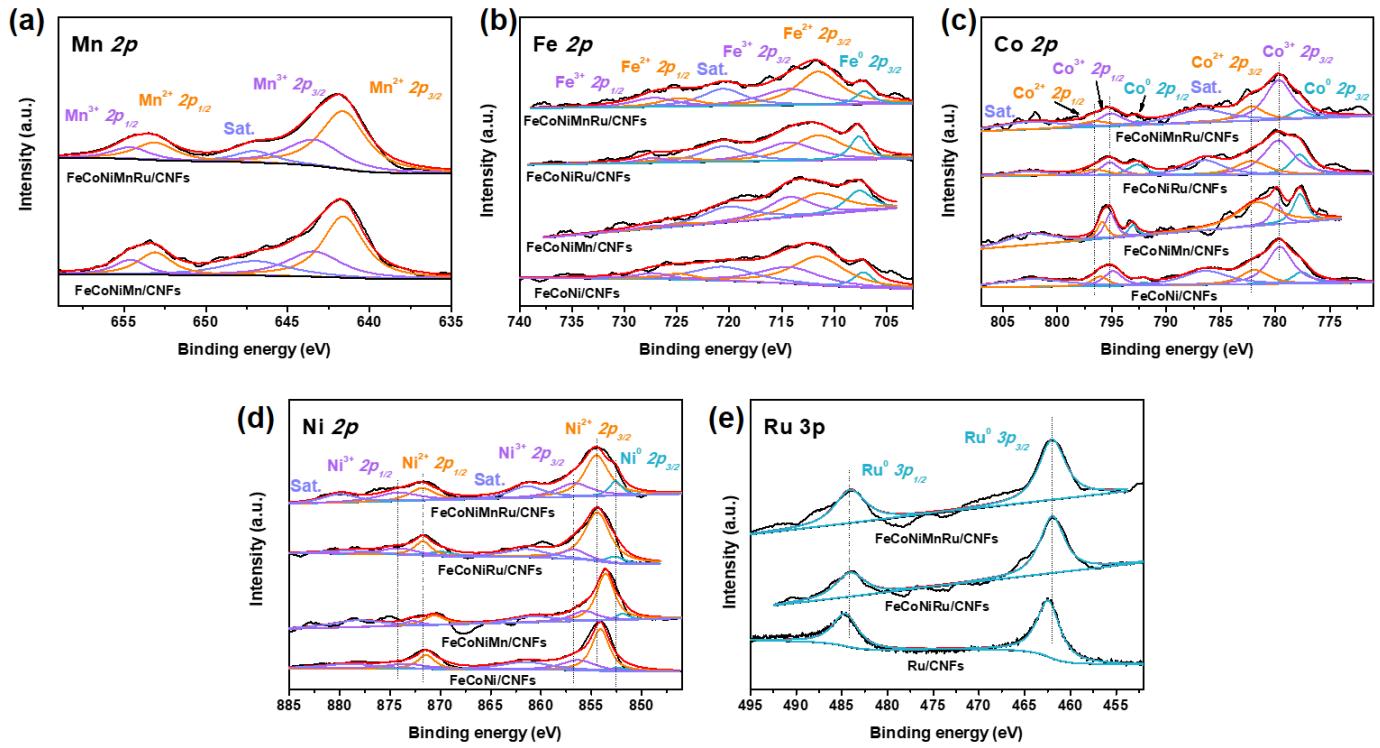
**Supplementary Figure 14.** WT-EXAFS spectra of (a) Co and Ru foil, (b) Co and Ru in FeCoNiMnRu/CNFs.



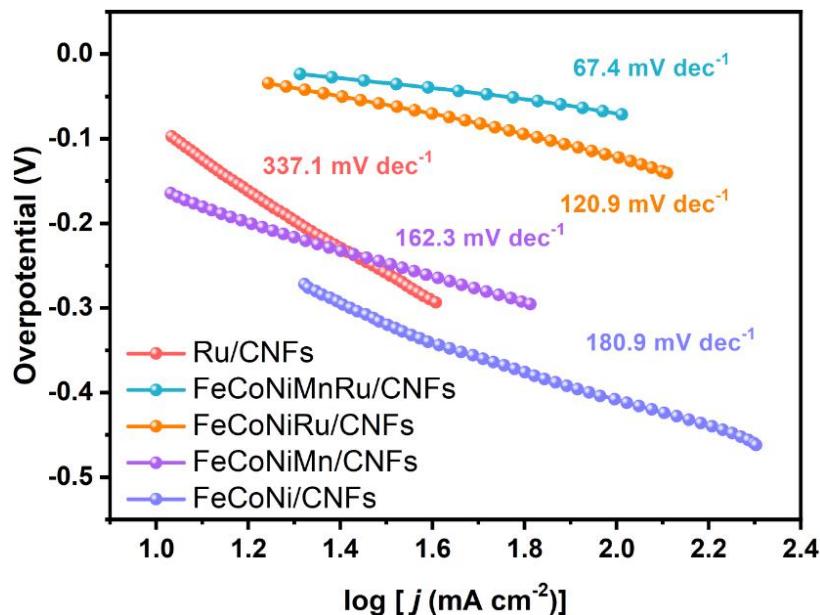
**Supplementary Fig. 15.** XPS survey spectrum of the as-prepared FeCoNiMnRu/CNFs.



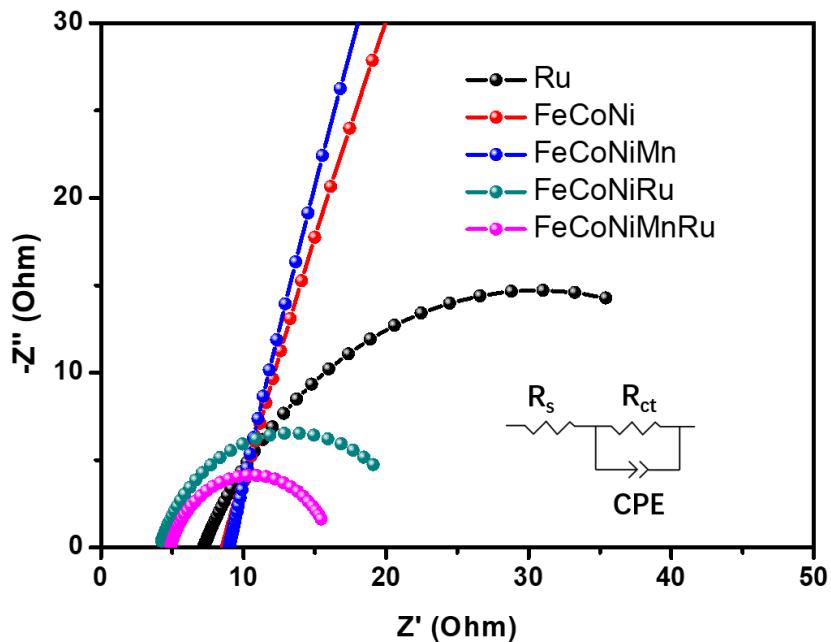
**Supplementary Fig. 16.** High-resolution O 1s XPS spectrum of FeCoNiMnRu/CNFs.



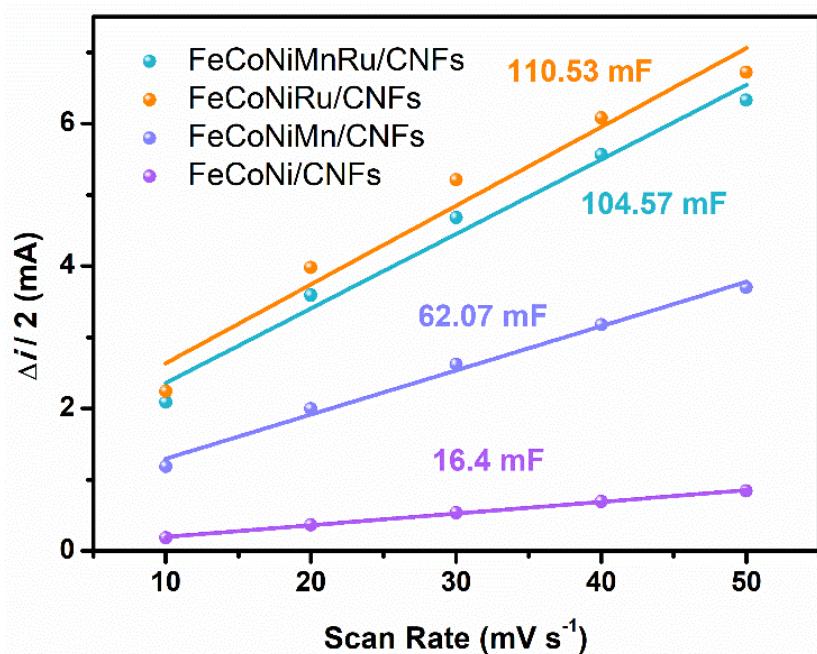
**Supplementary Fig. 17.** High-resolution (a) Mn 2p XPS spectrum of FeCoNiMn/CNFs and FeCoNiMnRu/CNFs, (b) Fe 2P, (c) Co 2p, (d) Ni 2p XPS spectra of FeCoNi/CNFs, FeCoNiMn/CNFs, FeCoNiRu/CNFs, and FeCoNiMnRu/CNFs, (e) Ru 3p XPS spectrum of Ru/CNFs, FeCoNiRu/CNFs, and FeCoNiMnRu/CNFs.



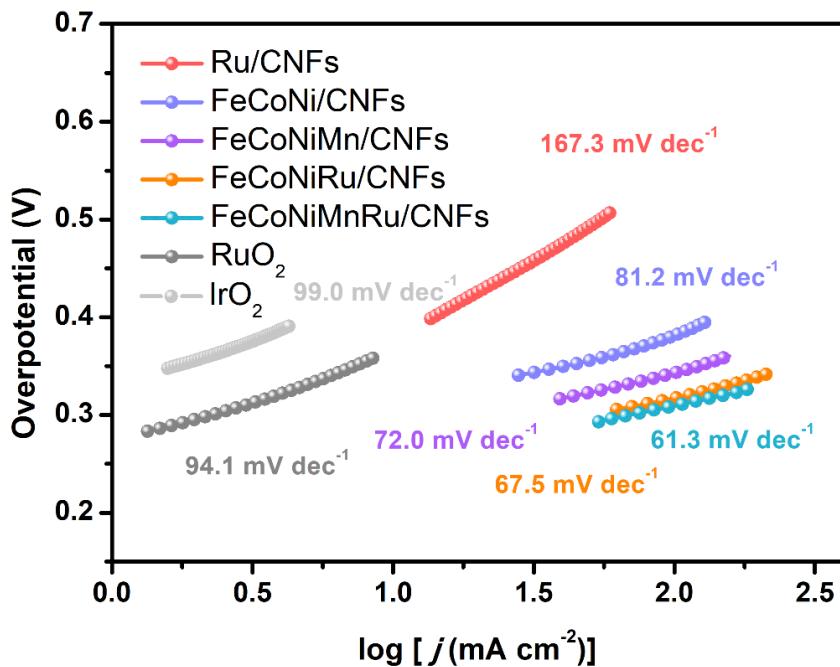
**Supplementary Fig. 18.** Tafel slopes of the as-prepared electrocatalysts for HER.



