

# ***Supplemental Information for: Water Electrolysis: From textbook knowledge to the latest scientific strategies and industrial developments***

Marian Chatenet<sup>a</sup>, Bruno G. Pollet<sup>b, c</sup>, Dario R. Dekel<sup>d, e</sup>, Fabio Dionigi<sup>f</sup>, Jonathan Deseure<sup>a</sup>, Pierre Millet<sup>g</sup>,

Richard D. Braatz<sup>h</sup>, Martin Z. Bazant<sup>h, i</sup>, Michael Eikerling<sup>j, k</sup>, Iain Staffell<sup>l</sup>,

Paul Balcombe<sup>m</sup>, Yang Shao-Horn<sup>n</sup>, and Helmut Schäfer<sup>o\*</sup>

## **Address:**

a: University Grenoble Alpes, CNRS, Grenoble-INP (Institute of Engineering, Univ. Grenoble Alpes), LEPMI 38000 Grenoble, France. [Orcid.org/0000-0002-9673-4775](https://orcid.org/0000-0002-9673-4775) (Marian Chatenet). [Orcid.org/0000-0002-2055-7653](https://orcid.org/0000-0002-2055-7653) (Jonathan Deseure).

b: Hydrogen Energy and Sonochemistry Research group, Department of Energy and Process Engineering, Faculty of Engineering, Norwegian University of Science and Technology (NTNU) NO-7491, Trondheim, Norway. [Orcid.org/0000-0002-4928-7378](https://orcid.org/0000-0002-4928-7378) (Bruno G. Pollet).

c: Green Hydrogen Lab, Hydrogen Research Institute (HRI), Université Du Québec à Trois-Rivières (UQTR), 3351 Boulevard des Forges, Trois-Rivières, Québec G9A 5H7, Canada.

d: The Wolfson Department of Chemical Engineering, Technion – Israel Institute of Technology, Haifa, 3200003, Israel. [Orcid.org/0000-0002-8610-0808](https://orcid.org/0000-0002-8610-0808) (Dario Dekel).

e: The Nancy & Stephen Grand Technion Energy Program (GTEP), Technion - Israel Institute of Technology, Haifa 3200003, Israel

f: Department of Chemistry, Chemical Engineering Division, Technical University Berlin, 10623, Berlin, Germany. [Orcid.org/0000-0002-0576-024X](https://orcid.org/0000-0002-0576-024X) (Fabio Dionigi)

g: Institut de Chimie Moléculaire et des Matériaux d'Orsay (UMR 8182), University of Paris-Saclay, rue du Doyen Georges Poitou, bâtiment 410, 91400, Orsay, France. [Orcid.org/0000-0002-0224-9868](https://orcid.org/0000-0002-0224-9868) (Pierre Millet).

h: Department of Chemical Engineering, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139, United States. [Orcid.org/0000-0003-4304-3484](https://orcid.org/0000-0003-4304-3484) (Richard D. Braatz). [Orcid.org/0000-0002-8200-4501](https://orcid.org/0000-0002-8200-4501) (Martin Z. Bazant).

i: Department of Mathematics, Massachusetts Institute of Technology, 77 Massachusetts Avenue, Cambridge, Massachusetts 02139, United States.

j: Chair of Theory and Computation of Energy Materials, Division of Materials Science and Engineering, RWTH Aachen University, Intzestraße 5, 52072 Aachen, Germany [Orcid.org/0000-0002-0764-8948](https://orcid.org/0000-0002-0764-8948) (Michael Eikerling).

k: Institute of Energy and Climate Research, IEK-13: Modelling and Simulation of Materials in Energy Technology, Forschungszentrum Jülich GmbH, 52425 Jülich, Germany.

l: Centre for Environmental Policy, Imperial College London, London, UK. [Orcid.org/0000-0003-1012-7075](https://orcid.org/0000-0003-1012-7075) (Iain Staffell).

m: Division of Chemical Engineering and Renewable Energy, School of Engineering and Material Science, Queen Mary University of London, London, UK. [Orcid.org/0000-0002-3490-0707](https://orcid.org/0000-0002-3490-0707) (Paul Balcombe).

n: Research Laboratory of Electronics and Department of Mechanical Engineering, Massachusetts Institute of Technology, Cambridge, Massachusetts 02139, United States. [Orcid.org/0000-0001-8714-2121](https://orcid.org/0000-0001-8714-2121) (Yang Shao-Horn).

o: Institute of Chemistry of New Materials, The Electrochemical Energy and Catalysis Group, University of Osnabrück, Barbarastrasse 7, 49076 Osnabrück, Germany. [Orcid.org/0000-0001-5906-3354](https://orcid.org/0000-0001-5906-3354) (Helmut Schäfer).

**Email:** [helmut.schaefer@uos.de](mailto:helmut.schaefer@uos.de)

**Table S1.** Chemical and mechanical properties of selected AEMs.

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
PTP-75	[(terphenyl piperidinium)-co-(oxindole terphenylylene)]	Piperidinium	2.10	48.6	31.4	4.9 <sup>b</sup>	45	29.2	6.8	466.7	20	+	1
PTP-85	[(terphenyl piperidinium)-co-(oxindole terphenylylene)]	Piperidinium	2.38	59.6	37.0	5.9 <sup>b</sup>	45	34.2	24.2	620.2	20	+	1
PTP-90	[(terphenyl piperidinium)-co-(oxindole terphenylylene)]	Piperidinium	2.52	64.4	39.7	7.4	45	36.5	25.2	777.6	20	+	1
FAA3-30 (OH-) (FumaTech) <sup>b</sup>	N/A	N/A	1.7-2.1	40	N/A	N/A	25-35	25-40	20-40	N/A	N/A	-	2
FAA3-50 (OH-) (FumaTech) <sup>b</sup>	N/A	N/A	1.59, 1.85	40	17	9	47-53	25-40	20-40	N/A	N/A	+	2, 3
QMSV-0.16	poly(ST-co-VBC	TMA	1.04	0.264	16.7	10.9	60	N/A	N/A	N/A	25	+	4
QMSV-0.33	poly(ST-co-VBC	TMA	2.14	6.80	127.4	15.9	60	N/A	N/A	N/A	25	+	4
PFOTFPh-TMA-C6	(fluorene-alt-tetrafluorophenylene	trimethylammonium (TMA)	3.2	156	122	97	20-30	26.8	8	N/A	70	+	5, 6
PFOTFPh-TMA-C8	(fluorene-alt-tetrafluorophenylene	trimethylammonium (TMA)	2.9	117	74 <sup>d</sup>	39	20-30	28.5	8	N/A	70	+	5, 6
PFOTFPh-TMA-C10	(fluorene-alt-tetrafluorophenylene	trimethylammonium (TMA)	2.7	101	67 <sup>d</sup>	33	20-30	40.9	10	N/A	70	+	5, 6
GT69 <sup>c</sup>	poly (norbornene) (PNB)	N,N,N',N'- tetramethyl-1,6-hexanediamine (TMHDA),	3.38	178	115	N/A	N/A	N/A	N/A	N/A	80	+	7
GT72-5 <sup>c</sup>	poly (norbornene) (PNB)	N,N,N',N'- tetramethyl-1,6-hexanediamine (TMHDA),	3.50	175	96	N/A	N/A	N/A	N/A	N/A	80	+	7
GT74 <sup>c</sup>	poly (norbornene) (PNB)	N,N,N',N'- tetramethyl-1,6-hexanediamine (TMHDA),	3.56	160	103	N/A	35 50	N/A	N/A	N/A	80	+	7
PAP-TP-85	poly(aryl piperidinium) hydroxide	piperidinium	2.37	193	65	12	10-25	67	117	N/A	95	+	8, 9, 10
PAP-TP-85-MQN	poly(aryl piperidinium) hydroxide	piperidinium	3.2	150 <sup>d</sup>	N/A	N/A	20	N/A	N/A	N/A	RT	+	8
PBI1-PVBC1-NMPD/OH	poly(vinylbenzyl chloride) (PVBC) cross-linked by	N-methylpiperidine	2.31	83	48	11	50	37.5 <sup>e</sup>	16.2 <sup>e</sup>	N/A	80	+	11

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
	polybenzimidazole (PBI) and quaternized by N-methylpiperidine (NMPD) <sup>i</sup>												
PAni-0.87	Polyaniline	3,30 - iminobis(N,Ndimethylpropylamine)	0.87	54.9	13.8	10.3		N/A	N/A	N/A	80	-	12
PAni-0.92	Polyaniline	3,30 - iminobis(N,Ndimethylpropylamine)	0.92	90.1	18.1	11.7		N/A	N/A	N/A	80	-	12
PAni-1.03	Polyaniline	3,30 - iminobis(N,Ndimethylpropylamine)	1.03	105	22.3	18.7	N/A	N/A	N/A	N/A	80	+	12
IPA <sup>b</sup>	N/A	N/A	1.27	32.4	23.1	15.6	N/A	N/A	N/A	N/A	80	+	12
AMB <sup>b</sup>	N/A	N/A	1.06	3.3	24.3	N/A	N/A	N/A	N/A	N/A	80	+	12
Tokuyama- A201 <sup>b</sup>	N/A	N/A	1.8	42	30	N/A	28 <sup>c</sup>	96.4 <sup>g</sup>	61.7 <sup>g</sup>	1123 <sup>g</sup>	(26,41)	+	2, 10, 13, 14
AEMION™-AF1-HNN8-50-X <sup>b</sup>	N/A	N/A	2.1-2.5	102	N/A	N/A	50	60 (dry I)	85-110 <sup>j</sup>	N/A	50	+	2, 15
AEMION™-AF1-HNN8-25-X <sup>b</sup>	N/A	N/A	2.1-2.5	131	N/A	N/A	25	60 (dry I)	85-110 <sup>g</sup>	N/A	50	+	2, 15
AEMION™-AF1-HNN5-50-X <sup>b</sup>	N/A	N/A	1.4-1.7	40	N/A	N/A	50	60 (dry I)	85-110 <sup>g</sup>	N/A	50	+	2
AF1-HNN5-25-X <sup>b</sup>	N/A	N/A	1.4-1.7	56	N/A	N/A	25	60 (dry I)	85 <sup>g</sup>	N/A	50	+	2, 15
SUSTAINION® Sustainion 37-50 <sup>b</sup>	Copolymer of styrene and vinyl benzyl chloride	Tetramethyl Imidazole	N/A	80 <sup>f</sup>	N/A	cracks when dry	50	cracks when dry	cracks when dry	N/A	30	+	2, 19, 16, 17, 18, 19, 20, 21, 22, 23
HDPE-AEM	High density polyethylene	TMA	2.44	214 <sup>g,i</sup>	155 <sup>g</sup>	38	29	35	283	N/A	80	-	24, 25
LDPE-AEM	Low density polyethylene	TMA	2.54	290 201 304	149	27	28	23	69	N/A	110 105 120	-	25, 24, 26, 27, 28, 29
C4-AEM	ETFE	pyrrolidinium	1.51	30	85	59	N/A	N/A	N/A	N/A	60	-	25, 30
SEBS-BTMA	polystyrene-b-poly(ethylene-co-butylene)-b-polystyrene (SEBS)	TMA	1.04	32	40	25.1	60	2.0	180.1	N/A	80	-	25, 31

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
SEBS-CH2-QA-1.5	SEBS	Long flexible alkyl spacer	1.35	56	48	29.6	60	3.3	345.7	N/A	80	-	25, 31
XL100-SEBS-C5-TMA-0.8	SEBS	TMA with different degree of functionalization and cation tether length	1.50	65	28	10	60	7.0	400	N/A	80	-	25, 32
1:1 DCPD:1	a tetraalkylammonium-functionalized norbornene with dicyclopentadiene (DCPD)	tetraalkylammonium	1.40	28	N/A	N/A	N/A	2.3	26	N/A	50	-	25, 33
P1.25-OH	poly(norbornene)	pendant trimethylammonium cations	1.25	177	82	N/A	N/A	N/A	N/A	N/A	80		25, 34
XL35-rPNB-X60-Y40	Poly(bromopropyl norbornene)-blockpoly(butyl norbornene) diblock copolymers	TMHDA	2.20	109	100	28	N/A	2.5	52	N/A	80	-	25, 35
HC-[1] <sub>498</sub> [2] <sub>200</sub>	PE	Imidazolium	1.69	134	115	17	N/A	N/A	N/A	N/A	80	-	25, 36
PNB-X62-Y38	The tetrablock copolymer, consisting of alternating butyl norbornene (BuNB) and bromopropyl norbornene (BPNB) blocks (two blocks each)	norbornene	2.21	102	71	N/A	50	N/A	N/A	N/A	80	-	25, 37
H22C9N	poly(olefin)s	trimethylamine with the alkyl group	1.43	70 <sup>1</sup>	177	40	N/A	N/A	N/A	N/A	80	-	25, 38
F20C9N	poly(olefin)s	trimethylamine with the alkyl group	1.21	91 <sup>1</sup>	109	26	70	N/A	N/A	N/A	80	-	25, 38
ATMPP	poly(phenylene)s	benzylic cations	2.39	18	156	N/A	N/A	N/A	N/A	N/A	22	-	25, 39
HTMA-DAPP	polyphenylene	hexamethyl trimethyl ammonium	2.6	120	58	N/A	26	>20	N/A	N/A	80	+	25, 40
QPAF (C6)-2	perfluoroalkylene and phenylene groups	TMA	1.14	96	45	N/A	50	24	218	N/A	80	-	25, 41
QP-QAF3	quinquephenylene and fluorene groups	pendant hexyltrimethylammonium	2.21	134	85	N/A	22	35	28	6	80	-	25, 42
PAImEE (12)	poly(arylimidazoliums)	Ethyl as alkyl chains	2.65	46 <sup>g</sup>	28 <sup>j</sup>	26.1	25	64	28.7	1075	80	+	25, 43
BPN1	biphenyl	trifluoromethyl	2.61	122	130	40	25	35	140	N/A	80	-	25, 44, 45
<i>p</i> -TPN1	Para- terphenyl	trifluoromethyl	2.15	81	43	6	25	24	20	N/A	80	-	25, 44, 45
<i>m</i> -TPN1	meta-terphenyl	trifluoromethyl	2.15	127	45	10	25	30	38	N/A	80	-	25, 44, 45
FLN-55	Quaternized poly(fluorene)s	TMA	2.50	120	180	60	30	N/A	N/A	N/A	80	-	25

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
PTPipQ6	poly(arylene)	piperidinium	2.04	111	44	N/A	60	N/A	N/A	N/A	80	-	25, 46
PTPipQ1	poly(arylene)	piperidinium	2.38	89	145	N/A	60	N/A	N/A	N/A	80	-	46
PVBC-MPy/15%PEK-cardo	poly(vinyl benzyl chloride) (PVBC-MPy) and poly(ether ketone-cardo)	methylpyrrolidinium	3.47	50.3	132	110	60	4.5	31.3	20.7	80	-	47
PVBC-MPy/25%PEK-cardo	(PVBC-MPy) and poly(ether ketone-cardo)	methylpyrrolidinium	3.10	37.7	59	61	60	9.5	33.3	30.9	80	-	47
PVBC-MPy/35%PEK-cardo	(PVBC-MPy) and poly(ether ketone-cardo)	methylpyrrolidinium	2.65	28.5	32	38	60	15.1	23.9	67.6	80	+	47
PVBC-MPy/45%PEK-cardo	(PVBC-MPy) and poly(ether ketone-cardo)	methylpyrrolidinium	2.34	15.4	24	26	60	22.1	19.3	110	80	-	47
PSF-TMA <sup>+</sup>	polysulfone (PSF)	trimethylammonium	2.05	30.5	N/A	N/A	40-80	N/A	N/A	N/A	50	+	48, 49
PSF-DMP <sup>+</sup>	polysulfone (PSF)	1,4-dimethylpiperazinium	1.51	14.4	230	N/A	40-80	N/A	N/A	N/A	50	-	49
xQAPS	PSF	Trimethylammonium Dimethyl diethyl ammonium as cross linker	1.34	60.5	N/A	4	N/A	N/A	N/A	N/A	80	+	50, 51, 52
PPO24-BIM (bromide form)	oly-(2,6-dimethyl-1,4-phenylene oxide)	mesityl-benzimidazole (BIM)	1.9	12	27	N/A	23	45.9	6.0	951	60	+	13
FAA3-PK-75 (bromide form) <sup>b</sup>	N/A	N/A	N/A	N/A	N/A	N/A	87	26.9	12.1	983	60	+	13, 53
PSEBS-CM-DABCO	polystyrene-block-poly(ethylene-ran-butylene)-block-polystyrene (PSEBS)	1,4-diazabicyclo[2.2.2]octane (DABCO)	0.76	75	N/A	N/A	100	N/A	N/A	N/A	30	+	54
PAEK-APMP75	Poly (arylene ether) ketone	1-(3-aminopropyl)-4-methylpiperazin	1.10	6.82	38	N/A	55	12.6	30.4	N/A	60	+	55
PAEK-APMP100	Poly (arylene ether) ketone	1-(3-aminopropyl)-4-methylpiperazin	1.32	9.94	48	N/A	55	9.7	41	N/A	60	+	55
QPDTB	three monomers of methacrylate:2-dimethylaminoethyl methacrylate (DMAEMA), 2,2,2-trifluoroethyl methacrylate (TFEMA), and butyl methacrylate (BMA)	Trimethyl amine	1.275	59	N/A	N/A	300	7.629	45.8	226	50	+	56, 57