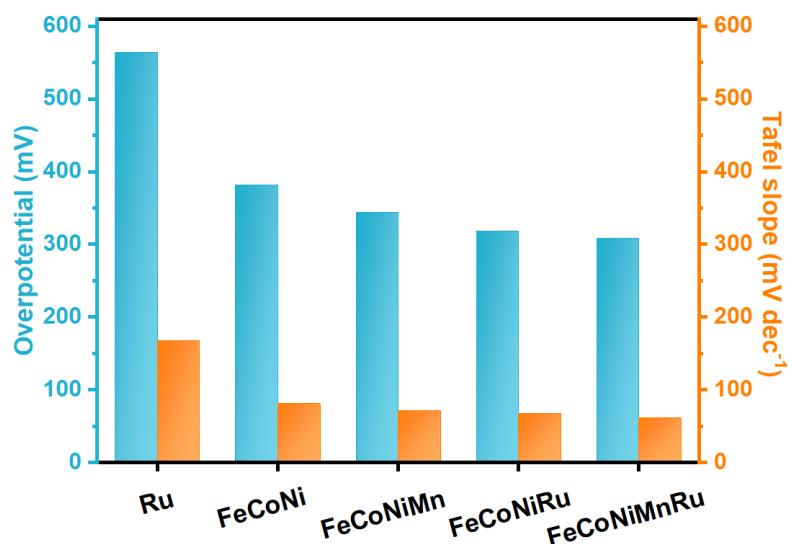
**Supplementary Fig. 19.** EIS spectra of the as-prepared electrocatalysts.



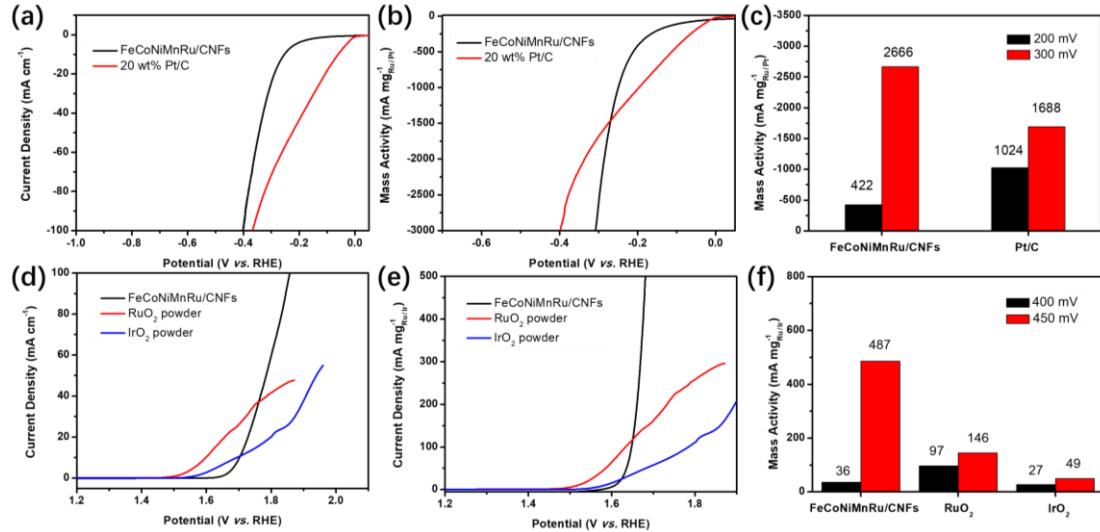
**Supplementary Fig. 20.** Double-layer capacitance per geometric area ( $C_{dl}$ ) of the as-prepared electrodes.



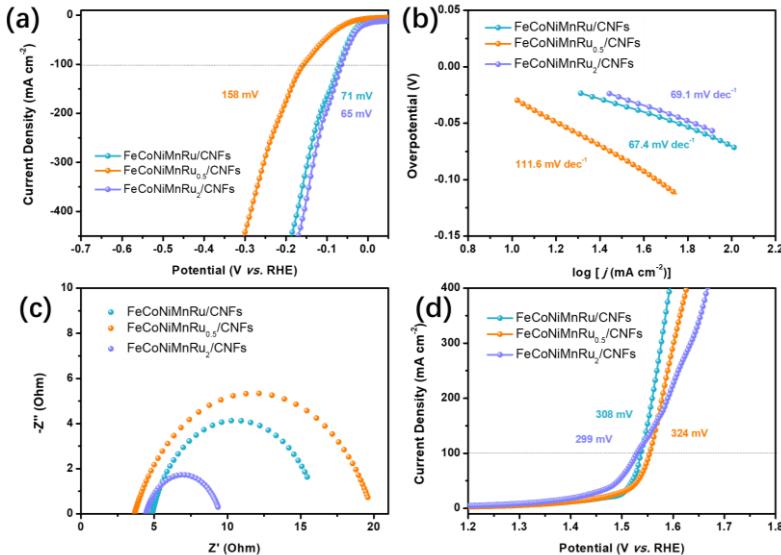
**Supplementary Fig. 21.** Tafel slopes of the as-prepared electrocatalysts for OER.



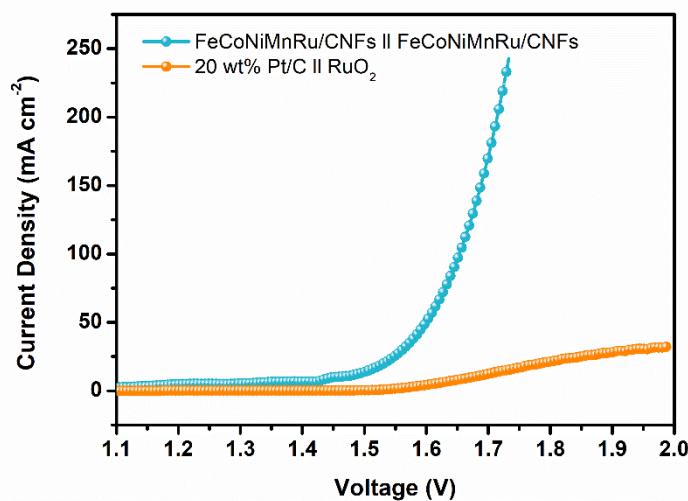
**Supplementary Fig. 22.** The corresponding histogram of OER overpotentials at  $100 \text{ mA cm}^{-2}$  and Tafel slopes obtained on FeCoNi/CNFs, FeCoNiMn/CNFs, FeCoNiRu/CNFs, FeCoNiMnRu/CNFs, and Ru/C in 1.0 M KOH electrolyte.



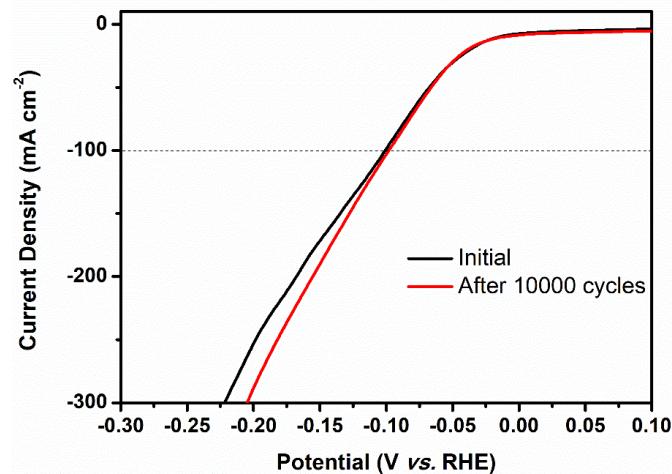
**Supplementary Fig. 23.** LSV curves of GCE supported FeCoNiMnRu/CNFs and 20 wt% Pt/C for HER normalized by (a) geometric area and (b) mass loading of noble metal (Ru or Pt). (c) The mass activities of FeCoNiMnRu/CNFs and 20 wt% Pt/C for HER at overpotentials of 200 and 300 mV. LSV curves of GCE supported FeCoNiMnRu/CNFs, RuO<sub>2</sub> powder, and IrO<sub>2</sub> powder for OER normalized by (d) geometric area and (e) mass loading of noble metal (Ru or Ir). (f) The mass activities of FeCoNiMnRu/CNFs, RuO<sub>2</sub> powder, and IrO<sub>2</sub> powder for OER at overpotentials of 400 and 450 mV.



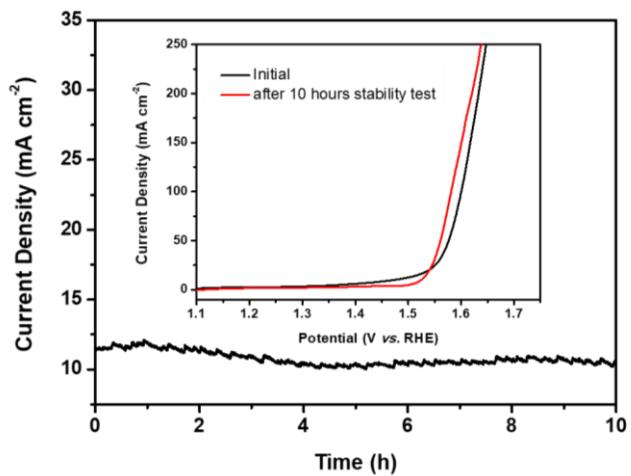
**Supplementary Fig. 24.** (a) LSV curves, corresponding (b) Tafel plots, and (c) EIS spectra of the as-prepared FeCoNiMnRu/CNFs, FeCoNiMnRu<sub>0.5</sub>/CNFs and FeCoNiMnRu<sub>2</sub>/CNFs used for HER in 1 M KOH. (d) LSV curves of the as-prepared FeCoNiMnRu/CNFs, FeCoNiMnRu<sub>0.5</sub>/CNFs and FeCoNiMnRu<sub>2</sub>/CNFs used for OER in 1 M KOH.



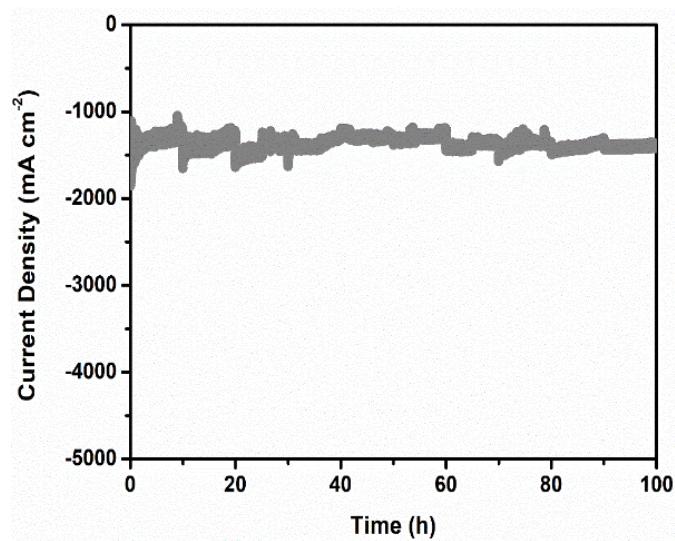
**Supplementary Fig. 25.** Polarization curves for overall water splitting by  $\text{FeCoNiMnRu/CNFs} \parallel \text{FeCoNiMnRu/CNFs}$  and  $20 \text{ wt\% Pt/C} \parallel \text{RuO}_2$  in a two-electrode configuration in 1 M KOH at a scan rate of 2 mV/s.



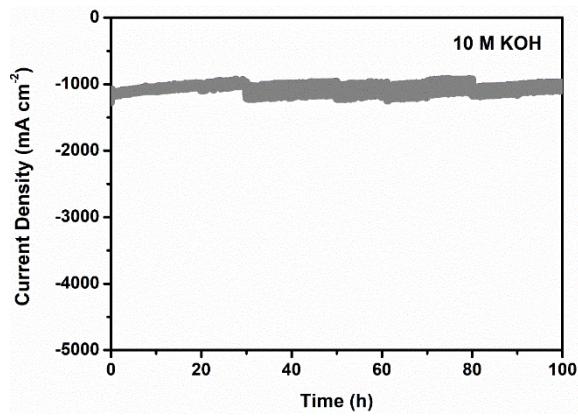
**Supplementary Fig. 26.** LSV curves of  $\text{FeCoNiMnRu/CNFs}$  before and after 10000 CV cycles.



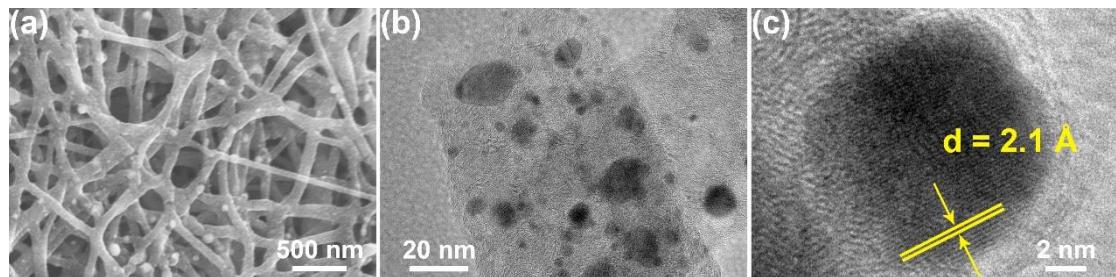
**Supplementary Fig. 27.** The chronoamperometric curve for FeCoNiMnRu/CNFs measured at 1.55 V vs. RHE for 10 h. Inset shows the LSV curves of FeCoNiMnRu/CNFs before and after stability test.



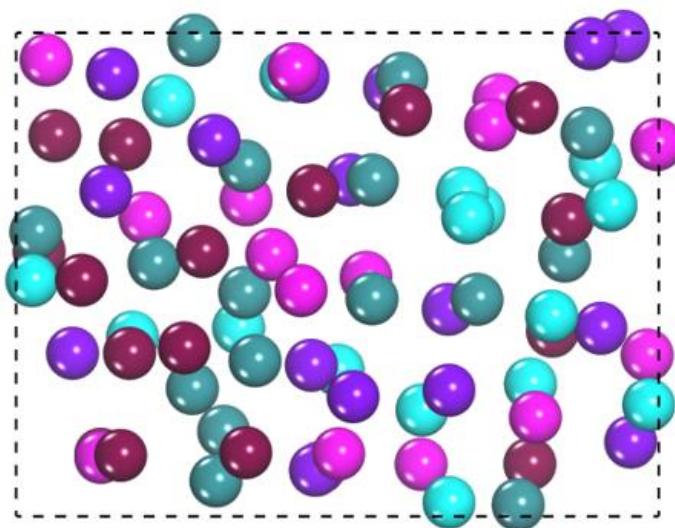
**Supplementary Fig. 28.** The chronoamperometric curve for FeCoNiMnRu/CNFs measured at 60 °C and -2.22 V vs. RHE for 100 h in 1 M KOH.



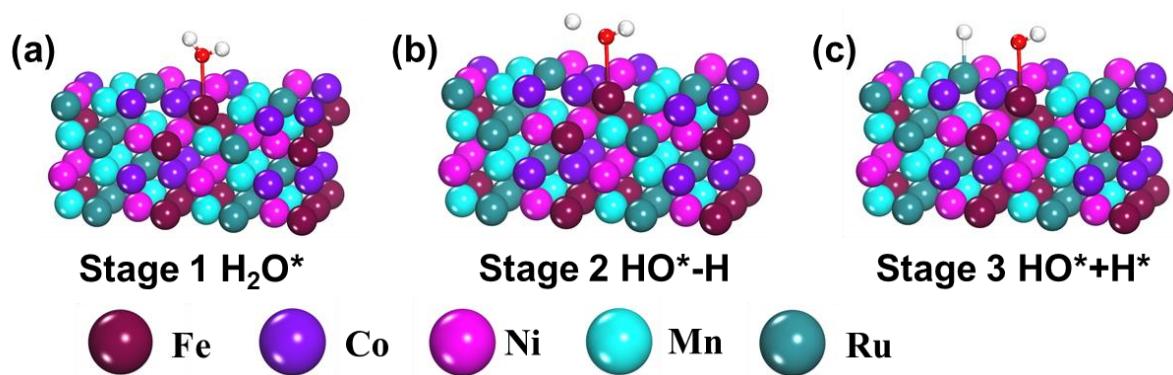
**Supplementary Fig. 29.** The chronoamperometric curve for FeCoNiMnRu/CNFs measured at -0.77 V vs. RHE for 100 h in 10 M KOH.



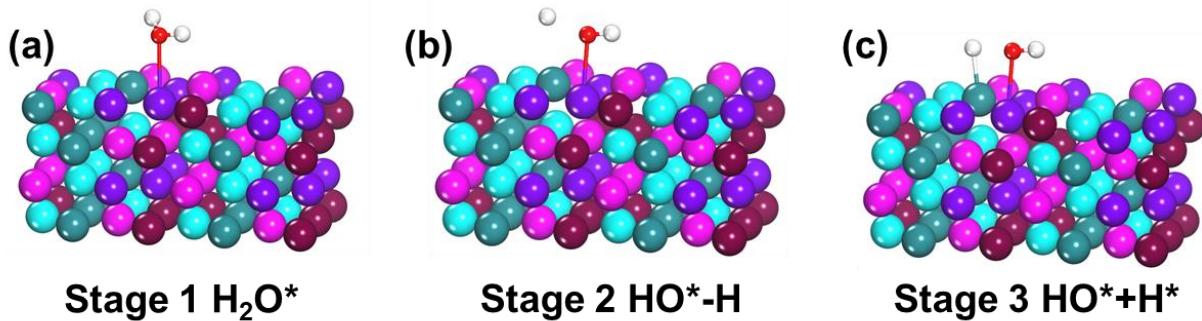
**Supplementary Fig. 30.** (a) FE-SEM, (b) TEM and (c) HRTEM images of FeCoNiMnRu/CNFs electrode after long-term stability test.



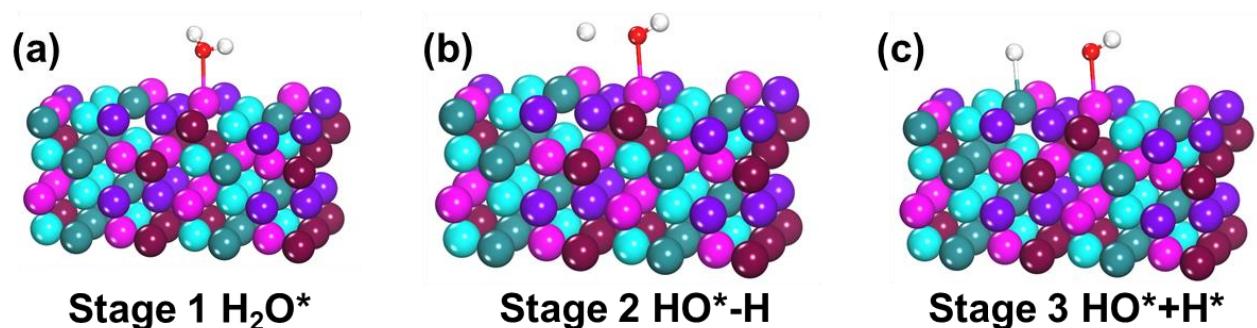
**Supplementary Fig. 31.** Chemical structure of FeCoNiMnRu HEA.



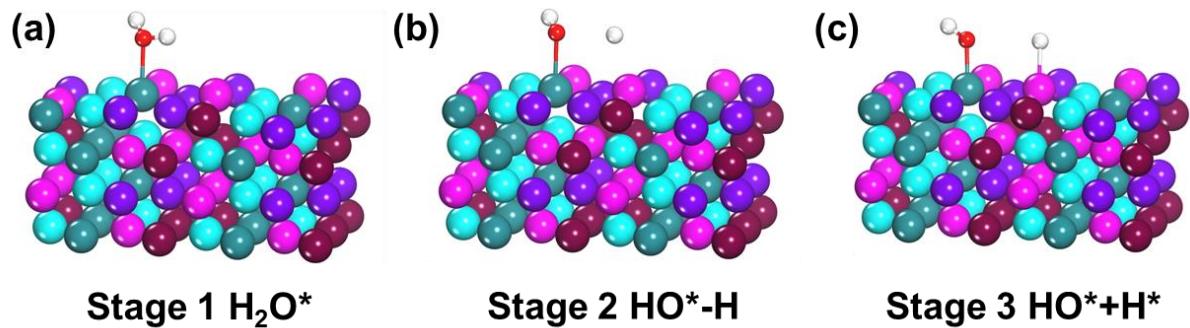
**Supplementary Fig. 32.** Atomic configurations on Fe site of FeCoNiMnRu HEA at the three stages (a-c) during the  $\text{H}_2\text{O}$  dissociation into  $\text{OH}^*$  and  $\text{H}^*$ .



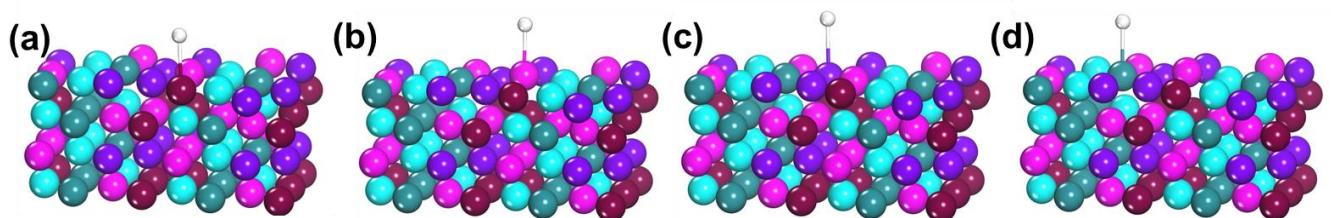
**Supplementary Fig. 33.** Atomic configurations on Co site of FeCoNiMnRu HEA at the three stages (a-c) during the  $\text{H}_2\text{O}$  dissociation into  $\text{OH}^*$  and  $\text{H}^*$ .