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
LDPE-g-VBC-Dab	Low density polyethylene (LDPE) grafted vinylbenzyl chloride (VBC)	,4-diazabicyclo(2.2.2)octane (Dabco)	1.5	25	81	N/A	N/A	N/A	N/A	N/A	60	+	58
SEBS-Pi-73%	Polystyrene-block-poly(ethylene-ran-butylene)-block-polystyrene	N-Methylpyridine	1.19	10.09	59.13	25.03	N/A	4.2	600	N/A	30	+	59
PBP-ASU-PPO	poly(biphenyl piperidinium) (PBP)/6-azaspiro[5.5]undecane (ASU)-functionalized polyphenyl ether (ASU-PP)	Piperidinium and ASU	2.61	128	120	21	50	N/A	N/A	N/A	80	+	60
Tokuyama- A901 <sub>b</sub>	N/A	N/A	1.8	N/A	N/A	N/A	9	N/A	N/A	N/A	N/A	+	61
FAA-3-PP-75	N/A	N/A	N/A	38	140	13-15	80	N/A	N/A	N/A	60	+	62
SEBS	polystyrene-b-poly(ethylene/butylene)-b-polystyrene	Trimethyl amine	1.9	140	N/A	N/A	120	N/A	N/A	N/A	50	+	63
PTFE+qPDTB-OH	Polytetrafluoroethylene (PTFE)	Quaternary ammonium Poly(DMAEMA-co-TFEMA-co-BMA)  (quaternary ammonium polymethacrylate)	1.02	34	N/A	146	30	10	10	406	50	+	64
mm-qPVBz/OH	Methylated melamine grafted poly vinyl benzylchloride (mm-qPVBz/Cl)	Amination with methylated melamine	N/A	27	N/A	N/A	70	12.1	14	142	60	+	65, 66
BNP1-100	poly(arylene)	N,N,N-trimethylpentan-1-ammonium	2.61	122	124	66	N/A	N/A	N/A	N/A	80	+	67
TPN1-100	poly(arylene)	N,N,N-trimethylpentan-1-ammonium	2.15	112	70	23	N/A	N/A	N/A	N/A	80	+	67
PAImEE(12)	poly(aryl)	Imidazolium	2.65	21.3	28.1	26.1	25	64.0	28.7	1075	22	+	43
PAImBB(14)	poly(aryl)	Imidazolium	2.3	8.5	12.2	14.0	20	65.2	20.4	1095	22	+	43
QPC-TMA	poly(carbazole)	polymer (poly(9-(6-(trimethylammonium bromide)hexyl)-9H-carbazole-co-1,1,1-trifluoroisopropane)	2.08 2.00	125	76	15	50	N/A	N/A	N/A	70	+	68
PVBC-MPy M2	poly(vinyl benzyl)	methylpyrrolidinium	2.02	32.8	28.5	23.8	80	9.4	16.5	N/A	80	+	69
PVBC-MPy M4	poly(vinyl benzyl)	methylpyrrolidinium	2.01	29.8	19.8	20.4	80	30.9	11.1	N/A	80	+	69

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
PVBC-MPy M6	poly(vinyl benzyl)	methylpyrrolidinium	1.97	30.5	21.5	17.0	80	20.6	8.9	N/A	80	+	69
MES-PBI (25 wt% KOH)	polybenzimidazole	poly(2,2'-(m-mesitylene)-5,5'-bibenzimidazole)	N/A	100	N/A	54	60	N/A	N/A	N/A	80	+	70
PPO-TMA	polyphenylene oxide (PPO)	trimethylamine (TMA)	2.1	52	104	N/A	N/A	6.3	1.9	N/A	70	+	71, 72
PPO-ABCO	polyphenylene oxide (PPO)	1-Azabicyclo[2.2.2]octane (ABCO)	1.9	39	147	N/A	N/A	10.3	2	N/A	70	+	71, 72
PISPVA46	poly(1-vinyl-3-imidazole-co-styrene) (PIS) co-poly(vinyl alcohol) (PVA)	imidazolium	1.65	90	101.1	18.6	55	15.5	316.5	N/A	60	+	73
PISPVA37	poly(1-vinyl-3-imidazole-co-styrene) (PIS) co-poly(vinyl alcohol) (PVA)	imidazolium	1.41	82	78.9	19.7	55	27.6	295.1	N/A	60	+	73
PISPVA28	poly(1-vinyl-3-imidazole-co-styrene) (PIS) co-poly(vinyl alcohol) (PVA)	imidazolium	1.13	74	49.7	18.7	55	30.8	216.9	N/A	60	+	73
QMter-co-Mpi-60%	ether-free polyarylene	piperidinium	1.65	8	37.15	11.89	N/A	N/A	N/A	N/A	30	-	74
QMter-co-Mpi-80%	ether-free polyarylene	piperidinium	2.10	21	49.39	17.77	N/A	N/A	N/A	N/A	30	-	74
QMter-co-Mpi-100%	ether-free polyarylene	piperidinium	2.42	37	78.92	29.00	N/A	11	24	N/A	30	+	74
BPi	PPO	piperidinium	1.94	18	29.0	9.5	50	32.7	2.7	N/A	20	-	75
SCPi	PPO	piperidinium	1.67	25	41.7	11	50	48.4	3.5	N/A	20	+	75
LSCPi	PPO	piperidinium	1.57	29	39.6	10.3	50	35.6	2.0	N/A	20	+	75
BTMA	PPO	Benzyltrimethyl ammonium	2.04	26	65.7	12	50	39.4	4.4	N/A	20	-	75
SCQA	PPO	side-chain-type	1.80	41	79.0	15	50	40.5	8.0	N/A	20	-	75
LSCQA	PPO	long side-chain-type	1.67	39	55.7	12	50	39.0	2.1	N/A	20	+	75
PSU-PVP75%	poly(arylene ether sulfone) and poly(vinylpyrrolidone)	pyrrolidone	N/A	N/A	N/A	N/A	A wide range	5	75	130	N/A	+	76
BPI-c-PVBC/OH 1:2	poly[2-2'-(m-phenylene)-5-5'-bibenzimidazole] co-poly(vinylbenzyl chloride) (PVBC)	DABCO	1.74	30	45	23	N/A	N/A	N/A	N/A	80	+	77



Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
BPI-c-PVBC/OH 1:3	poly[2-2'-(m-phenylene)-5-5'-bibenzimidazole] copoly(vinylbenzyl chloride) (PVBC)	DABCO	1.97	44	52	24	N/A	N/A	N/A	N/A	80	+	77
ABPBI-c-PVBC/OH 1:2	Poly(2,5-benzimidazole) (ABPBI) co - poly(vinylbenzyl chloride) (PVBC)	DABCO	1.70	56	51	90	N/A	N/A	N/A	N/A	90	+	77
C-PVaf-ABPBI	Polyvinyl alcohol (PVA) nanofiber/ABPBI	polyvinyl alcohol (PVA) nanofibers crosslinked with glutaraldehyde (GA)	N/A	41	55	17	30-45	5.0	6	1033	80	+	78
C-PVA-ABPBI	Polyvinyl alcohol (PVA) /ABPBI	polyvinyl alcohol (PVA) nanofibers crosslinked with glutaraldehyde (GA)	N/A	48	67	14	30-45	2.8	16	64	80	-	78
QPAF-4	perfluoroalkylene and fluorene	hexyltrimethylammonium groups	1.47	86.2	105	N/A	50	22.7	269	N/A	80	-	79
PP-BTMA	Aryl-Ether	Polyaromatics	2	10	116	N/A	20-40	N/A	N/A	N/A	30	-	80
PP-HTMA	Aryl-Ether	Polyaromatics	2.4	14	109	N/A	20-40	N/A	N/A	N/A	30	-	80
PPA-HTMA	Aryl-Ether	Polyaromatics	2.1	15	453	N/A	20-40	N/A	N/A	N/A	30	-	80
PSBFP-TMA	poly(2,2' - spirobifluorene-alt-1,3-phenylene)	Trimethyl ammonium	1.2	23.1	11.1	17	N/A	N/A	N/A	N/A	70	-	81
PSBFBP-TMA	poly(2,20 - spirobifluorene-alt-4,4' - biphenylene)	Trimethyl ammonium	2.3	86.2	24.5	38	N/A	N/A	N/A	N/A	70	-	81
F-PAE	polyaromatics	BTMA	2.7	46	99	N/A	N/A	30	9.8	N/A	80	+	82
ATM-PP	polyaromatics	BTMA	1.7	37	70	N/A	N/A	42	27	N/A	80	+	82
PBI linear	Polybenzimidazol	Polybenzimidazol	N/A	50	N/A	19	N/A	94	74	2554	22	+	83
CL PBI (crosslinked)	Polybenzimidazol	Polybenzimidazol	N/A	50	N/A	N/A	N/A	36	1.7	2907	22	+	83
Thermally cured PBI	Polybenzimidazol	Polybenzimidazol	N/A	50	N/A	47	N/A	139	59	2914	22	+	83

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
L-ABPBI (C <sub>KOH</sub> = 1.9 M)	ABPBI	ABPBI	N/A	7	37.8	50	90-120	N/A	N/A	N/A	22	+	84
LC-ABPBI (C <sub>KOH</sub> = 1.9 M)	ABPBI	ABPBI	N/A	8	33.8	50	90-120	N/A	N/A	N/A	22	+	84
PQDP-1	poly (quinquephenylene-co-diphenylene)	piperidinium	1.77	67.2	33	4.2	20-30	43	9.8	880	80	-	85
PQDP-3	poly (quinquephenylene-co-diphenylene)	piperidinium	2.44	140.5	90	36.5	20-30	41	16.6	980	80	-	85
SDQEO	PPO	alkoxyl- extender-containing dual quaternary ammonium	1.13	87.3	120	30	50	10.5	27.3	N/A	80	-	86
PFTP	poly(fluorenyl aryl piperidinium)	piperidinium	2.81	208	45	16	20	84.5	25.6	1580	80	-	87
m-PTP-20Q	oly(aryl piperidinium) (PAP)	piperidinium	3.06	144.2	51	6.25	60-70	N/A	N/A	N/A	80	-	88
O-PDQA-3	aryl-ether free poly(aryl piperidinium) (PBP)	piperidinium cations and ethylene oxide spacers	1.93	106	46.1	12.5	N/A	N/A	N/A	N/A	80	-	89
PP80N20	tetrakis(bromomethyl) monomers	spirocyclic QA	3.2	51.3	266	11	N/A	N/A	N/A	N/A	80	-	90
Q-CLP1	poly(benzimidazolium-imide)-	triazolium	1.36	84	58	17	N/A	15.37	7.80	N/A	80	-	91
BiPyBPPEEK-50%	polyetheretherketone (PEEK)	pyridine	3.51	36.99	16	4.4	N/A	66.0	4.5	2400	80	-	92
QPAES/QBGO-3.0	poly (arylene ether sulfone) (QPAES)	1,4-diazabicyclo [2,2,2]octane and 1,6-dibromohexane, and subsequently used to prepare multi-cationic oligomer brushes-decorated graphene oxide	1.68	58.7	103	13.6	60	32.4	9.1	1360	80	-	93
PAEK-HQACz-0.7	poly(arylene ether ketone) copolymers	Long alkyl densely quaternized carbazole derivative pendant	1.88	98.1	46.4	13.5	N/A	32.5	43.5	N/A	80	-	94
PBP-BOP-ASU 8%	poly(4-((1,1'-biphenyl)-4-yl)piperidine)(PBP)	long-chain 3-(3-(1-(8-bromooctyl) piperidin-4-yl) propyl)- 6-azaspiro[5.5] undecan-6-ium bromide(BOP-ASU)	2.65	117	140	32	N/A	N/A	N/A	N/A	80	-	95
PBP-ASU	poly(4-((1,1'-biphenyl)-4-yl)piperidine)(PBP)	ASU	2.82	91	132.98	26	N/A	N/A	N/A	N/A	80	-	95

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
CP2	poly(styrene-b-(ethylene-co-butylene)-b-styrene) (SEBS) poly(styrene-b-(ethylene-co-butylene)-b-styrene) (SEBS) poly(styrene-b-(ethylene-co-butylene)-b-styrene) (SEBS) Poly(styrene-b-(ethylene-co-butylene)-b-styrene) (SEBS)	branch polyethyleneimine (BPEI)	3.34	66.63	18.16	5.5	N/A	26.17	443	N/A	80	-	96
CP3	SEBS	BPEI	2.47	19.87	15.67	4.89	N/A	28.78	511	N/A	80	-	96
PPEEK:PEG 80:20	poly(ether ether ketone) (QPPEEK) and poly(ethylene glycol) (PEG) as the crosslinker	quatarnary phosphonium	1.01	102	139	20	100	13.7	48	338	80	-	97
MPyPPO	PPO	N-methylpyrrolidinium	1.73	84	92.9	23.9	N/A	15.52	25.45	N/A	80	-	98
qPBPTT-5	Poly(biphenyl N-methyl-4-piperidone 1,1,1-Trifluoroacetone 1,3,5-Triphenylbenzene)	piperidine	1.93	116.92	61.71	19.41	30	10.89	8.26	N/A	80	-	99
m-TPNPiQA	poly(terphenylene) backbone	tethered with piperidinium groups	2.66	68.7	2.54	52.25	50	N/A	N/A	N/A	80	-	100
HyAEM-MP-180	polychloromethylstyrene-b-polyethylene-b-polychloromethylstyrene (PCMS-b-PE-b-PCMS)	benzylmethylpiperidinium	2.1	179	33	16.5	N/A	23.9	29.4	39.0	80	-	101
PPO-DMP	PPO	six-membered dimethyl piperidinium	1.98	71.8	125.7	29.6	50	43.7	2.7	N/A	80	-	102
PPO-ASU	PPO	ASU	1.85	76.5	148.6	42.0	50	20.4	1.3	N/A	80	-	102
NPPO-2QA-1.85	Azide-modified PPO (NPPO)	Alkyne side chain precursor containing terminal doubleQA groups (TABB)	1.85	47.22	33.70	8.93	N/A	N/A	N/A	N/A	30	-	103
AEM-9.09	benzonorbornadiene derivative (BenzoNBD-Bis(ImBr <sup>+</sup> -Im <sup>+</sup> )) grafted	multi- imidazolium cations side-chains combined the rigid alkyl spacer and flexible alkoxy spacer	1.41	100.74	52	35.3	58	24.7	53.8	318.1	80	-	104
NAPAEK-Q-100	poly(arylene ether ketone)s	Naphthalene	1.46	74	24.3	4.5	N/A	44.43	6.39	2009	100	-	105