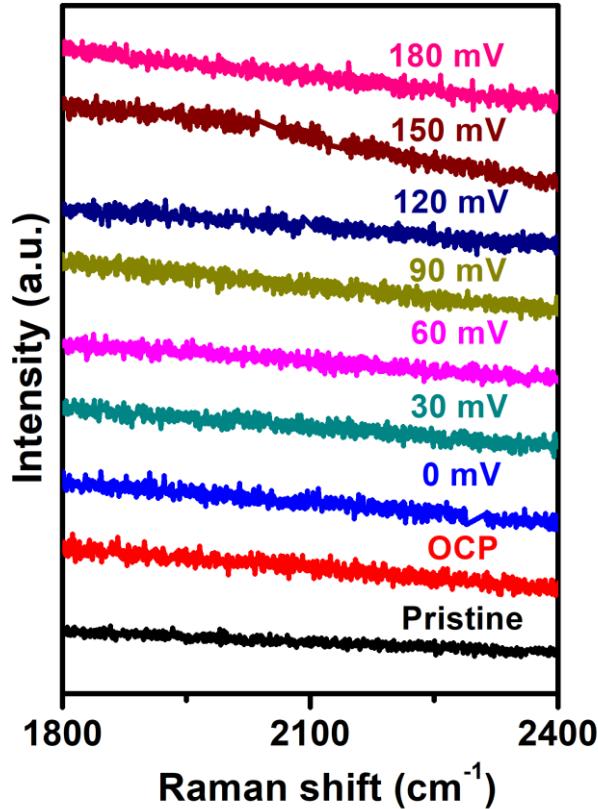
**Supplementary Fig. 34.** Atomic configurations on Ni site of FeCoNiMnRu HEA at the three stages (a-c) during the  $\text{H}_2\text{O}$  dissociation into  $\text{OH}^*$  and  $\text{H}^*$ .



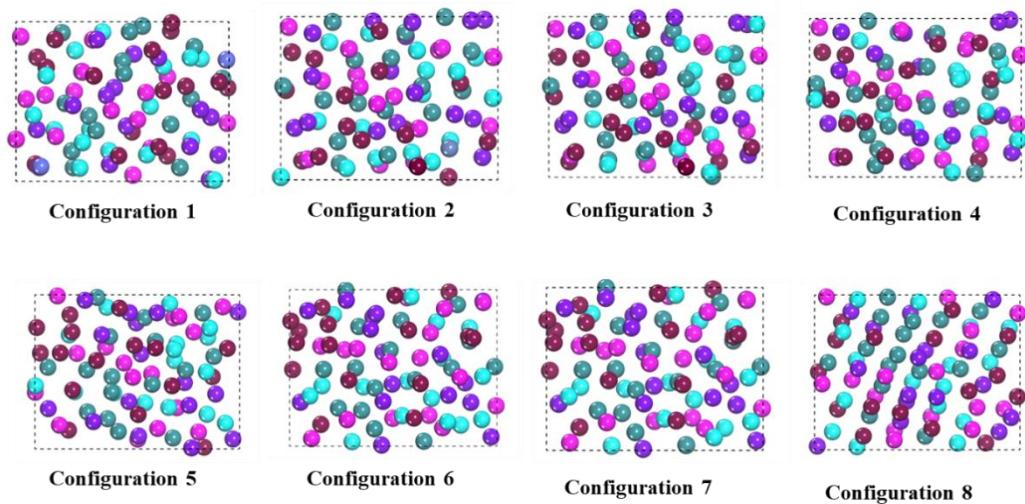
**Supplementary Fig. 35.** Atomic configurations on Ru site of FeCoNiMnRu HEA at the three stages (a-c) during the  $\text{H}_2\text{O}$  dissociation into  $\text{OH}^*$  and  $\text{H}^*$ .



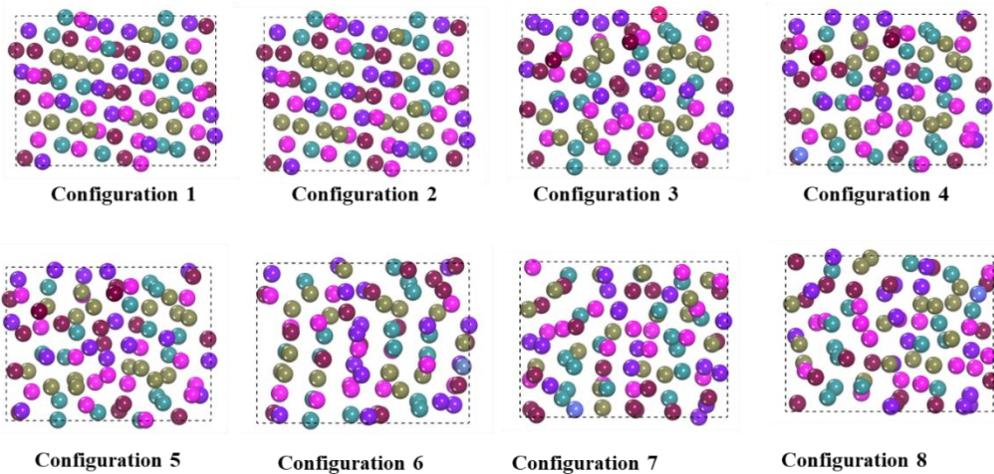
**Supplementary Fig. 36.** Atomic configurations on (a) Fe, (b) Ni, (c) Co, and (d) Ru sites of FeCoNiMnRu HEA at the stage of  $\text{H}^*$  absorption.



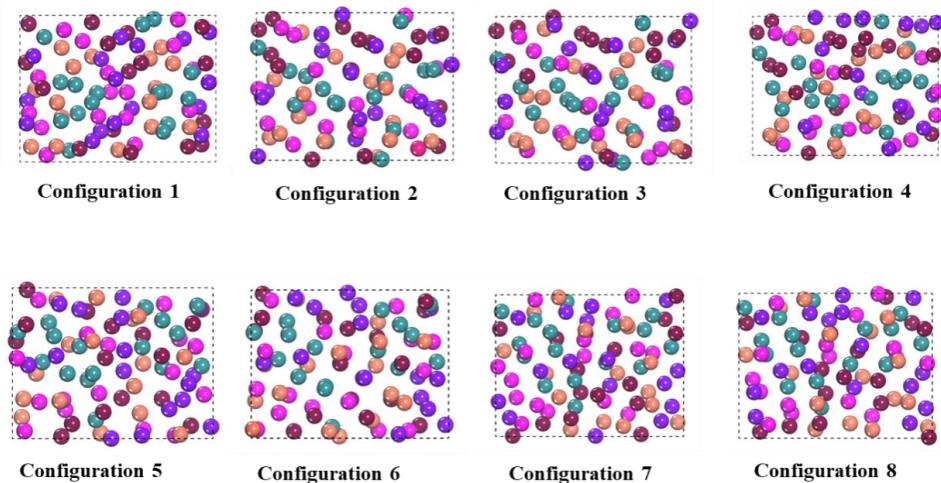
**Supplementary Fig. 37.** Operando electrochemical-Raman spectra collected by FeCoNiMn/CNFs during the HER process in 1.0 M electrolyte.



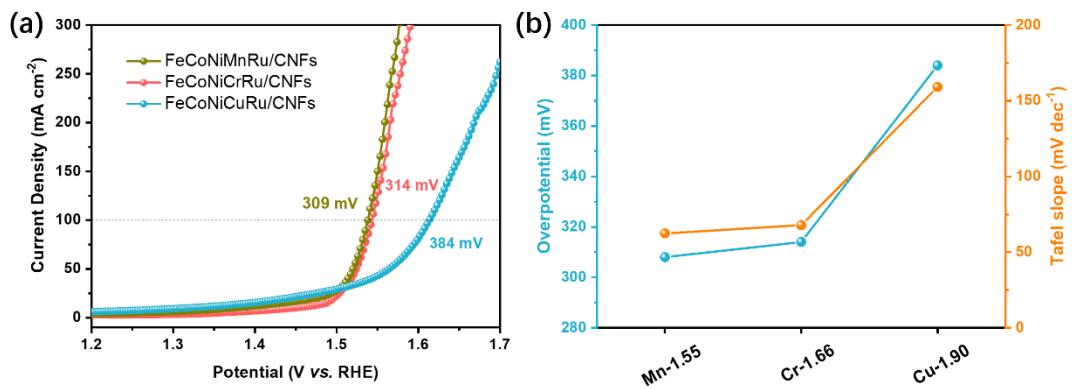
**Supplementary Fig. 38.** Different atomic configurations of FeCoNiMnRu HEA.



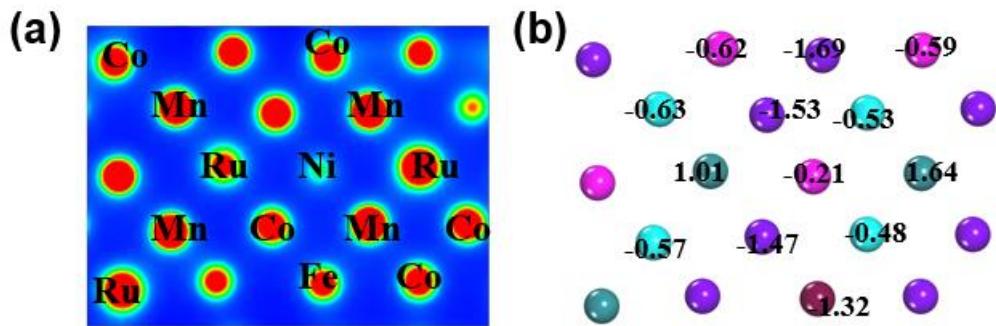
**Supplementary Fig. 39.** Different atomic configurations of FeCoNiCrRu HEA.



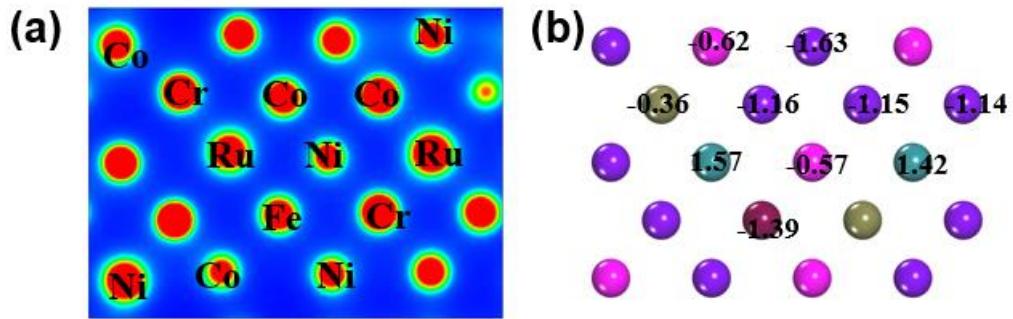
**Supplementary Fig. 40.** Different atomic configurations of FeCoNiCuRu HEA.



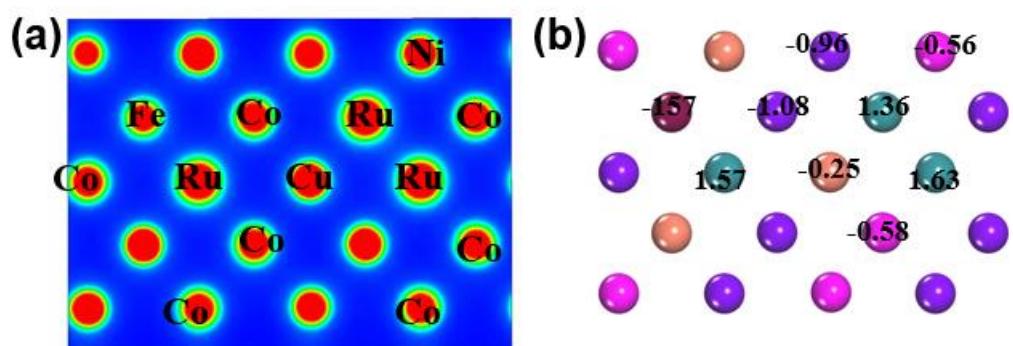
**Supplementary Fig. 41.** (a) OER polarization curves obtained for FeCoNiXRu/CNFs (X=Cr, Mn, and Cu) in 1.0 M KOH electrolyte. (b) Correlation between the OER overpotentials at 100 mA cm<sup>-2</sup>, Tafel slopes and the electronegativities of metals X (Cr, Mn, and Cu).