Membrane	Polymeric backbone	Functional group	Properties								Conductivity measurement temp. (°C)	Applied in EW	Refs.
			IEC <sup>a</sup>	Conductivity (mS/cm)	Water uptake (wt%)	Swelling ratio (%)	Thickness (µm)	Tensile strength (Mpa)	Elongation (%)	Young's modulus (MPa)			
VIB5/PMS2/PBI 0.5	N,N-butylvinylimidazolium with p-methylstyrene and polybenzimidazole,	Imidazolium	2.06	147	320.19 <sup>j</sup>	10.77 <sup>j</sup>	70	N/A	N/A	N/A	100	-	106
QPAE/GO-(APTS-c-PTMA) 0.7 wt%	quaternary ammonium functionalized graphene oxide (Q-GO) into quaternized poly(arylene ether) (QPAE) random copolymer containing 3-aminopropyl)triethoxysilane (APTS) and (3-bromopropyl)trimethyl ammonium bromide (PTMA)	PTMA	1.45	114.2	31.2	13.2	27	14.1	5.3	N/A	90	-	107
PDPF-DMP	PDPF	DMP	2.15	124	111	30	N/A	N/A	N/A	N/A	80	-	108
Q-PAES/PPO-55	Quaternized poly(arylene ether sulfone)/PPO belnded	triethylamine (TEA)	1.79	90.9	45.4	25.7		N/A	N/A	N/A	90	-	109
A-PEI-8	Alkalized poly(ether imide)	imidazolium	1.23	44.2	40.3	19.2	31	30.2	9.4	1400	90	-	110
PPO-22-3QA8F	a tri-quaternary ammonium side chain to the poly(phenylene oxide) (PPO) backbone	Terminal alkyne-containing 3-[(2-perfluorooctyl)ethoxy]prop-1-yne	1.58	83	10.2	3.7	20-30	N/A	N/A	N/A	90	-	111
QN-PAEK/rGO5.0wt %	poly(arylene ether ketone) (PAEK)/ reduced graphene oxide (rGO)	bearing fluorenyl group	1.32	116.8	79	26	30	40.8	4.9	400	90	-	112
QPAEK-CN-0.5	PAEK with various g-C <sub>3</sub> N <sub>4</sub> nanosheetsconten	alkaline quaternary ammonium groups and amine/imine groups	1.31	38.6	30.2	6.1	N/A	35.3	86	1370	90	-	113
GT82-5	poly (norbornene) (PNB)	N,N,N',N'- tetramethyl-1,6-hexanediamine (TMHDA),	3.84	212	122	N/A	N/A	N/A	N/A	N/A	80	-	114
PPO5-4QPip-2.6	PPO	piperidinium	2.6	221	115	N/A	50	N/A	N/A	N/A	80	-	115
QPAF-DMBA	quaternized aromatic/perfluoroaklyl copolymer (QPAF)	dimethylbutylamine (DMBA)	1.33	152	53	N/A	45	N/A	N/A	N/A	80	-	116
T20NC6NC5N	PPO	hexyl and pentyl spacers	2.52	176	135	18	100	N/A	N/A	N/A	80	-	
PBP-20Q4	PBP	Piperidinium	3.64	155	242	42	60	35	45	N/A	80	-	117
BeC30%-P	polyvinyl alcohol	grafted bis-crown ether	3.51	235	133						80	-	118

**Table S2.** *Ex-situ* alkaline stability data of selected AEMs.

Membrane	Temperature [°C]	OH <sup>-</sup> conc. [M]	Durability or half life time [h]	%degradation	Measurement Method	Applied in EW	Ref.
PTP-75	80	1	934	41	Conductivity	+	1
	80	1	934	47	NMR		
PTP-85	80	1	934	40	Conductivity	+	1
	80	1	934	36	NMR		
PTP-90	80	1	934	73	Conductivity	+	1
	80	1	934	62	NMR		
PAni-0.87	80	1	48	4	Dry weight ionic, conductivity, IEC	-	12
PAni-0.92	80	1	48	4		-	12
PAni-1.03	80	1	48	4		+	12
PFOTPh-TMA-C6	80	8	168	13.8	conductivity	+	6
HDPE-AEM	80	RH ¼ 100% N <sub>2</sub> atmosphere	500	8	conductivity	-	25, 24
LDPE-AEM	80	RH ¼ 100% N <sub>2</sub> atmosphere	500	6.2	conductivity	-	25, 24, 26
C4-AEM	80	1	672	13	IEC		25
SEBS-BTMA	60	1	360	13.6	conductivity		25
SEBS-CH2-QA-1.5	60	1	360	7.7	conductivity		25
XL100-SEBS-C5-TMA-0.8	80	1	500	2.4	conductivity		25
P1.25-OH	80	0.1	239	53	conductivity		25
XL35-rPNB-X60-Y40	80	1	576	1.4	conductivity		25
HC-[1]498 [2]200	80	1	720	4.2	conductivity		25
PNB-X62-Y38	80	1	1200	0.8	conductivity		25
H22C9N	80	1	500	9.6	conductivity		25
F20C9N	80	1	500	8.9	conductivity		25
ATMPP	22	4	336	33	conductivity		25
HTMA-DAPP	80	4	336	4.6	conductivity		40, 25
	80	4	336	0	IEC		40, 25
	80	4	3000	39	conductivity		25
	80	4	3000	8	IEC		25
	80	0.5	11,160	72	conductivity		25

Membrane	Temperature [°C]	OH <sup>-</sup> conc. [M]	Durability or half life time [h]	%degradation	Measurement Method	Applied in EW	Ref.
QPAF (C6)-2	60	1	400	90	conductivity		25
QP-QAF3	80	1	1000	15	conductivity		25
PAImEE (12)	80	10	240	6.1	NMR		43, 25
BPN1	95	1	1440	8	IEC		25
<i>m</i> -TPN1	95	1	1440	2	Conductivity		25, 119
FLN-55	80	1	500	2	IEC		25
PTPipQ6	90	2	720	64	NMR		25
PAP-TP-85	100	1	2000	3	IEC		9, 25
PTPipQ1	90	2	360	5	IEC		46
PVBC-MPy/15%PEK-cardo	60	1	432	15	Conductivity	-	47
	80	1	432	9			
	60	6	432	16			
PVBC-MPy/25%PEK-cardo	60	1	432	5	Conductivity	-	47
	80	1	432	15			
	60	6	432	20			
PVBC-MPy/35%PEK-cardo	60	1	432	9	Conductivity	+	47
	80	1	432	15			
	60	6	432	19			
PVBC-MPy/40%PEK-cardo	60	1	432	7	Conductivity	-	47
	80	1	432	11			
	60	6	432	25			
PST-TMA <sup>+</sup>	60	1	168	10	NMR	+	49
	60	1	168	37	Conductivity		
	60	2	168	44	Conductivity		
	60	6	168	13	NMR		
PST-DMP <sup>+</sup>	60	1	168	25	NMR	+	49
	60	1	168	49	Conductivity		
	60	2	168	67	Conductivity		
	60	6	168	38	NMR		
PPO24-BIM (bromide form)	80	1	336	65	Conductivity	+	13
FAA-30	80	1	336	8	Conductivity	+	13
Tokuyama A201	80	1	336	0	Conductivity	+	13
PSEBS-CM-DABCO	50	3.8	168	13	Conductivity	+	54

Membrane	Temperature [°C]	OH <sup>-</sup> conc. [M]	Durability or half life time [h]	%degradation	Measurement Method	Applied in EW	Ref.
	50	3.8	168	8	IEC		
	60	3.8	168	47	Conductivity		
	60	3.8	168	33	IEC		
PAEK-APMP75	60	5.3	672	15	IEC	+	55
PAEK-APMP100	60	5.3	672	25	IEC	+	55
SEBS-Pi-73%	80	1	576	25	Conductivity	+	59
				22	IEC		
PBP-ASU-PPO	80	1	2000	13.6	IEC	+	60
BPN1-100	95	1	1440	11	IEC	+	25, 67
	95	1	1440	8	NMR		
TPN1-100	95	1	1440	9	IEC	+	25, 44
	95	1	1440	2.3	NMR		
PAImBB(14)	80	10	240	2.3	NMR	+	43
QPC-TMA	80	1	1000	0	NMR IEC	+	68
PVBC-MPy M2	80	1	600	13	Conductivity	+	69
PVBC-MPy M4	80	1	600	22	Conductivity	+	69
PVBC-MPy M6	80	1	600	18	Conductivity	+	69
PISPVA46	60	0.5	240	56.2	Conductivity	+	73
PISPVA37	60	0.5	240	44.9	Conductivity	+	73
PISPVA28	60	0.5	240	39.7	Conductivity	+	73
QMter-co-Mpi- 100%	60	1	43	6	IEC	+	74
	60	1	43	8	Conductivity		
BPi	80	1	560	48	Conductivity	-	75
SCPi	80	1	560	57	Conductivity	+	75
LSCPi	80	1	560	2	Conductivity	+	75
SCQA	80	1	560	71	Conductivity	-	75
LSCQA	80	1	560	16	Conductivity	+	75
BPI-c-PVBC/OH 1:2	60	1	480	10	IEC	+	77
BPI-c-PVBC/OH 1:3	60	1	480	16	IEC	+	77
ABPBI-c-PVBC/OH 1:2	60	1	480	7	IEC	+	77
QPAF-4	80	1	1000	0	Conductivity	-	78
PSBFP-TMA	80	1	168	0	Conductivity	-	81
PSBFBP-TMA	80	1	168	0	Conductivity	-	81
F-PAE	80	0.5	2	61% backbone	NMR	+	82

Membrane	Temperature [°C]	OH <sup>-</sup> conc. [M]	Durability or half life time [h]	%degradation	Measurement Method	Applied in EW	Ref.
				12% benzylic position			
ATM-PP	80	0.5	2	0	NMR	+	82
PPO-TMA	60	1	720	70	IEC	+	71, 72
	60	6	720	83	IEC		
PPO-ABCO	60	1	720	74	IEC	+	71, 72
PPO-ABCO	60	6	720	92	IEC		
PBI linear	85	6	4224	100	Conductivity	+	83
CL PBI (crosslinked)	85	6	4224	50	Conductivity	+	83
Thermally cured PBI	85	6	4224	0	Conductivity	+	83
MES-PBI (25 wt% KOH)	88	9.5	4968	3	Relative mass	+	70
PQDP-1	80	1	720	3	NMR	-	85
SDQEO	60	1	192	25	Conductivity	-	86
PFTP	80	1	2000	4	NMR	-	120
PFTP	80	5	2000	20	NMR	-	120
m-PTP-20Q	80	2	1600	10.55	Conductivity	-	88
O-PDQA-3	80	2	1080	4	Conductivity	-	89
O-PDQA-3	80	2	1080	2.8	IEC	-	89
PP80N20	80	1	1000	30.8	Conductivity	-	90
PP80N20	80	5	1000	41.1	Conductivity	-	90
Q-CLP1	80	2	300	46	Conductivity	-	121
BiPyBPPEEK-50%	22	1	750	11	Conductivity	-	92
QPAES/QBGO-3.0	60	1	240	27	Conductivity	-	93
PAEK-HQACz-0.7	22	4	168	3.9	Conductivity	-	94
PBP-BOP-ASU 8%	80	2	1400	15.59	NMR	-	95
PBP-ASU	80	2	1400	11.37	NMR	-	95
CP3	60	2	480	20	NMR	-	96
PPEEK:PEG 80:20	80	1	400	15	Conductivity	-	97
MPyPPO	60	1	720	42	Conductivity	-	98
PPO-TMA	22	0.6 ( $\lambda=4$ )	646	8	NMR	-	122
qPBPTT-5	80	1	480	13.29	Conductivity	-	99
m-TPNPiQA	80	5	240	6	IEC	-	100
HyAEM-MP-180	22	9	168	50	IEC	-	101
NPPO-2QA-1.85	60	1	168	28	NMR	-	103
AEM-9.09	60	1	504	50	Conductivity	-	104



Membrane	Temperature [°C]	OH <sup>-</sup> conc. [M]	Durability or half life time [h]	%degradation	Measurement Method	Applied in EW	Ref.
ETFE-AEM	80	1	168	12	IEC	-	123
NAPAEK-Q-100	22	4	168	0.7	IEC	-	105
VIB5/PMS2/PBI0.5	25	2	204	25	Conductivity	-	106
QPAE/GO-(APTS-c-PTMA) 0.7 wt%	90	2	480	20	IEC	-	107
PDPF-DMP	90	2	2400	8.4	NMR	-	108
Q-PAES/PPO-55	50	2	1000	14.7	IEC	-	109
A-PEI-8	90	1	200	72.3	Conductivity	-	110
PPO-22-3QA8F	80	1	504	49	Conductivity	-	111
QN-PAEK/rGO5.0wt%	70	2	600	25	Conductivity	-	112
QPAEK-CN-0.5	60	1	240	11	Conductivity	-	113
GT82-15	80	1	1000	1.43	Conductivity	-	114
PPO5-4QPip-2.1	90	1	240	14	Conductivity, IEC	-	115
QPAF-DMBA	60	1	1000	42	Conductivity	-	116
T20NC6NC5N	80	1	500	10	Conductivity	-	
PBP-20Q4	80	2	1800	8.2	Conductivity	-	117

**Table S3.** AEMs and their performance in AEMWE cells using liquid electrolyte.

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
CuCoO <sub>3</sub>	Tokuyama A201	28	Ni/CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> /C	PTFE	1% K <sub>2</sub> CO <sub>3</sub> /KHCO <sub>3</sub>	1.9	0.47	50	10
IrO <sub>2</sub>	Fumatech FAA-3-PK-75	75	Pt/C	PTFE	0.5 M KOH	1.8	0.48	50	53
IrO <sub>2</sub>	Fumatech FAA-3-PK-75	75	Pt/C	PTFE	0.5 M KOH	1.8	0.61	60	53
IrO <sub>2</sub>	Fumatech FAA-3-PK-75	75	Pt/C	PTFE	0.5 M KOH	1.8	0.625	70	53
IrO <sub>2</sub>	Fumatech FAA-3-PK-75	75	Pt/C	PTFE	0.5 M KOH	1.8	0.87	80	53
IrO <sub>2</sub>	Fumatech FAA-3-PK-75	75	Pt/C	PTFE	0.5 M KOH	1.8	0.98	90	53
Li <sub>0.21</sub> Co <sub>2.79</sub> O <sub>4</sub>	QPDTB	50	Ni	QPDTB	0.2 M KOH	2.2	0.3	20	56, 124
Li <sub>0.21</sub> Co <sub>2.79</sub> O <sub>4</sub>	QPDTB	50	Ni	QPDTB	0.2 M KOH	2.05	0.3	40	56, 124
Ni/CPs	Tokuyama A201	28	Ni/CPs	—	1 M KOH	1.9	0.15	50	125
NiFe <sub>2</sub> O <sub>4</sub>	Tokuyama A201	28	NiFeCo	5% Nafion	1 M KOH	2.21	2.13	60	126
NiFe <sub>2</sub> O <sub>4</sub>	AEMION	38	NiFeCo	5% Nafion	1 M KOH	2.26	2.13	60	126
NiFe <sub>2</sub> O <sub>4</sub>	Sustanion	50	NiFeCo	5% Nafion	1 M KOH	2.13	2.13	60	126
CuCoO <sub>3</sub>	LDPE-g-VBC	60	Ni/CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> /C	PTFE	1% K <sub>2</sub> CO <sub>3</sub> /KHCO <sub>3</sub>	2.1	0.46	50	58
IrO <sub>2</sub>	SEBS-Pi	60	Pt/C		1 M KOH	2.0	0.4	50	31, 59
IrO <sub>2</sub>	SEBS-Pi	60	Pt/C		1 M KOH	2.0	0.5	80	31, 59
Pd/TNTA web	Tokuyama A201	28	Pt/C	PTFE	2 M NaOH	2	2	80	127
CuCoO <sub>x</sub> (on Ni foam) Acta's 3030	Tokuyama A201	28	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C on carbon paper Acta's 4030	I2	1% K <sub>2</sub> CO <sub>3</sub>	1.91	0.4	60	62

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, $\mu\text{m}$	Cathode catalyst						
CuCoO <sub>x</sub> (on Ni foam)	Fumatech FAA-3	40	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C on carbon paper	I2	1% K <sub>2</sub> CO <sub>3</sub>	1.91	0.4	60	62
CuCoO <sub>x</sub> (on Ni foam)	Fumatech FAA-3-PP-75	80	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C on carbon paper	I2	1% K <sub>2</sub> CO <sub>3</sub>	1.99	0.4	60	62
CuCoO <sub>x</sub> (on Ni foam)	Tokuyama A901	10	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C on carbon paper	I2	1% K <sub>2</sub> CO <sub>3</sub>	2.1	0.5	50	61
NiCo <sub>2</sub> O <sub>4</sub>	Polyethylene based radiation grafted		Pt	Polystyrene- <i>b</i> -poly(ethylene/butylene)- <i>b</i> -polystyrene	0.1 M KOH	1.65	0.1	60	63
NiAl	HMT-PMBI	34	NiAlMo	—	1 M KOH	2.1	2	60	128
IrO <sub>2</sub>	PTP-90	45	Pt/C		1 M NaOH	2.2	0.91	55	129
IrO <sub>2</sub>	PTP-90	45	Pt/C		1 M NaOH	2.2	1	75	129
IrO <sub>2</sub>	PTP-85	45	Pt/C		1 M NaOH	2.2	0.83	55	129
IrO <sub>2</sub>	PTP-75	45	Pt/C		1 M NaOH	2.2	0.76	55	129
IrO <sub>2</sub>	Sustainion 37-50	50	Pt/C		1 M KOH	1.63	1	60	16
NiFe	Sustainion 37-50	50	NiFeCo		1 M KOH	1.9	1	60	16
RANEYs-type-Ni	m-PBI	40	RANEYs-type-Ni Mo		24 wt% KOH	1.8	1.7	80	130
PtRu	PAImEE	13	Pt/C	PAI	6 M KOH	2.3	0.4	80	43
IrO <sub>2</sub>	HTMA-DAPP	26	PtRu/C	9 wt% TMA	0.1 M NaOH	1.8	0.95	60	40
NiFe	HTMA-DAPP	26	PtRu/C	20 wt% TMA	0.1 M NaOH	1.8	3.2	60	40
NiFe	HTMA-DAPP	26	PtRu/C	20 wt% TMA	1 M NaOH	1.8	5.3	60	40
IrO <sub>2</sub>	QPC-TMA	50	Pt/C	QPC-TMA	1 M KOH	1.9	3.5	70	68
IrO <sub>2</sub>	GT74	50	Pt/C	GT-18	3 wt% KOH	1.59	0.1	50	7
IrO <sub>2</sub>	GT74	50	Pt/C	GT-32	3 wt% KOH	1.62	0.1	50	7
IrO <sub>2</sub>	GT74	50	Pt/C	GT-0	3 wt% KOH	1.69	0.1	50	7
IrO <sub>2</sub>	GT74	50	Pt/C	GT-100	3 wt% KOH	1.83	0.1	50	7
IrO <sub>2</sub>	GT74	50	Pt/C	GT-75	3 wt% KOH	2.08	0.1	50	7

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
IrO <sub>2</sub>	GT74	50	Pt/C	GT-82	3 wt% KOH	2.11	0.1	50	7
NiFe	PVBC- MPy M2	80	NiMo	-----	1 M KOH	1.9	0.5	80	69
NiFe	PVBC- MPy M4	80	NiMo	-----	1 M KOH	1.9	0.4	80	69
NiFe	PVBC- MPy M6	80	NiMo	-----	1 M KOH	1.9	0.6	80	69
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion 37-50		NiFeCo	PFOTFPh-TMA C6	1 M KOH	2.09	0.94	50	17
IrO <sub>2</sub>	PFOTFPh-TMA C6	20-30	Pt/C	PFOTFPh-TMA C8	1 M KOH	1.77	1	80	5, 6
IrO <sub>2</sub>	PFOTFPh-TMA C8	20-30	Pt/C	PFOTFPh-TMA C10	1 M KOH	1.79	1	80	5, 6
IrO <sub>2</sub>	PFOTFPh-TMA C10	20-30	Pt/C	PFOTFPh-TMA C6	1 M KOH	1.84	1	80	5, 6
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion™ X37-50	50	NiFeCo		1 M KOH	1.9	1	60	18
IrO <sub>2</sub>	Sustainion™ X37-50	50	Pt/C		1 M KOH	1.63	1	60	18
CuCoO <sub>x</sub>	Tokuyama A201	28	Pt/C	AS-4	0.1 wt% K <sub>2</sub> CO <sub>3</sub>	2.22	0.8	50	14
CuCoO <sub>x</sub>	Tokuyama A201	28	Pt/C	AS-4	1 wt% K <sub>2</sub> CO <sub>3</sub>	2.05	0.8	50	14
CuCoO <sub>x</sub>	Tokuyama A201	28	Pt/C	AS-4	10 wt% K <sub>2</sub> CO <sub>3</sub>	1.87	0.8	50	14
CuCoO <sub>x</sub>	Tokuyama A201	28	Pt/C	AS-4	0.01 M KOH	2.06	0.8	50	14
Ir black	AF1-HNN8-25 Aemion™	25	Pt/C	Aemion™ AP1-HNN8	1 M KOH	1.75	1	50	15
Ir black	AF1-HNN8-50 Aemion™	50	Pt/C	Aemion™ AP1-HNN8	1 M KOH	1.8	1	50	15
Ir black	AF1-HNN5-25 Aemion™	25	Pt/C	Aemion™ AP1-HNN8	1 M KOH	1.81	1	50	15
Ir black	AF1-HNN5-50 Aemion™	50	Pt/C	Aemion™ AP1-HNN8	1 M KOH	1.92	1	50	15
Ir black	AF1-HNN8-25 Aemion™	25	Pt/C	Aemion™ AP1-HNN8	0.1 M KOH	1.85	1	50	15
Ir black	AF1-HNN8-50 Aemion™	50	Pt/C	Aemion™ AP1-HNN8	0.1 M KOH	1.9	1	50	15
Ir black	AF1-HNN5-25 Aemion™	25	Pt/C	Aemion™ AP1-HNN8	0.11 M KOH	1.98	1	50	15
Ir black	AF1-HNN5-50 Aemion™	50	Pt/C	Aemion™ AP1-HNN8	0.1 M KOH	2.13	1	50	15

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
Ir black	Fumatech FAA-3-PE-30	20-30	NiMo	Fumion FAA-3-solute-10	1 M KOH	1.9	1	50	131
Ir black	Fumatech FAA-3-PE-30	20-30	Pt/C	Fumion FAA-3-solute-10	1 M KOH	1.8	1	50	131
CuCoOx	Mg-Al LDH	300	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C	PTFE	0.1 M NaOH	2.1	0.16	60	132
CuCoOx	Mg-Al LDH	300	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C	PTFE	0.1 Na <sub>2</sub> CO <sub>3</sub>	2.1	0.1	60	132
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion®37-50	50	NiFeCo	Nafion	1 M KOH	1.9	1	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Fumatech FAS-50	50	NiFeCo	Nafion	1 M KOH	1.9	0.5	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Fumatech FAPQ	68-82	NiFeCo	Nafion	1 M KOH	1.9	0.17	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Neosepta ACM	110	NiFeCo	Nafion	1 M KOH	1.9	0.05	60	19
NiFe <sub>2</sub> O <sub>4</sub>	AMI 7001	18000	NiFeCo	Nafion	1 M KOH	1.9	0.15	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Celazole®PBI		NiFeCo	Nafion	1 M KOH	1.9	0.07	60	19
IrO <sub>2</sub>	Tokuyama A201	28	Pt/C	PTFE	0.5 M KOH	2	1.31	50	133
IrO <sub>2</sub>	Sustainion™		Pt	PTFE	1 M KOH	1.9	4.6	60	134
Plain nickel foam	mes-PBI	64	Plain nickel foam		25 wt% KOH	2.3	0.7	80	70
Plain nickel foam	mes-PBI	60	Plain nickel foam		15 wt% KOH	2.3	0.42	80	70
Plain nickel foam	mes-PBI	62	Plain nickel foam		5 wt% KOH	2.3	0.01	80	70
Plain nickel foam	m-PBI	40	Plain nickel foam		20 wt% KOH	2.3	0.9	80	70
Ni foam	L-ABPBI (linear)	90-120	Ni foam		1.9 M KOH	2	0.155	50	84
Ni foam	C-ABPBI (cross-linked)	90-120	Ni foam		1.9 M KOH	2	0.18	50	84
Ni foam	L-ABPBI	90-120	Ni foam	PTFE	1.9 M KOH	2	0.18	70	84
Ni foam	C-ABPBI	90-120	Ni foam	PTFE	1.9 M KOH	2	0.22	70	84
Ni	L-PBI	90-120	Ni	PTFE	30 wt% KOH	2	0.15	80	83
Ni	C-PBI	90-120	Ni	PTFE	30 wt% KOH	2	0.12	80	83

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
CE-CCO, CuCo-oxide on nickel foam	X37-50 Grade T	50	Pt/C	PTFE	1 M KOH	1.8	1.39	45	22
IrO <sub>2</sub>	Sustainion X37-50 Grade T	50	Pt/C	PTFE	1 M KOH	1.8	0.9	45	22
Cu <sub>0.5</sub> Co <sub>2.5</sub> O <sub>4</sub>	Sustainion X37-50 Grade T	50	Pt/C	PTFE	1 M KOH	1.8	1.3	45	23
IrO <sub>2</sub>	Sustainion X37-50 Grade T	50	Pt/C	PTFE	1 M KOH	1.8	1.03	45	23
Ni foam	Sustainion X37-50 Grade T	50	Pt/C	PTFE	1 M KOH	1.8	0.35	45	23
IrO <sub>2</sub>	PISPVA46	55	Pt/C	PTFE	0.5 M KOH	2	0.55	60	135
IrO <sub>2</sub>	PISPVA37	55	Pt/C	PTFE	0.5 M KOH	2	0.41	60	135
IrO <sub>2</sub>	PISPVA28	55	Pt/C	PTFE	0.5 M KOH	2	0.15	60	135
IrO <sub>2</sub>	QMter-co-Mpi		Pt/C		0.6 M KOH	2	0.25	50	74
IrO <sub>2</sub>	QMter-co-Mpi		Pt/C		1 M KOH	2	0.31	50	74
NiCo <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	100	NiFe <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	1 wt% KOH	2	0.065	40	54
NiCo <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	100	NiFe <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	5 wt% KOH	2	0.105	40	54
NiCo <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	100	NiFe <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	10 wt% KOH	2	0.128	40	54
NiCo <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	100	NiFe <sub>2</sub> O <sub>4</sub>	PSEBS-CM-DABCO	15 wt% KOH	2	0.15	40	54
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion® 37-50	50	NiFeCo		1 M KOH	1.9	1	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Fumatech FAS-50	50	NiFeCo		1 M KOH	1.9	0.5	60	19
NiFe <sub>2</sub> O <sub>4</sub>	Fumatech FAPQ	75	NiFeCo		1 M KOH	1.9	0.16	60	19
NiFe <sub>2</sub> O <sub>4</sub>	AMI-7001	450	NiFeCo		1 M KOH	1.9	0.11	60	19
NiFe-LDH/NF	PVBC-MPy/35%PEK-cardo	60	MoNi/NF	PTFE	1 M KOH	2	0.5	60	47
IrO <sub>2</sub>	Fumatech FAA-3-50	50	Pt/C		1 M KOH	2	0.63	60	136

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, $\mu\text{m}$	Cathode catalyst						
g-CN-CNF-800	Fumatech FAA-3-50	50	Pt/C		1 M KOH	2	0.98	60	136
FeNiMo-based	Sustainion X37-50 Grade T	50	NiMo-based	Nafion	1 M KOH	1.57	1	80	20
FeNiMo-based	Sustainion X37-50 Grade T	50	NiMo-based	Nafion	1 M KOH	1.62	1	60	20
FeNiMo-based	Sustainion X37-50 Grade T	50	NiMo-based	Nafion	1 M KOH	1.68	1	40	20
FeNiMo-based	Sustainion X37-50 Grade T	50	NiMo-based	Nafion	1 M KOH	1.69	0.6	20	20
APS NiAl-AA	NEOSEPTA (Astom)	220	APS NiAlMo-AA		1 M KOH	1.88	0.5	65	137
APS NiAl-AA	NEOSEPTA (Astom)	220	APS NiAlMo-AA		0.7 M KOH	1.93	0.5	65	137
APS NiAl-AA	NEOSEPTA (Astom)	220	APS NiAlMo-AA		0.5 M KOH	1.95	0.5	65	137
APS NiAl-AA	NEOSEPTA (Astom)	220	APS NiAlMo-AA		0.3 M KOH	2	0.5	65	137
APS NiAl-AA	NEOSEPTA (Astom)	220	APS NiAlMo-AA		0.1 M KOH	2.14	0.5	65	137
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion Grade T, PTFE reinforced	50	Raney nickel		1 M KOH	1.8	0.84	60	21
NiFe <sub>2</sub> O <sub>4</sub>	Sustainion X37-50	50	Raney nickel		1 M KOH	1.8	0.74	60	21
NiFe	PSU-PVP	120	NiMo		20 wt% KOH	1.9	0.5	80	76
Ni-doped FeOOH	Sustainion X37-50 Grade T	50	Pt/C	Nafion	1 M KOH	1.7	0.92	50	138
Ni-doped FeOOH	Sustainion X37-50 Grade T	50	Pt/C	Nafion	1 M KOH + 0.5 M NaCl	1.7	0.74	50	138
Ni-doped FeOOH	Sustainion X37-50 Grade T	50	Pt/C	Nafion	1 M KOH + seawater	1.7	0.73	50	138
IrO <sub>2</sub>	Sustainion X37-50 Grade T	50	Pt/C	PTFE	1 M KOH + seawater	1.7	0.47	50	138
Ni foam	PBI-c-PVBC/OH (1:3)	24	Ni foam		1 M KOH	1.98	0.38	50	77

Membrane electrode assembly (MEA) components				Ionomer/ binder	Feed type	Cell Voltage (V)	Current density (A/cm <sup>2</sup> )	Cell temperature (°C)	Ref.
Anode catalyst	Membrane type	Membrane thickness, $\mu\text{m}$	Cathode catalyst						
Ni foam	ABPBI-c-PVBC/OH (1:2)	90	Ni foam		1 M KOH	1.98	0.38	50	77
Ni foam	C-PVAF-ABPBI	30-45	Ni foam		15 wt% KOH	2	0.24	50	139
Ni foam	C-PVAF-ABPBI	30-45	Ni foam		15 wt% KOH	2	0.27	70	139
Graphene	Selemion AMV	120	Graphene		Water	2	0.09	30	140
NiCo <sub>2</sub> O <sub>4</sub>	AMB-commercial (now Pure Water Technologies)		NiCo <sub>2</sub> O <sub>4</sub>		1 M KOH	2	0.25	50	12
NiCo <sub>2</sub> O <sub>4</sub>	IPA- commercial (CSIR-CSMCRI)		NiCo <sub>2</sub> O <sub>4</sub>		1 M KOH	2	0.16	50	12
NiCo <sub>2</sub> O <sub>4</sub>	PAni-1.03	100	NiCo <sub>2</sub> O <sub>4</sub>		1 M KOH	2	0.4	50	12
PGM	BPN1-100		PGM	AS4	0.5 M NaOH	2.1	0.4	50	67
PGM	TPN1-100		PGM	AS4	0.5 M NaOH	2.35	0.36	50	67
NiMn <sub>2</sub> O <sub>4</sub>	Fumatech FAA3-50	50	Pt/KB		1 M KOH	2	0.24	40	3
NiMn <sub>2</sub> O <sub>4</sub>	Fumatech FAA3-50	50	Pt/KB		1 M KOH	2	0.37	50	3
NiMn <sub>2</sub> O <sub>4</sub>	Fumatech FAA3-50	50	Pt/KB		1 M KOH	2	0.47	60	3
NiMn <sub>2</sub> O <sub>4</sub>	Fumatech FAA3-50	50	Pt/KB		1 M KOH	2	0.53	80	3
IrO <sub>2</sub>	HTMA-DAPP	50	PtRu/C	HTMA-DAPP	0.05 M KOH	2	0.7	60	141
IrO <sub>2</sub>	HTMA-DAPP	50	PtRu/C	HTMA-DAPP	0.1 M KOH	2	1.22	60	141
IrO <sub>2</sub>	HTMA-DAPP	50	PtRu/C	HTMA-DAPP	0.5 M KOH	2	1.4	60	141
IrO <sub>2</sub>	HTMA-DAPP	50	PtRu/C	HTMA-DAPP	1 M KOH	2	1.74	60	141