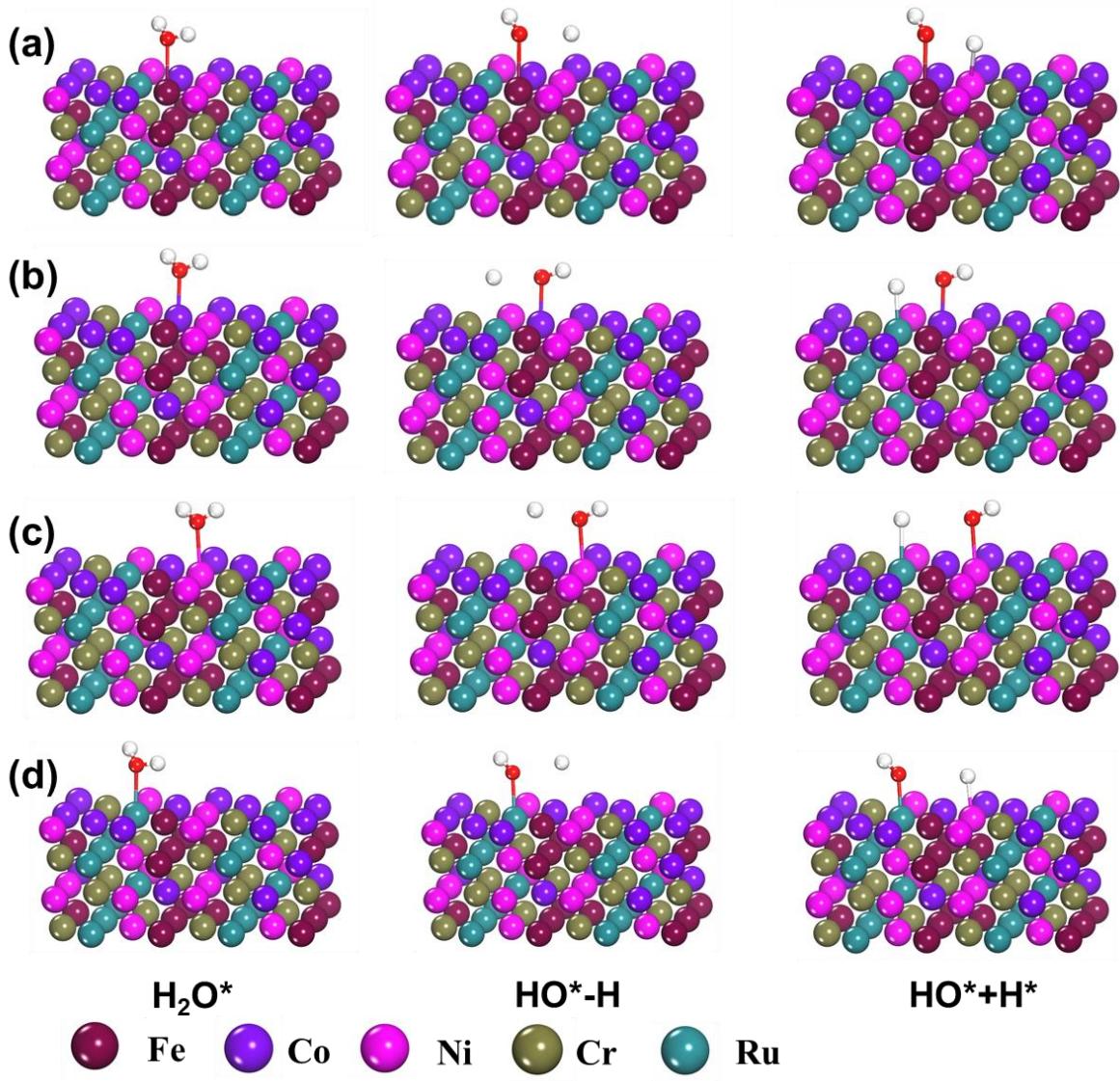
**Supplementary Fig. 42.** (a) Charge density difference analysis and (b) Bader charge analysis of FeCoNiMnRu system.



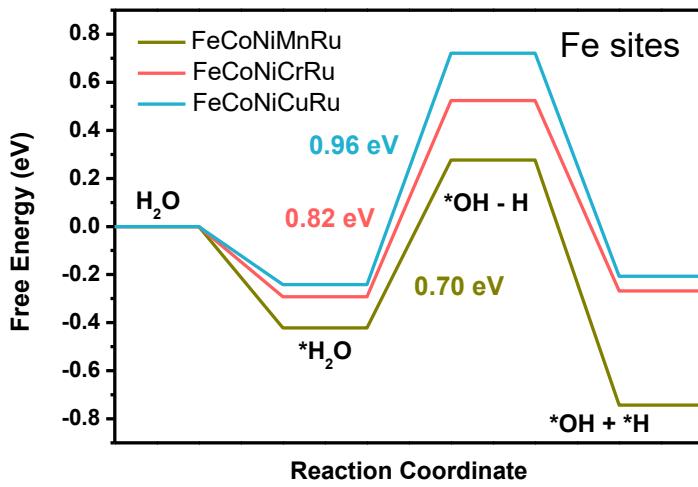
**Supplementary Fig. 43.** (a) Charge density difference analysis and (b) Bader charge analysis of FeCoNiCrRu system.



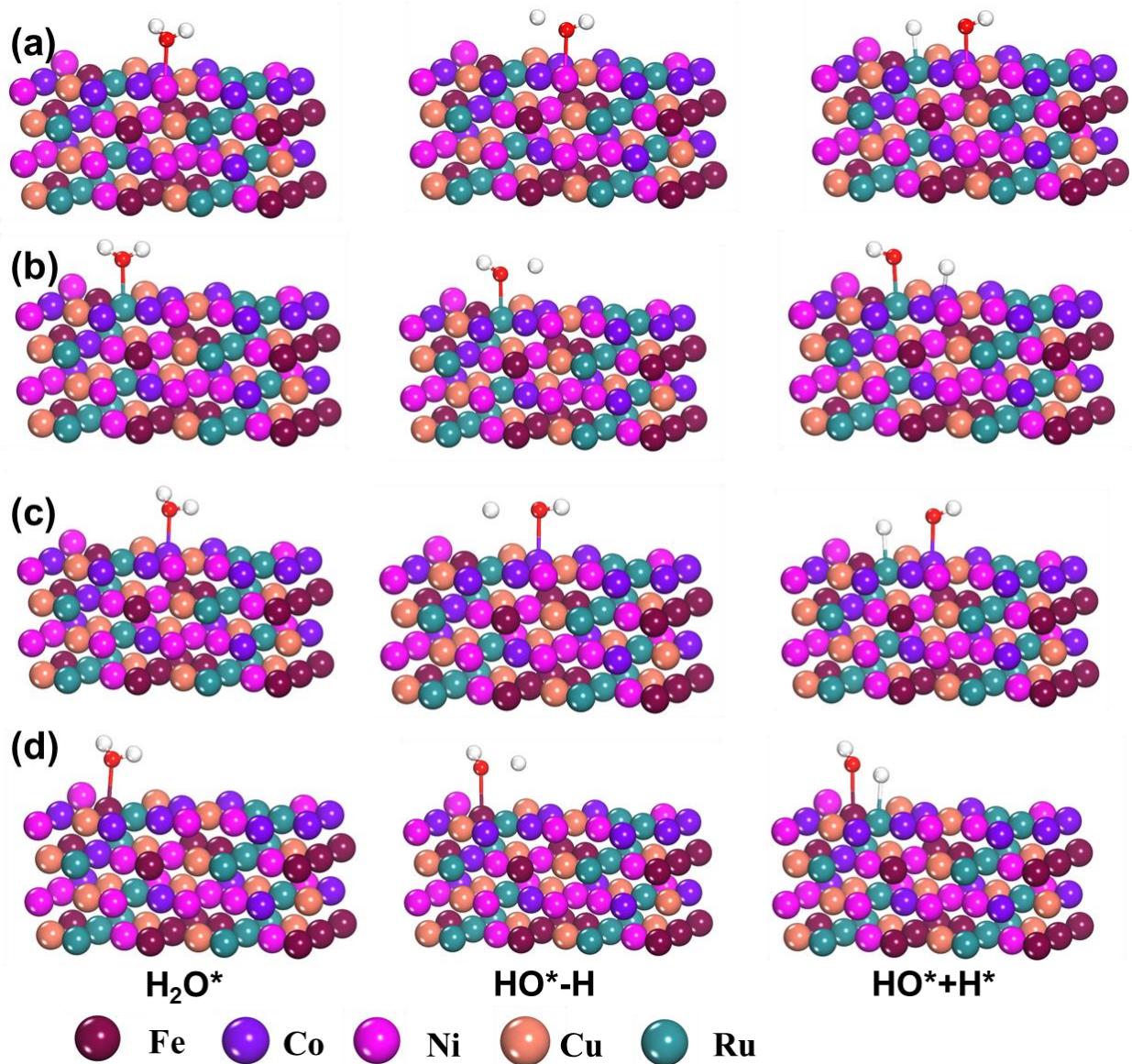
**Supplementary Fig. 44.** (a) Charge density difference analysis and (b) Bader charge analysis of FeCoNiCuRu system.



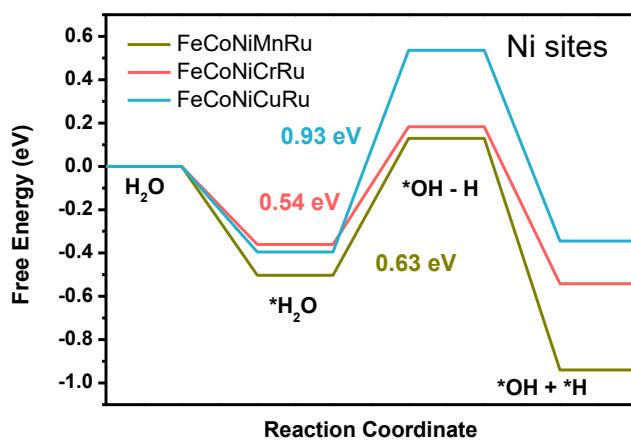
**Supplementary Fig. 45.** The atomic configurations on (a) Fe, (b) Co, (c) Ni, and (d) Ru catalytic sites of FeCoNiCrRu HEA during the water dissociation.