**Table S4.** AEMs and their performance in AEMWE cells fed with pure water (no liquid electrolyte).

Membrane electrode assembly (MEA) components				Ionomer/binder	Feed type	Cell Voltage (V)	Current density A cm <sup>-2</sup>	Cell temp. °C	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
Pb <sub>2</sub> Ru <sub>2</sub> O <sub>6.5</sub>	PSF-TMA	50	Pt black	PSF-TMA <sup>+</sup> Cl <sup>-</sup>	Water	1.8	0.4	50	48
Ni-Fe	xQAPS	70	Ni-Mo	xQAPS	Water	1.85	0.4	70	52
NiCoO <sub>x</sub> :Fe	Fumatech FAA-3	130	Pt/C	FAA-3	Water	2.0	0.37	50	142
IrO <sub>2</sub>	Tokuyama A201	28	Pt black	As-4	Water	1.6	0.1	50	148
Cu <sub>0.7</sub> CO <sub>2.3</sub> O <sub>4</sub>	QPDTB	50	Nano Ni	Poly(DMAEMA- <i>co</i> -TFEMA- <i>co</i> -BM)	Water	1.9	0.1	22	57
IrO <sub>2</sub>	SEBS-Pi	60	Pt/C		Water	2	0.1	50	31, 59
Ce <sub>0.2</sub> MnFe <sub>1.8</sub> O <sub>4</sub>	Fumatech FAA-3-PK-130	110-130	Pt on Ti	—	Water	1.8	0.3	—	143
NiCo <sub>2</sub> O <sub>4</sub> on steel mesh	Homemade, not specified		Acta 4030	Homemade, not specified	Water	1.86	0.4	60	144
Cu <sub>x</sub> Co <sub>3-x</sub> O <sub>4</sub>	PTFE-qPDTB-OH-	25	Pt/C	q-ammonium polymethacrylate	Water	2.0	0.4	22	64
IrO <sub>2</sub>	HTMA-DAPP	26	PtRu/C	9 wt% TMA	Water	1.8	0.48	60	40
IrO <sub>2</sub>	HTMA-DAPP	26	PtRu/C	9 wt% TMA	Water	1.8	0.79	85	40
NiFe	HTMA-DAPP	26	NiMo/C	20 wt% TMA	Water	1.8	0.9	85	40
NiFe	HTMA-DAPP	26	PtRu/C	20 wt% TMA	Water	1.8	2.7	85	40
NiFe-BTC-GNPs MOF	FAA-3-PK-130	130	MoNi <sub>4</sub> /MoO <sub>2</sub>		Water	1.85	0.54	70	145
IrO <sub>2</sub>	Fumatech FAA-3-50	50	Pt/C	FAA-3-B	Water	1.85	0.12	50	146
IrO <sub>2</sub>	QPC-TMA	50	Pt/C	QPC-TMA	Water	1.9	0.3	70	68
IrO <sub>2</sub>	PFOTFPh-TMA C6	20-30	Pt/C	PFOTFPh-TMA C8	Water	1.94	1	80	5, 6
IrO <sub>2</sub>	PFOTFPh-TMA C8	20-30	Pt/C	PFOTFPh-TMA C10	Water	1.95	1	80	5
IrO <sub>2</sub>	PFOTFPh-TMA C10	20-30	Pt/C	PFOTFPh-TMA C6	Water	2.14	1	80	5, 6
CuCoO <sub>x</sub>	Tokuyama A201	28	Pt/C	AS-4	Water	2.2	0.1	50	14

Membrane electrode assembly (MEA) components				Ionomer/binder	Feed type	Cell Voltage (V)	Current density A cm <sup>-2</sup>	Cell temp. °C	Ref.
Anode catalyst	Membrane type	Membrane thickness, μm	Cathode catalyst						
Cu <sub>0.7</sub> Co <sub>2.3</sub> O <sub>4</sub>	mm-qPVBz/Cl-	70	Ni nano powder	qPVB/Cl-	Water	2.19	0.1	25	65
Cu <sub>0.7</sub> Co <sub>2.3</sub> O <sub>4</sub>	mm-qPVBz/Cl-	70	Ni nano powder	qPVB/Cl-	Water	2.05	0.1	40	65
Cu <sub>0.7</sub> Co <sub>2.3</sub> O <sub>4</sub>	mm-qPVBz/Cl-	70	Ni nano powder	qPVB/Cl-	Water	1.99	0.1	55	65
CuCoOx	Mg-Al LDH	300	Ni/(CeO <sub>2</sub> -La <sub>2</sub> O <sub>3</sub> )/C	PTFE	Water	2.1	0.02	60	132
IrO <sub>2</sub>	PPO-TMA	40	Pt/C		Water	1.8	0.24	50	72
IrO <sub>2</sub>	PPO-ABCO	40	Pt/C		Water	1.8	0.075	50	72
IrO <sub>2</sub>	PISPVA46	55	Pt/C	PTFE	Water	2	0.06	60	135
IrO <sub>2</sub>	QMter-co-Mpi		Pt/C		Water	2	0.06	50	74
IrO <sub>2</sub>	LSCPi	50	Pt/C	LSCPi	Water	1.8	0.3	50	75
IrO <sub>2</sub>	SCPi	50	Pt/C	SCPi	Water	1.8	0.2	50	75
Graphene	Selemion AMV	120	Graphene		Water	2	0.09	30	140
NiMPL-PTL	Sustainion X37-50 Grade T	50	NiMPL-PTL		Water	1.9	0.5	60	147
IrO <sub>2</sub>	HTMA-DAPP	50	PtRu/C	HTMA-DAPP	Water	2	0.44	60	141

**Table S5.** AEMs and their performance stability in operando AEMWE cells.

Membrane	Anode catalyst	Cathode catalyst	Electrolyte	Temperature (°C)	Current density (A cm <sup>-2</sup> )	Durability (h)	Ref.
QMter-co-Mpi	IrO <sub>2</sub>	Pt/C	5.6 wt% KOH	50	0.2	500 (2.1 V)	74
Tokuyama A201	IrO <sub>2</sub>	Pt	Water	50	0.2	535 (2.25 V)	148
Tokuyama A201	Acta's 3030	Acta's 4030	1 wt% K <sub>2</sub> CO <sub>3</sub> /KHCO <sub>3</sub>	43	0.47	1000 (2.05 V)	10
PSF-TMA	Pb <sub>2</sub> Ru <sub>2</sub> O <sub>6.5</sub>	Pt	Pure water	50	0.2	6 (133.3 mV h <sup>-1</sup> )	48
xQAPS	Ni-Fe	Ni-Mo	Water	70	0.4	8 (1.8 V)	52
mes-PBI	Ni foam	Ni foam	25 wt% KOH	80	0.2	72 (1.95 V)	70
PSEBS-CM-DABCO	NiCo <sub>2</sub> O <sub>4</sub>	NiFe <sub>2</sub> O <sub>4</sub>	10 wt% KOH	50	0.3	150 (2.267 V)	54
Sustainion® 37-50	NiFe <sub>2</sub> O <sub>4</sub>	NiFeCo	5.6 wt% KOH	60	1	2000 (1.9 V)	19
Fumatech FAS-50	NiFe <sub>2</sub> O <sub>4</sub>	NiFeCo	5.6 wt% KOH	60	1	200 (400 μV h <sup>-1</sup> )	19
PAImEE(12)	Pt/C	Pt/C	33.6 wt% KOH	60	0.4	48 (2.6 V)	43
LSCPi	IrO <sub>2</sub>	Pt/C	Water	50	0.2	35 (16.3 μV h <sup>-1</sup> )	75
PVBC-MPy/35%PEK-cardo	NiFe-LDH/NF	MoNi/NF	5.6 wt% KOH	60	0.5	46 (2.0 V)	47
SEBS-Pi	IrO <sub>2</sub>	Pt/C	5.6 wt% KOH	50	0.4	100 (2.08 V)	59
Fumatech FAA-3	Acta's 3030	Acta's 4030	1 wt% K <sub>2</sub> CO <sub>3</sub>	60	0.5	31 (2.04 V)	62
Fumatech FAA-3-PP-75	Acta's 3030	Acta's 4030	1 wt% K <sub>2</sub> CO <sub>3</sub>	60	0.5	200 (2.38 μV h <sup>-1</sup> )	62
NEOSEPTA	APS NiAl-AA	APS NiAlMo-AA	1 M KOH	65	0.5	112 (350 μV h <sup>-1</sup> )	137
Sustainion Grade T	NiFe <sub>2</sub> O <sub>4</sub>	Raney nickel	1 M KOH	60	1	12180 (0.7 μV h <sup>-1</sup> )	21
Sustainion X37-50	NiFe <sub>2</sub> O <sub>4</sub>	Raney nickel	1 M KOH	60	1	10100 (0.7 μV h <sup>-1</sup> )	21
Fumatech FAS-50	NiFe <sub>2</sub> O <sub>4</sub>	Raney nickel	1 M KOH	60	1	140 (655 μV h <sup>-1</sup> )	21
AF1-HNN8-25X	IrO <sub>2</sub>	Pt/C	1 M KOH	50	0.5	17 (2390 μV h <sup>-1</sup> )	15
ATM-PP	IrO <sub>2</sub>	Pt/C		50	0.2	2100	82
PSU-PVP	NiFe	NiMo	20 wt% KOH	80	0.5	700	76
PTP-85	IrO <sub>2</sub>	Pt/C	1 M NaOH	55	0.4	120 (1.61 mV h-1)	129
Sustainion X37-50	NiFe <sub>2</sub> O <sub>4</sub>	NiFeCo	1 M KOH	50	0.2	1000	17
X37-50 Grade T	Ni-doped FeOOH	Pt/C	1 M KOH + seawater	50	0.5	15 (1.72 V)	138
HTMA-DAPP	NiFe	PtRu/C	Water	60	0.2	170 (2.1 V)	40
PFOTPh-TMA C8	IrO <sub>2</sub>	Pt/C	Water	80	0.2	130	5
PFOTPh-TMA C10	IrO <sub>2</sub>	Pt/C	1 M KOH	80	0.2	150	5
Sustainion X37-50 Grade T	CCO-11	Pt/C	1 M KOH	45	0.4	100	23

## References:

- <sup>1</sup> X. Hu, Y. Huang, L. Liu, Q. Ju, X. Zhou, X. Qiao, Z. Zheng, N. Li, Piperidinium functionalized aryl ether-free polyaromatics as anion exchange membrane for water electrolyzers: Performance and durability, *J. Memb. Sci.* 621 (2021) 118964. <https://doi.org/10.1016/j.memsci.2020.118964>.
- <sup>2</sup> D. Henkensmeier, M. Najibah, C. Harms, J. Zitka, J. Hnat, K. Bouzek, Overview: State-of-the Art Commercial Membranes for Anion Exchange Membrane Water Electrolysis, *J. Electrochem. Energy Convers. Storage.* 18 (2021). <https://doi.org/10.1115/1.4047963>.
- <sup>3</sup> A. Carbone, S.C. Zignani, I. Gatto, S. Trocino, A.S. Arico, Assessment of the FAA3-50 polymer electrolyte in combination with a NiMn<sub>2</sub>O<sub>4</sub> anode catalyst for anion exchange membrane water electrolysis, *Int. J. Hydrogen Energy.* 45 (2020) 9285–9292. <https://doi.org/10.1016/j.ijhydene.2020.01.150>.
- <sup>4</sup> S. Vengatesan, S. Santhi, S. Jeevanantham, G. Sozhan, Quaternized poly (styrene-co-vinylbenzyl chloride) anion exchange membranes for alkaline water electrolyzers, *J. Power Sources.* 284 (2015) 361–368. <https://doi.org/10.1016/j.jpowsour.2015.02.118>.
- <sup>5</sup> R. Soni, S. Miyanishi, H. Kuroki, T. Yamaguchi, Pure Water Solid Alkaline Water Electrolyzer Using Fully Aromatic and High-Molecular-Weight Poly(fluorene- alt-tetrafluorophenylene)-trimethyl Ammonium Anion Exchange Membranes and Ionomers, *ACS Appl. Energy Mater.* 4 (2021) 1053–1058. <https://doi.org/10.1021/acsaem.0c01938>.
- <sup>6</sup> S. Miyanishi, T. Yamaguchi, Highly conductive mechanically robust high M-w polyfluorene anion exchange membrane for alkaline fuel cell and water electrolysis application, *Polym. Chem.* 11 (2020) 3812–3820. <https://doi.org/10.1039/d0py00334d>.
- <sup>7</sup> G. Huang, M. Mandal, N.U. Hassan, K. Groenhout, A. Dobbs, W.E. Mustain, P.A. Kohl, Ionomer Optimization for Water Uptake and Swelling in Anion Exchange Membrane Electrolyzer: Oxygen Evolution Electrode, *J. Electrochem. Soc.* 167 (2020) 164514. <https://doi.org/10.1149/1945-7111/abcde3>.
- <sup>8</sup> J.W. Xiao, A.M. Oliveira, L. Wang, Y. Zhao, T. Wang, J.H. Wang, B.P. Setzler, Y.S. Yan, Water-Fed Hydroxide Exchange Membrane Electrolyzer Enabled by a Fluoride-Incorporated Nickel-Iron Oxyhydroxide Oxygen Evolution Electrode, *ACS Catal.* 11 (2021) 264–270. <https://doi.org/10.1021/acscatal.0c04200>.
- <sup>9</sup> J. Wang, Y. Zhao, B.P. Setzler, S. Rojas-Carbonell, C. Ben Yehuda, A. Amel, M. Page, L. Wang, K. Hu, L. Shi, S. Gottesfeld, B. Xu, Y. Yan, Poly(aryl piperidinium) membranes and ionomers for hydroxide exchange membrane fuel cells, *Nat. Energy.* 4 (2019) 392–398. <https://doi.org/10.1038/s41560-019-0372-8>.

- <sup>10</sup> C.C. Pavel, F. Cecconi, C. Emiliani, S. Santuccioli, A. Scaffidi, S. Catanorchi, M. Comotti, Highly Efficient Platinum Group Metal Free Based Membrane-Electrode Assembly for Anion Exchange Membrane Water Electrolysis, *Angew. CHEMIE-INTERNATIONAL Ed.* 53 (2014) 1378–1381. <https://doi.org/10.1002/anie.201308099>.
- <sup>11</sup> W.T. Lu, Z.Z. Yang, H. Huang, F. Wei, W.H. Li, Y.H. Yu, Y.G. Gao, Y.H. Zhou, G. Zhang, Piperidinium-Functionalized Poly(Vinylbenzyl Chloride) Cross-linked by Polybenzimidazole as an Anion Exchange Membrane with a Continuous Ionic Transport Pathway, *Ind. Eng. Chem. Res.* 59 (2020) 21077–21087. <https://doi.org/10.1021/acs.iecr.0c04548>.
- <sup>12</sup> M. Bhushan, M. Mani, A.K. Singh, A.B. Panda, V.K. Shahi, Self-standing polyaniline membrane containing quaternary ammonium groups loaded with hollow spherical NiCo<sub>2</sub>O<sub>4</sub> electrocatalyst for alkaline water electrolyser, *J. Mater. Chem. A.* 8 (2020) 17089–17097. <https://doi.org/10.1039/d0ta05373b>.
- <sup>13</sup> A. Marinkas, I. Strużyńska-Piron, Y. Lee, A. Lim, H.S. Park, J.H. Jang, H.J. Kim, J. Kim, A. Maljusch, O. Conradi, D. Henkensmeier, Anion-conductive membranes based on 2-mesitylbenzimidazolium functionalised poly(2,6-dimethyl-1,4-phenylene oxide) and their use in alkaline water electrolysis, *Polymer (Guildf)*. 145 (2018) 242–251. <https://doi.org/10.1016/j.polymer.2018.05.008>.
- <sup>14</sup> H. Ito, N. Kawaguchi, S. Someya, T. Munakata, N. Miyazaki, M. Ishida, A. Nakano, Experimental investigation of electrolytic solution for anion exchange membrane water electrolysis, *Int. J. Hydrogen Energy.* 43 (2018) 17030–17039. <https://doi.org/10.1016/j.ijhydene.2018.07.143>.
- <sup>15</sup> P. Fortin, T. Khoza, X.Z. Cao, S.Y. Martinsen, A.O. Barnett, S. Holdcroft, High-performance alkaline water electrolysis using Aemion (TM) anion exchange membranes, *J. Power Sources.* 451 (2020) 227814. <https://doi.org/10.1016/j.jpowsour.2020.227814>.
- <sup>16</sup> J.J. Kaczur, H.Z. Yang, Z.C. Liu, S.A. Sajjad, R.I. Masel, Carbon Dioxide and Water Electrolysis Using New Alkaline Stable Anion Membranes, *Front. Chem.* 6 (2018). <https://doi.org/10.3389/fchem.2018.00263>.
- <sup>17</sup> A.K. Niaz, A. Akhtar, J.Y. Park, H.T. Lim, Effects of the operation mode on the degradation behavior of anion exchange membrane water electrolyzers, *J. Power Sources.* 481 (2021) 229093. <https://doi.org/10.1016/j.jpowsour.2020.229093>.
- <sup>18</sup> Z. Liu, S.D. Sajjad, Y. Gao, J. Kaczur, R. Masel, An Alkaline Water Electrolyzer with Sustainion™ Membranes: 1 A/cm<sup>2</sup> at 1.9V with Base Metal Catalysts, *ECS Trans.* 77 (2017) 71–73. <https://doi.org/10.1149/07709.0071ecst>.

- <sup>19</sup> Z. Liu, S.D. Sajjad, Y. Gao, H. Yang, J.J. Kaczur, R.I. Masel, The effect of membrane on an alkaline water electrolyzer, *Int. J. Hydrogen Energy*. 42 (2017) 29661–29665. <https://doi.org/10.1016/j.ijhydene.2017.10.050>.
- <sup>20</sup> P.Z. Chen, X.L. Hu, High-Efficiency Anion Exchange Membrane Water Electrolysis Employing Non-Noble Metal Catalysts, *Adv. Energy Mater.* 10 (2020) 2002285. <https://doi.org/10.1002/aenm.202002285>.
- <sup>21</sup> B. Motealleh, Z. Liu, R.I. Masel, J.P. Sculley, Z. Richard Ni, L. Meroueh, Next-generation anion exchange membrane water electrolyzers operating for commercially relevant lifetimes, *Int. J. Hydrogen Energy*. 46 (2021) 3379–3386. <https://doi.org/10.1016/j.ijhydene.2020.10.244>.
- <sup>22</sup> Y.S. Park, J. Yang, J. Lee, M.J. Jang, J. Jeong, W.S. Choi, Y. Kim, Y. Yin, M.H. Seo, Z. Chen, S.M. Choi, Superior performance of anion exchange membrane water electrolyzer: Ensemble of producing oxygen vacancies and controlling mass transfer resistance, *Appl. Catal. B Environ.* 278 (2020) 119276. <https://doi.org/10.1016/j.apcatb.2020.119276>.
- <sup>23</sup> M.J. Jang, J. Yang, J. Lee, Y.S. Park, J. Jeong, S.M. Park, J.Y. Jeong, Y. Yin, M.H. Seo, S.M. Choi, K.H. Lee, Superior performance and stability of anion exchange membrane water electrolysis: PH-controlled copper cobalt oxide nanoparticles for the oxygen evolution reaction, *J. Mater. Chem. A*. 8 (2020) 4290–4299. <https://doi.org/10.1039/c9ta13137j>.
- <sup>24</sup> L. Wang, X. Peng, W.E. Mustain, J.R. Varcoe, Radiation-grafted anion-exchange membranes: The switch from low- to high-density polyethylene leads to remarkably enhanced fuel cell performance, *Energy Environ. Sci.* 12 (2019) 1575–1579. <https://doi.org/10.1039/c9ee00331b>.
- <sup>25</sup> S. Adhikari, M.K. Pagels, J.Y. Jeon, C. Bae, Ionomers for electrochemical energy conversion & storage technologies, *Polymer*. 211 (2020) 123080. <https://doi.org/10.1016/j.polymer.2020.123080>.
- <sup>26</sup> L. Wang, M. Bellini, H.A. Miller, J.R. Varcoe, A high conductivity ultrathin anion-exchange membrane with 500+ h alkali stability for use in alkaline membrane fuel cells that can achieve 2 W cm<sup>-2</sup> at 80 °C, *J. Mater. Chem. A*. 6 (2018) 15404–15412. <https://doi.org/10.1039/c8ta04783a>.
- <sup>27</sup> A. Zhegur-Khais, F. Kubanek, U. Krewer, D.R. Dekel, Measuring the true hydroxide conductivity of anion exchange membranes, *J. Memb. Sci.* 612 (2020) 118461. <https://doi.org/10.1016/j.memsci.2020.118461>.
- <sup>28</sup> J.C. Douglin, J.R. Varcoe, D.R. Dekel, A high-temperature anion-exchange membrane fuel

- cell, *J. Power Sources Adv.* 5 (2020) 100023. <https://doi.org/10.1016/j.powera.2020.100023>.
- <sup>29</sup> J.C. Douglin, R.K. Singh, S. Haj, S. Li, J. Biemolt, N. Yan, J.R. Varcoe, G. Rothenberg, D. Dekel, A High-Temperature Anion-Exchange Membrane Fuel Cell with a Critical Raw Material-free Cathode, *Chem. Eng. J. Adv.* 8 (2021) 100153. <https://doi.org/10.1016/j.cej.2021.100153>.
- <sup>30</sup> J. Ponce-González, I. Ouachan, J.R. Varcoe, D.K. Whelligan, Radiation-induced grafting of a butyl-spacer styrenic monomer onto ETFE: The synthesis of the most alkali stable radiation-grafted anion-exchange membrane to date, *J. Mater. Chem. A.* 6 (2018) 823–827. <https://doi.org/10.1039/c7ta10222d>.
- <sup>31</sup> C. Xiao Lin, X. Qin Wang, E. Ning Hu, Q. Yang, Q. Gen Zhang, A. Mei Zhu, Q. Lin Liu, Quaternized triblock polymer anion exchange membranes with enhanced alkaline stability, *J. Memb. Sci.* 541 (2017) 358–366. <https://doi.org/10.1016/j.memsci.2017.07.032>.
- <sup>32</sup> J.Y. Jeon, S. Park, J. Han, S. Maurya, A.D. Mohanty, D. Tian, N. Saikia, M.A. Hickner, C.Y. Ryu, M.E. Tuckerman, S.J. Paddison, Y.S. Kim, C. Bae, Synthesis of Aromatic Anion Exchange Membranes by Friedel-Crafts Bromoalkylation and Cross-Linking of Polystyrene Block Copolymers, *Macromolecules.* 52 (2019) 2139–2147. <https://doi.org/10.1021/acs.macromol.8b02355>.
- <sup>33</sup> T.J. Clark, N.J. Robertson, H.A. Kostalik IV, E.B. Lobkovsky, P.F. Mutolo, H.D. Abruña, G.W. Coates, A ring-opening metathesis polymerization route to alkaline anion exchange membranes: Development of hydroxide-conducting thin films from an ammonium-functionalized monomer, *J. Am. Chem. Soc.* 131 (2009) 12888–12889. <https://doi.org/10.1021/ja905242r>.
- <sup>34</sup> S.C. Price, X. Ren, A.M. Savage, F.L. Beyer, Synthesis and characterization of anion-exchange membranes based on hydrogenated poly(norbornene), *Polym. Chem.* 8 (2017) 5708–5717. <https://doi.org/10.1039/c7py01084b>.
- <sup>35</sup> W. Chen, M. Mandal, G. Huang, X. Wu, G. He, P.A. Kohl, Highly Conducting Anion-Exchange Membranes Based on Cross-Linked Poly(norbornene): Ring Opening Metathesis Polymerization, *ACS Appl. Energy Mater.* 2 (2019) 2458–2468. <https://doi.org/10.1021/acsaem.8b02052>.
- <sup>36</sup> W. You, E. Padgett, S.N. MacMillan, D.A. Muller, G.W. Coates, Highly conductive and chemically stable alkaline anion exchange membranes via ROMP of trans-cyclooctene derivatives, *Proc. Natl. Acad. Sci. U. S. A.* 116 (2019) 9729–9734. <https://doi.org/10.1073/pnas.1900988116>.
- <sup>37</sup> M. Mandal, G. Huang, P.A. Kohl, Anionic multiblock copolymer membrane based on vinyl

- addition polymerization of norbornenes: Applications in anion-exchange membrane fuel cells, *J. Memb. Sci.* 570–571 (2019) 394–402.  
<https://doi.org/10.1016/j.memsci.2018.10.041>.
- <sup>38</sup> L. Zhu, X. Peng, S.L. Shang, M.T. Kwasny, T.J. Zimudzi, X. Yu, N. Saikia, J. Pan, Z.K. Liu, G.N. Tew, W.E. Mustain, M. Yandrasits, M.A. Hickner, High Performance Anion Exchange Membrane Fuel Cells Enabled by Fluoropoly(olefin) Membranes, *Adv. Funct. Mater.* 29 (2019). <https://doi.org/10.1002/adfm.201902059>.
- <sup>39</sup> M.R. Hibbs, Alkaline stability of poly(phenylene)-based anion exchange membranes with various cations, *J. Polym. Sci. Part B Polym. Phys.* 51 (2013) 1736–1742.  
<https://doi.org/10.1002/polb.23149>.
- <sup>40</sup> D. Li, E.J. Park, W. Zhu, Q. Shi, Y. Zhou, H. Tian, Y. Lin, A. Serov, B. Zulevi, E.D. Baca, C. Fujimoto, H.T. Chung, Y.S. Kim, Highly quaternized polystyrene ionomers for high performance anion exchange membrane water electrolyzers, *Nat. Energy.* 5 (2020) 378–385.  
<https://doi.org/10.1038/s41560-020-0577-x>.
- <sup>41</sup> H. Ono, J. Miyake, S. Shimada, M. Uchida, K. Miyatake, Anion exchange membranes composed of perfluoroalkylene chains and ammonium-functionalized oligophenylenes, *J. Mater. Chem. A* 3 (2015) 21779–21788. <https://doi.org/10.1039/c5ta06454f>.
- <sup>42</sup> R. Akiyama, N. Yokota, K. Miyatake, Chemically Stable, Highly Anion Conductive Polymers Composed of Quinquephenylene and Pendant Ammonium Groups, *Macromolecules.* 52 (2019) 2131–2138. <https://doi.org/10.1021/acs.macromol.8b02199>.
- <sup>43</sup> J. Fan, S. Willdorf-Cohen, E.M. Schibli, Z. Paula, W. Li, T.J.G. Skalski, A.T. Sergeenko, A. Hohenadel, B.J. Frisken, E. Magliocca, W.E. Mustain, C.E. Diesendruck, D.R. Dekel, S. Holdcroft, Poly(bis-arylimidazoliums) possessing high hydroxide ion exchange capacity and high alkaline stability, *Nat. Commun.* 10 (2019) 2306. <https://doi.org/10.1038/s41467-019-10292-z>.
- <sup>44</sup> S. Noh, J.Y. Jeon, S. Adhikari, Y.S. Kim, C. Bae, Molecular Engineering of Hydroxide Conducting Polymers for Anion Exchange Membranes in Electrochemical Energy Conversion Technology, *Acc. Chem. Res.* 52 (2019) 2745–2755.  
<https://doi.org/10.1021/acs.accounts.9b00355>.
- <sup>45</sup> W.-H. Lee, E.J. Park, J. Han, D.W. Shin, Y.S. Kim, C. Bae, Poly(terphenylene) Anion Exchange Membranes: The Effect of Backbone Structure on Morphology and Membrane Property, *ACS Macro Lett.* 6 (2017) 566–570.  
<https://doi.org/10.1021/acsmacrolett.7b00148>.
- <sup>46</sup> J.S. Olsson, T.H. Pham, P. Jannasch, Poly(arylene piperidinium) Hydroxide Ion Exchange



- Membranes: Synthesis, Alkaline Stability, and Conductivity, *Adv. Funct. Mater.* 28 (2018) 1702758. <https://doi.org/10.1002/adfm.201702758>.
- <sup>47</sup> H. Li, M.R. Kraglund, A.K. Reumert, X. Ren, D. Aili, J. Yang, Poly(vinyl benzyl methylpyrrolidinium) hydroxide derived anion exchange membranes for water electrolysis, *J. Mater. Chem. A* 7 (2019) 17914–17922. <https://doi.org/10.1039/c9ta04868e>.
- <sup>48</sup> J. Parrondo, C.G. Arges, M. Niedzwiecki, E.B. Anderson, K.E. Ayers, V. Ramani, Degradation of anion exchange membranes used for hydrogen production by ultrapure water electrolysis., *RSC Adv.* 4 (2014) 9875–9879. <https://doi.org/10.1039/c3ra46630b>.
- <sup>49</sup> C.G. Arges, J. Parrondo, G. Johnson, A. Nadhan, V. Ramani, Assessing the influence of different cation chemistries on ionic conductivity and alkaline stability of anion exchange membranes, *J. Mater. Chem.* 22 (2012) 3733. <https://doi.org/10.1039/c2jm14898f>.
- <sup>50</sup> J. Pan, Y. Li, L. Zhuang, J. Lu, Self-crosslinked alkaline polymer electrolyte exceptionally stable at 90 °C, *Chem. Commun.* 46 (2010) 8597–8599. <https://doi.org/10.1039/c0cc03618h>.
- <sup>51</sup> J. Pan, C. Chen, L. Zhuang, J. Lu, Designing Advanced Alkaline Polymer Electrolytes for Fuel Cell Applications, *Acc. Chem. Res.* 45 (2012) 473–481.
- <sup>52</sup> L. Xiao, S. Zhang, J. Pan, C. Yang, M. He, L. Zhuang, J. Lu, First implementation of alkaline polymer electrolyte water electrolysis working only with pure water, *Energy Environ. Sci.* 5 (2012) 7869–7871. <https://doi.org/10.1039/c2ee22146b>.
- <sup>53</sup> A. Lim, H. juhn Kim, D. Henkensmeier, S. Jong Yoo, J. Young Kim, S. Young Lee, Y.E. Sung, J.H. Jang, H.S. Park, A study on electrode fabrication and operation variables affecting the performance of anion exchange membrane water electrolysis, *J. Ind. Eng. Chem.* 76 (2019) 410–418. <https://doi.org/10.1016/j.jiec.2019.04.007>.
- <sup>54</sup> J. Hnát, M. Plevová, J. Žitka, M. Paidar, K. Bouzek, Anion-selective materials with 1,4-diazabicyclo[2.2.2]octane functional groups for advanced alkaline water electrolysis, *Electrochim. Acta.* 248 (2017) 547–555. <https://doi.org/10.1016/j.electacta.2017.07.165>.
- <sup>55</sup> N. Lee, D.T. Duong, D. Kim, Cyclic ammonium grafted poly (arylene ether ketone) hydroxide ion exchange membranes for alkaline water electrolysis with high chemical stability and cell efficiency, *Electrochim. Acta.* 271 (2018) 150–157. <https://doi.org/10.1016/j.electacta.2018.03.117>.
- <sup>56</sup> X. Wu, K. Scott, A Li-doped Co<sub>3</sub>O<sub>4</sub> oxygen evolution catalyst for non-precious metal alkaline anion exchange membrane water electrolyzers, *Int. J. Hydrogen Energy.* 38 (2013) 3123–3129. <https://doi.org/10.1016/j.ijhydene.2012.12.087>.

- <sup>57</sup> X. Wu, K. Scott, A polymethacrylate-based quaternary ammonium OH<sup>-</sup> ionomer binder for non-precious metal alkaline anion exchange membrane water electrolyzers, *J. Power Sources*. 214 (2012) 124–129. <https://doi.org/10.1016/j.jpowsour.2012.03.069>.
- <sup>58</sup> M. Faraj, M. Boccia, H. Miller, F. Martini, S. Borsacchi, M. Geppi, A. Pucci, New LDPE based anion-exchange membranes for alkaline solid polymeric electrolyte water electrolysis, *Int. J. Hydrogen Energy*. 37 (2012) 14992–15002. <https://doi.org/10.1016/j.ijhydene.2012.08.012>.
- <sup>59</sup> X.D. Su, L. Gao, L. Hu, N.A. Qaisrani, X.M. Yan, W.J. Zhang, X.B. Jiang, X.H. Ruan, G.H. He, Novel piperidinium functionalized anionic membrane for alkaline polymer electrolysis with excellent electrochemical properties, *J. Memb. Sci.* 581 (2019) 283–292. <https://doi.org/10.1016/j.memsci.2019.03.072>.
- <sup>60</sup> N. Chen, C. Lu, Y. Li, C. Long, H. Zhu, Robust poly(aryl piperidinium)/N-spirocyclic poly(2,6-dimethyl-1,4-phenyl) for hydroxide-exchange membranes, *J. Memb. Sci.* 572 (2019) 246–254. <https://doi.org/10.1016/j.memsci.2018.10.067>.
- <sup>61</sup> I. Vincent, A. Kruger, D. Bessarabov, Hydrogen Production by water Electrolysis with an Ultrathin Anion-exchange membrane (AEM), *Int. J. Electrochem. Sci.* 13 (2018) 11347–11358. <https://doi.org/10.20964/2018.12.84>.
- <sup>62</sup> I. Vincent, A. Kruger, D. Bessarabov, Development of efficient membrane electrode assembly for low cost hydrogen production by anion exchange membrane electrolysis, *Int. J. Hydrogen Energy*. 42 (2017) 10752–10761. <https://doi.org/10.1016/j.ijhydene.2017.03.069>.
- <sup>63</sup> G. Gupta, K. Scott, M. Mamlouk, Performance of polyethylene based radiation grafted anion exchange membrane with polystyrene-*b*-poly (ethylene/butylene)-*b*-polystyrene based ionomer using NiCo<sub>2</sub>O<sub>4</sub> catalyst for water electrolysis, *J. Power Sources*. 375 (2018) 387–396. <https://doi.org/10.1016/j.jpowsour.2017.07.026>.
- <sup>64</sup> X. Wu, K. Scott, F. Xie, N. Alford, A reversible water electrolyser with porous PTFE based OH<sup>-</sup> conductive membrane as energy storage cells, *J. Power Sources*. 246 (2014) 225–231. <https://doi.org/10.1016/j.jpowsour.2013.07.081>.
- <sup>65</sup> Y.C. Cao, X. Wu, K. Scott, A quaternary ammonium grafted poly vinyl benzyl chloride membrane for alkaline anion exchange membrane water electrolyzers with no-noble-metal catalysts, *Int. J. Hydrogen Energy*. 37 (2012) 9524–9528. <https://doi.org/10.1016/j.ijhydene.2012.03.116>.
- <sup>66</sup> Y.C. Cao, X. Wang, M. Mamlouk, K. Scott, Preparation of alkaline anion exchange polymer membrane from methylated melamine grafted poly(vinylbenzyl chloride) and its fuel cell