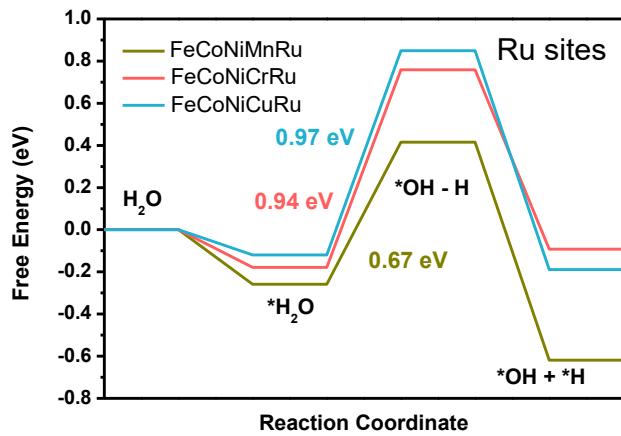
**Supplementary Fig. 46.** Reaction energy profile of water dissociation on Fe sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



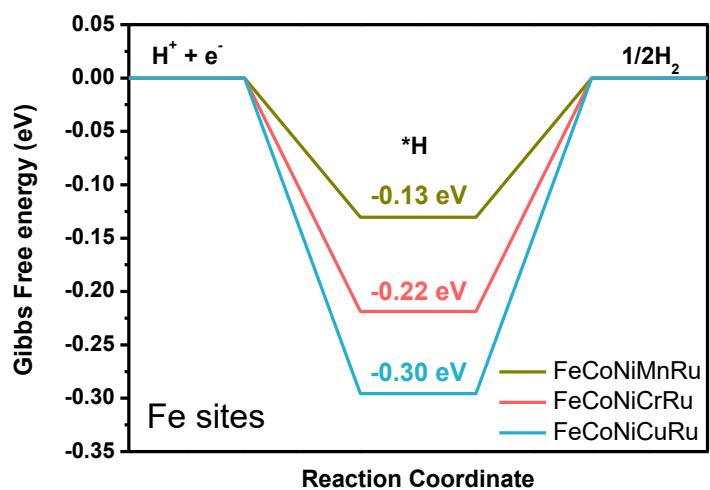
**Supplementary Fig. 47.** The atomic configurations on (a) Fe, (b) Co, (c) Ni, and (d) Ru catalytic sites of FeCoNiCuRu HEA during the water dissociation.



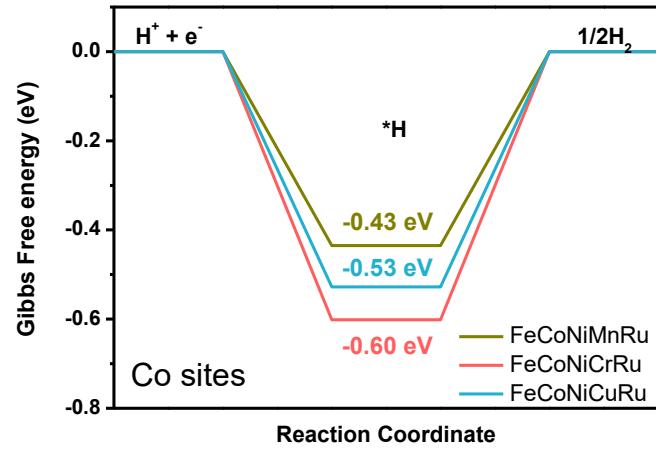
**Supplementary Fig. 48.** Reaction energy profile of water dissociation on Ni sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



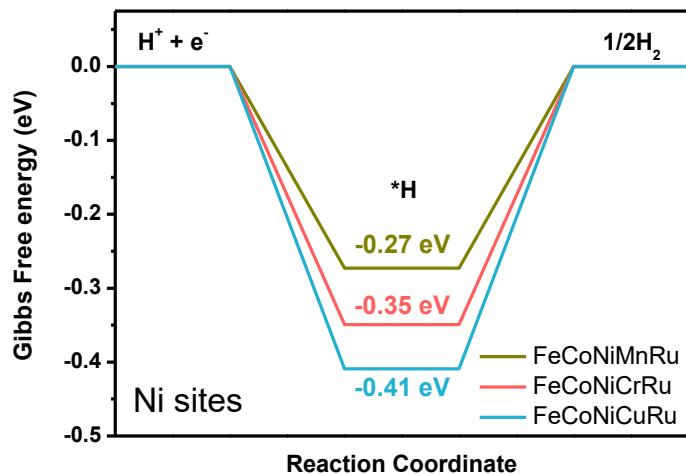
**Supplementary Fig. 49.** Reaction energy profile of water dissociation on Ru sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



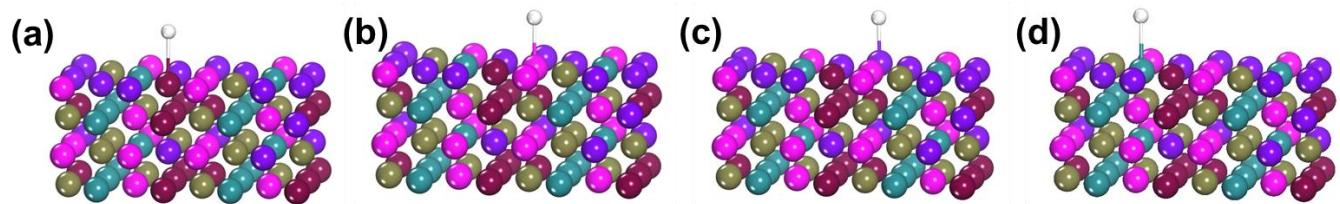
**Supplementary Fig. 50.** Gibbs free energy ( $\Delta G_{\text{H}^*}$ ) profiles on Fe sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



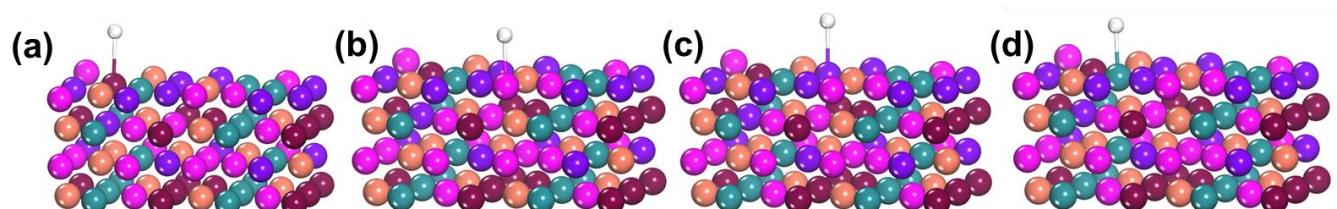
**Supplementary Fig. 51.** Gibbs free energy ( $\Delta G_{H^*}$ ) profiles on Co sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



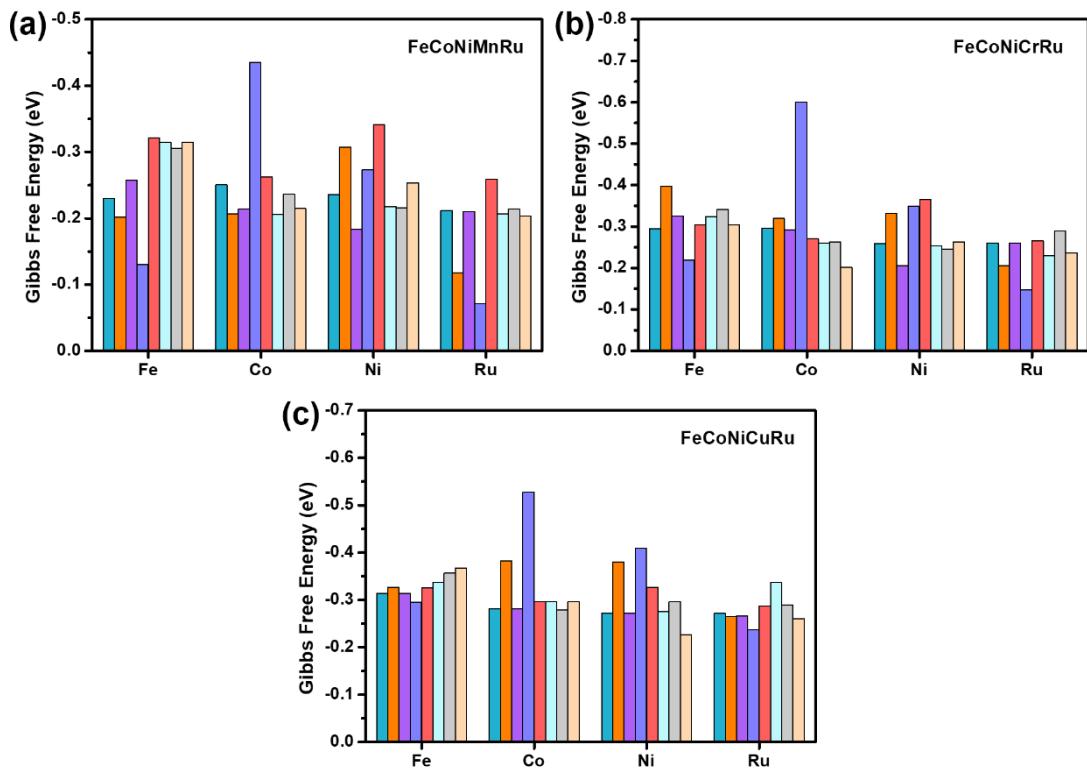
**Supplementary Fig. 52.** Gibbs free energy ( $\Delta G_{H^*}$ ) profiles on Ni sites of FeCoNiMnRu, FeCoNiCrRu, and FeCoNiCuRu HEA surfaces.



**Supplementary Fig. 53.** The atomic configurations on (a) Fe, (b) Ni, (c) Co, and (d) Ru catalytic sites of FeCoNiCrRu for H\* absorption.



**Supplementary Fig. 54.** The atomic configurations on (a) Fe, (b) Ni, (c) Co, and (d) Ru catalytic sites of FeCoNiCuRu for H\* absorption.



**Supplementary Fig. 55.** The summarized values of  $\Delta G_{H^*}$  on Fe-, Co-, Ni-, and Ru-sites in (a) FeCoNiMnRu, (b) FeCoNiCrRu, and (c) FeCoNiCuRu with different atomic configurations.

**Supplementary Table 1.** The mass loadings of Fe, Co, Ni, Mn, and Ru in FeCoNiMnRu/CNFs measured by ICP-OES and the corresponding atomic percentages.

FeCoNiMnRu	Fe	Co	Ni	Mn	Ru
Mass loading (wt. %)	2.97	3.05	2.96	1.78	4.38
Atomic percentage (at. %)	23.00	22.38	21.82	14.03	18.77

**Supplementary Table 2.** The mass loadings of Fe, Co, Ni, Cr, and Ru in FeCoNiCrRu/CNFs measured by ICP-OES and the corresponding atomic percentages.

FeCoNiMnRu	Fe	Co	Ni	Cr	Ru
Mass loading (wt. %)	2.57	2.82	2.67	2.68	4.07
Atomic percentage (at. %)	19.92	20.71	19.67	22.29	17.41

**Supplementary Table 3.** The mass loadings of Fe, Co, Ni, Cu, and Ru in FeCoNiCuRu/CNFs measured by ICP-OES and the corresponding atomic percentages.

FeCoNiMnRu	Fe	Co	Ni	Cu	Ru
Mass loading (wt. %)	2.85	2.95	2.57	2.54	4.13
Atomic percentage (at. %)	22.63	22.21	19.38	17.68	18.10

**Supplementary Table 4.** EXAFS fitting parameters at the Co and Ru K-edge for various samples.

Sample	Shell	CN <sup>a</sup>	R(Å) <sup>b</sup>	σ <sup>2</sup> (Å <sup>2</sup> ) <sup>c</sup>	ΔE <sub>0</sub> (eV) <sup>d</sup>	R factor
Co K-edge ( $S_0^2=0.811$ )						
Co foil	Co-Co	12*	2.49	0.0063	6.9	0.0023
CoO	Co-O	6.0	2.10	0.0106	1.5	0.0053
	Co-Co	12.0	3.01	0.0080		
	Co-O	6.2	1.91	0.0035		
Co <sub>3</sub> O <sub>4</sub>	Co-Co	6.1	2.86	0.0048	-8.1	0.0079
	Co-Co	7.4	3.36	0.0048		
FeCoNiMn Ru/CNFs before stability test	Co-Co	7.2	2.51	0.0066	4.5	0.0059
	Co-Ru	1.4	2.49	0.0066		
FeCoNiMn Ru/CNFs after stability test	Co-Co	6.9	2.50	0.0085	3.0	0.0071
	Co-Ru	0.3	2.52	0.0085		
Ru K-edge ( $S_0^2=0.816$ )						
Ru foil	Ru-Ru	12*	2.67	0.0033	6.6	0.0071
	Ru-O	6.0	1.98	0.0035		
RuO <sub>2</sub>	Ru-Ru	3.0	3.12	0.0045	2.4	0.0085
	Ru-Ru	11.7	3.56	0.0045		

FeCoNiMn Ru/CNFs before stability test	Ru-Co	6.7	2.52	0.0049	-3.0	0.0059
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**Supplementary Table 5.** The summarized binding energies for Ru/CNFs, FeCoNi/CNFs, FeCoNiMn/CNFs, FeCoNiRu/CNFs, and FeCoNiMnRu/CNFs.

	Ru/CNFs	FeCoNi/CNFs	FeCoNiMn/CNFs	FeCoNiRu/CNFs	FeCoNiMnRu/CNFs
Mn <sup>2+</sup> 2p <sub>3/2</sub>	/	/	641.6	/	641.6
Mn <sup>2+</sup> 2p <sub>1/2</sub>	/	/	653.1	/	653.1
Mn <sup>3+</sup> 2p <sub>3/2</sub>	/	/	643.4	/	643.4
Mn <sup>3+</sup> 2p <sub>1/2</sub>	/	/	654.6	/	654.6
Fe <sup>0</sup> 2p <sub>3/2</sub>	/	707.2	707.6	707.6	707.1
Fe <sup>2+</sup> 2p <sub>3/2</sub>	/	711.5	711.5	711.5	711.5
Fe <sup>2+</sup> 2p <sub>1/2</sub>	/	724.8	724.8	724.8	724.8
Fe <sup>3+</sup> 2p <sub>3/2</sub>	/	714.3	714.3	714.3	714.3
Fe <sup>3+</sup> 2p <sub>1/2</sub>	/	727.3	727.3	727.3	727.2
Co <sup>0</sup> 2p <sub>3/2</sub>	/	777.6	777.7	777.8	777.8
Co <sup>0</sup> 2p <sub>1/2</sub>	/	792.1	793.1	792.7	792.7
Co <sup>2+</sup> 2p <sub>3/2</sub>	/	781.9	781.7	782.2	782.2
Co <sup>2+</sup> 2p <sub>1/2</sub>	/	796.1	759.9	796.6	796.6
Co <sup>3+</sup> 2p <sub>3/2</sub>	/	779.6	779.8	779.7	779.7
Co <sup>3+</sup> 2p <sub>1/2</sub>	/	794.9	795.1	795.2	795.1
Ni <sup>0</sup> 2p <sub>3/2</sub>	/	852.4	851.9	852.6	852.6
Ni <sup>0</sup> 2p <sub>1/2</sub>	/	869.7	/	869.9	869.9
Ni <sup>2+</sup> 2p <sub>3/2</sub>	/	854.1	853.6	854.4	854.5
Ni <sup>2+</sup> 2p <sub>1/2</sub>	/	871.4	870.6	871.7	871.8
Ni <sup>3+</sup> 2p <sub>3/2</sub>	/	856.3	855.7	856.7	856.7
Ni <sup>3+</sup> 2p <sub>1/2</sub>	/	873.4	872.8	874.1	874.2
Ru <sup>0</sup> 3p <sub>3/2</sub>	462.4	/	/	462.0	462.0
Ru <sup>0</sup> 3p <sub>1/2</sub>	484.6	/	/	484.1	484.1

**Supplementary Table 6.** Summarized overpotentials at 10 mA cm<sup>-2</sup> ( $\eta_{10}$ ) and Tafel slopes of electrocatalysts for HER in 1.0 M KOH.

Electrocatalysts	$\eta_{10}$ (mV)	Tafel slope (mV dec <sup>-1</sup> )	Ref
<b>FeCoNiMnRu/CNFs</b>	<b>5</b>	<b>67.4</b>	<b>This work</b>
CuAlNiMoFe	9.7	60	1
FeCoNiAlTi	88.2	40.1	2
NiFeMoCoCr	172	66	3
Ru/CoO NRs	55	72	4
FeNiCoCrMn-G	210	105	5
PdPtCuNiP	32	37.4	6
(Ru-Co)O <sub>x</sub>	44.1	23.5	7
Ru <sub>1</sub> Ni <sub>1</sub> -NCNFs	35	30	8
RFNOH-10	13	30.0	9
RuO <sub>2</sub> /NiO/NF	22	31.7	10
Ru-NiFe-P	44	80	11
Ru-MoS <sub>2</sub> /CC	41	114	12
Ru-NiCo-LDH	28	42	13
NiFe-MOF-5	163	139	14
Ru/NF-2	10	34	15
ECM@Ru	83	59	16
ES-Ru-ZIF-900	21	64	17
Ru/PC	21	46.6	18
Ru-RuO <sub>2</sub> /MoO <sub>3</sub> CNRs-350	9.2	37	19
Ru/Ni(OH) <sub>2</sub> /NF	25	47	20
Ru <sub>1.0</sub> /NF	47	60	21
Ru@NGT/Ni foam	45	81	22
Ru-Co <sub>3</sub> O <sub>4</sub> -NiO-NF	44	53.9	23
Ru-CoV-LDH/NF	32	36.4	24

**Supplementary Table 7.** Summarized overpotentials at 10 mA cm<sup>-2</sup> ( $\eta_{10}$ ) and Tafel slopes of electrocatalysts for OER in 1.0 M KOH.

Electrocatalysts	$\eta_{10}$ (mV)	Tafel slope (mV dec <sup>-1</sup> )	Ref
<b>FeCoNiMnRu/CNFs</b>	<b>145</b>	<b>61.3</b>	<b>This work</b>
FeCoNiPB/(FeCoNi) <sub>3</sub> O <sub>4-x</sub>	229	146	25
MnFeCoNi HEA	302	83.7	26
Fe <sub>29</sub> Co <sub>27</sub> Ni <sub>23</sub> Si <sub>9</sub> B <sub>12</sub> ribbon	230	85	27
CoCrFeNiMo-20 Mg	220	59.0	28
MCPS	288	27.7	29
La(CrMnFeCo <sub>2</sub> Ni)O <sub>3</sub> HEPO	325	51.2	30
CoCuFeMoOOH@Cu	199	48.8	31
FeNiCoCrMn-G	229	40	5
HEF7	292	39	32
AlCrCuFeNi	270	77.5	33
(Ru-Co)O <sub>x</sub>	171.2	60.8	7
Ru <sub>1</sub> Ni <sub>1</sub> -NCNFs	290	/	8
RuO <sub>2</sub> /NiO/NF	250	50.5	10
Ru-NiFe-MOF/NF	205	50	34
Ru <sub>SA</sub> CoFe <sub>2</sub> /G	180	51	35
NiFe-MOF-5	168	42	14
Ru/NF-2	330	62	15
Ru-CoV-LDH/NF	230	81.2	24
Ru-NiFe LDH-F/NF	230.0	50.2	36
Co(S)-Fe(S)	355	62	37
Ni:Pi-Fe/NF	220	37	38
Ni <sub>2</sub> P/(NiFe) <sub>2</sub> P(O) NAs	150	60	39
Cu <sub>0.81</sub> Co <sub>2.19</sub> O <sub>4</sub>	290	65.6	40
NiFe/NF	191	44.1	41

**Supplementary Table 8.** Summarized overpotentials at 10 mA cm<sup>-2</sup> ( $\eta_{10}$ ) and Tafel slopes of electrocatalysts for OWS in 1.0 M KOH.

Electrocatalysts	cell voltage @ 10 mA cm <sup>-2</sup>	Ref
<b>FeCoNiMnRu/CNFs</b>	<b>1.452</b>	<b>This work</b>
FeNiCoCrMn-G	1.63	5
(Ru-Co)O <sub>x</sub>	1.488	7
Ru <sub>1</sub> Ni <sub>1</sub> -NCNFs	1.564	8
RuO <sub>2</sub> /NiO/NF	1.5	10
Ru-NiFe-P	1.47	11
NiFe-MOF-5	1.57	14
Ru/NF-2	1.56	15
Ru/Ni(OH) <sub>2</sub> /NF	1.50	20
Ru-CoV-LDH/NF	1.50	24

**Supplementary Table 9.**  $\Delta G_{H^*}$  of Fe-, Co-, Ni-, Ru-sites in FeCoNiMnRu with various atomic configurations.

Structure	Site	H Free Energy (eV)
Configuration 1	Fe	-0.22954
	Ni	-0.23571
	Co	-0.25067
	Ru	-0.21156
Configuration 2	Fe	-0.20215
	Ni	-0.30751
	Co	-0.20661
	Ru	-0.11792
Configuration 3	Fe	-0.25743
	Ni	-0.18365
	Co	-0.21365
	Ru	-0.21036
Configuration 4	Fe	-0.13054
	Ni	-0.27281
	Co	-0.43812
	Ru	-0.07115
Configuration 5	Fe	-0.32153
	Ni	-0.34089
	Co	-0.26261
	Ru	-0.25932
Configuration 6	Fe	-0.31425
	Ni	-0.21751
	Co	-0.20561
	Ru	-0.206913
Configuration 7	Fe	-0.305413
	Ni	-0.215841
	Co	-0.236172
	Ru	-0.213653
Configuration 8	Fe	-0.31423
	Ni	-0.25361
	Co	-0.21454
	Ru	-0.20361

**Supplementary Table 10.**  $\Delta G_{\text{H}^*}$  of Fe-, Co-, Ni-, Ru-sites in FeCoNiCrRu with various atomic configurations.