- performance, *J. Mater. Chem.* 21 (2011) 12910–12916. <https://doi.org/10.1039/c1jm12068a>.
- <sup>67</sup> E.J. Park, C.B. Capuano, K.E. Ayers, C. Bae, Chemically durable polymer electrolytes for solid-state alkaline water electrolysis, *J. Power Sources.* 375 (2018) 367–372. <https://doi.org/10.1016/j.jpowsour.2017.07.090>.
- <sup>68</sup> M.S. Cha, J.E. Park, S. Kim, S.-H. Han, S.-H. Shin, S.H. Yang, T.-H. Kim, D.M. Yu, S. So, Y.T. Hong, S.J. Yoon, S.-G. Oh, S.Y. Kang, O.-H. Kim, H.S. Park, B. Bae, Y.-E. Sung, Y.-H. Cho, J.Y. Lee, Poly(carbazole)-based anion-conducting materials with high performance and durability for energy conversion devices, *Energy Environ. Sci.* 13 (2020) 3633–3645. <https://doi.org/10.1039/d0ee01842b>.
- <sup>69</sup> H. Li, N. Yu, F. Gellrich, A.K. Reumert, M.R. Kraglund, J. Dong, D. Aili, J. Yang, Diamine crosslinked anion exchange membranes based on poly(vinyl benzyl methylpyrrolidinium) for alkaline water electrolysis, *J. Memb. Sci.* 633 (2021). <https://doi.org/10.1016/j.memsci.2021.119418>.
- <sup>70</sup> D. Aili, A.G. Wright, M.R. Kraglund, K. Jankova, S. Holdcroft, J.O. Jensen, Towards a stable ion-solvating polymer electrolyte for advanced alkaline water electrolysis, *J. Mater. Chem. A.* 5 (2017) 5055–5066. <https://doi.org/10.1039/c6ta10680c>.
- <sup>71</sup> C.G. Arges, L. Wang, J. Parrondo, V. Ramani, Best Practices for Investigating Anion Exchange Membrane Suitability for Alkaline Electrochemical Devices: Case Study Using Quaternary Ammonium Poly(2,6-dimethyl 1,4-phenylene)oxide Anion Exchange Membranes, *J. Electrochem. Soc.* 160 (2013) F1258–F1274. <https://doi.org/10.1149/2.049311jes>.
- <sup>72</sup> J. Parrondo, V. Ramani, Stability of Poly(2,6-dimethyl 1,4-phenylene)Oxide-Based Anion Exchange Membrane Separator and Solubilized Electrode Binder in Solid-State Alkaline Water Electrolyzers, *J. Electrochem. Soc.* 161 (2014) F1015–F1020. <https://doi.org/10.1149/2.0601410jes>.
- <sup>73</sup> H.J. Park, S.Y. Lee, T.K. Lee, H.-J. Kim, Y.M. Lee, N3-butyl imidazolium-based anion exchange membranes blended with Poly(vinyl alcohol) for alkaline water electrolysis, *J. Memb. Sci.* 611 (2020) 118355. <https://doi.org/10.1016/j.memsci.2020.118355>.
- <sup>74</sup> X.M. Yan, X. Yang, X.D. Su, L. Gao, J. Zhao, L. Hu, M.T. Di, T.T. Li, X.H. Ruan, G.H. He, Twisted ether-free polymer based alkaline membrane for high-performance water electrolysis, *J. Power Sources.* 480 (2020) 228805. <https://doi.org/10.1016/j.jpowsour.2020.228805>.
- <sup>75</sup> X.M. Chu, Y. Shi, L. Liu, Y.D. Huang, N.W. Li, Piperidinium-functionalized anion exchange membranes and their application in alkaline fuel cells and water electrolysis, *J.*

- Mater. Chem. A. 7 (2019) 7717–7727. <https://doi.org/10.1039/c9ta01167f>.
- <sup>76</sup> D. Aili, M.R. Kraglund, J. Tavecchi, C. Chatzichristodoulou, J.O. Jensen, Polysulfone-polyvinylpyrrolidone blend membranes as electrolytes in alkaline water electrolysis, *J. Memb. Sci.* 598 (2020) 117674. <https://doi.org/10.1016/j.memsci.2019.117674>.
- <sup>77</sup> R.E. Coppola, D. Herranz, R. Escudero-Cid, N. Ming, N.B. D’Accorso, P. Ocon, G.C. Abuin, Polybenzimidazole-crosslinked-poly(vinyl benzyl chloride) as anion exchange membrane for alkaline electrolyzers, *Renew. Energy.* 157 (2020) 71–82. <https://doi.org/10.1016/j.renene.2020.04.140>.
- <sup>78</sup> R.E. Coppola, F.N. Molinari, N.B. D’Accorso, G.C. Abuin, Polyvinyl alcohol nanofibers reinforced with polybenzimidazole: Facile preparation and properties of an anion exchange membrane, *Polym. Adv. Technol.* 32 (2021) 3505–3514. <https://doi.org/10.1002/pat.5361>.
- <sup>79</sup> H. Ono, T. Kimura, A. Takano, K. Asazawa, J. Miyake, J. Inukai, K. Miyatake, Robust anion conductive polymers containing perfluoroalkylene and pendant ammonium groups for high performance fuel cells, *J. Mater. Chem. A.* 5 (2017) 24804–24812. <https://doi.org/10.1039/c7ta09409d>.
- <sup>80</sup> S. Maurya, C.H. Fujimoto, M.R. Hibbs, C. Narvaez Villarrubia, Y.S. Kim, Toward Improved Alkaline Membrane Fuel Cell Performance Using Quaternized Aryl-Ether Free Polyaromatics, *Chem. Mater.* 30 (2018) 2188–2192. <https://doi.org/10.1021/acs.chemmater.8b00358>.
- <sup>81</sup> S. Miyanishi, T. Yamaguchi, Highly durable spirobifluorene-based aromatic anion conducting polymer for a solid ionomer of alkaline fuel cells and water electrolysis cells, *J. Mater. Chem. A.* 7 (2019) 2219–2224. <https://doi.org/10.1039/c8ta08400a>.
- <sup>82</sup> Y.K. Choe, C. Fujimoto, K.S. Lee, L.T. Dalton, K. Ayers, N.J. Henson, Y.S. Kim, Alkaline stability of benzyl trimethyl ammonium functionalized polyaromatics: A computational and experimental study, *Chem. Mater.* 26 (2014) 5675–5682. <https://doi.org/10.1021/cm502422h>.
- <sup>83</sup> D. Aili, M.K. Hansen, R.F. Renzaho, Q.F. Li, E. Christensen, J.O. Jensen, N.J. Bjerrum, Heterogeneous anion conducting membranes based on linear and crosslinked KOH doped polybenzimidazole for alkaline water electrolysis, *J. Memb. Sci.* 447 (2013) 424–432. <https://doi.org/10.1016/j.memsci.2013.07.054>.
- <sup>84</sup> L.A. Diaz, J. Hnát, N. Heredia, M.M. Bruno, F.A. Viva, M. Paidar, H.R. Corti, K. Bouzek, G.C. Abuin, Alkali doped poly (2,5-benzimidazole) membrane for alkaline water electrolysis: Characterization and performance, *J. Power Sources.* 312 (2016) 128–136. <https://doi.org/10.1016/j.jpowsour.2016.02.032>.

- <sup>85</sup> Y. Yang, T. Jiang, L. Li, S. Zhou, H. Fang, X. Li, H. Wei, Y. Ding, Chemo-stable poly(quinquephenylene-co-diphenylene piperidinium) ionomers for anion exchange membrane fuel cells, *J. Power Sources*. 506 (2021) 230184. <https://doi.org/10.1016/j.jpowsour.2021.230184>.
- <sup>86</sup> L. Li, J. Wang, M. Hussain, L. Ma, N.A. Qaisrani, S. Ma, L. Bai, X. Yan, X. Deng, G. He, F. Zhang, Side-chain manipulation of poly (phenylene oxide) based anion exchange membrane: Alkoxy extender integrated with flexible spacer, *J. Memb. Sci.* 624 (2021) 119088. <https://doi.org/10.1016/j.memsci.2021.119088>.
- <sup>87</sup> N. Chen, H.H. Wang, S.P. Kim, H.M. Kim, W.H. Lee, C. Hu, J.Y. Bae, E.S. Sim, Y.C. Chung, J.H. Jang, S.J. Yoo, Y. Zhuang, Y.M. Lee, Poly(fluorenyl aryl piperidinium) membranes and ionomers for anion exchange membrane fuel cells, *Nat. Commun.* 12 (2021) 2367. <https://doi.org/10.1038/s41467-021-22612-3>.
- <sup>88</sup> C. Long, Z. Wang, H. Zhu, High chemical stability anion exchange membrane based on poly(aryl piperidinium): Effect of monomer configuration on membrane properties, *Int. J. Hydrogen Energy*. 46 (2021) 18524–18533. <https://doi.org/10.1016/j.ijhydene.2021.02.209>.
- <sup>89</sup> J. Zhang, K. Zhang, X. Liang, W. Yu, X. Ge, M.A. Shehzad, Z. Ge, Z. Yang, L. Wu, T. Xu, Self-aggregating cationic-chains enable alkaline stable ion-conducting channels for anion-exchange membrane fuel cells, *J. Mater. Chem. A*. 9 (2021) 327–337. <https://doi.org/10.1039/d0ta11011f>.
- <sup>90</sup> X. Qiao, X. Wang, S. Liu, Y. Shen, N. Li, The alkaline stability and fuel cell performance of poly(N-spirocyclic quaternary ammonium) ionenes as anion exchange membrane, *J. Memb. Sci.* 630 (2021) 119325. <https://doi.org/10.1016/j.memsci.2021.119325>.
- <sup>91</sup> K. Firouz Tadavani, A. Abdolmaleki, M.R. Molavian, M. Zhiani, New Strategy Based on Click Reaction for Preparation of Cross-Linked Poly(Benzimidazolium-Imide) as an Anion-Exchange Membrane with Improved Alkaline Stability, *Ind. Eng. Chem. Res.* 60 (2021) 7097–7110. <https://doi.org/10.1021/acs.iecr.1c00071>.
- <sup>92</sup> Y. Yuan, T. Zhang, Z. Wang, Preparation of an Anion Exchange Membrane by Pyridine-Functionalized Polyether Ether Ketone to Improve Alkali Resistance Stability for an Alkali Fuel Cell, *Energy and Fuels*. 35 (2021) 3360–3367. <https://doi.org/10.1021/acs.energyfuels.0c03428>.
- <sup>93</sup> Y. Lu, L. Liu, Y. Pu, Y. Liu, N. Li, Z. Hu, S. Chen, Towards performance improved anion exchange membrane: Cross-linking with multi-cations oligomer modified graphene oxide, *Int. J. Hydrogen Energy*. 46 (2021) 23855–23867. <https://doi.org/10.1016/j.ijhydene.2021.04.167>.

- <sup>94</sup> D. Liu, L. Lin, Y. Xie, J. Pang, Z. Jiang, Anion exchange membrane based on poly(arylene ether ketone) containing long alkyl densely quaternized carbazole derivative pendant, *J. Memb. Sci.* 623 (2021) 119079. <https://doi.org/10.1016/j.memsci.2021.119079>.
- <sup>95</sup> F. Wang, Y. Li, C.H. Li, H. Zhu, Preparation and study of spirocyclic cationic side chain functionalized polybiphenyl piperidine anion exchange membrane, *J. Memb. Sci.* 620 (2021) 118919. <https://doi.org/10.1016/j.memsci.2020.118919>.
- <sup>96</sup> Y. Xiao, M. Zhang, D. Dong, Z. Yang, Y. Cao, K. Wang, M. Fan, Preparation of Branch Polyethyleneimine (BPEI) Crosslinked Anion Exchange Membrane Based on Poly(styrene-*b*-(ethylene-co-butylene)-*b*-styrene) (SEBS), *Macromol. Mater. Eng.* 306 (2021) 2000693. <https://doi.org/10.1002/mame.202000693>.
- <sup>97</sup> M. Kumari, J.C. Douglin, D.R. Dekel, Crosslinked quaternary phosphonium-functionalized poly(ether ether ketone) polymer-based anion-exchange membranes, *J. Memb. Sci.* 626 (2021) 119167. <https://doi.org/10.1016/j.memsci.2021.119167>.
- <sup>98</sup> M.I. Khan, X. Li, J. Fernandez-Garcia, M.H. Lashari, A. Ur Rehman, N. Elboughdiri, L. Kolsi, D. Ghernaout, Effect of Different Quaternary Ammonium Groups on the Hydroxide Conductivity and Stability of Anion Exchange Membranes, *ACS Omega.* 6 (2021) 7994–8001. <https://doi.org/10.1021/acsomega.0c05134>.
- <sup>99</sup> L. Bai, L. Ma, L. Li, A. Zhang, X. Yan, F. Zhang, G. He, Branched, Side-Chain Grafted Polyarylpiperidine Anion Exchange Membranes for Fuel Cell Application, *ACS Appl. Energy Mater.* 4 (2021) 6957–6967. <https://doi.org/10.1021/acsaem.1c01037>.
- <sup>100</sup> X. Wang, C. Lin, Y. Gao, R.G.H. Lammertink, Anion exchange membranes with twisted poly(terphenylene) backbone: Effect of the N-cyclic cations, *J. Memb. Sci.* 635 (2021) 119525. <https://doi.org/10.1016/j.memsci.2021.119525>.
- <sup>101</sup> N.C. Buggy, Y.F. Du, M.C. Kuo, K.A. Ahrens, J.S. Wilkinson, S. Seifert, E.B. Coughlin, A.M. Herring, A Polyethylene-Based Triblock Copolymer Anion Exchange Membrane with High Conductivity and Practical Mechanical Properties, *ACS Appl. Polym. Mater.* 2 (2020) 1294–1303. <https://doi.org/10.1021/acsaem.9b01182>.
- <sup>102</sup> H.J. Park, X. Chu, S.P. Kim, D. Choi, J.W. Jung, J. Woo, S.Y. Baek, S.J. Yoo, Y.C. Chung, J.G. Seong, S.Y. Lee, N. Li, Y.M. Lee, Effect of N-cyclic cationic groups in poly(phenylene oxide)-based catalyst ionomer membranes for anion exchange membrane fuel cells, *J. Memb. Sci.* 608 (2020) 118183. <https://doi.org/10.1016/j.memsci.2020.118183>.
- <sup>103</sup> Y. Wang, D. Zhang, X. Liang, M.A. Shehzad, X. Xiao, Y. Zhu, X. Ge, J. Zhang, Z. Ge, L. Wu, T. Xu, Improving fuel cell performance of an anion exchange membrane by terminal

- pending bis-cations on a flexible side chain, *J. Memb. Sci.* 595 (2020) 117483. <https://doi.org/10.1016/j.memsci.2019.117483>.
- <sup>104</sup> C. Cheng, X. He, S. Huang, F. Zhang, Y. Guo, Y. Wen, B. Wu, D. Chen, Novel self-cross-linked multi-imidazolium cations long flexible side chains triblock copolymer anion exchange membrane based on ROMP-type polybenzoxorbornadiene, *Int. J. Hydrogen Energy.* 45 (2020) 19676–19690. <https://doi.org/10.1016/j.ijhydene.2020.04.276>.
- <sup>105</sup> Z. Liu, X. Li, K. Shen, P. Feng, Y. Zhang, X. Xu, W. Hu, Z. Jiang, B. Liu, M.D. Guiver, Naphthalene-based poly(arylene ether ketone) anion exchange membranes, *J. Mater. Chem. A.* 1 (2013) 6481–6488. <https://doi.org/10.1039/c3ta10355b>.
- <sup>106</sup> A. Ouadah, H. Xu, T. Luo, S. Gao, Z. Zhang, Z. Li, C. Zhu, Synthesis of novel copolymers based on: P -methylstyrene, N, N -butylvinylimidazolium and polybenzimidazole as highly conductive anion exchange membranes for fuel cell application, *RSC Adv.* 7 (2017) 47806–47817. <https://doi.org/10.1039/c7ra06394f>.
- <sup>107</sup> J.Y. Chu, K.H. Lee, A.R. Kim, D.J. Yoo, Improved electrochemical performance of composite anion exchange membranes for fuel cells through cross linking of the polymer chain with functionalized graphene oxide, *J. Memb. Sci.* 611 (2020) 118385. <https://doi.org/10.1016/j.memsci.2020.118385>.
- <sup>108</sup> T.H. Pham, A. Allushi, J.S. Olsson, P. Jannasch, Rational molecular design of anion exchange membranes functionalized with alicyclic quaternary ammonium cations, *Polym. Chem.* 11 (2020) 6953–6963. <https://doi.org/10.1039/d0py01291b>.
- <sup>109</sup> S.H. Kim, K.H. Lee, J.Y. Chu, A.R. Kim, D.J. Yoo, Enhanced hydroxide conductivity and dimensional stability with blended membranes containing hyperbranched pae/linear ppo as anion exchange membranes, *Polymers (Basel).* 12 (2020) 3011. <https://doi.org/10.3390/polym12123011>.
- <sup>110</sup> B.H. Oh, A.R. Kim, D.J. Yoo, Profile of extended chemical stability and mechanical integrity and high hydroxide ion conductivity of poly(ether imide) based membranes for anion exchange membrane fuel cells, *Int. J. Hydrogen Energy.* 44 (2019) 4281–4292. <https://doi.org/10.1016/j.ijhydene.2018.12.177>.
- <sup>111</sup> Y. Li, J. Zhang, H. Yang, S. Yang, S. Lu, H. Wei, Y. Ding, Boosting the performance of an anion exchange membrane by the formation of well-connected ion conducting channels, *Polym. Chem.* 10 (2019) 2822–2831. <https://doi.org/10.1039/c9py00011a>.
- <sup>112</sup> J.Y. Chu, K.H. Lee, A.R. Kim, D.J. Yoo, Graphene-mediated organic-inorganic composites with improved hydroxide conductivity and outstanding alkaline stability for anion exchange membranes, *Compos. Part B Eng.* 164 (2019) 324–332.