Structure	Site	H Free Energy (eV)
Configuration 1	Fe	-0.29541
	Ni	-0.25816
	Co	-0.29615
	Ru	-0.26035
Configuration 2	Fe	-0.39654
	Ni	-0.33263
	Co	-0.31951
	Ru	-0.20592
Configuration 3	Fe	-0.32542
	Ni	-0.20654
	Co	-0.29234
	Ru	-0.26015
Configuration 4	Fe	-0.21879
	Ni	-0.34942
	Co	-0.60138
	Ru	-0.14727
Configuration 5	Fe	-0.30413
	Ni	-0.36581
	Co	-0.27156
	Ru	-0.26571
Configuration 6	Fe	-0.32474
	Ni	-0.25415
	Co	-0.26027
	Ru	-0.22941
Configuration 7	Fe	-0.341552
	Ni	-0.245913
	Co	-0.263512
	Ru	-0.289521
Configuration 8	Fe	-0.30355
	Ni	-0.26304
	Co	-0.20148
	Ru	-0.23651

**Supplementary Table 11.**  $\Delta G_{\text{H}^*}$  of Fe-, Co-, Ni-, Ru-sites in FeCoNiCuRu with various atomic configurations.

Structure	Site	H Free Energy (eV)
Configuration 1	Fe	-0.3136
	Ni	-0.2715
	Co	-0.2815
	Ru	-0.2715
Configuration 2	Fe	-0.32695
	Ni	-0.37994
	Co	-0.38255
	Ru	-0.26513
Configuration 3	Fe	-0.31361
	Ni	-0.27156
	Co	-0.28154
	Ru	-0.26556
Configuration 4	Fe	-0.29566
	Ni	-0.40929
	Co	-0.52791
	Ru	-0.23708
Configuration 5	Fe	-0.32551
	Ni	-0.32692
	Co	-0.29647
	Ru	-0.28691
Configuration 6	Fe	-0.33691
	Ni	-0.27571
	Co	-0.29617
	Ru	-0.33654
Configuration 7	Fe	-0.35691
	Ni	-0.29657
	Co	-0.27891
	Ru	-0.28952
Configuration 8	Fe	-0.36694
	Ni	-0.22656
	Co	-0.29651
	Ru	-0.26015

## Supplementary reference

1. Yao, R. et al. Nanoporous surface high-entropy alloys as highly efficient multisite electrocatalysts for nonacidic hydrogen evolution reaction. *Adv. Funct. Mater.* **31**, 2009613 (2021).
2. Jia, Z. et al. A novel multinary intermetallic as an active electrocatalyst for hydrogen evolution. *Adv. Mater.* **32**, 2000385 (2020).
3. Zhang, G. et al. High entropy alloy as a highly active and stable electrocatalyst for hydrogen evolution reaction. *Electrochim. Acta* **279**, 19-23 (2018).
4. Guo, J. et al. High electrocatalytic hydrogen evolution activity on a coupled Ru and CoO hybrid electrocatalyst. *J. Energy Chem.* **37**, 143-147 (2019).
5. Nguyen, T. X., Su, Y., Lin, C., Ruan, J. & Ting, J. A new high entropy glycerate for high performance oxygen evolution reaction. *Adv. Sci.* **8**, 2002446 (2021).
6. Jia, Z. et al. A self-supported high-entropy metallic glass with a nanospunge architecture for efficient hydrogen evolution under alkaline and acidic conditions. *Adv. Funct. Mater.* **2101586** (2021).
7. Wang, C. & Qi, L. Heterostructured inter-doped ruthenium–cobalt oxide hollow nanosheet arrays for highly efficient overall water splitting. *Angew. Chem. Int. Ed.* **59**, 1–7 (2020).
8. Li, M. et al. RuNi nanoparticles embedded in N-doped carbon nanofibers as a robust bifunctional catalyst for efficient overall water splitting. *Adv. Sci.* **7**, 1901833 (2020).
9. Xiao, X. et al. In situ growth of Ru nanoparticles on  $(\text{Fe}, \text{Ni})(\text{OH})_2$  to boost hydrogen evolution activity at high current density in alkaline media. *Small Methods* **4**, 1900796 (2020).
10. Liu, J. et al. NiO as a bifunctional promoter for  $\text{RuO}_2$  toward superior overall water splitting. *Small* **14**, 1704073 (2018).
11. Qu, M. et al. Regulating electron density of NiFe-P nanosheets electrocatalysts by a trifle of Ru for high-efficient overall water splitting. *Appl. Catal. B: Environ.* **263**, 118324 (2020).
12. Wang, D., Li, Q., Han, C., Xing, Z. & Yang, X. Single-atom ruthenium based catalyst for enhanced hydrogen evolution. *Appl. Catal. B: Environ.* **249**, 91-97 (2019).
13. Li, D. et al. Boosting hydrogen evolution activity in alkaline media with dispersed ruthenium clusters in NiCo-layered double hydroxide. *Electrochem. Commun.* **101**, 23-27 (2019).
14. Mou, Q. et al. A bimetal hierarchical layer structure MOF grown on Ni f foam as a bifunctional catalyst for the OER and HER. *Inorg. Chem. Front.* **8**, 2889 (2021).
15. Pei, Y. et al. Interface engineering with ultralow ruthenium loading for efficient water splitting. *ACS Appl. Mater. Interfaces* **12**, 36177–36185 (2020).
16. Zhang, H., Zhou, W., Lu, X. F., Chen, T. & Lou, X. W. Implanting isolated Ru atoms into edge-rich carbon matrix for efficient electrocatalytic hydrogen evolution. *Adv. Energy Mater.* **10**, 2000882 (2020).
17. Fan, L. et al. Ru nanoparticles encapsulated in ZIFs-derived porous N-doped hierarchical carbon nanofibers for enhanced hydrogen evolution reaction. *Catal. Sci. Technol.* **10**, 7302–7308 (2020).
18. Ding, R. et al. Salt template-assisted in situ construction of Ru nanoclusters and porous carbon: excellent catalysts toward hydrogen evolution, ammonia-borane hydrolysis, and 4-nitrophenol reduction. *Green Chem.* **22**, 835–842 (2020).
19. Fan, L. et al. Electrospun Ru– $\text{RuO}_2/\text{MoO}_3$  carbon nanorods with multi-active components: a Pt-like catalyst for the hydrogen evolution reaction. *Chem. Commun.* **56**, 739-742 (2020).
20. Chen, Q., Yang, X., Hou, C., Li, K. & Chen, Y. Inlay of ultrafine Ru nanoparticles into a self-supported  $\text{Ni}(\text{OH})_2$  nanoarray for hydrogen evolution with low overpotential and enhanced kinetics. *J. Mater. Chem. A* **7**, 11062–11068 (2019).
21. Xia, J. et al. Low-cost porous ruthenium layer deposited on Ni foam as a highly active universal-pH electrocatalyst for hydrogen evolution reaction. *ChemSusChem* **12**, 2780-2787 (2019).
22. Barman, B. K. et al. In situ decoration of ultrafine Ru nanocrystals on N-doped graphene tube and their applications as oxygen reduction and hydrogen evolution catalyst. *ACS Appl. Energy Mater.* **2**, 7330–7339 (2019).
23. Zhang, H. et al. Synergistic engineering of morphology and electronic structure in constructing metal-organic framework-derived Ru doped cobalt-nickel oxide heterostructure towards efficient alkaline hydrogen evolution reaction. *Chem. Eng. J.*

- 426, 131300 (2021).
24. Li, W., Feng, B., Yi, L., Li, J. & Hu, W. Highly efficient alkaline water splitting with Ru-doped Co-V layered double hydroxide nanosheets as bifunctional electrocatalyst. *ChemSusChem* **14**, 730-737 (2021).
25. Wei, R. et al. Defect-rich FeCoNiPB/(FeCoNi)<sub>3</sub>O<sub>4-x</sub> high-entropy composite nanoparticles for oxygen evolution reaction: impact of surface activation. *Appl. Surf. Sci.* **549**, 149327 (2021).
26. Dai, W., Lu, T. & Pan, Y. Novel and promising electrocatalyst for oxygen evolution reaction based on MnFeCoNi high entropy alloy. *J. Power Sources* **430**, 104-111 (2019).
27. Wang, H. et al. Nanostructured amorphous Fe<sub>29</sub>Co<sub>27</sub>Ni<sub>23</sub>Si<sub>9</sub>B<sub>12</sub> high-entropy-alloy: an efficient electrocatalyst for oxygen evolution reaction. *J. Mater. Sci. Technol.* **68**, 191-198 (2021).
28. Tang, J., Xu, J. L., Ye, Z. G., Li, X. B. & Luo, J. M. Microwave sintered porous CoCrFeNiMo high entropy alloy as an efficient electrocatalyst for alkaline oxygen evolution reaction. *J. Mater. Sci. Technol.* **79**, 171-177 (2021).
29. Ding, Z. et al. High entropy intermetallic–oxide core–shell nanostructure as superb oxygen evolution reaction catalyst. *Adv. Sustainable Syst.* **4**, 1900105 (2020).
30. Nguyen, T. X., Liao, Y., Lin, C., Su, Y. & Ting, J. Advanced high entropy perovskite oxide electrocatalyst for oxygen evolution reaction. *Adv. Funct. Mater.* **31**, 2101632 (2021).
31. Zhang, L., Cai, W. & Bao, N. Top-level design strategy to construct an advanced high-entropy Co–Cu–Fe–Mo (oxy)hydroxide electrocatalyst for the oxygen evolution reaction. *Adv. Mater.* **33**, 2100745 (2021).
32. Sukkurji, P. A. et al. Mechanochemical synthesis of novel rutile-type high entropy fluorides for electrocatalysis. *J. Mater. Chem. A* **9**, 8998–9009 (2021).
33. Liu, L., Li, N., Han, M., Han, J. & Liang, H. Scalable synthesis of nanoporous high entropy alloys for electrocatalytic oxygen evolution. *Rare Met.* (2021).
34. Lin, Y., Zhao, L., Wang, L. & Gong, Y. Ruthenium-doped NiFe-based metal–organic framework nanoparticles as highly efficient catalysts for the oxygen evolution reaction. *Dalton Trans.* **50**, 4280–4287 (2021).
35. Lee, J. et al. Stabilizing the OOH\* intermediate via pre-adsorbed surface oxygen of a single Ru atom-bimetallic alloy for ultralow overpotential oxygen generation. *Energy Environ. Sci.* **13**, 5152-5164 (2020).
36. Wang, Y. et al. Interfacial synergy between dispersed Ru sub-nanoclusters and porous NiFe layered double hydroxide on accelerated overall water splitting by intermediate modulation. *Nanoscale* **12**, 9669–9679 (2020).
37. Maccato, C. et al. Plasma-assisted synthesis of Co<sub>3</sub>O<sub>4</sub>-based electrocatalysts on Ni foam substrates for the oxygen evolution reaction. *Adv. Mater. Interfaces* **8**, 2100763 (2021).
38. Li, Y. & Zhao, C. Iron-doped nickel phosphate as synergistic electrocatalyst for water oxidation. *Chem. Mater.* **28**, 5659–5666 (2016).
39. Xi, W. et al. Oxygen-doped nickel iron phosphide nanocube arrays grown on Ni foam for oxygen evolution electrocatalysis. *Small* **14**, 1802204 (2018).
40. Choi, W. et al. Three-dimensional honeycomb-like Cu<sub>0.81</sub>Co<sub>2.19</sub>O<sub>4</sub> nanosheet arrays supported by Ni foam and their high efficiency as oxygen evolution electrodes. *ACS Appl. Mater. Interfaces* **10**, 38663–38668 (2018).
41. Jin, J. et al. Exceptional electrocatalytic oxygen evolution efficiency and stability from electrodeposited NiFe alloy on Ni foam. *Electrochim. Acta* **299**, 567-574 (2019).