<https://doi.org/10.1016/j.compositesb.2018.11.084>.

- <sup>113</sup> Y. Lu, X. Pan, N. Li, Z. Hu, S. Chen, Improved performance of quaternized poly(arylene ether ketone)s/graphitic carbon nitride nanosheets composite anion exchange membrane for fuel cell applications, *Appl. Surf. Sci.* 503 (2020) 144071. <https://doi.org/10.1016/j.apsusc.2019.144071>.
- <sup>114</sup> M. Mandal, G. Huang, N.U. Hassan, X. Peng, T. Gu, A.H. Brooks-Starks, B. Bahar, W.E. Mustain, P.A. Kohl, The Importance of Water Transport in High Conductivity and High-Power Alkaline Fuel Cells, *J. Electrochem. Soc.* 167 (2020) 054501. <https://doi.org/10.1149/2.0022005jes>.
- <sup>115</sup> H.S. Dang, P. Jannasch, High-Performing Hydroxide Exchange Membranes with Flexible Tetra-Piperidinium Side Chains Linked by Alkyl Spacers, *ACS Appl. Energy Mater.* 1 (2018) 2222–2231. <https://doi.org/10.1021/acsaem.8b00294>.
- <sup>116</sup> A.M.A. Mahmoud, A.M.M. Elsaghier, K. Otsuji, K. Miyatake, High Hydroxide Ion Conductivity with Enhanced Alkaline Stability of Partially Fluorinated and Quaternized Aromatic Copolymers as Anion Exchange Membranes, *Macromolecules.* 50 (2017) 4256–4266. <https://doi.org/10.1021/acs.macromol.7b00401>.
- <sup>117</sup> N. Chen, C. Lu, Y. Li, C. Long, Z. Li, H. Zhu, Tunable multi-cations-crosslinked poly(arylene piperidinium)-based alkaline membranes with high ion conductivity and durability, *J. Memb. Sci.* 588 (2019) 117120. <https://doi.org/10.1016/j.memsci.2019.05.044>.
- <sup>118</sup> C. Shang, Z. Wang, L. Wang, J. Wang, Preparation and characterization of a polyvinyl alcohol grafted bis-crown ether anion exchange membrane with high conductivity and strong alkali stability, *Int. J. Hydrogen Energy.* 45 (2020) 16738–16750. <https://doi.org/10.1016/j.ijhydene.2020.04.134>.
- <sup>119</sup> H.A.K. Iv, T.J. Clark, N.J. Robertson, P.F. Mutolo, J.M. Longo, G.W. Coates, Solvent Processable Tetraalkylammonium-Functionalized Polyethylene for Use as an Alkaline Anion Exchange Membrane, 43 (2010) 7147–7150. <https://doi.org/10.1021/ma101172a>.
- <sup>120</sup> Y. Liu, J. Wang, Y. Yang, T.M. Brenner, S. Seifert, Y. Yan, M.W. Liberatore, A.M. Herring, Anion transport in a chemically stable, sterically bulky  $\alpha$ -C modified imidazolium functionalized anion exchange membrane, *J. Phys. Chem. C.* 118 (2014) 15136–15145. <https://doi.org/10.1021/jp5027674>.
- <sup>121</sup> K. Firouz Tadavani, A. Abdolmaleki, M.R. Molavian, M. Zhiani, New Strategy Based on Click Reaction for Preparation of Cross-Linked Poly(Benzimidazolium-Imide) as an Anion-Exchange Membrane with Improved Alkaline Stability, *Ind. Eng. Chem. Res.* 60 (2021) 7097–7110. <https://doi.org/10.1021/acs.iecr.1c00071>.



- <sup>122</sup> S. Willdorf-Cohen, A.N. Mondal, D.R. Dekel, C.E. Diesendruck, Chemical stability of poly(phenylene oxide)-based ionomers in an anion exchange-membrane fuel cell environment, *J. Mater. Chem. A*. 6 (2018) 22234–22239. <https://doi.org/10.1039/C8TA05785K>.
- <sup>123</sup> L. Wang, J.J. Brink, Y. Liu, A.M. Herring, J. Ponce-González, D.K. Whelligan, J.R. Varcoe, Non-fluorinated pre-irradiation-grafted (peroxidated) LDPE-based anion-exchange membranes with high performance and stability, *Energy Environ. Sci.* 10 (2017) 2154–2167. <https://doi.org/10.1039/c7ee02053h>.
- <sup>124</sup> X. Wu, K. Scott,  $\text{Cu}_x\text{Co}_{3-x}\text{O}_4$  ( $0 \leq x < 1$ ) nanoparticles for oxygen evolution in high performance alkaline exchange membrane water electrolyzers, *J. Mater. Chem.* 21 (2011) 12344–12351. <https://doi.org/10.1039/c1jm11312g>.
- <sup>125</sup> S.H. Ahn, B.S. Lee, I. Choi, S.J. Yoo, H.J. Kim, E.A. Cho, D. Henkensmeier, S.W. Nam, S.K. Kim, J.H. Jang, Development of a membrane electrode assembly for alkaline water electrolysis by direct electrodeposition of nickel on carbon papers, *Appl. Catal. B Environ.* 154–155 (2014) 197–205. <https://doi.org/10.1016/j.apcatb.2014.02.021>.
- <sup>126</sup> I. V Pushkareva, A.S. Pushkarev, S.A. Grigoriev, P. Modisha, D.G. Bessarabov, Comparative study of anion exchange membranes for low-cost water electrolysis, *Int. J. Hydrogen Energy*. 45 (2020) 26070–26079. <https://doi.org/10.1016/j.ijhydene.2019.11.011>.
- <sup>127</sup> Y.X. Chen, A. Lavacchi, H.A. Miller, M. Bevilacqua, J. Filippi, M. Innocenti, A. Marchionni, W. Oberhauser, L. Wang, F. Vizza, Nanotechnology makes biomass electrolysis more energy efficient than water electrolysis, *Nat. Commun.* 5 (2014) 4036. <https://doi.org/10.1038/ncomms5036>.
- <sup>128</sup> L. Wang, T. Weissbach, R. Reissner, A. Ansar, A.S. Gago, S. Holdcroft, K.A. Friedrich, High Performance Anion Exchange Membrane Electrolysis Using Plasma-Sprayed, Non-Precious-Metal Electrodes, *ACS Appl. Energy Mater.* 2 (2019) 7903–7912. <https://doi.org/10.1021/acsaem.9b01392>.
- <sup>129</sup> X. Hu, Y.D. Huang, L. Liu, Q. Ju, X.X. Zhou, X.Q. Qiao, Z.F. Zheng, N.W. Li, Piperidinium functionalized aryl ether-free polyaromatics as anion exchange membrane for water electrolyzers: Performance and durability, *J. Memb. Sci.* 621 (2021) 118964. <https://doi.org/10.1016/j.memsci.2020.118964>.
- <sup>130</sup> M.R. Kraglund, M. Carmo, G. Schiller, S.A. Ansar, D. Aili, E. Christensen, J.O. Jensen, Ion-solvating membranes as a new approach towards high rate alkaline electrolyzers, *Energy Environ. Sci.* 12 (2019) 3313–3318. <https://doi.org/10.1039/c9ee00832b>.

- <sup>131</sup> A.Y. Faid, A.O. Barnett, F. Seland, S. Sunde, Highly Active Nickel-Based Catalyst for Hydrogen Evolution in Anion Exchange Membrane Electrolysis, *Catalysts*. 8 (2018). <https://doi.org/10.3390/catal8120614>.
- <sup>132</sup> L. Zeng, T.S. Zhao, Integrated inorganic membrane electrode assembly with layered double hydroxides as ionic conductors for anion exchange membrane water electrolysis, *Nano Energy*. 11 (2015) 110–118. <https://doi.org/10.1016/j.nanoen.2014.10.019>.
- <sup>133</sup> M.K. Cho, H.Y. Park, H.J. Lee, H.J. Kim, A. Lim, D. Henkensmeier, S.J. Yoo, J.Y. Kim, S.Y. Lee, H.S. Park, J.H. Jang, Alkaline anion exchange membrane water electrolysis: Effects of electrolyte feed method and electrode binder content, *J. Power Sources*. 382 (2018) 22–29. <https://doi.org/10.1016/j.jpowsour.2018.02.025>.
- <sup>134</sup> R.I. Masel, Z. Liu, S.D. Sajjad, Anion Exchange Membrane Electrolyzers Showing 1 A/cm(2) at Less Than 2 V, *ECS Trans*. 75 (2016) 1143–1146. <https://doi.org/10.1149/07514.1143ecst>.
- <sup>135</sup> H.J. Park, S.Y. Lee, T.K. Lee, H.J. Kim, Y.M. Lee, N3-butyl imidazolium-based anion exchange membranes blended with Poly alcohol) for alkaline water, *J. Memb. Sci.* 611 (2020) 118355. <https://doi.org/10.1016/j.memsci.2020.118355>.
- <sup>136</sup> J.E. Park, M.J. Kim, M.S. Lim, S.Y. Kang, J.K. Kim, S.H. Oh, M. Her, Y.H. Cho, Y.E. Sung, Graphitic carbon nitride-carbon nanofiber as oxygen catalyst in anion-exchange membrane water electrolyzer and rechargeable metal-air cells, *Appl. Catal. B-Environmental*. 237 (2018) 140–148. <https://doi.org/10.1016/j.apcatb.2018.05.073>.
- <sup>137</sup> F. Razmjooei, A. Farooqui, R. Reissner, A.S. Gago, S.A. Ansar, K.A. Friedrich, Elucidating the Performance Limitations of Alkaline Electrolyte Membrane Electrolysis: Dominance of Anion Concentration in Membrane Electrode Assembly, *ChemElectroChem*. 7 (2020) 3951–3960. <https://doi.org/10.1002/celec.202000605>.
- <sup>138</sup> Y.S. Park, J. Lee, M.J. Jang, J. Yang, J. Jeong, J. Park, Y. Kim, M.H. Seo, Z. Chen, S.M. Choi, High-performance anion exchange membrane alkaline seawater electrolysis, *J. Mater. Chem. A*. 9 (2021) 9586–9592. <https://doi.org/10.1039/d0ta12336f>.
- <sup>139</sup> R.E. Coppola, F.N. Molinari, N.B. D’Accorso, G.C. Abuin, Polyvinyl alcohol nanofibers reinforced with polybenzimidazole: Facile preparation and properties of an anion exchange membrane, *Polym. Adv. Technol.* 32 (2021) 3505–3514. <https://doi.org/10.1002/pat.5361>.
- <sup>140</sup> J.D. Joe, D.B.S. Kumar, P. Sivakumar, Production Of Hydrogen By Anion Exchange Membrane Using AWE, *Int. J. Sci. Technol. Res.* 3 (2014) 38–42.

- <sup>141</sup> J. Liu, Z. Kang, D. Li, M. Pak, S.M. Alia, C. Fujimoto, G. Bender, Y.S. Kim, A.Z. Weber, Elucidating the Role of Hydroxide Electrolyte on Anion-Exchange-Membrane Water Electrolyzer Performance, *J. Electrochem. Soc.* 168 (2021) 054522. <https://doi.org/10.1149/1945-7111/ac0019>.
- <sup>142</sup> D. Xu, M.B. Stevens, M.R. Cosby, S.Z. Oener, A.M. Smith, L.J. Enman, K.E. Ayers, C.B. Capuano, J.N. Renner, N. Danilovic, Y. Li, H. Wang, Q. Zhang, S.W. Boettcher, Earth-Abundant Oxygen Electrocatalysts for Alkaline Anion-Exchange-Membrane Water Electrolysis: Effects of Catalyst Conductivity and Comparison with Performance in Three-Electrode Cells, *ACS Catal.* 9 (2019) 7–15. <https://doi.org/10.1021/acscatal.8b04001>.
- <sup>143</sup> T. Pandiarajan, L.J. Berchmans, S. Ravichandran, Fabrication of spinel ferrite based alkaline anion exchange membrane water electrolyzers for hydrogen production, *RSC Adv.* 5 (2015) 34100–34108. <https://doi.org/10.1039/c5ra01123j>.
- <sup>144</sup> L. Zeng, T.S. Zhao, R.H. Zhang, J.B. Xu, NiCo<sub>2</sub>O<sub>4</sub> nanowires@MnOx nanoflakes supported on stainless steel mesh with superior electrocatalytic performance for anion exchange membrane water splitting, *Electrochem. Commun.* 87 (2018) 66–70. <https://doi.org/10.1016/j.elecom.2018.01.002>.
- <sup>145</sup> P. Thangavel, M. Ha, S. Kumaraguru, A. Meena, A.N. Singh, A.M. Harzandi, K.S. Kim, Graphene-nanoplatelets-supported NiFe-MOF: High-efficiency and ultra-stable oxygen electrodes for sustained alkaline anion exchange membrane water electrolysis, *Energy Environ. Sci.* 13 (2020) 3447–3458. <https://doi.org/10.1039/d0ee00877j>.
- <sup>146</sup> J.E. Park, S.Y. Kang, S.H. Oh, J.K. Kim, M.S. Lim, C.Y. Ahn, Y.H. Cho, Y.E. Sung, High-performance anion-exchange membrane water electrolysis, *Electrochim. Acta.* 295 (2019) 99–106. <https://doi.org/10.1016/j.electacta.2018.10.143>.
- <sup>147</sup> F. Razmjooei, T. Morawietz, E. Taghizadeh, E. Hadjixenophontos, L. Mues, M. Gerle, B.D. Wood, C. Harms, A.S. Gago, S.A. Ansar, K.A. Friedrich, Increasing the performance of an anion-exchange membrane electrolyzer operating in pure water with a nickel-based microporous layer, *Joule* 5 (2021) 1776–1799. <https://doi.org/10.1016/j.joule.2021.05.006>.
- <sup>148</sup> Y. Leng, G. Chen, A.J. Mendoza, T.B. Tighe, M.A. Hickner, C.Y. Wang, Solid-state water electrolysis with an alkaline membrane, *J. Am. Chem. Soc.* 134 (2012) 9054–9057. <https://doi.org/10.1021/ja302439z>.